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Role of OH radicals in the formation of oxygen molecules following vacuum ultraviolet photodissociation of amorphous solid water

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Photodesorption of O2(X 3Σg−) and O2(a 1Δg) from amorphous solid water at 90 K has been studied following photoexcitation within the first absorption band at 157 nm. Time-of-flight and rotational spectra of O2 reveal the translational and internal energy distributions, from which production mechanisms are deduced. Exothermic and endothermic reactions of OH+O(3P) are proposed as plausible formation mechanisms for O2(X 3Σg− and a 1Δg). To examine the contribution of the O(3P)+O(3P) recombination reaction to the O2 formation following 157 nm photolysis of amorphous solid water, O2 products following 193 nm photodissociation of SO2 adsorbed on amorphous solid water were also investigated. © 2010 American Institute of Physics. [doi:10.1063/1.3474999]

I. INTRODUCTION

The effect of radiation on water ice has intrigued many scientists in the fields of interstellar chemistry and planetary ice science as well as reaction dynamics since water is the predominant component of interstellar icy grain mantles in dense molecular clouds and small solar system bodies such as comets.1–3 Oxygen molecule is known to be a product when water ice is irradiated with photons, electrons, or with energetic ions.4–6 Westley et al.7,8 observed desorption of H2, O2, and H2O by a quadrupole mass spectrometer (QMS) during vacuum ultraviolet (VUV) irradiation of water ice at 35–100 K with mainly Lyman-α photons (λ ≈ 121 nm). Öberg et al.9 used a VUV lamp (118 nm ≤ λ ≤ 177 nm) to irradiate water ice at 18–100 K, and detected OH, H2, O2, and H2O as desorbing species by QMS. These experimental studies indicate that VUV photodissociation of amorphous solid water (ASW) leads to secondary reactions that result in molecular oxygen formation on or in ASW.

Various primary processes are energetically possible following photoexcitation of water ice at λ ≤ 130 nm,10–12 whereas photodissociation of H2O in the first absorption band of water ice (130–165 nm) involves mainly two primary processes.13,14

H2O + hν → H + OH,  

H2O + hν → H2 + O(1D).  

Measurements of the translational and internal energy distributions of the photoproducts generated from photodissociation of ASW allow the assessment of possible secondary reactions on/in ASW from reaction dynamics point of view. The translational and internal energy distributions of OH(v=0,1) radicals and O(1D) and O(3P) atoms were previously measured following 157 nm photodissociation of ASW.15–17 For the O(3P) production, two different formation mechanisms were proposed: the exothermic recombination reaction of OH and the photodissociation of OH on the ASW surface.17

\[
\text{OH} + \text{OH} \rightarrow \text{H}_2\text{O} + \text{O}(3\text{P}), \quad \Delta H = -67 \text{ kJ mol}^{-1}, \quad (3)
\]

\[
\text{OH} + \text{h} \nu \rightarrow \text{H} + \text{O}(3\text{P}). \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \quad \q
distributions of O₂ formed via highly exothermic recombination reaction of two O(³P), i.e., reaction (7), which was proposed as a source for O₂(X ³Σg⁻) formation in VUV photon and electron irradiations of ASW. The ³P + ³P + M → O₂(X ³Σg⁻) + M,  ΔH = 498. (7)

II. EXPERIMENTAL

A. Apparatus and preparation of ice

A photodissociation study of ASW at 90 K was carried out in an ultrahigh vacuum chamber, which was equipped with two turbomolecular pumps in tandem, a pulsed molecular beam, an excimer laser, and a dye laser. Experimental details are described elsewhere. A vacuum chamber was evacuated to a base pressure of 10⁻⁶ Torr by turbomolecular pumps. ASW was prepared with the backfilling deposition of water vapor for 60 min on an optically flat sapphire disk substrate, sputter coated with a thin polycrystalline film of Au(111) cooled by liquid nitrogen at 90 K. The exposure was typically 1500 L (1 L = 1 × 10⁻⁶ Torr s). This exposure resulted in the formation of roughly 500 ML of H₂O on the substrate if we adopt the reported experimental conversion factor of 1 ML deposition by 3 L exposure, which is sufficiently thick if we adopt the reported experimental conversion factor of 1 ML deposition by 3 L exposure, 26 which was adopted as a source for O₂(X ³Σg⁻) formation in VUV photon and electron irradiations of ASW.

