KINETIC ANALYSIS OF THE MIXED POTENTIAL, THE REACTION RESISTANCE AND THE CORROSION CURRENT

By

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(Received March 20, 1975)

Abstract

General expressions for the reaction resistance at the mixed potential, \( V_r \), and for the concentration dependences of \( V_r \) and the corrosion current are derived and applied to the systems where the concentration polarizations are present.

The condition subjected to \( V_r \) was that (i) \( V_r \) is far from the equilibrium potentials of the respective reaction, \( V_{t\text{-eq}}, t=1,2,\ldots \), and (ii) \( V_r \) is far from some of \( V_{t\text{-eq}} \)'s, \( t=1,2,\ldots,f \), and close to the rest, \( t=f+1,\ldots \).

Resulting formulae were then extended to the cases of (a) a diffusion-controlled anodic (or cathodic) reaction (b) a cathodic (or anodic) reaction with concentration polarization (c) surface process-controlled anodic and cathodic reactions. Special case where the diffusing species are common for both of the anodic and cathodic reactions is also discussed.

Introduction

In the previous paper\(^{1}\), the present authors developed the general expressions for the concentration dependence of the mixed potential, \( V_r \), and the reaction resistance at \( V_r \), and then discussed the systems where no concentration polarization is present. These general expressions cover all relations reported for limiting cases\(^2,3\).

It is now desirable to derive the concentration dependence of the corrosion current and also to apply the above general treatment to the system where the concentration polarizations are present as will be the case in many systems. However, reported treatments on such a system seem to be rather qualitative or limited under certain conditions. Stern\(^3\) discussed only the shape of the polarization curve. Recently, Herblin et
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al\textsuperscript{40} analyzed the mixed potential of the two-reaction system where one of the two reactions is under diffusion control, and showed the method for obtaining the Tafel constant of the non-diffusion controlled reaction from the behaviour of $V_r$. Their treatment, however, is based upon the assumption that the mixed potential is far apart from the equilibrium potentials of the respective reactions and is not applicable to other more complicated cases, e.g., no allowance for neglecting the reverse rate of the respective reactions, and the presence of more than two electrode reactions.

The purpose of the present paper is to extend our previous general treatment to the system where the concentration polarizations are present. Expressions for the concentration dependence of $V$, the concentration dependence of the corrosion current, and the reaction resistance at $V_r$ are obtained in a general manner and will be discussed for several cases which will be most probable to occur.

**Fundamental**

The mixed potential, $V$, will be determined in general by several electrochemical reactions of the type,

$$aA + bB + \cdots = xX + yY + \cdots + ne^-,$$

some of them proceeding in the anodic direction and the others in the cathodic direction, respectively. These reactions will be symbolized in the text as

$$\sum m_{ts}M_s - n_t e^- = 0, \quad t=1, 2, \cdots,$$

where $m_{ts}$ represents the stoichiometric coefficient of the $t$-th reaction with respect to the species, $M_s$, being positive or negative in accordance with $M_s$ being the reduced ($M_s = A, B, \cdots$) or oxidized species ($M_s = X, Y, \cdots$), respectively, and $n_t$ designates the number of electron involved in the $t$-th reaction.

The net rate of the respective reactions, $I_t$, will be a function of the electrode potential, $V$, and the concentrations, $C_s$’s,

$$I_t = f_t(V, C_1, C_2, \cdots, C_s), \quad t=1, 2, \cdots.$$  \hspace{1cm} (3)

By denoting the net rate of the $t$-th reaction at $V_r$ as $I_{tr}$, the condition for the mixed potential is expressed as\textsuperscript{5}

$$\sum I_{tr} = 0, \quad t=1, 2, \cdots.$$  \hspace{1cm} (4)
This is the most fundamental relation on which the following treatments are based.

