



Title	Structural determination of vanillin, isovanillin and ethylvanillin by means of gas electron diffraction and theoretical calculations
Author(s)	Egawa, Toru; Kameyama, Akiyo; Takeuchi, Hiroshi
Citation	Journal of Molecular Structure, 794(1-3), 92-102 https://doi.org/10.1016/j.molstruc.2006.01.042
Issue Date	2006-08-07
Doc URL	http://hdl.handle.net/2115/14773
Type	article (author version)
Additional Information	There are other files related to this item in HUSCAP. Check the above URL.
File Information	JMS794-1-3.pdf (本文)



[Instructions for use](#)

Structural determination of vanillin, isovanillin and ethylvanillin by means of gas electron diffraction and theoretical calculations

Toru Egawa^{*}, Akiyo Kameyama, Hiroshi Takeuchi

*Division of Chemistry, Graduate School of Science, Hokkaido University,
Sapporo 060-0810, Japan*

(Received)

Abstract

The molecular structures of vanillin (4-hydroxy-3-methoxybenzaldehyde), isovanillin (3-hydroxy-4-methoxybenzaldehyde) and ethylvanillin (3-ethoxy-4-hydroxybenzaldehyde) were determined by means of gas electron diffraction. Among them, vanillin and ethylvanillin have a vanilla odor but isovanillin smells differently. The nozzle temperatures were 125 °C, 173 °C and 146 °C, for vanillin, isovanillin and ethylvanillin, respectively. The results of MP2 and B3LYP calculations with the 6-31G** basis set were used as supporting information. The MP2 calculations predicted that vanillin and isovanillin have two stable conformers and ethylvanillin has four stable conformers. The electron diffraction data were found to be consistent with these conformational compositions. The determined structural parameters (r_g and \angle_α) of vanillin are as follows: $\langle r(\text{C}-\text{C})_{\text{ring}} \rangle = 1.397(4) \text{ \AA}$; $r(\text{C}_1-\text{C}_{\text{aldehyde}}) = 1.471(\leftarrow) \text{ \AA}$; $r(\text{C}_3-\text{O}_{\text{Me}}) = 1.374(9) \text{ \AA}$; $r(\text{C}_4-\text{O}_{\text{H}}) = 1.361(\leftarrow) \text{ \AA}$; $r(\text{O}-\text{C}_{\text{Me}}) = 1.428(\leftarrow) \text{ \AA}$; $r(\text{C}=\text{O}) = 1.214(8) \text{ \AA}$; $\langle r(\text{C}-\text{H}) \rangle = 1.110(11) \text{ \AA}$; $r(\text{O}-\text{H}) = 0.991(\leftarrow) \text{ \AA}$; $\angle \text{C}_6-\text{C}_1-\text{C}_2 = 120.6(2)^\circ$; $\angle \text{C}_1-\text{C}_2-\text{C}_3 = 118.8(\leftarrow)^\circ$; $\angle \text{C}_1-\text{C}_6-\text{C}_5 = 120.1(\leftarrow)^\circ$; $\angle \text{C}_2-\text{C}_1-\text{C}_{\text{aldehyde}} = 122.7(18)^\circ$; $\angle \text{C}_1-\text{C}=\text{O} = 119.4(16)^\circ$; $\angle \text{C}_4-\text{C}_3-\text{O}_{\text{Me}} = 112.2(12)^\circ$; $\angle \text{C}_3-\text{C}_4-\text{O}_{\text{H}} = 119.1(\leftarrow)^\circ$; $\angle \text{C}_3-\text{O}-\text{C} = 121.7(29)^\circ$. Those of isovanillin are as follows: $\langle r(\text{C}-\text{C})_{\text{ring}} \rangle = 1.402(4)$

^{*} Corresponding author. Phone: +81-11-706-3506; Fax: +81-11-706-4924.

E-mail address: egawa@sci.hokudai.ac.jp

Å; $r(\text{C}_1\text{-C}_{\text{aldehyde}}) = 1.479(\leftarrow)$ Å; $r(\text{C}_4\text{-O}_{\text{Me}}) = 1.369(9)$ Å; $r(\text{C}_3\text{-O}_{\text{H}}) = 1.357(\leftarrow)$ Å; $r(\text{O-C}_{\text{Me}}) = 1.422(\leftarrow)$ Å; $r(\text{C=O}) = 1.221(9)$ Å; $\langle r(\text{C-H}) \rangle = 1.114(14)$ Å; $r(\text{O-H}) = 0.995(\leftarrow)$ Å; $\angle \text{C}_6\text{-C}_1\text{-C}_2 = 120.2(3)^\circ$; $\angle \text{C}_1\text{-C}_2\text{-C}_3 = 119.0(\leftarrow)^\circ$; $\angle \text{C}_1\text{-C}_6\text{-C}_5 = 119.9(\leftarrow)^\circ$; $\angle \text{C}_2\text{-C}_1\text{-C}_{\text{aldehyde}} = 124.6(25)^\circ$; $\angle \text{C}_1\text{-C=O} = 121.3(24)^\circ$; $\angle \text{C}_3\text{-C}_4\text{-O}_{\text{Me}} = 114.4(12)^\circ$; $\angle \text{C}_4\text{-C}_3\text{-O}_{\text{H}} = 121.2(\leftarrow)^\circ$; $\angle \text{C}_4\text{-O-C} = 123.8(26)^\circ$. Those of ethylvanillin are as follows: $\langle r(\text{C-C})_{\text{ring}} \rangle = 1.397(6)$ Å; $r(\text{C}_1\text{-C}_{\text{aldehyde}}) = 1.471(\leftarrow)$ Å; $r(\text{C}_3\text{-O}_{\text{Et}}) = 1.365(13)$ Å; $r(\text{C}_4\text{-O}_{\text{H}}) = 1.352(\leftarrow)$ Å; $r(\text{O-C}_{\text{Et}}) = 1.427(\leftarrow)$ Å; $r(\text{C-C})_{\text{Et}} = 1.494(21)$ Å; $r(\text{C=O}) = 1.206(9)$ Å; $\langle r(\text{C-H}) \rangle = 1.109(10)$ Å; $r(\text{O-H}) = 0.990(\leftarrow)$ Å; $\angle \text{C}_6\text{-C}_1\text{-C}_2 = 120.2(3)^\circ$; $\angle \text{C}_1\text{-C}_2\text{-C}_3 = 118.4(\leftarrow)^\circ$; $\angle \text{C}_1\text{-C}_6\text{-C}_5 = 119.7(\leftarrow)^\circ$; $\angle \text{C}_2\text{-C}_1\text{-C}_{\text{aldehyde}} = 121.7(21)^\circ$; $\angle \text{C}_1\text{-C=O} = 128.8(22)^\circ$; $\angle \text{C}_4\text{-C}_3\text{-O}_{\text{Et}} = 112.8(14)^\circ$; $\angle \text{C}_3\text{-C}_4\text{-O}_{\text{H}} = 119.6(\leftarrow)^\circ$; $\angle \text{C}_3\text{-O-C} = 115.1(27)^\circ$; $\angle \text{O-C-C}_{\text{Et}} = 102.7(28)^\circ$. Angle brackets denote average values; parenthesized values are the estimated limits of error (3σ) referring to the last significant digit; left arrows in the parentheses mean that these parameters are bound to the preceding one.

: Vanillin; Isovanillin; Ethylvanillin; Molecular structure; Gas electron diffraction; MP2 calculations; DFT calculations

1. Introduction

Recently, we have been focusing on the structure determination of some bioactive compounds including odorant molecules and have investigated the geometrical structures and conformational properties of minty compounds, menthol and carvone, as well as isomenthol, that is the non-minty isomer of menthol, by means of gas electron diffraction [1, 2]. The conformation of carvone has been further studied by laser-jet spectroscopy [3]. Vanillin (4-hydroxy-3-methoxybenzaldehyde) and related molecules, isovanillin (3-hydroxy-4-methoxybenzaldehyde) and ethylvanillin (3-ethoxy-4-hydroxybenzaldehyde), have been chosen as the second group of targets of

this project in the present study (see Fig. 1). Vanillin is included in the vanilla bean extract and has a characteristic vanilla odor and ethylvanillin has a stronger vanilla odor than vanillin. On the other hand, isovanillin has almost no, if any, odor of vanilla.

The recognition of vanillin and ethylvanillin by their olfactory receptors has been investigated by Touhara and co-workers recently [4]. They identified the receptor of eugenol (a compound with a odor of clove), mOR-EG, and that of ethylvanillin, mOR-EV, and measured the sensitivity of various odorant molecules that have the structural resemblance to ethylvanillin and eugenol, to these receptors. Their results showed that vanillin and isovanillin are recognized by mOR-EG and mOR-EV differently. Vanillin is more sensitive than ethylvanillin to mOR-EG, but the order is reversed for their sensitivity with mOR-EV. No other compound than vanillin and ethylvanillin was recognized by the ethylvanillin receptor, mOR-EV [4].

Although it has been pointed out that the stereochemical structure of odorant molecule is the essential factor in the molecular recognition by olfactory receptors, most of the discussions are not based on the experimentally determined geometrical structures of the odorant molecules, and it is expected that the reliable molecular structures of them will contribute to the investigation of the molecular recognition by the receptors.