ΔH (kJ mol⁻¹)

O(³P) + O(³P) + M → O₂(X ³Σg⁻) + M,  ΔH = 498. (7)

III. SPECTRAL ANALYSIS

A. Simulation of 2+1 REMPI spectra of oxygen molecules

1. O₂(X ³Σg⁻, ν = 0, 1, 2, 3)

The spectra were simulated using the program PGOPHER, 34 a program simulating the rotational structure accompanying multiphoton electronic transitions. The simulated spectra were used to estimate the maximum values of the rotational temperatures of the oxygen products. Spectroscopic parameters for the X ³Σg⁻ (ν = 0, 1) and C ³Πg (ν = 2) states of O₂ are summarized in Table I. For spectral simulations of 2+1 REMPI via the sequence bands of the
C $^3\text{II}_g(v'=v''+2)\rightarrow X ^3\Sigma_g^-(v'')$ transitions for $v''=1, 2, \text{ and } 3$, the literature constants for $C ^3\text{II}_g(v'=2)$ and $X ^3\Sigma_g^-(v''=1)$ were adopted$^{35-39}$ since no information is available on the vibrationally excited levels. Although White and Beuhler$^{30}$ reported some assignments of the $2+1$ REMPI spectrum via the $C ^3\text{II}_g(v'=2)\rightarrow X ^3\Sigma_g^-(v''=0)$ transition, the Rydberg-valence state interactions with the repulsive $H ^3\text{II}_g$ valence state resulted in highly perturbed rotational structures and large predissociation linewidths in the triplet manifold of the excited states. $O_2(X ^3\Sigma_g^-,v=0)$ via the spin forbidden REMPI $d ^3\text{II}_g(v'=2)-X ^3\Sigma_g^-(v''=0)$ transition was also included in the spectral simulation,$^{31,35}$ and the contribution to the observed $O_2(X ^3\Sigma_g^-,v=0)$ REMPI signals was found to be very small, less than 3%.

Rotational energy, $E_{\text{rot}}$, is characterized $\langle E_{\text{rot}} \rangle = k_B T_{\text{rot}}$, where $k_B$ and $T_{\text{rot}}$ are the Boltzmann constant and the rotational temperature, respectively.

2. $O_2(a ^1\Delta_g,v=0)$

Morrill et al.$^{32}$ reported the complex rotational structures in the REMPI spectra via the five $d ^3\text{II}_g(v'=0-4)$ states due to the strong perturbations between $d ^3\text{II}_g$ Rydberg and $H ^3\text{II}_g$ valence states. The spectral segmenting is the result of avoided crossings between the rotational level manifolds of $d ^3\text{II}_g(v'=1)$ and $H ^3\text{II}_g(v'=5,6,7).$ Unfortunately, we could not estimate the rotational temperature of the $O_2(a ^1\Delta_g,v=0)$ products because of these strong interations.

B. Simulation of time-of-flight spectra of desorbed oxygen molecules

Time-of-flight (TOF) spectra of photoproducts were taken as a function of time, $t$, between the photolysis and REMPI laser pulses using a delay generator (Stanford Research) to investigate the translational energies of desorbing photoproducts. The measured TOF spectra, $S(t,T_{\text{trans}})$, were fitted to a sum of flux-weighted Maxwell–Boltzmann (MB) distributions, $S_{\text{MB}}$, with coefficients $a_i$, which are defined by a translational temperature, $T_{\text{trans}}$. Details regarding the simulation of such TOF spectra have been reported previously,$^{25}$

$$S(t,T_{\text{trans}}) = \sum a_i S_{\text{MB}}(t,T_{\text{trans}}).$$

(8)

$$S_{\text{MB}}(t,T_{\text{trans}}) = r^3 r^{-4} \exp[-mr^2/2k_B T_{\text{trans}}^2],$$