When the concentration of any species is changed, the mixed potential will shift so that condition (4) is satisfied. Hence, the resulting change in \( I_r \) is expressed from Eq. (3) as

\[
dI_r = \left( \frac{\partial I_r}{\partial V_r} \right)_{C_s} dV_r + \sum_t \left( \frac{\partial I_r}{\partial C_{s'}} \right)_{V_r, C_s} dC_{s'}, \quad t = 1, 2, \ldots
\]

where the subscript \( C_s \) stands for the concentration of all species concerned and \( C_s \neq C_s' \). Since the summation of \( dI_r \) equals zero, we have from the above equation,

\[
dV_r = -\frac{\sum_t \left[ \sum_s \left( \frac{\partial I_r}{\partial \ln C_s} \right)_{V_r, C_s'} \right] d\ln C_s}{\sum_t \left( \frac{\partial I_r}{\partial V_r} \right)_{C_s}}.
\]

Thus, the concentration dependence of \( V_r \) with respect to a species, \( M_s \), is given in general form as

\[
I_s = \left( \frac{\partial V_r}{\partial \ln C_s} \right)_{C_s'} = -\frac{\sum_t \left( \frac{\partial I_r}{\partial \ln C_s} \right)_{V_r, C_s'}}{\sum_t \left( \frac{\partial I_r}{\partial V_r} \right)_{C_s}}.
\]

Let us consider next the polarization of the electrode with a small current, \( dI \). In such a case, we have

\[
dI = \sum_t dI_t = \sum_t \left[ \frac{\partial I_t}{\partial V} \right]_{V=V_r} dV
\]

and hence the reaction resistance, \( \rho \), at \( V_r \) is given in general form as

\[
\rho = \left[ \frac{\partial V}{\partial I} \right]_{V=V_r} = \frac{1}{\sum_t \left[ \frac{\partial I_t}{\partial V} \right]_{V=V_r}}.
\]

Reciprocal of \( \rho \) is equal to the denominator of the expression for \( I_s \), Eq. (7).

The total current which flows in the forward direction at \( V_r \), is equal to that in the reverse direction \( i.e. \),

\[
\sum_t I_t^+ = \sum_t I_t^- = I_m,
\]

where \( I_t^+ \) or \( I_t^- \) is a forward or backward unidirectional current of the \( t \)-th reaction at \( V_r \) and \( I_m \) will be called the corrosion current. Variation of \( I_m \) caused by the change of \( C_s \) is given as
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\[ dI_m = \sum_i \left[ \frac{\partial I_{im}/\partial C_s}{\partial V_r} dC_s + \frac{\partial I_{im}/\partial V_r}{\partial C_s} dV_r \right] \]

\[ = \sum_i \left[ \frac{\partial I_{im}/\partial C_s}{\partial V_r} dC_s + \frac{\partial I_{im}/\partial V_r}{\partial C_s} dV_r \right]. \]

Thus, the concentration dependence of \( I_m \) with respect to a species, \( M_s \), is given by taking into account Eq. (7) as

\[ \sigma_s = \frac{\partial I_{im}}{\partial \ln C_s} = \sum_i \left( \frac{\partial I_{im}^+}{\partial \ln C_s} \right)_{fr, c_s} + \tau_s \sum_i \left( \rho_{im}^+ \right)^{-1} \quad \text{(9)} \]

where \( \rho_{im}^+ = \frac{\partial I_{im}^+}{\partial V_r} \left| C_{s} \right| \) and \( \rho_{im}^- = \frac{\partial I_{im}^-}{\partial V_r} \left| C_{s} \right| \). The expression (7) for \( \tau_s \) is also derived from the above two expressions, (9), considering that \( I_{tr} = I_{im}^+ - I_{im}^- \).

The above three fundamental expressions, (7), (8) and (9), are valid in any systems with no assumptions and save to a large extent the awkward calculations required when we attempt to obtain \( \tau_s, \sigma_s \) and \( \rho \) directly from the rate equations by using only condition (4).

Application

In this section, we apply the general expressions of \( \tau_s, \rho \) and \( \sigma_s \) to the systems where the concentration polarizations are present.

