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EFFECT OF OXYGEN FUGACITY ON FASSAITIC PYROXENE

by

Kosuke Onuma

(with 9 text-figures)

Abstract

Stability of fassaitic pyroxene in the system $CaMgSi_2O_6-CaFeAlSiO_6-CaA_2SiO_6-CaTiA_1O_6$ is discussed on the basis of published data. In the system $CaMgSi_2O_6-CaFeAlSiO_6$ the stability field of clinopyroxene decreases with an increase in $CaFeAlSiO_6$ content and a decrease in oxygen fugacity at constant temperature. The stability field of clinopyroxene in the system $CaMgSi_2O_6-CaAlFeSiO_4-CaAl_2SiO_6$ is strongly influenced by oxygen fugacity at low and high pressure and decreases with respect to $CaFeAlSiO_6$ component. Titanium content of the clinopyroxene does not seem to be affected by the change of oxygen fugacity when magnetite is stable. These data are probably useful in evaluating the condition under which natural fassaitic pyroxenes are formed.

Introduction

The clinopyroxenes in undersaturated alkalic rocks and also in some metamorphic rocks are rich in $CaO$, $Al_2O_3$, $Fe_2O_3$, and $TiO_2$, and most of them plot on the wollastonite-rich region beyond the diopside-hedenbergite join in the pyroxene quadrilateral. The main components of such pyroxenes are $CaMgSi_2O_6$ (Di), $CaFeSi_3O_6$ (Hd), $CaTiAl_2O_6$ (Ti-pyroxene component, Tp), $CaAl_2SiO_6$ (Ca-Tschoermak’s component, CaTs), $CaFe^3+AlSiO_6$ (ferri-aluminum Tschermak’s component, FATs), $NaFe^3+Si_2O_6$ (Ac), and $Mg_2Si_2O_6$ (En) (Onuma and Yagi, 1971). According to Deer et al. (1978), the name fassaite is used to Al ($Fe^3+$)-rich clinopyroxenes in which most of the M2 positions are occupied by Ca, and the introduction of the trivariant cations into M1 is compensated almost entirely by the replacement of Si by Al in tetrahedral site. From this view point the clinopyroxenes containing $CaFe^3+AlSiO_6$ and $CaAl_2SiO_6$ components are fassaitic pyroxene or fassaite. Onuma and Yagi (1975) and Onuma et al. (1981) emphasized the significance of the system Di-CaTs-FATs-Tp for understanding the crystallization of pyroxene in alkali rocks.

Since these pyroxenes contain $CaFe^3+AlSiO_6$ and $CaFe^2+Si_2O_6$, their composition and evolutionary trend must be significantly affected by the variation of oxygen fugacity of the magma from which the pyroxenes crystallize. Therefore, for evaluating the effect of oxygen fugacity on the fassaitic pyroxenes, the author has been engaged in the experimental study on the systems involving FATs component with co-workers. In this paper are reviewed the results of the experimental studies at controlled oxygen fugacity in our laboratory.

Effect of Oxygen Fugacity on Stability of Fassaite

The FATs component, which is regarded as a main component of $Fe^3+$-rich fassaite, is a stable compound with clinopyroxene structure $C2/c$ at 1 atm in air below 1250°C (Hijikata, 1968; Huckenholz et al., 1974), and there is a complete series of solid solution between

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Onuma diopside and FATs under this condition (Hijikata & Onuma, 1969). Oba and Onuma (1978) made an experiment on this system with varying oxygen fugacity to clarify the effect of oxygen fugacity on the stability of Fe$^{3+}$-rich fassaite, and found that the stability field of pyroxene is restricted at low $fO_2$. Text-fig. 1 shows the change of clinopyroxene one-phase

**Text-fig. 1** Pyroxene one-phase field in the Di-FATs system at low $fO_2$. From data of Oba and Onuma (1978).

**Text-fig. 2** Stability field of the Di-FATs pyroxene for $fO_2$ at 1100°C. From data of Oba and Onuma (1978).
FASSAITIC PYROXENE