As for vanillin and isovanillin, X-ray diffraction studies of crystal have been reported [5, 6]. However, the solid-phase structures are not suitable for the study of the structure-function relationship of bioactive molecules, because they are subject to the distortion caused by the packing effect. In the present study, the molecular structures and conformation of the title compounds have been investigated by means of gas electron diffraction and theoretical calculations in order to provide the "distortion-free" structural parameters.

2. Experimental

The samples of vanillin, isovanillin and ethylvanillin with purity of 99% were

purchased from Aldrich Chemical Co. and were used without further purification. Electron diffraction patterns were recorded on 8 in. \times 8 in. Kodak projector slide plates with an apparatus equipped with an r^3 -sector [7]. The camera distance was about 244 mm to cover the s -range sufficient for the molecules of this size. These samples have insufficient vapor pressure for the electron diffraction experiment at room temperature, in spite of their distinct smells. Therefore, the samples were heated by using the nozzle reported in Ref. [8]. The acceleration voltage of incident electrons was about 37 kV and the electron wavelength was calibrated to the r_a (C=S) distance of CS₂ (1.5570 Å) [9]. Other experimental conditions are summarized in Table 1. The photographic plates were developed for 4.5 min. in a Dektol developer diluted 1:1. The photometry process was described in detail elsewhere [10]. The experimental intensities and backgrounds are available as Supplementary Information (Table S1).

Elastic atomic scattering factors were calculated as described in Ref. [11], and inelastic ones were taken from Ref. [12]. The experimental molecular scattering intensities are shown in Fig. 2 with the final calculated ones. A diagonal weight matrix was used in the least-squares analysis on the molecular scattering intensities.

3. Theoretical calculations

Possible conformers. The planar skeleton can be assumed for the case of vanillin and isovanillin, and hence, conformational variation is brought about by the combination of the internal rotation by 180° of the three single bonds, C_{ring}-CHO, C_{ring}-OMe and C_{ring}-OH. The geometrical optimizations of the various conformers of vanillin at the MP2(frozen core)/6-31G** level of theory revealed that the energies of the two conformers shown in Fig. 1 are lower than those of other conformers by more than 20 kJ/mol. This is fully consistent with the results of the calculations by using the MP2 and B3LYP methods with the similar basis set, 6-31G* reported by Velcheva and Stamboliyska recently [13]. These results seem reasonable because the methoxy and

hydroxyl groups of these conformers are oriented so that an intramolecular hydrogen bond is formed between the hydroxyl H and the methoxy O atoms. Therefore the possibility of the other conformers was ruled out and only these conformers were considered in the following analysis. They are labeled *s-cis* and *s-trans* according to the orientation of the aldehyde group with respect to the methoxy group (see Fig. 1). The conformers of isovanillin have been labeled similarly. In the case of ethylvanillin, both of the *s-cis* and *s-trans* forms are further classified according to the orientation of the ethyl C–C bond (*trans* and *gauche*) caused by the internal rotation of the ethyl group. Therefore, there can be four conformers in total and they are labeled *s-cis-trans*, *s-trans-trans*, *s-cis-gauche* and *s-trans-gauche*. All the possible conformers are shown in Fig. 1.

MP2 and B3LYP calculations. Geometry optimizations of the conformers shown in Fig. 1 were carried out at the MP2(frozen core)/6-31G** level of theory. Program GAUSSIAN 98 [14] was used. The obtained geometrical parameters and energies are listed in Tables 2 to 4. The relative abundance of the conformers at the nozzle temperature for each molecule was estimated by using the obtained energies as listed in these Tables. It is worth noting that the energy difference between the *s-trans* and *s-cis* conformers of isovanillin (0.08 kJ/mol) is significantly smaller than the corresponding values of vanillin and ethylvanillin (4.73 to 5.04 kJ/mol). This will be discussed later.

The vibrational calculations for all the possible conformers shown in Fig.1 were carried out with the B3LYP method and the 6-31G** basis set. The obtained Cartesian force constants were used for the following normal vibration analysis.

4. Analyses

Normal vibration analysis. The Cartesian force constants obtained by the B3LYP/6-31G** calculations were transformed into the force constants, f_{ij} , for the

internal coordinates. The theoretical f_{ij} 's of vanillin were modified by the scaling method so as to reproduce the experimental vibrational wavenumbers obtained from the IR spectrum measured in the CCl_4 solution [15]. The linear scaling formula f_{ij} (scaled) = $(c_i c_j)^{1/2} f_{ij}$ (unscaled) was used where c_i is a scale factor [16]. The definitions of internal coordinates with the resultant scale factors are listed in Table S2 of Supplementary Information. The observed and calculated vibrational wavenumbers are listed in Table S3 of Supplementary Information. The obtained scale factors, c_i , of vanillin were used to calculate the force constants, f_{ij} , of isovanillin and ethylvanillin. The scale factor c of 0.92 was assumed for the ethoxy group related internal coordinates of ethylvanillin, for which there is no corresponding value of vanillin.

Analysis of electron diffraction data. The following treatments were adopted on the results of the MP2/6-31G** geometry optimizations in order to reduce the number of adjustable parameters: (1) planar skeletons were assumed with the exception of the *s-cis-gauche* and *s-trans-gauche* conformers of ethylvanillin, for which the planarity was not assumed for the ethyl group; (2) the differences among the ring C–C bonds and the C₁–C₇ single bonds were set equal to their theoretical values; (3) the differences among the three C–O single bonds were set equal to their theoretical values; (4) the differences among all the C–H and O–H bonds were set equal to their theoretical values; (5) the difference between the C_{ring}–C_{ring}–O_{methoxy/ethoxy} and C_{ring}–C_{ring}–O_{hydroxyl} angles was set equal to their theoretical values; (6) the differences among the three angles in the ring, C₂–C₁–C₆, C₁–C₂–C₃ and C₁–C₆–C₅, were set equal to their theoretical values; (7) the C–C(=O)–H angle in the aldehyde group was assumed to be equal to that of benzaldehyde, 115.1° [17]; (8) the C_{ring}–O–H angle was assumed to be equal to that of phenol, 106.4° [18]; (9) the C_{ring}–H bonds were assumed to bisect the C–C–C angle; (10) the angles in the methyl or ethyl group containing the C–H bond were set equal to their theoretical values. For the *s-cis-gauche* and *s-trans-gauche* conformers of ethylvanillin, the additional treatments were adopted as follows: (11) the ethyl torsional angle,

C₃–O₉–C₁₀–C₁₁, was set equal to their theoretical values; (12) the torsional angle, C₂–C₃–O₉–C₁₀, was set equal to their theoretical values. The assumption (11) had to be adopted because a reasonable convergence of the least-squares fitting could not be obtained when this dihedral angle was varied, contrary to our initial expectation to determine it. The independent parameters and the constraints are summarized in Tables 5 to 7.

Mean amplitudes, l , and shrinkage corrections, $r_a - r_\alpha$ [19], were calculated from the above-mentioned scaled force constants. The model of small-amplitude vibrations was adopted. The mean amplitudes were adjusted in groups. The groups were separated according to the r_a distances of the atomic pairs. The differences among the mean amplitudes in each group were fixed at the calculated values. Table S4 of Supplementary Information list the mean amplitudes with the corresponding r_a distances. The anharmonicity parameters, κ [20], for bonded atom pairs were estimated in a diatomic approximation, $\kappa = (a/6)l^4$ [21], where the Morse parameter, a , was assumed to be 2.0 Å⁻¹. Those for nonbonded atom pairs were assumed to be zero.

5. Results and discussions

All the effort to determine the *s-cis* over *s-trans* abundance ratio of vanillin (trying various set of constraints, for example) were in vain: the uncertainty of the abundance (3σ) exceeded the absolute value of the abundance of the *s-trans*, or the abundance of the *s-trans* resulted in a negative value. So, it was decided to fix the *s-cis* over *s-trans* ratio of vanillin as well as isovanillin at the theoretical value obtained from the MP2/6-31G** calculations (81/19 for vanillin and 51/49 for isovanillin, see Tables 2 and 3). On the other hand, a significant result was obtained in determining the *trans* vs. *gauche* abundance ratio of ethylvanillin when the ratios, *s-cis-trans* over *s-trans-trans* and *s-cis-gauche* over *s-trans-gauche*, were fixed to the theoretical values listed in Table 4 (45/12 and 35/8, respectively) and then the ratio (*s-cis-trans* + *s-trans-trans*) over

(*s-cis-gauche* + *s-trans-gauche*) was varied. This treatment is equivalence of assuming the energy differences caused by the positional change of the aldehyde group of ethylvanillin from the *s-cis* to *s-trans* to be equal to the MP2 values, as in the case of vanillin and isovanillin. The obtained abundances of the *s-cis-trans*, *s-trans-trans*, *s-cis-gauche* and *s-trans-gauche* conformers of ethylvanillin are 35%, 9%, 45% and 11%, respectively. The uncertainty for the abundance of *s-cis-trans* + *s-trans-trans* (44% in total) is $\pm 27\%$.