We first assume that one of the reactants of each reaction \( d(t) \), is of small concentration and that one of the products of each reaction, \( D(t) \), accumulates on the surface in large excess of the bulk concentration. Applying Fick’s 1st law to the respective diffusion processes of the \( t \)-th reaction, we have the following relation at the steady state,

\[ I_{tr} = k_{a(t)} \left[ C_{a(t)} - C_{a(t)}(s) \right] = k_{D(t)} \left[ C_{D(t)}(s) - C_{D(t)} \right], \]

where \( k_{a(t)} \) or \( k_{D(t)} \) is the constant including the factor for the difference in the stoichiometric coefficients in the chemical equation, and \( C_{a(t)}(s) \) or \( C_{D(t)}(s) \) is the surface concentration of \( d(t) \) or \( D(t) \). From the above equations,

\[ C_{a(t)}(s) = C_{a(t)}(1 - I_{tr}/I_{L}) \]
\[ C_{D(t)}(s) = C_{D(t)}(1 + I_{tr}/I_{L}), \]

where \( I_{L}^+ = k_{a(t)} C_{a(t)} \) and \( I_{L}^- = k_{D(t)} C_{D(t)} \).
When the reaction occurs in the reverse direction, \( d(t) \) accumulates on the surface and \( D(t) \) is of small concentration at the surface, since \( I_r \) is negative in this case.

Secondly, we assume the following rate equations

\[
I_r = I_{r+} - I_{r-}
\]

\[
= k^+ \tilde{I} C_s^{\tilde{d}'} \left[ C_d(t) (1 - I_{r+} / I_{r+}) \right]^{\tilde{d}_{d(t)}} \exp \left( \alpha^+_r V_r F/RT \right)
\]

\[ - k^- \tilde{I} C_s^{\tilde{d}''} \left[ C_d(t) (1 + I_{r-} / I_{r-}) \right]^{\tilde{d}_{D(t)}} \exp \left( -\alpha^-_r V_r F/RT \right),
\]

(10)

where \( s^' \equiv d(t) \) or \( D(t) \), \( k^' \)'s the rate constants, \( \tilde{z}^+_s \) or \( \tilde{z}^-_s \) the reaction order with respect to \( M_s \), \( \tilde{z}^+_{d(t)} \) or \( \tilde{z}^-_{d(t)} \) that of \( d(t) \) or \( D(t) \), and \( \alpha^+_r \) or \( \alpha^-_r \) the Tafel constant of the respective unidirectional rates of the \( t \)-th reaction.

When the diffusion rate is sufficiently large, \( I_{r+} / I_{r+} \) or \( I_{r-} / I_{r-} \) becomes zero and the surface concentrations equal the bulk ones. This is the case which we have treated previously\(^1\). On the other hand, when the concentration of, say, \( d(t) \) is so small that the diffusion step of \( d(t) \) governs the reaction rate, \( I_r \) becomes \( I_{r+} \).

**Case (1); All of \( d(t) \)'s and \( D(t) \)'s are different species.**

Each derivatives in Eqs. (7) and (8) will be easily obtained from rate equation (10). Thus, we obtain finally the following relations for the concentration dependences of \( V_r \) with respect to the species, \( M_s \), \( d(t) \) and \( D(t) \),

\[
\tau_s = -\left[ \sum_i \left( I_{r+} \tilde{z}^+_{d(t)} - I_{r-} \tilde{z}^-_{d(t)} \right) / \delta_i \right] \times \rho
\]

(11)

\[
\tau_{d(+)} = -\left[ I_{r+} \tilde{z}^+_{d(+)}/(1 - \Delta d) \right] \times \rho
\]

(12)

\[
\tau_{D(\omega)} = \left[ I_{r+} \tilde{z}^-_{D(\omega)}/(1 + \Delta D) \right] \times \rho
\]

(13)

\[
1/\rho = (F/RT) \sum_i \left[ (I_{r+} \alpha^+_r + I_{r-} \alpha^-_r) / \delta_i \right]
\]