Field with a change in $fO_2$. The one-phase field decreases with increasing temperature at constant $fO_2$ and decreases considerably with decreasing $fO_2$ at constant temperature as also shown in Text-fig. 2; 27 wt% at $10^{-9}$ and 14 wt% at $10^{-10}$ at 1100°C and 8 wt% at $10^{-11}$, 1050°C. Two-phase field of Cpx + Mel is present to the higher temperature and lower $fO_2$ side of one-phase field. As demonstrated in Text-fig. 3, the pyroxene one-phase field is stable at least above the $fO_2$ defined by MW buffer, while the assemblage Cpx + Mel is present in the field where Wu is stabilized. Therefore, the melilite may contain $Ca_2Fe^{2+}Si_2O_7$ component as expected from the equation, $CaMgSi_2O_6 + CaFe^{3+}AlSiO_6 = Ca_2Fe^{2+}Si_2O_7 + MgSiO_3 + 1/2Al_2O_3 + 3/4O_2$, and clinopyroxene would become to contain Mg-Tschermak’s component. At low $fO_2$ clinopyroxene becomes hedenbergitic due to the presence of $Fe^{2+}$. In the Di-poor portion the following reactions may take place and spinel is present.

$$2CaMgSi_2O_6 + 6CaFe^{3+}AlSiO_6 = 4Ca_2Fe^{2+}Si_2O_7 + MgSiO_3$$

$$+ Fe^{2+}SiO_3 + Al_2O_3 + MgAl_2O_4 + FeAl_2O_4 + 3/2O_2$$

pyroxene
spinel

$$3CaFe^{3+}AlSiO_6 = Ca_2Fe^{2+}Si_2O_7 + CaSiO_3 + 3/2FeAl_2O_4 + 1/6Fe_3O_4$$

Thus, $CaFe^{2+}Si_2O_6$ component is expected in the clinopyroxene of the magnetite stability field. Text-fig. 3 shows the stability of clinopyroxenes with various compositions in the Di-FATs system over temperature and $fO_2$. The clinopyroxene poorer in Di than 70 wt% is not stable at the temperature and $fO_2$ ranges studied, and therefore the broken lines are drawn tentatively. Text-fig. 3 indicates that the stability field of clinopyroxene for temperature and $fO_2$ decreases with an increase in the FATs content of clinopyroxene.

Text-fig. 3 Stability of the Di-FATs pyroxene for $fO_2$ and temperature. The pyroxene is stable above curves.
The stability of FATs-pyroxene at high pressure has been studied by Ohashi and Hariya (1975 a, b). The FATs-pyroxene is stable up to about 43 kbar and 1300°C under the condition where hematite is stabilized, and above this pressure and temperature decomposes into garnet + oxide as suggested by Hijikata and Onuma (1969). On the other hand, the stability field of this pyroxene decreases with decreasing $f_{O_2}$, and the clinopyroxene decomposes into Cpx + Gar + Sp at least 6 kbar at the $f_{O_2}$ defined by $Fe_2O_3 - Fe_2O_4$ and $MnO - Mn_3O_4$ buffers. The observations mentioned above indicate that there is a complete series of solid solution between Di and FATs from 1 atm to 45 kbar under the condition where hematite is stabilized, and that the stability of fassaitic pyroxene of Di-FATs system is more sensitive for change of $f_{O_2}$ than pressure.

The phase relation in the system Di-FATs-CaTs at 1 atm in air was studied by Onuma and Yagi (1975) and Onuma et al. (1981). The phase relation at liquidus is shown in Text-fig. 4. Clinopyroxene, anorthite, spinel, and magnetite are present as primary phase. In the liquidus diagram there are two points showing four-phase assemblage; one at 1250°C, shows the liquid coexisting with Cpx + An + Sp and the other, at 1270°C, the liquid coexisting with Cpx + Mt + Sp. These points, however, are neither eutectic nor piercing points because of the nature of the six-component system at liquidus temperature. At subsolidus temperatures there are present one-phase field of pyroxene, Cpx + An + Mel, and Cpx + An + Mel + Sp, and the limit of one-phase field of pyroxene is shown with a heavy broken line in Text-fig. 4.