The MP2/6-31G** calculations predicted the relative stability for the two conformers of isovanillin quite differently from that of vanillin. The calculated energy difference between the *s-trans* and *s-cis* conformers of vanillin is about 5 kJ/mol and that of isovanillin is less than 0.1 kJ/mol (see Tables 2 and 3) and consequently, the predicted abundance ratio of the conformers for the latter is about 1:1. Although these conformational properties are not supported experimentally but totally theoretical, it is worthwhile to try giving them a qualitative (and speculative) explanation by means of a quite simple model. According to the atomic charge distribution obtained from the natural population analysis [22], the electric dipole of the aldehyde group of vanillin and isovanillin lies commonly on the C₇=O₈ bond with the positive charge end at the position of the C₇ atom. On the other hand, the electric dipole of the rest of the atoms in vanillin has a comparable magnitude and it lies approximately on the C₃-C₄ bond with the positive end at the position of the C₃ atom independently of the conformation (*s-cis* and *s-trans*). Therefore, the two electric dipoles of the *s-cis* conformer are directed in a counter parallel manner so that this conformer is expected to have lower energy than the *s-trans* conformer, whose electric dipoles are directed perpendicularly. For isovanillin, however, the positive end of the electric dipole is at the position of the C₄ atom and hence the relative relationship of the two dipoles of the *s-cis* conformer is parallel and that of the *s-trans* is perpendicular, anticipating the higher energy for the *s-cis* conformer than the *s-trans*. Of course, this explanation is too simplified, because this model predicts that

the conformational stability of the *s-cis* and *s-trans* conformers of isovanillin is opposite to that of vanillin, whereas the actual calculated energy for the *s-trans* of isovanillin is still slightly higher than that of the *s-cis*. However, more sophisticated explanation is out of the scope of the present work, and above all, the conformational properties of these molecules should be further investigated with spectroscopic methods.

Table 8 lists the obtained structural parameters for the *s-cis* conformer of vanillin and isovanillin and those for the *s-cis-trans* conformer of ethylvanillin. Experimental radial distribution curves with residuals are shown in Fig. 3. The resultant *R*-factors¹ were 0.044, 0.054 and 0.047 for vanillin, isovanillin and ethylvanillin, respectively. The indices of resolution, *k*,² for them were 0.94±0.02, 0.86±0.02 and 0.88±0.02, respectively. The correlation matrices are listed in Table S5 of Supplementary Information. It was common to all the three molecules that only one element of the correlation matrix, *r*(C–C) vs. *r*(C–O), had the absolute value larger than 0.9.

The theoretical structural parameters obtained from the MP2/6-31G** and B3LYP/6-31G** calculations are compared with the experimental ones in Table 8. For the bond lengths, the C–O and C–C bonds showed satisfactory agreement between the experimental and theoretical values, but the calculations failed to reproduce the C–H/O–H and C=O bond lengths. As for the bond angles, disagreement was found for some bond angles such as ∠C–O–C, ∠C₂–C₁–C₇ and ∠C–C=O. It should be also noted that there is no significant difference between the MP2 and DFT structural parameters in regard to the agreement with the experimental results.

For vanillin and isovanillin, the structural parameters obtained from the X-ray diffraction studies [5, 6] are compared in Table 9. There are some amounts of

¹ *R*-factor is defined as $R = \{ \sum_i W_i (\Delta sM(s)_i)^2 / \sum_i W_i (sM(s)_i^{\text{obs}})^2 \}^{1/2}$, where $\Delta sM(s)_i = sM(s)_i^{\text{obs}} - sM(s)_i^{\text{calc}}$ and W_i is a diagonal element of the weight matrix.

² *k* is defined as $sM(s)^{\text{obs}} = k sM(s)^{\text{calc}}$.

differences between the gas-phase and crystal values of many structural parameters. Among them, the most characteristic ones are found for the C–C=O and C–O–C angles. The values of the C–C=O angle for the crystal, $126.1\pm 0.2^\circ$ and $125.3\pm 0.5^\circ$ for vanillin and isovanillin, respectively, are significantly larger than the corresponding gas-phase values, $119.4\pm 1.6^\circ$ and $121.1\pm 2.4^\circ$. The large C–C=O angles in the crystal might be due to the intermolecular hydrogen bond between the C=O and O–H groups.

Finally, the structural parameters obtained in the present study are compared with those of the related molecules, benzaldehyde [17], anisole [23] and phenol [18], in Table 10. Most of the corresponding parameters are in moderate agreement with each other. Therefore it can be concluded that the mutual influence between the side chain groups, aldehyde, methoxy, hydroxyl and ethoxy, on the structural parameters is not very large.

6. Supplementary information

Tables of the leveled total intensities and the backgrounds, definitions of the internal coordinates with the scale factors, observed and calculated vibrational wavenumbers of vanillin, mean amplitudes, and the correlation matrices are deposited with B.L.L.D. as publication No. SUP .

Acknowledgements

The authors thank the Research Center for Computational Science, Okazaki Japan, for the use of the Fujitsu VPP5000 computer and the Library Program GAUSSIAN 98. T. E. thanks Shigehiro Konaka, Professor Emeritus of Hokkaido University, for his critical reading of the manuscript.

References

- [1] T. Egawa, Y. Kachi, T. Takeshima, H. Takeuchi, S. Konaka, *J. Mol. Struct.*, 658

- (2003) 241.
- [2] T. Egawa, M. Sakamoto, H. Takeuchi, S. Konaka, *J. Phys. Chem. A*, 107 (2003) 2757.
- [3] M. Mineyama, T. Egawa, *J. Mol. Struct.*, 734 (2005) 61.
- [4] K. Kajiya, K. Inaki, M. Tanaka, T. Haga, H. Kataoka, K. Touhara, *J. Neurosci.*, 21 (2001) 6018.
- [5] R. Velavan, P. Sureshkumar, K. Sivakumar, S. Natarajan, *Acta Cryst.*, C51 (1995) 1131.
- [6] F. Iwasaki, *Chem. Lett.*, (1973) 227.
- [7] S. Konaka, M. Kimura, 13th Austin Symposium on Gas Phase Molecular Structure, 12-14 March, The University of Texas, Austin, TX, 1990, S21.
- [8] N. Kuze, M. Ebizuka, H. Fujiwara, H. Takeuchi, T. Egawa, S. Konaka, G. Fogarasi, *J. Phys. Chem. A*, 102 (1998) 2080.
- [9] A. Tsuboyama, A. Murayama, S. Konaka, M. Kimura, *J. Mol. Struct.*, 118 (1984) 351.
- [10] N. Kuze, H. Fujiwara, H. Takeuchi, T. Egawa, S. Konaka, G. Fogarasi, *J. Phys. Chem. A*, 103 (1999) 3054.
- [11] M. Kimura, S. Konaka, M. Ogasawara, *J. Chem. Phys.*, 46 (1967) 2599.
- [12] C. Tavard, D. Nicolas, M. Rouault, *J. Chim. Phys. Phys.-Chim. Biol.*, 64 (1967) 540.
- [13] E. A. Velcheva, B. A. Stamboliyska, *Spectrochim. Acta*, A60 (2004) 2013.
- [14] Gaussian 98, Revision A.9, M. J. Frisch, G. W. Trucks, H. B. Schlegel, G. E. Scuseria, M. A. Robb, J. R. Cheeseman, V. G. Zakrzewski, J. A. Montgomery Jr., R. E. Stratmann, J. C. Burant, S. Dapprich, J. M. Millam, A. D. Daniels, K. N. Kudin, M. C. Strain, O. Farkas, J. Tomasi, V. Barone, M. Cossi, R. Cammi, B. Mennucci, C. Pomelli, C. Adamo, S. Clifford, J. Ochterski, G. A. Petersson, P. Y. Ayala, Q. Cui, K. Morokuma, D. K. Malick, A. D. Rabuck, K. Raghavachari, J. B. Foresman, J.

- Cioslowski, J. V. Ortiz, A. G. Baboul, B. B. Stefanov, G. Liu, A. Liashenko, P. Piskorz, I. Komaromi, R. Gomperts, R. L. Martin, D. J. Fox, T. Keith, M. A. Al-Laham, C. Y. Peng, A. Nanayakkara, M. Challacombe, P. M. W. Gill, B. Johnson, W. Chen, M. W. Wong, J. L. Andres, C. Gonzalez, M. Head-Gordon, E. S. Replogle, J. A. Pople, Gaussian, Inc., Pittsburgh PA, 1998.
- [15] SDBSWeb : <http://www.aist.go.jp/RIODB/SDBS/>(Oct., 2000).
- [16] J. E. Boggs, in I. Hargittai and M. Hargittai (Ed.), *Stereochemical Applications of Gas-phase Electron Diffraction, Part B*, VCH, New York, 1988, Chapter 10.
- [17] K. B. Borisenko, C. W. Bock, I. Hargittai, *J. Phys. Chem.*, 100 (1996) 7426.
- [18] G. Portalone, G. Shultz, A. Domenicano, I. Hargittai, *Chem. Phys. Lett.*, 197 (1992) 482.
- [19] K. Kuchitsu, S. J. Cyvin, in S. J. Cyvin (Ed.), *Molecular Structures and Vibrations*, Elsevier, Amsterdam, 1972, Chapter 12.
- [20] K. Kuchitsu, *Bull. Chem. Soc. Jpn.*, 40 (1967) 505.
- [21] K. Kuchitsu, L. S. Bartell, *J. Chem. Phys.*, 35 (1961) 1945.
- [22] NBO, Version 3.1, E. D. Glendening, A. E. Reed, J. E. Carpenter, F. Weinhold,
- [23] H. M. Seip, R. Seip, *Acta Chem. Scand.*, 27 (1973) 4024.

