INTRODUCTION

The contribution of methane (CH₄) to global warming is estimated to be about 20% (IPCC, 2007). Enormous quantities of carbon occur in marine sediments in the form of CH₄ (Kvenvolden, 1988) and at least some CH₄ in marine sediments seeps into the ocean and thus into the atmosphere (Etiope and Milkov, 2004; Judd, 2004; Mau et al., 2007). The mechanisms and fate of this seeping as well as its quantification are poorly constrained. In particular, CH₄ emitted from deep seafloor sources has been ignored to date in estimating atmospheric CH₄ budgets (Etiope and Milkov, 2004; Judd, 2004). Indeed, most CH₄ would be oxidized within bottom water if vertical eddy diffusion in the water column is the only process for transporting CH₄ from water depths greater than 2000 m to surface (Merothewet al., 1985; Paull et al., 1995; Heesch et al., 2003; Sauter et al., 2006). In addition, episodic huge discharges of CH₄ due to massive decomposition of
Seafloor methane hydrates have been suggested as possible causes of sudden climate changes in the geological past (MacDonald, 1990; Dickens et al., 1997; Bains et al., 1999; Svensen et al., 2004). Besides having environmental significance, seafloor CH4 bubbles might also be a possible geohazard and have a resource significance (Sills and Wheeler, 1992).

In 1997, gas bubbling with large-scale chemosynthetic communities and carbonate crusts including chimney-like structures were found on top of Kuroshima Knoll (depth of ca. 640 m), located in the forearc region of the Ryukyu island arc (Matsumoto et al., 1998) (Figs. 1 and 2). Subsequent studies revealed that the bubbles were composed mostly of CH4 (Machiyama et al., 2001). Based on both δ13C and δ18O values of the carbonate crusts, Takeuchi et al. (2007) concluded that methane hydrate layers had developed under the area. The CH4-rich fluid derived from the partial dissociation of the methane hydrate produced both the bubbles and the carbonate crusts (Takeuchi et al., 2007). They assumed that subtle changes in the marine environment were responsible for the dissociation of methane hydrate. If this assumption is correct, the area is an important field for studying the dissociation process for seafloor methane hydrate and discussing future possible changes in seafloor CH4 flux to the ocean and atmosphere. However, we had little evidence for the existence of methane hydrate beneath the seafloor of the knoll and for the progression of dissociation. In addition, the influence of CH4 discharge on the surrounding marine environment was poorly constrained.

In this study, we collected samples of both gas bubbles and bottom fluid from the active gas seeping site of Kuroshima Knoll (Fig. 2) using the WHATS gas-tight sampler (Water and Hydrothermal-fluid Atsuryoku Tigh

![Fig. 1. Locations of Kuroshima Knoll (square), Taketomi spring (circle), and submarine Quaternary volcanic features (▲) (modified from Machiyama et al., 2001). The detailed bottom topography in the square (the area around Kuroshima Knoll) is presented in Fig. 2.](image_url)
Deep-sea seeping methane bubbles in the Ryukyu forearc

(367x164) Sampler) (Tsunogai et al., 2003; Saegusa et al., 2006) attached to the vehicle Shinkai 2000 to determine the origin of CH$_4$ and discuss the possible sub-seafloor dissociation of methane hydrate using geochemical tracers. Knowledge of the origin of CH$_4$ (i.e., biogenic from acetate fermentation or CO$_2$ reduction, thermogenic, inorganic) and the other coexisting gases would provide information on the processes and environments responsible for their generation. Thus seepage gases could be used as a tracer for sub-seafloor hydrocarbon reservoirs and of secondary processes, such as migration through geodynamic processes (Price and Schoell, 1995; Etiope et al., 2009), microbial alteration (Head et al., 2003; Kinnaman et al., 2007), and gas hydrate formation/decomposition (Sassen et al., 2001; Milkov et al., 2004).