(9)

where $r$ is the flight distance to the REMPI probe region. In these calculations we assume that signals come from a disk area (UV photoradiation area) with a radius of 6 mm. The MB distribution, $P_{\text{MB}}(E_{\text{trans}})$, as a function of translational energy, $E_{\text{trans}}$, is characterized by the averaged translational energy, $\langle E_{\text{trans}} \rangle = 2k_B T_{\text{trans}}^{2}$.

$$P_{\text{MB}}(E) = (k_B T_{\text{trans}}^{-2})^{-2}E_i \exp[-E_i/k_B T_{\text{trans}}].$$

(10)

Conversion from the $E_i$ distribution to the TOF distribution was performed using the Jacobian given by Zimmerman and Ho.$^{41}$ The angular distribution of the molecules photodesorbed from the ice surface was assumed to follow a cos$^n$ $\theta$ function, where $\theta$ is the surface polar coordinate. $n=0$ was adopted in the present best-fit procedures.$^{25,42}$

IV. RESULTS

A. Kinetic and rotational energy distributions of $O_2$ following 157 nm photodissociation of ASW

1. $O_2(X ^3\Sigma_g^-,v=0)$

Figure 1(a) shows a 2+1 REMPI spectrum following 157 nm photolysis of ASW at a fixed delay $t=1.5$ $\mu$s that corresponds to the peak TOF, as shown in Fig. 2. Highly perturbed rotational structures were observed in Fig. 1(a). By comparison with spectral simulation for $O_2(X ^3\Sigma_g^-,v=0)$ shown in Fig. 1(e), $T_{\text{rot}}$ was estimated to be $\sim 2000$ K at $t$ = $1.5$ $\mu$s. Figure 2 shows a typical TOF spectrum at 286.127 nm, which is characterized by $T_{\text{trans}}=4500 \pm 1000$ K and 500 $\pm 200$ K. These results are summarized in Table II. Changing the REMPI probe wavelength to other wavelengths, no discernible change in the TOF profiles was ob-
TABLE II. Translational and rotational temperatures of oxygen molecules following 157 nm photolysis of ASW and 193 nm photolysis of SO$_2$ on ASW at 90 K.

<table>
<thead>
<tr>
<th></th>
<th>Translational temperature $T_{\text{trans}}$ (K)$^*$</th>
<th>Averaged translational energy $\langle E_{\text{trans}}\rangle$ (kJ mol$^{-1}$)</th>
<th>Rotational temperature $T_{\text{rot}}$ (K)</th>
<th>Averaged rotational energy $\langle E_{\text{rot}}\rangle$ (kJ mol$^{-1}$)</th>
<th>Reaction mechanism</th>
</tr>
</thead>
<tbody>
<tr>
<td>O$_2$(X $3\Sigma^{-}_g$, $v=0$) from ASW</td>
<td>$4500 \pm 1000$ (70%)</td>
<td>$75 \pm 17$</td>
<td>$\sim 2000$ $^b$</td>
<td>$\sim 17$</td>
<td>OH+O($^3P$)</td>
</tr>
<tr>
<td></td>
<td>$500 \pm 200$ (30%)</td>
<td>$8 \pm 3$</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>O$_2$(a $^1\Delta_g$, $v=0$) from ASW</td>
<td>$2500 \pm 500$ $^c$</td>
<td>$42 \pm 8$</td>
<td></td>
<td></td>
<td>OH+O($^3P$)</td>
</tr>
<tr>
<td></td>
<td>$250 \pm 100$ $^d$</td>
<td>$4 \pm 2$</td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>O$_2$(X $3\Sigma^{-}_g$, $v=1$–3) from SO$_2$ on ASW$^e$</td>
<td>$3000 \pm 1000$ (55%)</td>
<td>$50 \pm 17$</td>
<td>$\sim 8000$ $^b$</td>
<td>$\sim 67$</td>
<td>O($^1P$)+O($^3P$)</td>
</tr>
<tr>
<td></td>
<td>$500 \pm 200$ (45%)</td>
<td>$8 \pm 3$</td>
<td></td>
<td></td>
<td></td>
</tr>
</tbody>
</table>

$^a$Percentage in the parenthesis is a contributions of each temperature component.
$^b$Measured at time-of-flight=1.5 $\mu$s.
$^c$Each contribution depends on REMPI wavelength. See text for details.
$^d$Not estimated. See text for details.
$^e$With the relative ratios of 0.2: 1: 1 for $v=1$:2:3.