(14)

where

\[
\delta_i = 1 + \delta^+ + \delta^-,
\]

\[
\delta^+ = \Delta^+ \tilde{z}^+_{d(+)}/(1 - \Delta^+),
\]

\[
\delta^- = \Delta^- \tilde{z}^-_{D(\omega)}/(1 + \Delta^-)
\]

and

\[
\begin{align*}
\Delta^+ &= I_{r+}/I_{r+}, & \Delta^+ &= I_{r+}/I_{r+}, & \Delta^- &= I_{r-}/I_{r-}, \\
\Delta^- &= I_{r-}/I_{r-}.
\end{align*}
\]

(15)
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The above equations, (11) to (14), are more general than those derived previously since the concentration polarizations are taken into account. \( \sigma_i \) becomes

\[
\sigma_i = \sum_i \left[ I_i^+ z_i^+ + \left( \frac{F}{RT} \right) I_i^- \alpha_i^+ \right] - \sum_i \left\{ \frac{z_{id(t)}^+}{(1 - D_{z+})} \left[ I_i^+ z_i^+ - I_i^- z_i^- + \left( \frac{F}{RT} \right) \left( I_i^- \alpha_i^+ + I_i^+ \alpha_i^- \right) \right] \right\}.
\]

\[
\sigma_{d(t)} = \frac{(1 + \delta_i^-) I_i^+ z_{id(t)}^+}{1 - \delta_i^{-}} + \left( \frac{r_{d(t)}^+}{RT} \right) \sum_i \left[ I_i^+ \alpha_i^+ - \frac{\delta_i^{-} (I_i^- \alpha_i^+ + I_i^+ \alpha_i^-)}{\delta_i^{-}} \right].
\]

When the diffusion rates are sufficiently large, i.e., when \( \delta_i \)'s approach zero, \( \delta_i \) of Eq. (15) becomes unity and hence we have the previous equations:

\[
\gamma_s = - \sum_i (I_i^+ z_i^+ - I_i^- z_i^-) \times \rho
\]

\[
\gamma_{d(t)} = - I_i^+ z_{id(t)}^+ \rho, \quad r_{D(t)} = I_i^+ z_{D(t)}^+ \rho
\]

\[
a_i^- \text{ of Eq. (16) becomes at } \delta_i = 0, \text{i.e., at } \delta_i = 1, \delta_i^+ = \delta_i^- = 0,
\]

\[
\sigma_i = \sum_i \left[ I_i^+ z_i^+ + \gamma_s (F/RT) \alpha_i^+ I_i^+ \right] \sum_{i} \alpha_i^+ I_{i}\!
\]

\[
\sigma_{d(t)} = I_i^+ z_{id(t)}^+ + \gamma_{d(t)} (F/RT) \sum_i \alpha_i^+ I_{i}\!
\]

\[
\sigma_{D(t)} = r_{D(t)} (F/RT) \sum_i \alpha_i^+ I_{i}\!
\]

Now let us consider the two-reaction system, where the mixed potential is determined by two reactions, one of which is the anodic \( (t=1) \) and the other is the cathodic reaction \( (t=2) \). In the two-reaction system, Eq. (4) gives

\[
I_{1t} = - I_{2t} = I_m.
\]

(i): When \( V_t \) is far apart from the equilibrium potentials of the respective reactions \( (t=1, 2) \),

* The term, \( \delta_i^{-}(1 + \delta_i^-) I_i^+ z_{id(t)}^+ / \delta_i \), of \( \sigma_{d(t)} \) becomes \( I_i^+ z_{id(t)}^+ \) at \( \delta_i = 0 \) from the expressions for \( \delta_i^+, \delta_i^- \) and \( \delta_i \) of Eq. (15).
Fig. 1. Open circuit systems treated in the text as an example under the condition of $|V_r-V_{t=\text{eq}}| \gg RT/\alpha_1^+ F$, $RT/\alpha_1^- F$, $t=1$ (anodic), $2$ (cathodic).
(a) no concentration polarizations, (b) concentration polarization for $t=2$, and (c) diffusion control for $t=2$.