Besides the gas bubbles and seeping fluid at the seafloor, the effluent plume in the water column was sampled in and around the knoll to evaluate the distribution and the fate of the CH$_4$ bubbles in the ocean water column, with special emphasis on the influence of aerobic CH$_4$ oxidation on the fate of CH$_4$ in the water column. The distribution of the concentration and stable carbon isotopic composition ($\delta^{13}C$) of CH$_4$ in a CH$_4$-enriched plume are good tracers for determining the fate of CH$_4$ in a water column, especially for quantitatively determining the progress of microbial oxidation (Tsunogai et al., 1998, 2000, 2005; Nishimura et al., 1999). Indeed, CH$_4$ oxidation consumes $^{12}CH_4$ preferentially, leaving the residual CH$_4$ enriched in $^{13}CH_4$ (Silverman and Oyama, 1968), while $\delta^{13}C_C$ does not vary through eddy diffusion. In addition, it was shown that the kinetic isotope effect (KIE) $\alpha$ is around 1.005 at the bottom of the water column (Tsunogai et al., 2000). $\alpha$ is defined as the ratio of the relative reaction rates of molecules containing different isotopes:

$$\alpha = \frac{k_{13}}{k_{12}}$$  \hspace{1cm} (1)

where $k_{13}$ and $k_{12}$ are the respective rate constants for

$$^{12}CH_4 + 2O_2 \rightarrow ^{13}CO_2 + 2H_2O \hspace{1cm} (2)$$

and

$$^{12}CH_4 + 2O_2 \rightarrow ^{12}CO_2 + 2H_2O. \hspace{1cm} (3)$$

Fig. 2. The bottom topography of Kuroshima Knoll and the locations of the gas seeping site (○) and the sampling stations for hydrocasts (○, Stn. 01-0; □, Stn. 01-I; △, Stn. 01-2) (modified from Machiyama et al., 2001).
Using the distribution of $\delta^{13}C_{CH_4}$ in a CH$_4$-enriched plume and the determined KIE, we traced the progress of CH$_4$ oxidation in the CH$_4$-enriched plume and determined whether the main cause of the CH$_4$ decrease in the water column was eddy diffusion (dilution) or oxidation.

**GEOLOGICAL BACKGROUND**

The Kuroshima Knoll (Figs. 1 and 2) is about 26 km south of the island of Ishigaki in the forearc region of the southern Ryukyu arc. The knoll has a flat table for its top at a depth of 1000 to 630 m, while the foot part reaches to a depth of 2400 m. On the basis of geological observations on the mudstones, sandstones, and mud cores recovered from the knoll, Cenozoic sediments, such as the Pliocene Shimajiri Group, are assumed to cover most of its surface (Matsumoto et al., 1998; Takeuchi et al., 2007).

The Cenozoic of the Ryukyu arc is divided into the lower Eocene Wano and Kayo Formations, upper Eocene Miyara Group, lower Miocene Yaeyama Group, Sonai conglomerate of unknown age, middle Miocene Takarajima Group, Pliocene Shimajiri Group, middle to upper Pleistocene Rykyu Group and Holocene coastal and terrestrial deposits (Nakagawa, 1983). The region of the southern Ryukyu arc, including the islands of Ishigaki and Iriomote (Fig. 1), can be characterized by volcanic activity in the Miocene, such as the 21 Ma old granitic intrusions in Ishigaki (Kawano and Ueda, 1966) and volcanic breccia, tuff breccia, and lava flows along with small dikes in Iriomote (Monden, 1968; Tiba and Saito, 1974). On the other hand, the Pliocene Shimajiri Group consists mostly of fossiliferous sandstone and siltstone. Except for the backarc region (Fig. 1), no volcanic activity has been found in the southern Ryukyu forearc area from the Pliocene to the present (Nakagawa, 1983).

Large-scale colonies of *Calyptogena kawamurai* and *Bathymodiolus* spp. (both living and dead) (Okutani et al., 2004), calcareous crusts, many gravels of sand chimneys, and emissions of gas bubbles have been found on the seafloor in diving surveys (Matsumoto et al., 1998; Machiyama et al., 2001). Active microbial oxidation of CH$_4$ (both aerobic and anaerobic) has been found in sediments just beneath the seafloor, especially in areas that can be characterized by slow, diffusive seeping of CH$_4$ gas or CH$_4$-enriched fluid (Inagaki et al., 2004).