We were unable to estimate a rotational temperature for the $T_{\text{trans}}=500$ K component because of weak REMPI spectrum intensity.

2. O$_2$(a $^1\Delta_g$, $v=0$)

Figures 3(a) and 3(b) show REMPI spectra following 157 nm photolysis of ASW, which were recorded at (a) $t=2.0$ $\mu$s and (b) 10.0 $\mu$s. Figure 4 shows TOF spectra at the two different REMPI wavelengths, (a) 329.965 nm for a high rotational level, and (b) 331.215 nm for a low rotational level. The TOF spectra are reproduced by two MB distributions with $T_{\text{trans}}=2500 \pm 500$ K and 250 $\pm 100$ K with different contributions: (a) 80% and 20% at 329.965 nm and (b) 15% and 85% at 331.215 nm, respectively.

B. Evolution of O$_2$ signal intensity as a function of 157 nm photolirradiation time in the ASW photolysis

1. O$_2$(X $3\Sigma^{-}_g$, $v=0$)

The solid black line in Fig. 5(a) shows the time evolution of the O$_2$(X $3\Sigma^{-}_g$, $v=0$) signal intensity at 286.127 nm recorded at $t=1.5$ $\mu$s as a function of 157 nm irradiation time on ASW. For comparison purpose, Fig. 5(a) includes the time evolution of photodesorbed O($^3P_2$) signal intensity that was recorded at the peak of TOF, i.e., $t=4.5$ $\mu$s.$^{17}$ The behaviors of O$_2$(X $3\Sigma^{-}_g$, $v=0$) and O($^3P_2$) signals are in accordance to each other. Previously reported time evolution of the OH signal due to the H$_2$O$_2$ photoprotect on the 157 nm photoirradiated ASW surface was also included in Fig. 5(a), which reflects the concentration of photogenerated H$_2$O$_2$ on the ice surface.$^{15,43}$ The O$_2$(X $3\Sigma^{-}_g$, $v=0$) REMPI signal increases faster than that of the H$_2$O$_2$ accumulated photolytically on the water ice surface.

2. O$_2$(a $^1\Delta_g$, $v=0$)

Figures 5(b) and 5(c) show time evolution curves of O$_2$(a $^1\Delta_g$, $v=0$), which were recorded at (b) $t=2.0$ $\mu$s for O$_2$(T$_{\text{trans}}=2500$ K) at a high rotational level REMPI wavelength of 329.965 nm, and (c) $t=10.0$ $\mu$s for O$_2$(T$_{\text{trans}}=250$ K) at a low rotational level REMPI wavelength of 331.215 nm. Figures 5(b) and 5(c) also show the time evolution of the secondary photoprodut H$_2$O$_2$. The curve for O$_2$($T_{\text{trans}}=2500$ K) matches to that for H$_2$O$_2$. In addition, O$_2$(T$_{\text{trans}}=250$ K) signals were observed promptly after 157 nm irradiation of H$_2$O$_2$/H$_2$O mixed ices, suggesting that the source of O$_2$(T$_{\text{trans}}=2500$ K) is photochemically produced H$_2$O$_2$ by 157 nm irradiation on ASW.

On the other hand, the O$_2$($T_{\text{trans}}=250$ K) component appeared promptly upon 157 nm irradiation of ASW, as shown in Fig. 5(c). These results suggest that the source of the
Oxygen molecules following photolysis of ice


C. O₂ products following 193 nm photolysis of SO₂/H₂O mixed ice

193 nm irradiation of neat ASW produced no measurable REMPI signals of O₂(1S, v=0) and O₂(1D, v=0). However, as SO₂ was deposited on ASW, m/z=32 signal was observed at 285.038 nm and its intensity increased, as shown in Fig. 6, i.e., O₂ was photodesorbed via reaction (7) following 193 nm photolysis of adsorbed SO₂ on ASW. Figure 1(b) shows a REMPI spectrum of O₂ following the 193 nm photoirradiation at t=1.5 μs that corresponds to the peak TOF of O₂ in Fig. 7. The TOF spectrum in Fig. 7 is characterized by Tₜₚₙ₉=3000±1000 K and 500±200 K. These results are summarized in Table II. O₂(a₁Δg, v=0) failed to be detected at 329–332 nm.