Text-fig. 4 Liquidus diagram of the Di-FATs-CaTs system in air at 1 atm (Onuma et al., 1981). Broken line shows the limit of pyroxene one-phase field at subsolidus.
An experimental study for the Di-CaTs system at 1 atm has made by Schairer and Yoder (1970), who demonstrated that the clinopyroxene one-phase field attains 12 wt% CaTs at about 1250°C. The substitution of CaTs in diopside, according to Clark et al. (1962), is favored by high pressure. The lowest stability of CaTs was found by Hays (1966) and Hijikata and Yagi (1967) to be located at about 11 kbar and 1100°C. Below this temperature and pressure the clinopyroxene breaks down into Cpx + An + Sp. Hijikata (1973) determined the P-T stability of clinopyroxene in the Di-CaTs system and demonstrated that the lowest stability of Di₅₀CaTs₅₀ and Di₁₂₅CaTs₇₅ is 5.8 kbar, 975°C and 8.8 kbar, 1010°C, respectively. These data show that the stability field of clinopyroxene increase with increasing pressure.

On the basis of the data stated above, we can estimate the stability field of fassaitic pyroxene over pressure and fO₂ in the Di-CaTs-FATs system. The stability field is strongly influenced by fO₂ at low and high pressure and decrease with respect to the FATs component with decreasing fO₂ as shown in Text-fig. 5. The clinopyroxene is stable in the region to the left of the curve. There exists only small clinopyroxene one-phase field at 10⁻¹¹ atm. The stability field of pyroxene, however, expands from the Di-FATs line toward the CaTs apex, indicating that mostly the CaTs component is affected by the change of pressure.

Although the earlier works (Oba and Onuma, 1978; Schairer and Yoder, 1970; Hays, 1966; Hijikata and Yagi, 1967) did not give the chemical compositions of the pyroxenes, since the composition of pyroxene in the Di-FATs-CaTs system (Onuma et al., 1981) and
also in the Di-CaTs system (Yang, 1975) varies beyond the one-phase field, the stability limit
does not necessarily mean the solubility limit of the FATs and CaTs components in
pyroxene. Nevertheless, it is expected that pressure and $fO_2$ give arise the same effect on the
composition of solid solution as on the stability limit. It is therefore concluded that FATs
content of fassaite pyroxene is influenced by $fO_2$ and is independent of pressure, while
CaTs content is effected by pressure.

Ti-Fassaite

The phase relation in the Di-FATs-Tp system at 1 atm in air was determined by Akasaka
and Onuma (1979). A liquidus diagram is shown in Text-fig. 6. Clinopyroxene, forsterite,
perovskite, magnetite, spinel, and hibonite are present. An unknown phase was encountered.
It was first found by Hijikata and Onuma (1969) in the Di-FATs system and named "phase
X". Although there are four points showing a four-phase assemblage, these points are again
neither eutectic nor piercing points, because the system belongs to the seven-component
system Fe-O-CaO-MaO-Al$_2$O$_3$-TiO$_2$-SiO$_2$ at liquidus temperature.

In the 10 wt% Tp section, a complete series of solid solution of pyroxene is present at
subsolidus temperatures. In the portion more than 10% Tp, the pyroxene one-phase field is
no more stable, and perovskite or Pv + Mel, + An appears in addition to the pyroxene solid
solution. The limit of pyroxene one-phase field is shown as broken line in Text-fig. 6.

Akasaka (1981) made an experimental study on this system at $10^{-11}$ atm $fO_2$. No
one-phase field of clinopyroxene was observed, and the fields of Cpx + Mel, Cpx + Mel + Sp,
and Cpx + Mel + Sp + An were present instead at subsolidus temperatures in the
compositions with 10% Tp. Akasaka (1981) analysed the clinopyroxenes and the melilite by
Mössbauer spectra method as well as microprobe, and demonstrate that the clinopyroxenes

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Text-fig. 6 Liquidus diagram of
the Di-FATs-Tp system in air
at 1 atm (Akasaka and
Onuma, 1979). Broken line
shows the limit of pyroxene
one-phase field.
contain 9 wt% CaFe$^{2+}$Si$_2$O$_6$ as well as FATs component, supporting the estimated reaction of Oba and Onuma (1978). The melilite was also confirmed to contain at least 35.5 wt% Ca$_2$Fe$^{3+}$Si$_2$O$_7$. He also showed that Ti content of the clinopyroxene is not too much affected by the change of $f$O$_2$ and the clinopyroxenes crystallizing from the compositions with 10 wt% Tp also contain about 10 wt% Tp (3.4 wt% TiO$_2$).

**Significance of the Di-CaTs-FATs-Tp System to the Natural Pyroxene**

The significance of the Di-CaTs-FATs-Tp system to the Natural fassaitic pyroxene has been discussed by Akasaka and Onuma (1979), Onuma and Akasaka (1980), and Onuma et al. (1981). The main points in these discussion are briefly stated here.