On the basis of the $^{13}$C-depleted carbon isotopic ratios (from −4.1 to −43.2‰ VPDB) and the $^{3}$He-enrichment (from +2.5 to +7.9‰ VPDB) of the carbonates, Takeuchi et al. (2007) proposed that methane hydrate layers could have developed under Kuroshima Knoll. The CH$_4$-rich,
than the cause of the dissociation of methane hydrate. Other some subtle changes in the marine environment as possi-
70 m layer of the knoll, Takeuchi (JAMSTEC) during the NT02-08 cruise in June 2002,
Because methane hydrate is theoretically stable in the top
using the WHATS (Tsunogai (JAMSTEC) during the NT02-08 cruise in June 2002,
above the seafloor of the gas seepage site) were collected
by the human occupied vehicle
under the seafloor of the knoll.
bles and carbonates on the seafloor (Takeuchi 2006). The sampler comprised four 150 cm³ stainless steel
connected to the inlet. The motor-driven arm on the rail is
distributed inside the submersible cabin.
ple fluid/gas from the inlet into the cylinder. After com-
used to open one of the four cylinders. By pumping out
seawater that had filled the cylinder, we introduced sam-
ple fluid/gas from the inlet into the cylinder. After com-
pletely replacing the seawater with sample, the valves at
both ends of the cylinder are tightly closed by moving
the motor-driven arm again. At this stage, the next cylin-
der is ready for sampling. The entire operation was con-
ducted inside the submersible cabin.
In dives for sampling bottom fluid, a titanium inlet
tube was used for the sample inlet, while a specially de-
dsigned funnel-shaped inlet was used for sampling gas
bubbles (Saegusa et al., 2006). Because we could not rec-
ognize fluid seeping from Kuroshima Knoll, we took
seawater less than 30 cm above the seafloor, in areas
where active fluid seeping was anticipated, such as areas
covered by dense Bathymodiolus mussels. A temperature
probe was attached to the titanium inlet tube to measure
the temperature of the incoming bottom fluid. The bot-
tom fluid samples, however, showed temperatures close
to that of the ambient seawater (about 8.0°C), so that we
could not clarify the difference from the ambient seawater.
In addition to samples taken by the WHATS, samples of
gas bubbles were collected by non-gas-tight, cylinder

### Table 2. Contents and stable carbon isotopic compositions of CH₄ in the bottom fluid samples, together with their sampling locations, depths, and heights from the seafloor