V. DISCUSSION

The 157 nm photolysis involves two primary processes; reactions (11) and (12),

\[
\text{H}_2\text{O}(\text{ads}) + h\nu(157 \text{ nm}) \rightarrow \text{H} + \text{OH}(v = 0 \text{ and } 1), \quad (11)
\]

\[
\text{H}_2\text{O}(\text{ads}) + h\nu(157 \text{ nm}) \rightarrow \text{H}_2 + \text{O}^1(\text{D}). \quad (12)
\]

In the case of 157 nm photolysis (Eₚʰₒₜₐₜ₉=762 \text{ kJ mol}^{-1}), the available energies for reaction (11) is Eₐₜ₉(11) = 220 \text{ kJ mol}^{-1} for OH(v=0), and Eₐₜ₉(12)=38 \text{ kJ mol}^{-1}. O^1(\text{D}) atoms via reaction (12) were successfully detected following 157 nm photolysis of ASW, but would play only a minor role in the present study, considering the small yield (≤0.01 for 145–185 nm) in the gas phase, and high reactivity of O^1(\text{D}) with parent H₂O molecules by collisions with ASW to produce OH or H₂O₂.

Our previous study showed that O^3(\text{P}) atoms were photodesorbed immediately after 157 nm irradiation started, and are mainly formed by recombination of two OH radicals moving on the ASW surface, i.e., reaction (13),

\[
\text{OH} + \text{OH} \rightarrow \text{H}_2\text{O} + \text{O}^3(\text{P}), \quad \Delta\text{H} = -67 \text{ kJ mol}^{-1}. \quad (13)
\]

Considering reactions (11) and (13), we will discuss three possible O₂ desorption mechanisms; exothermic and endo-

FIG. 5. (a) Time evolutions of O₂(1S, v=0) (black line) at t=1.5 μs at 286.127 nm and O³(P) (gray line) from ASW as a function of 157 nm irradiation time. The O³(P) signal was recorded at the peak time-of-flight of signal, t=4.5 μs. (b) Time evolutions of the O₂(a₁Δg, v=0) at t=2.0 μs (red line) at 329.065 nm and (c) t=10.0 μs (blue line) at 331.215 nm from ASW as a function of 157 nm irradiation time. Open circles in (a)-(c) represent the time evolution of H₂O₂ photoproducts accumulated on ASW (Refs. 15 and 43). (d) Expanded time evolution curves of the O₂(1S, v=0) (black line), O₂(a₁Δg, v=0) (blue line), and O³(P) (gray line) signal intensity as a function of 157 nm irradiation time on ASW. The signal intensity of O₂(a₁Δg, v=0) is adjusted for comparison.

FIG. 6. Time evolution curve of the vibrationally excited O₂(1S) REMPI signal at 285.038 nm as a function of SO₂ exposure time at 157 nm at t =1.5 μs.

FIG. 7. Time-of-flight spectrum of O₂ following 193 nm photolysis of SO₂ adsorbed on ASW at 90 K at 285.038 nm. The solid curves are fits to the data with Tₜₚₙ₉=3000 and 500 K.
thermic reactions of OH with O(3P), i.e., reactions (14) and (15), and two O(3P) recombination reaction, i.e., reaction (16),

\[ \Delta H \ (kJ \ mol^{-1}) \]

\[ \text{OH} + \text{O}(3P) \rightarrow \text{O}_2(3\Sigma_g^-) + \text{H}, \quad -68, \quad (14) \]

\[ \text{OH} + \text{O}(3P) \rightarrow \text{O}_2(a \ 1\Delta_g) + \text{H}, \quad 26, \quad (15) \]

\[ \text{O}(3P) + \text{O}(3P) + \text{M} \rightarrow \text{O}_2(3\Sigma_g^+) + \text{M}, \quad -498. \quad (16) \]