The unidirectional rates, $I_{r1}$ and $I_{r2}$, of the anodic and cathodic reactions are neglected and we have $I_{r1}=I_{r1}^+$ and $I_{r2}=-I_{r2}^-$. This is the case called "the mixed control" (Fig. 1). Now, $A_{+1}$ and $A_{-2}$ of Eq. (15) equal $A_{+1}^+$ and $-A_{-2}^-$, respectively. Thus, Eqs. (11) to (15) become

$$\gamma_s = -\frac{RT}{F} \cdot \frac{z_{2a}^+/\delta_1 - z_{2b}^-/\delta_2}{\alpha_1^+/\delta_1 + \alpha_2^-/\delta_2}$$

$$\gamma_{A_{+1}} = -\frac{RT}{F} \cdot \frac{z_{2a}(1-A_{+1}^+)}{\alpha_1^+/\delta_1 + \alpha_2^-/\delta_2}$$

$$\gamma_{A_{-2}} = \frac{RT}{F} \cdot \frac{z_{2b}(1-A_{-2}^-)}{\alpha_1^+/\delta_1 + \alpha_2^-/\delta_2}$$

$$1/\rho = (F/RT)(\gamma_{\text{in}}(\alpha_1^+/\delta_1 + \alpha_2^-/\delta_2))$$

$$(18)$$

$$(19)$$

$$(20)$$
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From Eq. (16), \( \sigma \)'s become

\[
\sigma_s = I_m \left( z_{ls}^+ + (F/RT) \gamma_s \alpha_s^+ \right) / \delta_1
\]

\[
\sigma_{d(1)} = I_m \left[ z_{d(1)}^+/ \left( 1 - D_{d(1)}^+ \right) + (F/RT) \gamma_{d(1)} \alpha_{d(1)}^+ \right] / \delta_1
\]

\[
\sigma_{D(2)} = I_m \left( F/RT \right) \gamma_{D(2)} \alpha_{D(2)}^+ / \delta_1,
\]

(22)

(a) \( D_1 = 0 \); When both of the mass transfer steps are sufficiently fast so that \( D_1 \)'s of Eq. (15) equal practically zero (Fig. 1, (a)), \( \delta_1 \)'s of Eq. (21) become unity. Equations (19) and (20), will then be

\[
\gamma_s = - \frac{(RT/F) (z_{ls}^- - z_{ls}^-)/ \delta_1}{\delta_1}
\]

\[
\gamma_{d(1)} = - \frac{(RT/F) z_{d(1)}^+/(\alpha_{d(1)}^+ + \alpha_{d(1)}^-)}{\delta_1}
\]

\[
\gamma_{D(2)} = \frac{(RT/F) z_{D(2)}^-/(\alpha_{D(2)}^+ + \alpha_{D(2)}^-)}{\delta_1}
\]

\[
1/\rho = (F/RT) I_m (\alpha_{ls}^+ + \alpha_{ls}^-)/\delta_1,
\]

(23)

which we have already derived. \( \sigma \)'s of Eq. (22) become

\[
\sigma_s = I_m \left[ z_{ls}^- \alpha_{ls}^- + z_{ls}^- \alpha_{ls}^+ \right] / \delta_1
\]

\[
\sigma_{d(1)} = I_m \left[ z_{d(1)}^- \alpha_{d(1)}^- \right] / \delta_1
\]

\[
\sigma_{D(2)} = I_m \left[ z_{D(2)}^- \alpha_{D(2)}^- \right] / \delta_1.
\]

(24)

The above expression of \( \sigma_s \) has been reported by DONAHUE.