Text-Fig. 7 shows the plots of metamorphic Ca-pyroxene from various localities in the Di-CaTs-FATs triangle. Most pyroxenes fall in the clinopyroxene stability field below 6 kbar at higher $f$O$_2$ where hematite or magnetite is stabilized, or at least in the field of Cpx + An + Mel, where the Al and Fe$^{3+}$ contents of pyroxene vary continuously at 1 atm in air. This indicates that it is possible for metamorphosed fassait to be formed even at low pressure (at least below 6 kbar and even at 1 atm) if $f$O$_2$ is high enough to stabilize hematite or magnetite. As stated before, pyroxenes contain a considerable amount of FATs at high pressure under the condition where hematite is stable. However, it is unlikely that magma undergoes such a high $f$O$_2$. If $f$O$_2$ is lowered, FATs-rich pyroxene becomes unstable and decomposes into CATs-rich pyroxene and some other phases at high pressure. The FATs-pyroxene decomposes into Cpx + Gar + Sp at 6 kbar, 1000°C at the $f$O$_2$ defined by
MnO-Mn$_2$O$_4$ buffer, where magnetite is stable (Ohashi and Hariya, 1975 b). According to Onuma and Hariya (unpublished data), a pyroxene having the composition of Di$_{70}$FAT$_{15}$-CaT$_{15}$ decomposes into aluminous pyroxene, garnet, and spinel at 10 kbar, 1000°C under the same condition. Judging from the data of Oba and Onuma (1978), this pyroxene may stable under this condition at 1 atm. These experimental results and chemical compositions of natural fassaitic pyroxenes suggest that the volcanic rocks including fassaitic pyroxene, such as melilitite, nephelinite, basanite, etc., formed at rather lower pressure (less than 10 kbar), if $fO_2$ is lower than that defined by MnO-Mn$_2$O$_4$ buffer, where magnetite is stable.

Text-fig 8 shows the plots of fassaitic pyroxenes from Hocheifel (Huckenholz et al., 1965 a, b, 1966) and Westeifel (Becker, 1977), West Germany in the Di-CaT$_6$-FAT$_5$ system. In general, the compositions of cores plot in the Di-richer portion, while those of rims in the Di-poorer potion, indicating that the pyroxenes gradually become rich in FATs as crystallization proceeds. The crystallization trends cross the stability limit of the pyroxene for $fO_2$. Since Fe$^{2+}$ is calculated as Ca(Mg, Fe$^{2+}$)$_2$Si$_2$O$_6$, an increase in FATs means a decrease in the hedenbergite component, in other wards, an increase in $fO_2$. For this reason, it is suggested that in the crystallization trends of these pyroxenes $fO_2$ increases as crystallization proceeds and temperature falls.

The plots of the clinopyroxene of Tahiti (Tracy and Robinson, 1977) and Hocheifel (Huckenholz, 1965) on the Di-FATs-Tp plane are shown in Text-fig 9, revealing two trends of clinopyroxene composition from core to rim; the trend of Tahiti pyroxene is approximately parallel to the Di-Tp join, indicating enrichment of Tp at constant FATs.
(about 8 wt%), whereas in the Hocheifel pyroxene both Tp and FATs contents increase as crystallization proceeds, and FATs attains 18 wt%. Tracy and Robinson (1977) considered that low fO2 is one cause of crystallization of pyroxenes extremely rich in Ti. The experimental results of the Di-FATs-Tp system in air however indicate that the pyroxene crystallizing in air can contain TiO2 as much as 8 wt%, suggesting that when oxide minerals are absent low fO2 is not a necessary condition to the entry of Ti into the pyroxene structure and the Ti-rich fassaitic pyroxenes are formed under equilibrium condition from the liquids rich in TiO2 and poor in SiO2. Based on the experimental results of the Di-FATs system at low fO2, however, it can be said that a low FATs content in the Tahiti pyroxene implies the crystallization at low fO2. Therefore, the difference in the pyroxene trend is probably due to the difference of fO2 at which the pyroxenes crystallized from the liquid, and many intermediate trends of clinopyroxenes in undersaturated alkalic rocks between the extreme cases shown in Text-fig. 9 are postulated, depending on the bulk chemistry and fO2 of the magma from which the pyroxene crystallized.

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