Table 1

Experimental conditions of gas electron diffraction experiments for vanillin, isovanillin and ethylvanillin

	Vanillin	Isovanillin	Ethylvanillin
Camera distance / mm	244.18	244.35	244.37
Nozzle temperature / K	398	446	419
Electron wavelength / Å	0.06334	0.06326	0.06326
Uncertainty in the scale factor / %	0.05	0.04	0.10
Background pressure during exposure / 10^{-6} Torr	5.0 – 6.0	4.1 – 4.8	2.6 – 3.0
Beam current / μA	1.24 - 1.27	1.41 - 1.48	1.38 – 1.39
Exposure time / s	65 – 74	68 – 96	65 – 70
Number of plates used	3	4	4
Range of s value / Å^{-1}	4.7 – 33.8	4.7 – 33.8	4.5 – 33.5

Table 2

Geometrical parameters and relative energies of the *s-cis* and *s-trans* conformers of vanillin obtained from the MP2/6-31G** calculations

Parameters ^a	<i>s-cis</i>	<i>s-trans</i>
Bond lengths / Å		
C ₁ –C ₂	1.406	1.405
C ₂ –C ₃	1.386	1.390
C ₃ –C ₄	1.414	1.410
C ₄ –C ₅	1.391	1.394
C ₅ –C ₆	1.395	1.392
C ₁ –C ₆	1.397	1.399
C ₁ –C ₇	1.472	1.474
C ₃ –O ₉	1.373	1.375
C ₄ –O ₁₁	1.361	1.361
O ₉ –C ₁₀	1.428	1.426
C ₇ =O ₈	1.229	1.228
C ₇ –H ₁₂	1.106	1.107
C ₂ –H ₁₃	1.081	1.082
C ₅ –H ₁₈	1.082	1.082
C ₆ –H ₁₉	1.084	1.082
C ₁₀ –H ₁₄	1.085	1.085
C ₁₀ –H _{15, 16}	1.090	1.091
O ₁₁ –H ₁₇	0.970	0.970
Bond angles / °		
C ₂ –C ₁ –C ₆	120.7	120.6

C ₁ -C ₂ -C ₃	118.9	119.2
C ₂ -C ₃ -C ₄	120.5	120.2
C ₃ -C ₄ -C ₅	120.2	120.1
C ₄ -C ₅ -C ₆	119.5	119.8
C ₁ -C ₆ -C ₅	120.2	120.0
C ₂ -C ₁ -C ₇	119.2	119.1
C ₆ -C ₁ -C ₇	120.1	120.3
C ₁ -C ₇ =O ₈	124.4	124.6
C ₂ -C ₃ -O ₉	126.5	126.6
C ₄ -C ₃ -O ₉	113.0	113.2
C ₃ -C ₄ -O ₁₁	119.9	120.0
C ₅ -C ₄ -O ₁₁	120.0	119.9
C ₃ -O ₉ -C ₁₀	116.3	116.5
C ₁ -C ₇ -H ₁₂	114.9	114.8
O ₈ =C ₇ -H ₁₂	120.7	120.6
C ₁ -C ₂ -H ₁₃	118.3	119.5
C ₃ -C ₂ -H ₁₃	122.8	121.3
C ₄ -C ₅ -H ₁₈	118.5	118.3
C ₆ -C ₅ -H ₁₈	122.0	121.8
C ₅ -C ₆ -H ₁₉	120.0	121.3
C ₁ -C ₆ -H ₁₉	119.8	118.7
C ₄ -O ₁₁ -H ₁₇	106.5	106.6
O ₉ -C ₁₀ -H ₁₄	105.7	105.7
O ₉ -C ₁₀ -H _{15, 16}	110.8	110.9
H ₁₄ -C ₁₀ -H _{15, 16}	110.0	109.8
H ₁₅ -C ₁₀ -H _{16,}	109.6	109.6

Dihedral angles / °		
C ₂ -C ₁ -C ₇ =O ₈	0.0	180.0
C ₆ -C ₁ -C ₇ =O ₈	180.0	0.0
C ₂ -C ₁ -C ₇ -H ₁₂	180.0	0.0
C ₆ -C ₁ -C ₇ -H ₁₂	0.0	180.0
C ₃ -O ₉ -C ₁₀ -H ₁₄	180.0	180.0
C ₃ -O ₉ -C ₁₀ -H ₁₅	60.9	61.0
C ₃ -O ₉ -C ₁₀ -H ₁₆	-60.9	-61.0
ΔE / kJ mol ⁻¹ ^b	0.000	4.892
Abundance / % ^c	81.4	18.6

^a See Fig. 1 for the atom numbering.

^b The absolute energy of the *s-cis* conformer is $-533.7807521 E_h$.

^c Estimated from the energy difference, ΔE , at 398 K.

Table 3

Geometrical parameters and relative energies of the *s-cis* and *s-trans* conformers of isovanillin obtained from the MP2/6-31G** calculations

Parameters ^a	<i>s-cis</i>	<i>s-trans</i>
Bond lengths / Å		
C ₁ –C ₂	1.403	1.402
C ₂ –C ₃	1.385	1.388
C ₃ –C ₄	1.414	1.410
C ₄ –C ₅	1.394	1.396
C ₅ –C ₆	1.398	1.395
C ₁ –C ₆	1.397	1.398
C ₁ –C ₇	1.475	1.476
C ₃ –O ₉	1.362	1.363
C ₄ –O ₁₀	1.373	1.374
O ₁₀ –C ₁₁	1.427	1.427
C ₇ =O ₈	1.228	1.228
C ₇ –H ₁₂	1.107	1.106
C ₂ –H ₁₃	1.082	1.084
C ₅ –H ₁₈	1.080	1.081
C ₆ –H ₁₉	1.084	1.082
C ₁₁ –H ₁₅	1.085	1.085
C ₁₁ –H _{16, 17}	1.090	1.090
O ₉ –H ₁₄	0.969	0.969
Bond angles / °		
C ₂ –C ₁ –C ₆	120.7	120.5

C ₁ -C ₂ -C ₃	119.4	119.7
C ₂ -C ₃ -C ₄	119.9	119.7
C ₃ -C ₄ -C ₅	120.8	120.7
C ₄ -C ₅ -C ₆	119.0	119.4
C ₁ -C ₆ -C ₅	120.3	120.1
C ₂ -C ₁ -C ₇	119.8	119.4
C ₆ -C ₁ -C ₇	119.6	120.0
C ₁ -C ₇ =O ₈	124.5	124.4
C ₂ -C ₃ -O ₉	120.3	120.2
C ₄ -C ₃ -O ₉	119.8	120.1
C ₃ -C ₄ -O ₁₀	113.0	113.1
C ₅ -C ₄ -O ₁₀	126.2	126.3
C ₄ -O ₁₀ -C ₁₁	116.8	116.8
C ₁ -C ₇ -H ₁₂	114.8	115.0
O ₈ =C ₇ -H ₁₂	120.7	120.7
C ₁ -C ₂ -H ₁₃	120.3	121.4
C ₃ -C ₂ -H ₁₃	120.3	118.9
C ₄ -C ₅ -H ₁₈	120.9	120.7
C ₆ -C ₅ -H ₁₈	120.1	119.9
C ₅ -C ₆ -H ₁₉	119.8	121.0
C ₁ -C ₆ -H ₁₉	119.9	118.9
C ₃ -O ₉ -H ₁₄	106.7	106.5
O ₁₀ -C ₁₁ -H ₁₅	105.7	105.6
O ₁₀ -C ₁₁ -H _{16,17}	110.9	110.9
H ₁₅ -C ₁₁ -H _{16,17}	109.8	109.8
H ₁₆ -C ₁₁ -H ₁₇	109.7	109.7

Dihedral angles / °		
C ₂ -C ₁ -C ₇ =O ₈	0.0	180.0
C ₆ -C ₁ -C ₇ =O ₈	180.0	0.0
C ₂ -C ₁ -C ₇ -H ₁₂	180.0	0.0
C ₆ -C ₁ -C ₇ -H ₁₂	0.0	180.0
C ₄ -O ₁₀ -C ₁₁ -H ₁₅	180.0	180.0
C ₄ -O ₁₀ -C ₁₁ -H ₁₆	-61.0	-61.1
C ₄ -O ₁₀ -C ₁₁ -H ₁₇	61.0	61.1
ΔE / kJ mol ⁻¹ ^b	0.000	0.079
Abundance / % ^c	50.5	49.5

^a See Fig. 1 for the atom numbering.

^b The absolute energy of the *s-cis* conformer is $-533.7785465 E_h$.

^c Estimated from the energy difference, ΔE , at 446 K.