<table>
<thead>
<tr>
<th>Sample ID</th>
<th>Latitude (N)</th>
<th>Longitude (E)</th>
<th>Depth* (m)</th>
<th>Height** (m)</th>
<th>CH₄ (μmol/kg)</th>
<th>δ¹³C(CH₄) (‰VPDB)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Samples taken by WHATS</td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>D1355W2</td>
<td>24°07.798'</td>
<td>124°11.547'</td>
<td>638</td>
<td>&lt;0.3</td>
<td>52</td>
<td>-33.3</td>
</tr>
<tr>
<td>D1355W4</td>
<td>24°07.798'</td>
<td>124°11.565'</td>
<td>641</td>
<td>&lt;0.3</td>
<td>19</td>
<td>-37.0</td>
</tr>
<tr>
<td>D1356W2</td>
<td>24°07.805'</td>
<td>124°11.540'</td>
<td>637</td>
<td>&lt;0.3</td>
<td>2.0 ×10⁴</td>
<td>-33.0</td>
</tr>
<tr>
<td>D1356W3</td>
<td>24°07.831'</td>
<td>124°11.553'</td>
<td>638</td>
<td>&lt;0.3</td>
<td>1.5 ×10⁴</td>
<td>-37.5</td>
</tr>
<tr>
<td>D1357W1</td>
<td>24°07.800'</td>
<td>124°11.536'</td>
<td>638</td>
<td>&lt;0.3</td>
<td>662</td>
<td>-35.0</td>
</tr>
<tr>
<td>D1357W3</td>
<td>24°07.800'</td>
<td>124°11.536'</td>
<td>638</td>
<td>&lt;0.3</td>
<td>1.4 ×10⁴</td>
<td>-37.8</td>
</tr>
<tr>
<td>D1357W4</td>
<td>24°07.810'</td>
<td>124°11.571'</td>
<td>641</td>
<td>&lt;0.3</td>
<td>2.3 ×10⁴</td>
<td>-38.3</td>
</tr>
<tr>
<td>D1360W2</td>
<td>24°07.811'</td>
<td>124°11.134'</td>
<td>643</td>
<td>&lt;0.3</td>
<td>1.1</td>
<td>-39.8</td>
</tr>
<tr>
<td>D1360W3</td>
<td>24°07.805'</td>
<td>124°11.535'</td>
<td>636</td>
<td>&lt;0.3</td>
<td>601</td>
<td>-33.7</td>
</tr>
<tr>
<td>D1361W4</td>
<td>24°07.809'</td>
<td>124°11.532'</td>
<td>635</td>
<td>&lt;0.3</td>
<td>55</td>
<td>-39.8</td>
</tr>
<tr>
<td>D1364W2</td>
<td>24°07.825'</td>
<td>124°11.561'</td>
<td>642</td>
<td>&lt;0.3</td>
<td>13</td>
<td>-31.3</td>
</tr>
<tr>
<td>D1364W4</td>
<td>24°07.794'</td>
<td>124°11.129'</td>
<td>642</td>
<td>&lt;0.3</td>
<td>1.5</td>
<td>-40.4</td>
</tr>
<tr>
<td>D1366W2</td>
<td>24°07.814'</td>
<td>124°11.346'</td>
<td>635</td>
<td>&lt;0.3</td>
<td>3.3 ×10⁴</td>
<td>-32.2</td>
</tr>
<tr>
<td>D1368W4</td>
<td>24°07.815'</td>
<td>124°11.352'</td>
<td>640</td>
<td>&lt;0.3</td>
<td>165</td>
<td>-40.3</td>
</tr>
<tr>
<td>Samples taken by Niskin</td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
<td></td>
</tr>
<tr>
<td>D1355N2</td>
<td>24°07.798'</td>
<td>124°11.547'</td>
<td>638</td>
<td>1.5</td>
<td>12</td>
<td>-38.2</td>
</tr>
<tr>
<td>D1356N1</td>
<td>24°07.797'</td>
<td>124°11.553'</td>
<td>638</td>
<td>1.5</td>
<td>1.3</td>
<td>-37.4</td>
</tr>
<tr>
<td>D1360N1</td>
<td>24°07.772'</td>
<td>124°11.527'</td>
<td>637</td>
<td>1.5</td>
<td>0.9</td>
<td>-38.7</td>
</tr>
<tr>
<td>D1360N2</td>
<td>24°07.814'</td>
<td>124°11.535'</td>
<td>636</td>
<td>1.5</td>
<td>0.9</td>
<td>-38.7</td>
</tr>
<tr>
<td>D1361N2</td>
<td>24°07.805'</td>
<td>124°11.535'</td>
<td>636</td>
<td>1.5</td>
<td>63</td>
<td>-38.1</td>
</tr>
</tbody>
</table>

*The seafloor depth.
**Approximate sampling height from the seafloor.

¹⁸O-enriched fluids likely derived from the partial disso-
ciation of the methane hydrate produced both the bub-
bles and carbonates on the seafloor (Takeuchi et al., 2007). Because methane hydrate is theoretically stable in the top
70 m layer of the knoll, Takeuchi et al. (2007) assumed
some subtle changes in the marine environment as possi-
ble cause of the dissociation of methane hydrate. Other
than the δ¹⁸O values of the carbonates, however, there is
little evidence for the existence of methane hydrate be-
neath the seafloor of the knoll.