A. \( \text{O}_2(3\Sigma_g, v=0) \) formation mechanisms

Figure 5(a) shows that the temporal signal behaviors of \( \text{O}_2(3\Sigma_g, v=0) \) and O(3P) are in fair agreement to each other, while OH signal intensity due to the H2O2 photoproducts on the 157 nm photoirradiated ASW surface increased slowly. These results indicate that \( \text{O}_2(3\Sigma_g, v=0) \) are not produced from the accumulated H2O2 photoproducts on the ASW surface.

On the assumptions that (a) a large amount of OH radicals are produced on the ASW surface after 157 nm laser shots on ASW, allowing successive surface reactions on the ASW surface to occur efficiently, and (b) the concentration of OH on ASW fed by primary photodissociation reaction (11) is much higher than that of O(3P) formed via secondary reaction (13), the appearance behavior of \( \text{O}_2(3\Sigma_g, v=0) \) signal intensity in Fig. 5(a) can be explained if the reaction schemes (17)–(21) proceeds on ASW,

\[ \text{H}_2\text{O} + h\nu(157 \text{ nm}) \rightarrow \text{H} + \text{OH}(v = 0 \text{ and } 1), \quad (17) \]

\[ \text{OH} + \text{OH} \rightarrow \text{H}_2\text{O}_2, \quad (18) \]

\[ \text{OH} + \text{OH} \rightarrow \text{H}_2\text{O} + \text{O}(3P), \quad (19) \]

\[ \text{OH} + \text{O}(3P) \rightarrow \text{O}_2(3\Sigma_g^-, v = 0) + \text{H}, \quad (20) \]

\[ \text{H}_2\text{O}_2 + h\nu(157 \text{ nm}) \rightarrow 2\text{OH}(v = 0 \text{ and } 1). \quad (21) \]

Furthermore, after successive 157 nm irradiation on ASW, OH radicals are accumulated on ASW since they can exist on/in water ice at 90 K.\(^3\) These adsorbed OH on ASW can be additional sources for O(3P) via reactions (19) and (22),\(^17\) and contribute to the \( \text{O}_2(3\Sigma_g, v=0) \) production, as shown in Fig. 5(a),

\[ E_{\text{avail}} \ (kJ \ mol^{-1}) \]

\[ \text{OH}(\text{ads}) + h\nu(157 \text{ nm}) \rightarrow \text{H} + \text{O}(3P), \quad 278. \quad (22) \]

The translational temperature of \( \text{O}_2(3\Sigma_g^-, v=0) \) is much higher than the substrate temperature of 90 K, as shown in Table 1. These results suggest that observed \( \text{O}_2(3\Sigma_g^-, v=0) \) originate only from the ASW surface, not from the bulk. Kimmel and their co-workers\(^35\) suggested that O2 formation occurred at or near the ASW/vacuum interface in the low-energy electron-stimulated desorption experiments. The electron-stimulated migration of OH or OH\(^-\) to the vacuum interface, where they react and produce O2, occurs via transport through the hydrogen bond network of the ASW.\(^35\)

Arasa et al. performed molecular dynamics simulations of the UV photodissociation of ASW at 10–90 K. They reported that OH desorbs into the vacuum only when they were generated from the top 3 ML of ASW at 90 K and its probability is <0.05 per absorbed UV photon, whereas most of OH are trapped on/in ASW.\(^40\) Anderson et al.\(^57,48\) predicted that OH radicals in bulk ASW formed following photodissociation of ASW move only less than 0.5 nm, while OH radicals formed in the top 3 ML of ASW were able to move up to more than 6 nm until they were trapped. This large mobility of OH on the top ASW makes secondary reactions of OH with adsorbed species possible. These models support our proposed reaction schemes.