Table 4

Geometrical parameters and relative energies of the *s-cis-trans*, *s-trans-trans*, *s-cis-gauche* and *s-trans-gauche* conformers of ethylvanillin obtained from the MP2/6-31G** calculations

Parameters ^a	<i>s-cis-trans</i>	<i>s-trans-trans</i>	<i>s-cis-gauche</i>	<i>s-trans-gauche</i>
Bond lengths / Å				
C ₁ –C ₂	1.406	1.405	1.406	1.405
C ₂ –C ₃	1.387	1.390	1.387	1.391
C ₃ –C ₄	1.414	1.411	1.415	1.411
C ₄ –C ₅	1.391	1.394	1.391	1.394
C ₅ –C ₆	1.395	1.392	1.395	1.391
C ₁ –C ₆	1.397	1.399	1.397	1.399
C ₁ –C ₇	1.472	1.474	1.472	1.474
C ₃ –O ₉	1.373	1.375	1.374	1.377
C ₄ –O ₁₂	1.361	1.362	1.360	1.361
O ₉ –C ₁₀	1.435	1.433	1.439	1.437
C ₁₀ –C ₁₁	1.510	1.510	1.515	1.516
C ₇ =O ₈	1.229	1.228	1.229	1.228
C ₇ –H ₁₃	1.106	1.107	1.106	1.108
C ₂ –H ₁₄	1.081	1.082	1.081	1.083
C ₅ –H ₂₁	1.082	1.082	1.082	1.082
C ₆ –H ₂₂	1.084	1.082	1.084	1.082
C ₁₀ –H ₁₅	1.093	1.093	1.091	1.092
C ₁₀ –H ₁₆	1.093	1.093	1.088	1.088
C ₁₁ –H ₁₇	1.089	1.089	1.090	1.090
C ₁₁ –H ₁₈	1.088	1.088	1.087	1.087

C ₁₁ -H ₁₉	1.088	1.088	1.088	1.088
O ₁₂ -H ₂₀	0.970	0.970	0.970	0.970

Bond angles / °

C ₂ -C ₁ -C ₆	120.7	120.6	120.8	120.7
C ₁ -C ₂ -C ₃	118.9	119.2	119.0	119.3
C ₂ -C ₃ -C ₄	120.5	120.2	120.2	120.0
C ₃ -C ₄ -C ₅	120.2	120.2	120.4	120.3
C ₄ -C ₅ -C ₆	119.5	119.8	119.5	119.8
C ₁ -C ₆ -C ₅	120.2	120.0	120.1	119.9
C ₂ -C ₁ -C ₇	119.2	119.1	119.1	119.1
C ₆ -C ₁ -C ₇	120.1	120.3	120.1	120.3
C ₁ -C ₇ =O ₈	124.4	124.6	124.4	124.6
C ₂ -C ₃ -O ₉	126.6	126.6	127.1	127.1
C ₄ -C ₃ -O ₉	113.0	113.2	112.7	112.9
C ₃ -C ₄ -O ₁₂	119.8	119.9	119.7	119.9
C ₅ -C ₄ -O ₁₂	120.0	119.9	119.9	119.8
C ₃ -O ₉ -C ₁₀	116.8	117.0	117.7	117.8
O ₉ -C ₁₀ -C ₁₁	107.0	107.0	111.8	111.9
C ₁ -C ₇ -H ₁₃	114.9	114.8	114.9	114.8
O ₈ =C ₇ -H ₁₃	120.7	120.6	120.7	120.6
C ₁ -C ₂ -H ₁₄	118.4	119.6	118.1	119.3
C ₃ -C ₂ -H ₁₄	122.7	121.3	122.9	121.4
C ₄ -C ₅ -H ₂₁	118.5	118.3	118.5	118.3
C ₆ -C ₅ -H ₂₁	122.0	121.9	122.0	121.9
C ₅ -C ₆ -H ₂₂	120.0	121.3	120.1	121.3
C ₁ -C ₆ -H ₂₂	119.8	118.7	119.8	118.8

C ₄ -O ₁₂ -H ₂₀	106.3	106.4	106.2	106.3
O ₉ -C ₁₀ -H ₁₅	109.2	109.3	109.4	109.6
O ₉ -C ₁₀ -H ₁₆	109.2	109.3	104.0	104.0
C ₁₁ -C ₁₀ -H ₁₅	111.5	111.4	112.1	112.1
C ₁₁ -C ₁₀ -H ₁₆	111.5	111.4	111.1	111.0
H ₁₅ -C ₁₀ -H ₁₆	108.3	108.4	108.1	107.8
C ₁₀ -C ₁₁ -H ₁₇	109.9	110.0	109.9	110.0
C ₁₀ -C ₁₁ -H ₁₈	110.4	110.3	111.3	111.3
C ₁₀ -C ₁₁ -H ₁₉	110.4	110.3	109.9	109.9
H ₁₇ -C ₁₁ -H ₁₈	108.6	108.7	108.4	108.5
H ₁₇ -C ₁₁ -H ₁₉	108.6	108.7	108.6	108.6
H ₁₈ -C ₁₁ -H ₁₉	108.9	108.9	108.7	108.5

Dihedral angles / °

C ₆ -C ₁ -C ₂ -C ₃	0.0	0.0	-0.0	-0.2
C ₇ -C ₁ -C ₂ -C ₃	180.0	180.0	179.7	179.8
C ₁ -C ₂ -C ₃ -C ₄	0.0	0.0	-0.1	0.1
C ₁ -C ₂ -C ₃ -O ₉	180.0	180.0	178.7	178.4
C ₂ -C ₃ -C ₄ -C ₅	0.0	0.0	0.2	0.0
C ₂ -C ₃ -C ₄ -O ₁₂	180.0	180.0	-180.0	180.0
O ₉ -C ₃ -C ₄ -C ₅	180.0	180.0	-178.7	-178.5
O ₉ -C ₃ -C ₄ -O ₁₂	0.0	0.0	1.1	1.4
C ₃ -C ₄ -C ₅ -C ₆	0.0	0.0	-0.2	-0.1
O ₁₂ -C ₄ -C ₅ -C ₆	180.0	180.0	180.0	180.0
C ₄ -C ₅ -C ₆ -C ₁	0.0	0.0	0.1	0.0
C ₅ -C ₆ -C ₁ -C ₂	0.0	0.0	0.0	0.1
C ₅ -C ₆ -C ₁ -C ₇	180.0	180.0	-179.7	-179.8

C ₂ -C ₁ -C ₇ =O ₈	0.0	180.0	-0.3	179.9
C ₆ -C ₁ -C ₇ =O ₈	180.0	0.0	179.4	-0.1
C ₂ -C ₃ -O ₉ -C ₁₀	0.0	0.0	7.9	11.7
C ₄ -C ₃ -O ₉ -C ₁₀	180.0	180.0	-173.3	-169.8
C ₃ -O ₉ -C ₁₀ -C ₁₁	180.0	180.0	75.8	73.8
C ₂ -C ₁ -C ₇ -H ₁₃	180.0	0.0	179.7	-0.1
C ₆ -C ₁ -C ₇ -H ₁₃	0.0	180.0	-0.6	179.9
C ₆ -C ₁ -C ₂ -H ₁₄	180.0	180.0	-180.0	179.9
C ₇ -C ₁ -C ₂ -H ₁₄	0.0	0.0	-0.2	-0.1
C ₄ -C ₃ -C ₂ -H ₁₄	180.0	180.0	179.9	-180.0
O ₉ -C ₃ -C ₂ -H ₁₄	0.0	0.0	-1.4	-1.7
C ₃ -O ₉ -C ₁₀ -H ₁₅	59.2	59.2	-49.0	-51.2
C ₃ -O ₉ -C ₁₀ -H ₁₆	-59.2	-59.2	-164.3	-166.3
O ₉ -C ₁₀ -C ₁₁ -H ₁₇	180.0	180.0	177.1	177.3
O ₉ -C ₁₀ -C ₁₁ -H ₁₈	-60.2	-60.2	-62.8	-62.5
O ₉ -C ₁₀ -C ₁₁ -H ₁₉	60.2	60.2	57.6	57.8
H ₁₅ -C ₁₀ -C ₁₁ -H ₁₇	-60.6	-60.6	-59.6	-59.1
H ₁₅ -C ₁₀ -C ₁₁ -H ₁₈	59.2	59.3	60.5	61.1
H ₁₅ -C ₁₀ -C ₁₁ -H ₁₉	179.6	179.6	-179.0	-178.6
H ₁₆ -C ₁₀ -C ₁₁ -H ₁₇	60.6	60.6	61.5	61.6
H ₁₆ -C ₁₀ -C ₁₁ -H ₁₈	-179.6	-179.6	-178.4	-178.2
H ₁₆ -C ₁₀ -C ₁₁ -H ₁₉	-59.2	-59.3	-58.0	-57.9
C ₃ -C ₄ -O ₁₂ -H ₂₀	0.0	0.0	0.3	0.4
C ₅ -C ₄ -O ₁₂ -H ₂₀	180.0	180.0	-179.9	-179.7
C ₃ -C ₄ -C ₅ -H ₂₁	180.0	180.0	179.9	-180.0
O ₁₂ -C ₄ -C ₅ -H ₂₁	0.0	0.0	0.1	0.1
C ₁ -C ₆ -C ₅ -H ₂₁	180.0	180.0	179.9	179.9

H ₂₁ -C ₅ -C ₆ -H ₂₂	0.0	0.0	0.0	0.1
C ₄ -C ₅ -C ₆ -H ₂₂	180.0	180.0	-179.8	-179.8
C ₂ -C ₁ -C ₆ -H ₂₂	180.0	180.0	179.9	179.9
C ₇ -C ₁ -C ₆ -H ₂₂	0.0	0.0	0.2	-0.0
$\Delta E / \text{kJ mol}^{-1}$ ^b	0.000	4.731	3.339	8.382
Abundance / % ^c	45.3	11.7	34.8	8.2

^a See Fig. 1 for the definitions of the atom numbering.