**Sampling**

Gas bubbles and bottom fluid (seawater taken just
above the seafloor of the gas seepage site) were collected
by the human occupied vehicle Shinkai 2000 of the Japan
Agency for Marine-Earth Science and Technology (JAMSTEC) during the NT02-08 cruise in June 2002, using the WHATS (Tsunogas et al., 2003; Saegusa et al., 2006). The sampler comprised four 150 cm³ stainless steel
sample cylinders, eight ball valves, a motor-driven arm,
a rail, a peristaltic pump, and a flexible Teflon tube con-
ected to the inlet. The motor-driven arm on the rail is
used to open one of the four cylinders. By pumping out
seawater that had filled the cylinder, we introduced sam-
ple fluid/gas from the inlet into the cylinder. After com-
pletely replacing the seawater with sample, the valves at
both ends of the cylinder are tightly closed by moving
the motor-driven arm again. At this stage, the next cylin-
der is ready for sampling. The entire operation was con-
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to that of the ambient seawater (about 8.0°C), so that we
could not clarify the difference from the ambient seawater.
In addition to samples taken by the WHATS, samples of
gas bubbles were collected by non-gas-tight, cylinder
samplers made of acrylic resin (noted as M-type samplers; Sakai et al., 1990; Gamo et al., 1991) and samples of bottom fluid were collected by Niskin samplers. The Niskin samplers were attached to the Shinkai 2000 and operated at the heights of 1.5 m from the seafloor. All the samples of gas bubbles and bottom fluid are listed in Tables 1 and 2, respectively.

The collected samples were immediately subsampled into several kinds of bottles in the onboard laboratory of the mother ship Natsushima. For the determinations of the major components in the gas bubbles (N2, O2, H2S, CH4, CO2, CO, and non-methane hydrocarbons), each sample in a stainless steel cylinder was expanded into a container (50 cm3) made of stainless steel with a ball valve at each end (Xu et al., 1995) and stored under inner pressures higher than atmospheric pressure. In addition, for the determination of noble gases in the bubbles, an aliquot of gas sample was expanded into a pre-evacuated line and subsampled in a lead glass bottle (50 cm3) under a pressure slightly higher than atmospheric pressure.

For measurements of dissolved gas components, especially CH4, in the samples of bottom fluid, each fluid sample taken in a stainless steel cylinder was first transferred into a pre-evacuated line (ca. 1,500 cm3 volume) and mixed with both mercury chloride (to remove dissolved H2S as HgS) and phosphoric acid (to reduce the pH of the fluid to less than 1). After waiting for gas exchange equilibrium between the gas and the liquid phase within the line, the pressure was measured for the gas phase and part of it was subsampled into a container (50 cm3) made of stainless steel with a bellows valve at each end.

Each sample of gas bubbles collected by the M-type bottom sampler was subsampled into a glass bottle (200 cm3) with a stopcock at each end through the downward displacement of pure water. Each sample of ambient bottom seawater collected by the Niskin sampler was slowly transferred into 65 cm3 glass vials, for the determination of the concentration and δ13C of CH4 (Tsunogai et al., 2000). After approximately 3-fold volume overflow to prevent air contamination, 0.5 cm3 of saturated HgCl2 solution (6 wt%) was slowly added to each vial as a preservative. The vial was then sealed with a butyl rubber stopper and stored in the dark at 4°C until analysis.

Samples of effluent plume water supplied from the gas seeping field were taken in around the knoll on 26 Oct. 2005 at stations 01-0, 01-1, and 01-2 (Fig. 2) by the R/V Tansei-maru (JAMSTEC). While recording serial hydrographic data using a CTD system (Falmouth Scientific, Inc.), the samples of effluent plume were collected with a 24-port carousel multisampler with 5 L Niskin bottles (General Oceanic Inc.). To analyze the concentration and δ13C of CH4, a water sample was slowly transferred into a 120 cm3 glass vial. After approximately 3-fold volume overflow to prevent air contamination, 1 cm3 of saturated HgCl2 solution (6 wt%) was slowly added as a preservative. To minimize air contamination with the addition of HgCl2 solution, the solution was degassed immediately prior to its addition. The vial was then sealed with a butyl rubber stopper and stored in the dark at 4°C until analysis. The actual sampling depth of each Niskin bottle was verified by analyzing the salinity and dissolved oxygen concentration for each bottle using standard methods (Carpenter, 1965).

**ANALYTICAL METHODS**

The concentrations and δ13C values of CO2, CH4, CO and non-methane hydrocarbons in the gas samples (including the gas phase samples extracted from the bottom fluid samples) were determined by continuous-flow isotope ratio mass spectrometry (CF-IRMS) at Hokkaido University (Tsunogai et al., 1999, 2000, 2002b; Ijiri et al., 2003; Komatsu et al., 2005). The concentrations of N2, O2, H2, and H2S in the gas samples were determined...
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through conventional methods (Sakai et al., 1976). The helium and neon contents and ⁴He/²⁰Ne ratio in the gas samples were measured with a modified-VG5400 (MS-III) noble gas mass spectrometer in the University of Tokyo with an accuracy better than 4%, following the procedure described by Aka et al. (2001), except for using the interlaboratory helium standard (HESJ, Matsuda et al., 2002) for the standard material of the ³He/²⁰Ne ratio.