B. \( \text{O}_2(a \ 1\Delta_g, v=0) \) formation mechanisms

The time evolution curve for \( \text{O}_2(a \ 1\Delta_g, T_{\text{trans}}=250 \text{ K}) \) recorded at \( t=10.0 \mu s \) appears promptly, as shown in Fig. 5(c). Furthermore, Fig. 5(d) shows that \( \text{O}_2(a \ 1\Delta_g, v=0), \) O(3P), and \( \text{O}_2(3\Sigma_g^-, v=0) \) were photodesorbed immediately after the 157 nm irradiation started. Since vibrationally excited OH(v = 1) are produced from primary photodissociation of H2O2,\(^15\) \( \text{O}_2(a \ 1\Delta_g, T_{\text{trans}}=250 \text{ K}) \) can be produced via reaction of vibrationally excited OH(v = 1) and O(3P), as Lunt et al. observed in the gas phase, i.e., reaction (23),\(^23\)

\[ \Delta H \ (kJ \ mol^{-1}) \]

\[ \text{OH}(v = 1) + \text{O}(3P) \rightarrow \text{O}_2(a \ 1\Delta_g, v = 0) + \text{H}, \quad -17. \quad (23) \]

On the other hand, Fig. 5(b) shows that \( \text{O}_2(a \ 1\Delta_g, T_{\text{trans}}=2500 \text{ K}) \) was not observed immediately after 157 nm photoirradiation, and the appearance behavior is in accordance with the concentration of photogenerated H2O2 on the ice surface. Furthermore, the \( \text{O}_2(a \ 1\Delta_g, T_{\text{trans}}=2500 \text{ K}) \) was observed promptly after 157 nm irradiation of the H2O2/H2O mixed ices. These results suggest that H2O2 photoproduction on ASW is a source for the \( \text{O}_2(a \ 1\Delta_g, T_{\text{trans}}=2500 \text{ K}) \) formation. Hama et al.\(^15\) previously reported that \( \text{OH}(v = 0 \text{ and } 1), T_{\text{trans}}=7500 \text{ K} \) are produced from the secondary photodissociation of H2O2 that is photolytically accumulated on the ASW surface by 157 nm irradiation, i.e., reaction (24),

\[ \text{H}_2\text{O}_2(\text{ads}) + h\nu(157 \text{ nm}) \]

\[ \rightarrow 2\text{OH}(v = 0 \text{ and } 1), T_{\text{trans}}=7500 \text{ K}, \quad (24) \]

where \( E_{\text{avail}}(24) \) is 493 kJ mol\(^{-1}\). These OH(v = 0 and 1, \( T_{\text{trans}}=7500 \text{ K} \) can proceed in endothermic reactions on ASW surface.\(^6\) Thus, \( \text{O}_2(a \ 1\Delta_g, T_{\text{trans}}=2500 \text{ K}) \) can be produced from the hot photoproducits such as OH from H2O2 via reaction (25),

\[ \text{OH}(v = 0 \text{ and } 1), T_{\text{trans}}=7500 \text{ K} + \text{O}(3P) \]

\[ \rightarrow \text{O}_2(a \ 1\Delta_g, T_{\text{trans}}=2500 \text{ K}) + \text{H}. \quad (25) \]

Lunt et al.\(^23\) reported 0.025 for the yield of \( \text{O}_2(a \ 1\Delta_g) \) formation via reaction of OH(v ≥ 1) with O(3P) in the gas phase, and the majority of the products are \( \text{O}_2(X \ 3\Sigma_g^-) \). Lique et al.\(^21\) pointed out a possibility of the nonadiabatic transition for the OH+O reaction. The potential energy surface of OH+O reaction corresponding to the electronically excited \( 2A' \) state correlates only with the \( \text{O}_2(a \ 1\Delta_g) + \text{H} \) product channel through the electronically excited state of the
HO$_2$(2$A^\pi$) complex, but nonadiabatic electronic relaxation of this complex to the lowest 2$A^\pi$ state can provide O$_2$(X 3$\Sigma_g^-$) +H product channel. 21 Our previous observations of vibrational distribution of OH($v=1/v'=0$) produced via reactions (17) and (24) were both 0.2$\pm$0.1. 15 These previously reported results suggest that the relative product ratio of O$_2(a^1\Delta_g,v=0)$ would be small in the present study.