^b The absolute energy of the *s-cis-trans* conformer is $-572.9685873 E_h$.

^c Estimated from the energy difference, ΔE , at 419 K.

Table 5

Structural constraints and independent parameters of vanillin

Parameters	<i>s-cis</i>	<i>s-trans</i>
Bond lengths (Å)		
C ₁ –C ₂	r_1	$r_1 - 0.001$
C ₂ –C ₃	$r_1 - 0.020$	$r_1 - 0.016$
C ₃ –C ₄	$r_1 + 0.008$	$r_1 + 0.005$
C ₄ –C ₅	$r_1 - 0.015$	$r_1 - 0.012$
C ₅ –C ₆	$r_1 - 0.011$	$r_1 - 0.014$
C ₁ –C ₆	$r_1 - 0.009$	$r_1 - 0.007$
C ₁ –C ₇	$r_1 + 0.067$	$r_1 + 0.068$
C ₃ –O ₉	r_2	$r_2 + 0.002$
C ₄ –O ₁₁	$r_2 - 0.013$	$r_2 - 0.012$
O ₉ –C ₁₀	$r_2 + 0.054$	$r_2 + 0.053$
C ₇ =O ₈	r_3	$r_3 - 0.001$
C ₇ –H ₁₂	r_4	$r_4 + 0.001$
C ₂ –H ₁₃	$r_4 - 0.025$	$r_4 - 0.024$
C ₅ –H ₁₈	$r_4 - 0.025$	$r_4 - 0.024$
C ₆ –H ₁₉	$r_4 - 0.022$	$r_4 - 0.024$
C ₁₀ –H ₁₄	$r_4 - 0.021$	$r_4 - 0.021$
C ₁₀ –H _{15, 16}	$r_4 - 0.016$	$r_4 - 0.015$
O ₁₁ –H ₁₇	$r_4 - 0.136$	$r_4 - 0.136$
Bond angles (°)		
C ₂ –C ₁ –C ₆	θ_1	$\theta_1 - 0.1$
C ₁ –C ₂ –C ₃	$\theta_1 - 1.8$	$\theta_1 - 1.6$

$C_1-C_6-C_5$	$\theta_1 - 0.5$	$\theta_1 - 0.7$
$C_2-C_1-C_7$	θ_2	$\theta_2 - 0.1$
$C_1-C_7=C_8$	θ_3	$\theta_3 + 0.2$
$C_4-C_3-O_9$	θ_4	$\theta_4 + 0.2$
$C_3-C_4-O_{11}$	$\theta_4 + 6.9$	$\theta_4 + 7.0$
$C_3-O_9-C_{10}$	θ_5	$\theta_5 + 0.2$
$C_1-C_7-H_{12}$	115.1	115.1
$C_1-C_2-H_{13}$	$180.0 - 0.5 (C_1-C_2-C_3)$	$180.0 - 0.5 (C_1-C_2-C_3)$
$O_9-C_{10}-H_{14}$	105.7	105.7
$O_9-C_{10}-H_{15, 16}$	110.8	110.9
$C_4-O_{11}-H_{17}$	106.4	106.4
$C_4-C_5-H_{18}$	$180.0 - 0.5 (C_4-C_5-C_6)$	$180.0 - 0.5 (C_4-C_5-C_6)$
$C_5-C_6-H_{19}$	$180.0 - 0.5 (C_1-C_6-C_5)$	$180.0 - 0.5 (C_1-C_6-C_5)$
Dihedral angles (°)		
$C_2-C_1-C_7=C_8$	0.0	180.0
$C_3-O_9-C_{10}-H_{15}$	60.9	61.0
$C_3-O_9-C_{10}-H_{16}$	- 60.9	- 61.0

See Fig. 1 for the atom numbering and the definitions of the conformers.

Table 6

Structural constraints and independent parameters of isovanillin

Parameters	<i>s-cis</i>	<i>s-trans</i>
Bond lengths (Å)		
C ₁ –C ₂	r_1	$r_1 - 0.001$
C ₂ –C ₃	$r_1 - 0.018$	$r_1 - 0.015$
C ₃ –C ₄	$r_1 + 0.011$	$r_1 + 0.007$
C ₄ –C ₅	$r_1 - 0.010$	$r_1 - 0.007$
C ₅ –C ₆	$r_1 - 0.006$	$r_1 - 0.009$
C ₁ –C ₆	$r_1 - 0.006$	$r_1 - 0.005$
C ₁ –C ₇	$r_1 + 0.072$	$r_1 + 0.073$
C ₃ –O ₉	r_2	$r_2 + 0.001$
C ₄ –O ₁₀	$r_2 + 0.011$	$r_2 + 0.012$
O ₁₀ –C ₁₁	$r_2 + 0.064$	$r_2 + 0.065$
C ₇ =O ₈	r_3	$r_3 + 0.000$
C ₇ –H ₁₂	r_4	$r_4 - 0.000$
C ₂ –H ₁₃	$r_4 - 0.025$	$r_4 - 0.023$
C ₅ –H ₁₈	$r_4 - 0.026$	$r_4 - 0.026$
C ₆ –H ₁₉	$r_4 - 0.023$	$r_4 - 0.025$
C ₁₁ –H ₁₅	$r_4 - 0.022$	$r_4 - 0.022$
C ₁₁ –H _{16,17}	$r_4 - 0.016$	$r_4 - 0.016$
O ₉ –H ₁₄	$r_4 - 0.137$	$r_4 - 0.137$
Bond angles (°)		
C ₂ –C ₁ –C ₆	θ_1	$\theta_1 - 0.1$
C ₁ –C ₂ –C ₃	$\theta_1 - 1.2$	$\theta_1 - 1.0$

$C_1-C_6-C_5$	$\theta_1 - 0.4$	$\theta_1 - 0.6$
$C_2-C_1-C_7$	θ_2	$\theta_2 - 0.3$
$C_1-C_7=O_8$	θ_3	$\theta_3 - 0.1$
$C_4-C_3-O_9$	θ_4	$\theta_4 + 0.3$
$C_3-C_4-O_{10}$	$\theta_4 - 6.8$	$\theta_4 - 6.7$
$C_4-O_{10}-C_{11}$	θ_5	$\theta_5 - 0.0$
$C_1-C_7-H_{12}$	115.1	115.1
$C_1-C_2-H_{13}$	$180.0 - 0.5 (C_1-C_2-C_3)$	$180.0 - 0.5 (C_1-C_2-C_3)$
$C_3-O_9-H_{14}$	106.4	106.4
$O_{10}-C_{11}-H_{15}$	105.7	105.6
$O_{10}-C_{11}-H_{16,17}$	110.9	110.9
$C_4-C_5-H_{18}$	$180.0 - 0.5 (C_4-C_5-C_6)$	$180.0 - 0.5 (C_4-C_5-C_6)$
$C_5-C_6-H_{19}$	$180.0 - 0.5 (C_1-C_6-C_5)$	$180.0 - 0.5 (C_1-C_6-C_5)$
Dihedral angles ($^\circ$)		
$C_2-C_1-C_7=C_8$	0.0	180.0
$C_4-O_{10}-C_{11}-H_{16}$	- 61.0	- 61.1
$C_4-O_{10}-C_{11}-H_{17}$	61.0	61.1

See Fig. 1 for the atom numbering and the definitions of the conformers.

Table 7

Structural constraints and independent parameters of ethylvanillin

Parameters	<i>s-cis-trans</i>	<i>s-trans-trans</i>	<i>s-cis-gauche</i>	<i>s-trans-gauche</i>
Bond lengths (Å)				
C ₁ –C ₂	r_1	$r_1 - 0.001$	$r_1 - 0.000$	$r_1 - 0.001$
C ₂ –C ₃	$r_1 - 0.019$	$r_1 - 0.016$	$r_1 - 0.019$	$r_1 - 0.015$
C ₃ –C ₄	$r_1 + 0.008$	$r_1 + 0.005$	$r_1 + 0.009$	$r_1 + 0.005$
C ₄ –C ₅	$r_1 - 0.015$	$r_1 - 0.012$	$r_1 - 0.014$	$r_1 - 0.012$
C ₅ –C ₆	$r_1 - 0.010$	$r_1 - 0.014$	$r_1 - 0.011$	$r_1 - 0.015$
C ₁ –C ₆	$r_1 - 0.009$	$r_1 - 0.007$	$r_1 - 0.009$	$r_1 - 0.007$
C ₁ –C ₇	$r_1 + 0.066$	$r_1 + 0.068$	$r_1 + 0.067$	$r_1 + 0.068$
C ₃ –O ₉	r_2	$r_2 + 0.002$	$r_2 + 0.001$	$r_2 + 0.003$
C ₄ –O ₁₂	$r_2 - 0.012$	$r_2 - 0.012$	$r_2 - 0.013$	$r_2 - 0.012$
O ₉ –C ₁₀	$r_2 + 0.062$	$r_2 + 0.060$	$r_2 + 0.066$	$r_2 + 0.064$
C ₁₀ –C ₁₁	r_3	$r_3 + 0.000$	$r_3 + 0.005$	$r_3 + 0.006$
C ₇ =O ₈	r_4	$r_4 - 0.001$	$r_4 - 0.000$	$r_4 - 0.001$
C ₇ –H ₁₃	r_5	$r_5 + 0.001$	r_5	$r_5 + 0.001$
C ₂ –H ₁₄	$r_5 - 0.025$	$r_5 - 0.024$	$r_5 - 0.025$	$r_5 - 0.024$
C ₅ –H ₂₁	$r_5 - 0.025$	$r_5 - 0.024$	$r_5 - 0.025$	$r_5 - 0.024$
C ₆ –H ₂₂	$r_5 - 0.022$	$r_5 - 0.024$	$r_5 - 0.022$	$r_5 - 0.024$
C ₁₀ –H ₁₅	$r_5 - 0.013$	$r_5 - 0.013$	$r_5 - 0.015$	$r_5 - 0.014$
C ₁₀ –H ₁₆	$r_5 - 0.013$	$r_5 - 0.013$	$r_5 - 0.018$	$r_5 - 0.018$
C ₁₁ –H ₁₇	$r_5 - 0.017$	$r_5 - 0.017$	$r_5 - 0.017$	$r_5 - 0.017$
C ₁₁ –H ₁₈	$r_5 - 0.019$	$r_5 - 0.019$	$r_5 - 0.019$	$r_5 - 0.019$
C ₁₁ –H ₁₉	$r_5 - 0.019$	$r_5 - 0.019$	$r_5 - 0.018$	$r_5 - 0.018$
O ₁₂ –H ₂₀	$r_5 - 0.136$	$r_5 - 0.136$	$r_5 - 0.136$	$r_5 - 0.136$