(b) \( D_1^+ = 0 \); When only the anodic reaction is controlled by the surface process (Fig. 1 (b)), Eqs. (19) to (22) become

\[
\gamma_s = - \frac{(RT/F) (z_{ls}^+ - z_{ls}^-)/ \delta_2}{\delta_2}
\]

\[
\gamma_{d(1)} = - \frac{(RT/F) z_{d(1)}^+/(\alpha_{d(1)}^+ + \alpha_{d(1)}^-)}{\delta_2}
\]

\[
\gamma_{D(2)} = \frac{(RT/F) z_{D(2)}^-/(1 - D_{D(2)}^-)}{\delta_2}
\]

\[
1/\rho = (F/RT) I_m (\alpha_{ls}^+ + \alpha_{ls}^-)/\delta_2,
\]

(25)

\[
\delta_2 = 1 + D_{D(2)}^- \left( 1 - D_{D(2)}^- \right)
\]

\[
\sigma_s = I_m \left[ z_{ls}^- - \alpha_{ls}^- (z_{ls}^- - z_{ls}^-)/ \delta_2 \right]
\]

\[
\sigma_{d(1)} = I_m \left[ z_{d(1)}^- \alpha_{d(1)}^- \right] / \delta_2
\]

\[
\sigma_{D(2)} = I_m \left[ z_{D(2)}^- \alpha_{D(2)}^- \right] / \delta_2.
\]

(c) \( D_1^+ = 0, D_{D(2)}^- = 1 \); When the rate of the anodic reaction is controlled by the surface process and that of the cathodic reaction is controlled by
which have been derived previously\(^1\).

(b) \( \Delta_{+1} = 0 \); When only the anodic reaction is governed by the surface process (Fig. 2, (b)), i.e., \( \Delta_{+1} = 0 \), Eqs. (29) to (31), become

\[
\begin{align*}
\gamma_s &= - \left[ I_{+} z_{+}^{1} + (I_{+} z_{-}^{1} - I_{-} z_{-}^{2}) / \delta_{1} \right] \times \rho \\
\gamma_{d(1)} &= - I_{+} z_{+}^{1} \rho \\
1/\rho &= (F/RT) \left[ I_{+} \alpha_{+}^{1} + (I_{+} \alpha_{+}^{2} + I_{-} \alpha_{-}^{2}) / \delta_{1} \right].
\end{align*}
\]

Equations for \( \gamma_{d(2)} \), \( \gamma_{d(3)} \) and \( \delta_{2} \) are the same as those of Eqs. (30) and (31).

(c) \( \Delta_{+1} = 1 \); When the anodic reaction rate is governed by the diffusion step of \( d(1) \) (Fig. 2, (c)), i.e., \( \Delta_{+1} = 1 \), \( \delta_{1} \) becomes infinite. Hence,

\[
\begin{align*}
\gamma_s &= - (RT/F) (I_{+} z_{+}^{1} - I_{-} z_{-}^{2}) / (I_{+} \alpha_{+}^{1} + I_{-} \alpha_{-}^{2}) \\
\gamma_{d(1)} &= - (RT/F) I_{+} / (I_{+} \alpha_{+}^{1} + I_{-} \alpha_{-}^{2}) \\
1/\rho &= (F/RT) (I_{+} \alpha_{+}^{1} + I_{-} \alpha_{-}^{2}) / \delta_{1}.
\end{align*}
\]

Equations for \( \gamma_{d(2)} \), \( \gamma_{d(3)} \) and \( \delta_{2} \) are given in Eqs. (30) and (31).

(d) \( \Delta_{+1} = 1 \), \( \Delta_{-1} = 0 \); In addition, let the diffusion rates of \( d(2) \) and \( D(2) \) be sufficiently large, i.e., \( \Delta_{+1} = 1 \) and \( \Delta_{-2} = 0 \) (Fig. 2, (d)). In this case, \( \delta_{1} \) is infinite and \( \delta_{2} \) equals unity. Hence,

\[
\begin{align*}
\gamma_s &= - (RT/F) (I_{+} z_{+}^{1} - I_{-} z_{-}^{2}) / (I_{+} \alpha_{+}^{1} + I_{-} \alpha_{-}^{2}) \\
\gamma_{d(1)} &= - (RT/F) I_{+} / (I_{+} \alpha_{+}^{1} + I_{-} \alpha_{-}^{2}) \\
\gamma_{d(3)} &= I_{+} / \delta_{1} \\
1/\rho &= (F/RT) (I_{+} \alpha_{+}^{1} + I_{-} \alpha_{-}^{2})
\end{align*}
\]

\( \sigma \)'s for (a)~(d) will be calculated from Eq. (32) by using respective values of \( \Delta \)'s.