C. Other possible O$_2$ formation mechanisms

To consider a contribution of highly exothermic recombination reaction of two O(3$P$), we measured the translational and internal energy distributions of O$_2$ following 193 nm photoexcitation of SO$_2$ adsorbed on ASW since O(3$P$) were efficiently produced from the 193 nm photolysis of SO$_2$. 49 As shown in Figs. 1(b), 6, and 7, 193 nm photolysis of SO$_2$ on ASW yielded O$_2$ products via subsequent reactions (26) and (27) and they were desorbed into the vacuum,

\[ E_{\text{avail}} \text{(kJ mol}^{-1}) \]

\[ \begin{align*}
\text{SO}_2(\text{ads}) + h\nu(193 \text{ nm}) & \rightarrow \text{SO} + \text{O}(3\text{P}), \quad 43 \tag{26} \\
\Delta H \text{(kJ mol}^{-1}) & \\
\text{O}(3\text{P}) + \text{O}(3\text{P}) + \text{M} & \rightarrow \text{O}_2(\text{X 3}\Sigma_g^-) + \text{M}, \quad -498. \tag{27}
\end{align*} \]

The available energy for reaction (26) at 193 nm ($E_{\text{photon}} =620 \text{ kJ mol}^{-1}$) was calculated using dilute solution phase thermodynamic data for the adsorbed SO$_2$(ads). 20 The REMPI spectrum of Fig. 1(b) may be simulated by a sum of the sequence bands of the $C^3\Pi_u(v' =v'+2)-X^3\Sigma_g^-(v)$ transitions mainly for $v'=2$ and 3 [Fig. 1(g)]. In addition, O$_2(a^1\Delta_g,v=0)$ was failed to be detected at 329–332 nm. These results indicate that highly exothermic reaction (27) yields mainly vibrationally excited O$_2$(X 3$\Sigma_g^-$, $v=2$ and 3), and plays only a minor role in the formation of O$_2$(X 3$\Sigma_g^-$, $v=0$) and O$_2(a^1\Delta_g,v=0)$ from neat ASW photolysis at 157 nm.

Hydroperoxyl radical, HO$_2$, is known to be photodissociated to OH and O(1$D$ or 3$P$) at excitation wavelengths of 248–147 nm in the gas phase. 50–52 Our previous studies showed no contribution of HO$_2$ photoproducts on ASW to production of OH and O(1$D$ and 3$P$) following 157 nm photolysis of ASW, suggesting that the amount of HO$_2$ photoproducts on the ice surface was negligible at 90 K. Therefore, the O$_2$ formation following reactions of HO$_2$+H or HO$_2$+OH are considered to be unlikely sources for the observed O$_2$ ejection from the ASW surface in the present experiments. 35,53,54 Theoretical and experimental studies reported that the probabilities for the photodissociation processes of HO$_2$ to H+O$_2$ fragments were estimated to be small, since this photoprocesses are prohibited by potential barriers. 52,55

VI. CONCLUSION

We have measured time-of-flight and rotational spectra of the photodesorbed O$_2$(X 3$\Sigma_g^-$, $v=0$) and O$_2(a^1\Delta_g,v=0)$ following pulsed 157 nm irradiation of ASW at 90 K, and the translational and internal energy distributions were obtained. Measured translational energies of O$_2$(X 3$\Sigma_g^-$, $v=0$) and O$_2(a^1\Delta_g,v=0)$ suggested that the observed O$_2$(X 3$\Sigma_g^-$, $v=0$) and O$_2(a^1\Delta_g,v=0)$ fragments originate only from the ASW surface, not from the ice bulk. The exothermic barrierless reaction of OH and O(3$P$) is proposed as the most plausible source for O$_2$(X 3$\Sigma_g^-$, $v=0$) formation. The reaction of translationally and internally excited OH with O(3$P$) plays a key role in the O$_2(a^1\Delta_g,v=0)$ formation. The contribution of recombination reaction of two O(3$P$) to O$_2$ formation was found to be negligible. O(3$P$) would react with OH much more frequently than they encounter another O(3$P$) to produce O$_2$ since concentration of primary product OH on ASW is much higher than that of O(3$P$) formed via secondary reaction on ASW.

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