Bond angles (°)

C ₂ -C ₁ -C ₆	θ_1	$\theta_1 - 0.1$	$\theta_1 + 0.1$	$\theta_1 - 0.1$
C ₁ -C ₂ -C ₃	$\theta_1 - 1.8$	$\theta_1 - 1.6$	$\theta_1 - 1.7$	$\theta_1 - 1.4$
C ₁ -C ₆ -C ₅	$\theta_1 - 0.5$	$\theta_1 - 0.7$	$\theta_1 - 0.6$	$\theta_1 - 0.8$
C ₂ -C ₁ -C ₇	θ_2	$\theta_2 - 0.1$	$\theta_2 - 0.1$	$\theta_2 - 0.1$
C ₁ -C ₇ =O ₈	θ_3	$\theta_3 + 0.2$	$\theta_3 - 0.0$	$\theta_3 + 0.2$
C ₄ -C ₃ -O ₉	θ_4	$\theta_4 + 0.2$	$\theta_4 - 0.3$	$\theta_4 - 0.1$
C ₃ -C ₄ -O ₁₂	$\theta_4 + 6.8$	$\theta_4 + 6.9$	$\theta_4 + 6.7$	$\theta_4 + 6.9$
C ₃ -O ₉ -C ₁₀	θ_5	$\theta_5 + 0.2$	$\theta_5 + 0.9$	$\theta_5 + 1.0$
O ₉ -C ₁₀ -C ₁₁	θ_6	$\theta_6 + 0.0$	$\theta_6 + 4.8$	$\theta_6 + 5.0$
C ₁ -C ₇ -H ₁₃	115.1	115.1	115.1	115.1
C ₁ -C ₂ -H ₁₄	180.0 - 0.5	180.0 - 0.5	180.0 - 0.5	180.0 - 0.5
	(C ₁ -C ₂ -C ₃)	(C ₁ -C ₂ -C ₃)	(C ₁ -C ₂ -C ₃)	(C ₁ -C ₂ -C ₃)
O ₉ -C ₁₀ -H ₁₅	109.2	109.3	109.4	109.6
O ₉ -C ₁₀ -H ₁₆	109.2	109.3	104.0	104.0
C ₁₀ -C ₁₁ -H ₁₇	109.9	110.0	109.9	110.0
C ₁₀ -C ₁₁ -H ₁₈	110.4	110.3	111.3	111.3
C ₁₀ -C ₁₁ -H ₁₉	110.4	110.3	109.9	109.9
C ₄ -O ₁₂ -H ₂₀	106.4	106.4	106.4	106.4
C ₄ -C ₅ -H ₂₁	180.0 - 0.5	180.0 - 0.5	180.0 - 0.5	180.0 - 0.5
	(C ₄ -C ₅ -C ₆)	(C ₄ -C ₅ -C ₆)	(C ₄ -C ₅ -C ₆)	(C ₄ -C ₅ -C ₆)
C ₅ -C ₆ -H ₂₂	180.0 - 0.5	180.0 - 0.5	180.0 - 0.5	180.0 - 0.5
	(C ₁ -C ₆ -C ₅)	(C ₁ -C ₆ -C ₅)	(C ₁ -C ₆ -C ₅)	(C ₁ -C ₆ -C ₅)

Dihedral angles (°)

C ₂ -C ₁ -C ₇ =C ₈	0.0	180.0	0.0	180.0
--	-----	-------	-----	-------

C ₂ -C ₃ -O ₉ -C ₁₀	0.0	0.0	7.9	11.7
C ₃ -O ₉ -C ₁₀ -C ₁₁	180.0	180.0	75.8	73.8
C ₃ -O ₉ -C ₁₀ -H ₁₅	59.2	59.2	- 49.0	- 51.2
C ₃ -O ₉ -C ₁₀ -H ₁₆	- 59.2	- 59.2	- 164.3	- 166.3
O ₉ -C ₁₀ - C ₁₁ -H ₁₇	180.0	180.0	177.1	177.3
O ₉ -C ₁₀ - C ₁₁ -H ₁₈	- 60.2	- 60.2	- 62.8	- 62.5
O ₉ -C ₁₀ - C ₁₁ -H ₁₉	60.2	60.2	57.6	57.8

See Fig. 1 for the atom numbering and the definitions of the conformers.

Table 8

Molecular structures of vanillin, isovanillin and ethylvanillin

Parameters ^a	Vanillin (<i>s-cis</i>)			Isovanillin (<i>s-cis</i>)			Ethylvanillin (<i>s-cis-trans</i>)		
	ED (r_g and \angle_α)	MP2	DFT ^d	ED (r_g and \angle_α)	MP2	DFT ^d	ED (r_g and \angle_α)	MP2	DFT ^d
Bond lengths (Å)									
C ₁ –C ₂	1.405 } 	1.406	1.410	1.407 } 	1.403	1.406	1.405 } 	1.406	1.410
C ₂ –C ₃	1.385	1.386	1.382	1.388	1.385	1.384	1.385	1.387	1.383
C ₃ –C ₄	1.413	1.414	1.418	1.418	1.414	1.417	1.413	1.414	1.418
C ₄ –C ₅	1.390 } (4)	1.391	1.393	1.397 } (4)	1.394	1.394	1.390 } (6)	1.391	1.393
C ₅ –C ₆	1.394	1.395	1.394	1.401	1.398	1.397	1.394	1.395	1.394
C ₁ –C ₆	1.396	1.397	1.398	1.400	1.397	1.397	1.396	1.397	1.398
<C–C> _{ring}	1.397	1.398	1.399	1.402	1.398	1.399	1.397	1.398	1.399
C ₁ –C ₇	1.471 } 	1.472	1.471	1.479 } 	1.475	1.475	1.471 } 	1.472	1.471
C _{ring} –O _{Me/Et}	1.374 } 	1.373	1.372	1.369 } 	1.373	1.368	1.365 } 	1.373	1.371
C _{ring} –O _H	1.361 } (9)	1.361	1.354	1.357 } (9)	1.362	1.359	1.352 } (13)	1.361	1.354
O–C _{Me/Et}	1.428 } 	1.428	1.423	1.422 } 	1.427	1.423	1.427 } 	1.435	1.433

C–C _{Et}								1.494 (21)	1.510	1.517
C ₇ =O ₈	1.214 (8)	1.229	1.219	1.221 (9)	1.228	1.217	1.206 (9)	1.229	1.219	
<C–H>	1.110 } (11)	1.088	1.093	1.114 } (14)	1.088	1.093	1.109 } (10)	1.089	1.094	
O–H	0.991 } (11)	0.970	0.971	0.995 } (14)	0.969	0.970	0.990 } (10)	0.970	0.971	
Bond angles (°)										
C ₂ –C ₁ –C ₆	120.6 } (2)	120.7	120.1	120.2 } (3)	120.7	120.2	120.2 } (3)	120.7	120.1	
C ₁ –C ₂ –C ₃	118.8 } (2)	118.9	119.5	119.0 } (3)	119.4	119.9	118.4 } (3)	118.9	119.6	
C ₁ –C ₆ –C ₅	120.1 } (2)	120.2	120.5	119.9 } (3)	120.3	120.4	119.7 } (3)	120.2	120.5	
C ₂ –C ₃ –C ₄	120.9 — ^e	120.5	120.2	121.9 — ^e	119.9	119.6	122.8 — ^e	120.5	120.1	
C ₃ –C ₄ –C ₅	119.6 — ^e	120.2	120.1	117.8 — ^e	120.8	120.7	116.8 — ^e	120.2	120.2	
C ₄ –C ₅ –C ₆	120.0 — ^e	119.5	119.5	121.2 — ^e	119.0	119.3	122.0 — ^e	119.5	119.5	
C ₂ –C ₁ –C ₇	122.7 (18)	119.2	119.5	124.6 (25)	119.8	120.1	121.7 (21)	119.2	119.6	
C ₁ –C ₇ =O ₈	119.4 (16)	124.4	124.9	121.3 (24)	124.5	125.0	128.8 (22)	124.4	124.9	

C-C _{ring} -O _{Me/Et}	112.2	} (12)	113.0	113.2	114.4	} (12)	113.0	113.4	112.8	} (14)	113.0	113.2
C-C _{ring} -O _H ^g	119.1		119.9	119.7	121.2		119.8	119.9	119.6		119.8	119.6
C _{ring} -O-C _{Me/Et}	121.7	(29)	116.3	117.9	123.8	(26)	116.8	118.6	115.1	(27)	116.8	118.4
O-C-C _{Et}									102.7	(28)	107.0	107.8
C ₁ -C ₇ -H	115.1	— ^h	114.9	114.4	115.1	— ^h	114.8	114.3	115.1	— ^h	114.9	114.4
C _{ring} -O-H	106.4	— ^h	106.5	107.2	106.4	— ^h	106.7	107.4	106.4	— ^h	106.3	107.1

^a See Fig. 1 for the atom numbering. Angle brackets denote average values.