Case (2); One of \( d(1) \)'s and \( D(1) \)'s is common among more than one reactions

This will be the case for the system of e.g., \( \text{H}_2\text{O}_2 \) which is common for the anodic and cathodic reactions, i.e., \( d(1) = D(2) = d \). For the sake of simplicity, we treat here only the two-reaction system. Let \( t = 1 \) and 2 be the anodic and the cathodic reactions as we did, each satisfying the condition (18). The respective rate equations are

\[
\begin{align*}
I_{+} &= I_{+}^{\text{H}} = k_{+}^{\text{H}} H \tilde{C}^{\text{H}} \left[ C_{a}(1 - I_{+}^{\text{H}} / I_{L}) \right]^{\text{H}} \exp (\alpha_{+}^{1} V, F/RT) \\
I_{-} &= -I_{-} = k_{-}^{\text{H}} H \tilde{C}^{\text{H}} \left[ C_{a}(1 - I_{-} / I_{L}) \right]^{\text{H}} \exp (-\alpha_{-}^{1} V, F/RT).
\end{align*}
\]
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where $I_L$ is the limiting current which is common for the reactions, $t=1$ and 2. $I_L$ is different from the diffusion limiting current of $d$, since the diffused $d$ is consumed by both reactions.

The concentration dependences of $V_r$ and $I_m$ are given from the expressions of (7) to (9) and Eq. (37) as

$$\gamma_s = -\frac{RT}{F} \cdot \frac{(z_{\alpha_1}^+ / \delta_1 - z_{\alpha_2}^- / \delta_2)}{(\alpha_1^+ / \delta_1 + \alpha_2^- / \delta_2)}$$

$$\gamma_d = -\frac{RT}{F(1-\Delta)} \cdot \frac{(z_{\alpha_1}^+ / \delta_1 - z_{\alpha_2}^- / \delta_2)}{(\alpha_1^+ / \delta_1 + \alpha_2^- / \delta_2)}$$

$$1/\rho = (F/RT) I_m(1-\Delta) (\alpha_1^+ / \delta_1 + \alpha_2^- / \delta_2)$$

$$\sigma_s = \left[ I_m(1-\Delta) / \delta_1 \right] \left[ z_{\alpha_1}^+ + (F/RT) \gamma_s \alpha_1^+ \right]$$

$$\sigma_d = \left[ I_m / \delta_1 \right] \left[ z_{\alpha_1}^+ + (F/RT) (1-\Delta) \gamma_d \alpha_1^+ \right]$$

$$\sigma_d = \left[ I_m / \delta_2 \right] \left[ z_{\alpha_2}^- - (F/RT) (1-\Delta) \gamma_d \alpha_2^- \right]$$

$$\delta_1 = 1 - (1 - z_{\alpha_1}^+) \Delta, \quad \delta_2 = 1 - (1 - z_{\alpha_2}^-) \Delta$$

$$\Delta = I_m / I_L, \quad I_m = I_{I_r} = I_{I_2}.$$

(a) $\Delta = 0$; When the diffusion rate of $d$ is sufficiently large, $\Delta$ becomes zero. Hence, the both reaction rates are controlled by the surface process. This case has been already treated and the above equations reproduce Eq. (23). $\sigma_s$ is the same as that of Eq. (23) and $\sigma_d$ is given as the sum of $\sigma_{d(1)}$ and $\sigma_{d(2)}$ by putting $d(1) = D(2) = d$.