^b Values in parentheses are estimated error limits (3σ) referring to the last significant digit.

^c r_e structure obtained from the MP2/6-31G** calculation.

^d r_e structure obtained from the B3LYP/6-31G** calculation.

^e Dependent parameter.

^f C₄-C₃-O₉ angle for vanillin and ethylvanillin, and C₃-C₄-O₁₀ angle for isovanillin.

^g C₃-C₄-O₁₁ angle for vanillin, C₄-C₃-O₉ angle for isovanillin and C₃-C₄-O₁₂ angle for ethylvanillin.

^h Assumed.

Table 9

Comparison of the gas-phase and crystalline structures of vanillin and isovanillin

Parameters ^a	Vanillin (<i>s-cis</i>)		Isovanillin (<i>s-trans</i>)	
	ED (r_g and \angle_α)	X-ray ^c	ED (r_g and \angle_α) ^b	X-ray ^d
Bond lengths (Å)				
C ₁ –C ₂	1.405 } 	1.400 (3)	1.406 } 	1.396 (6)
C ₂ –C ₃	1.385	1.370 (3)	1.391	1.377 (6)
C ₃ –C ₄	1.413	1.403 (3)	1.414	1.399 (6)
C ₄ –C ₅	1.390 } (4)	1.381 (3)	1.400 } (4)	1.385 (6)
C ₅ –C ₆	1.394	1.378 (4)	1.398	1.385 (6)
C ₁ –C ₆	1.396	1.379 (3)	1.402	1.385 (6)
C ₁ –C ₇	1.471 } 	1.459 (3)	1.480 } 	1.463 (6)
C _{ring} –O _{Me}	1.374 } 	1.360 (2)	1.369 } 	1.365 (6)
C _{ring} –O _H	1.361 } (9)	1.348 (3)	1.358 } (9)	1.358 (6)
O–C _{Me}	1.428 } 	1.430 (3)	1.422 } 	1.428 (6)
C ₇ =O ₈	1.214 (8)	1.205 (3)	1.221 (9)	1.212 (6)
Bond angles (°)				
C ₂ –C ₁ –C ₆	120.6 } 	119.6 (2)	120.1 } 	119.5 (5)
C ₁ –C ₂ –C ₃	118.8 } (2)	120.2 (2)	119.3 } (3)	121.0 (5)
C ₁ –C ₆ –C ₅	120.1 } 	120.6 (2)	119.7 } 	120.1 (5)
C ₂ –C ₃ –C ₄	120.9 — ^e	119.8 (2)	121.7 — ^e	118.9 (5)
C ₃ –C ₄ –C ₅	119.6 — ^e	119.9 (2)	117.7 — ^e	120.5 (5)
C ₄ –C ₅ –C ₆	120.0 — ^e	120.0 (2)	121.5 — ^e	119.9 (5)
C ₂ –C ₁ –C ₇	122.7 (18)	121.0 (2)	124.2 (25)	118.6 (5)

$C_1-C_7=O_8$	119.4	(16)	126.1	(2)	121.1	(24)	125.3	(5)
$C-C_{\text{ring}}-O_{\text{Me}}^f$	112.2	} (12)	114.2	(2)	114.5	} (12)	114.3	(5)
$C-C_{\text{ring}}-O_H^g$	119.1		121.8	(2)	121.5		122.4	(5)
$C_{\text{ring}}-O-C_{\text{Me}}$	121.7	(29)	117.3	(2)	123.8	(26)	117.4	(5)

^a See Fig. 1 for the atom numbering and the definitions of the conformers.

^b Present study. Values in parentheses are estimated error limits (3σ) referring to the last significant digit.

^c Ref. [5]. The average of the four nonequivalent molecules contained in the symmetry unit. Values in parentheses are estimated error limits referring to the last significant digit.

^d Ref. [6]. Values in parentheses are estimated error limits (3σ) referring to the last significant digit.

^e Dependent parameter.

^f $C_4-C_3-O_9$ angle for vanillin and $C_3-C_4-O_{10}$ angle for isovanillin.

^g $C_3-C_4-O_{11}$ angle for vanillin and $C_4-C_3-O_9$ angle for isovanillin.

Table 10

Comparison of the gas-phase structures of vanillin, isovanillin and related molecules determined by electron diffraction ^a

Parameters ^b	Vanillin ^c	Isovanillin ^c	Benzaldehyde ^d	Anisole ^e	Phenol ^f
Bond lengths (Å)					
$\langle C-C \rangle_{\text{ring}}$	1.397	1.402	1.397 (3)	1.399 (3)	1.399 (3)
$C_{\text{ring}}-C_{\text{aldehyde}}$	1.471	1.479	1.479 (4)		
C=O	1.214 (8)	1.221 (9)	1.212 (3)		
$C_{\text{ring}}-O_{\text{Me}}$	1.374	1.369		1.363 (15)	
$C_{\text{ring}}-O_{\text{H}}$	1.361 (9)	1.357 (9)			1.381 (4)
O- C_{Me}	1.428	1.422		1.425 (15)	
Bond angles (°)					
$C_2-C_1-C_{\text{aldehyde}}$	122.7 (18)	124.6 (25)	120.9 (6)		
$C_1-C_{\text{aldehyde}}=O$	119.4 (16)	121.3 (24)	123.6 (4)		
$C_1-C_{\text{aldehyde}}-H$	115.1 — ^g	115.1 — ^g	115.1 (16)		
$C-C_{\text{ring}}-O_{\text{Me}}$ ^h	112.2	114.4		116.0 — ^g	
$C-C_{\text{ring}}-O_{\text{H}}$ ⁱ	119.1	121.2			121.2 (12)
$C_{\text{ring}}-O-C_{\text{Me}}$	121.7 (29)	123.8 (26)		120.0 (20)	
$C_{\text{ring}}-O-H$	106.4 — ^g	106.4 — ^g			106.4 (37)

^a Values in parentheses are estimated error limits referring to the last significant

digit.

^b See Fig. 1 for the atom numbering.

^c r_g and \angle_α of the *s-cis* conformer (present study). Error limits are 3σ .

^d r_g and \angle_α (Ref. [17]).

^e r_g distances were calculated from the r_a distances and mean amplitudes reported in Ref. [23]. No shrinkage correction was applied for the bond angles.

^f r_g and \angle_α (Ref. [18]).

^g Assumed.

^h $C_4-C_3-O_9$ angle for vanillin, $C_3-C_4-O_{10}$ angle for isovanillin and the corresponding angle for anisole.

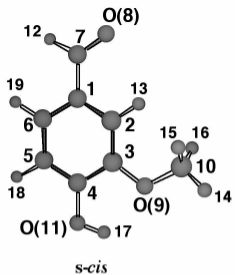
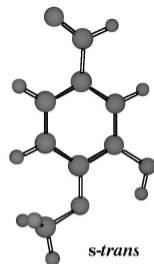
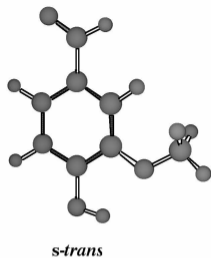
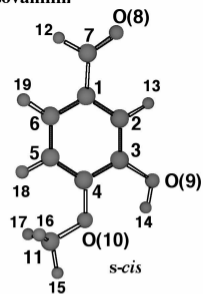
ⁱ $C_3-C_4-O_{11}$ angle for vanillin, $C_4-C_3-O_9$ angle for isovanillin and the corresponding angle for phenol.

Figure captions

Fig. 1. Molecular models and atom numberings for the possible conformers of vanillin, isovanillin and ethylvanillin.

Fig. 2. Experimental (open circles) and theoretical (solid curves) molecular scattering intensities of vanillin, isovanillin and ethylvanillin; $\Delta sM(s) = sM(s)^{\text{obs}} - sM(s)^{\text{calc}}$. The theoretical curves were calculated from the best-fit parameters.

Fig. 3. Experimental radial distribution curves of vanillin, isovanillin and ethylvanillin; $\Delta f(r) = f(r)^{\text{obs}} - f(r)^{\text{calc}}$. Distance distributions are indicated by vertical bars.

Vanillin**Isovanillin****Ethylvanillin**