(b) $\Delta = 1$; On the contrary, if the diffusion rate of $d$ is small and the both reactions are controlled by the diffusion, values of $\gamma_s$ and $\sigma_s$ become zero because of $z_{\alpha_1}^+ = z_{\alpha_2}^- = 0$ and $\Delta = 1$, whereas $\gamma_d$ becomes indefinite. The reaction resistance is infinite as expected from the independence of the limiting diffusion rate on the electrode potential. $\sigma_d$, i.e., the change of $I_m$ with $C_d$, becomes

$$\sigma_d = I_m \text{ or } \partial \ln I_m / \partial \ln C_d = 1.$$

Discussion

In the foregoing sections, we first developed three fundamental expressions, (7) to (9), for the concentration dependence of $V_r$, the concentration dependence of $I_m$, and the reaction resistance at $V_r$. These expressions
are unconditionally valid in any systems. If the rate equations of each reaction are available, one can estimate the behaviour of the mixed potential from these expressions. On the other hand, it may happen that some of the kinetic parameters are hardly determined by experiments. In such a case, the observed behaviours of the mixed potential and of \( I_m, i.e., \partial_s, \partial_i, \partial_2, \cdots, \partial_s, \partial_s, \cdots \), and the reaction resistance, may be used to determine these unknown parameters.

In the present analysis, we assumed the rate equation of the type given by Eq. (10). Equation (10) includes the concentration polarization of only one reactant and one product. This might be the most probable case and be taken general enough, for an introduction of other concentration polarization factors will not cause to any essential difficulty in the analysis except for the increase of terms in the expressions of \( \partial_s \) and \( \rho \).

We have treated above typical cases which may have practical importance. Some of them will be exemplified below.

i) No concentration polarization

**H\(_2\)O\(_2\)-Hg system**; The present authors\(^7\) studied the kinetics of the anodic and cathodic reactions on Hg in acidic and alkaline H\(_2\)O\(_2\) solutions and concluded that the mixed potential is determined by the anodic reaction, \( 2\text{Hg} + \text{SO}_4^{2-} \rightarrow \text{Hg}_2\text{SO}_4 + 2\text{e}^- \) (acidic) or \( \text{H}_2\text{O}_2 \rightarrow \text{O}_2 + 2\text{H}^+ + 2\text{e}^- \) (alkaline) and the cathodic reaction, \( \text{H}_2\text{O}_2 + 2\text{H}^+ + 2\text{e}^- \rightarrow 2\text{H}_2\text{O} \) (acidic) or \( \text{H}_2\text{O}_2 + 2\text{e}^- \rightarrow 2\text{OH}^- \) (alkaline). This was the case of the two-reaction system, each reaction satisfying the condition (18). Calculated concentration dependences of \( V_r \) from equation (23), \( \partial_{\text{H}_2\text{O}_2}, \partial_{\text{H}^+}, \partial_{\text{SO}_4^{2-}} \), are in excellent agreement with the observed ones\(^1\).

**H\(_2\)O\(_2\)-Pt system**; Ostrovidova et al.\(^6\) studied the anodic and cathodic reactions on Pt in alkaline H\(_2\)O\(_2\) solution. This is another example of the two-reaction system which satisfies the condition (18). They calculated the value of \( \partial_{\text{H}_2\text{O}_2} \) and \( \partial_{\text{H}^+} \) directly from their rate equations and showed good agreement with the observed ones.

It will be emphasized that, in principle, the behaviour of \( V_r \) must be analyzed by the rate equations obtained around the mixed potential. In the above two examples, however, the concentration dependences were calculated by using the kinetic parameters, determined experimentally in the potential range far from the mixed potential. The excellent agreement between the calculated and observed ones indicates that the rate equations observed at high anodic and cathodic polarizations are maintained around the mixed potential.
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ii) Concentration polarization

Very few work have been reported on the system where the concentration polarization is present. HERBELIN et al\(^4\) analyzed the mixed potential of the two-reaction system where one of the reactions is diffusion controlled. They derived first the equation of the mixed potential and then deduced that to \(\tau\) by differentiation. Such a procedure is not applicable to the cases where the formula of \(V_r\) is not given by an explicit function of the observed kinetic parameters.

References