RESULTS

Gas bubbles

Analytical results for the bubble samples are listed in Table 1 and presented in Figs. 3, 4, and 5. The samples of gas bubbles taken by the WHATS (D1365W1, D1365W2, W1365W3, and D1367W2 in Table 1) were composed mostly of CH₄ (C₁) having uniform δ¹³C values around −40‰ VPDB while the concentrations of N₂, O₂, and H₂S were small, less than 1 vol%. Only trace amounts of CO₂ and CO were detected; less than 100 ppmv and less than 3 ppmv respectively. Non-methane hydrocarbons were also present in trace amounts, having concentrations of 300 ± 20 ppmv for ethane (C₂), 21 ± 14 ppmv for propane (C₃), 1.6 ± 0.9 ppmv for isobutane (iso-C₄), 1.0 ± 0.5 ppmv for n-butane (n-C₄), 0.4 ± 0.04 ppmv for isopentane (iso-C₅), and 0.02 ± 0.002 ppmv for n-pentane (n-C₅). The C₁/(C₂+C₃) ratio is greater than 3400 (Fig. 3) and the C₂/C₃ ratio is greater than 9. The δ¹³C values of the non-methane hydrocarbons were characterized by ¹³C-enrichment relative to C₁, which gave −28.3 ± 0.4‰ VPDB for C₂, −28.5 ± 0.2‰ VPDB for n-C₄, −19.9 ± 0.0‰ VPDB for iso-C₄, −16.9 ± 1.5‰ VPDB for iso-C₅ (Fig. 5). Noble gases in the samples were characterized by both helium enrichment (9.6 to 13.1 ppmv) and high ⁴He/²⁰Ne ratios (from 15 to 545) in comparison with values for air, indicating that no air contamination or fractionation affected bubble samples taken by the WHATS. Thus the concentration and isotopic composition of gases represented those of the bubbles on the seafloor. On the other hand, the gas samples taken by the M-type sampler (D1357M2 and D1364M1 in Table 1) were rather different from those taken by the WHATS, having concentrations for C₁, N₂, and O₂ of 20 vol%, 55

Fig. 4. The helium isotopic compositions of the bubble samples plotted as a function of their ²³⁹⁰Ne ratios (●), together with those of the mantle (Lupton, 1983), crust (Lupton, 1985; Mamyrin and Tolstikhin, 1984), pore gases in deep sediments taken around the Japan islands during the Deep Sea Drilling Project Leg 87 (Sano and Wakita, 1987), and seafloor seeping gases at the Taketomi spring (Oomori et al., 1993). Hypothetical mixing lines between the seawater and each end-member (mantle, crust, sediments, and Kuroshima bubbles) are also plotted as well (see the main text for details).

Fig. 5. The carbon isotopic compositions of hydrocarbons in the bubble samples taken at Kuroshima Knoll (●: D1365W1, □: D1367W2) plotted as functions of the reciprocal of carbon number, together with compositions for the Sirius-1 Exmouth Plateau, offshore Australia, where anomalous carbon isotopic compositions have been reported (James and Burns, 1984).
vol% and 25 vol%, respectively. We concluded that air had significantly contaminated the samples, presumably during the transportation from the seafloor to the surface. Thus, we do not consider them in the discussions that follow. To determine the concentration and isotopic composition of a gas sample taken from a deep seafloor, the sample should be thus kept gas-tight during transportation.

**Bottom fluid**

Analytical results for the bottom fluid samples, which might represent a mixture of bottom seawater and seeping material in the knoll, are presented in Table 2. They show significant enrichment in CH₄, especially for samples taken just above the seafloor (with sampling heights less than 0.3 m from the seafloor) by the WHATS. The observed concentration of CH₄ almost corresponds to the concentration under a saturation condition of about 10⁵ µmol/kg for the pressure and the temperature at the seafloor depth (65 bar and 8°C, respectively). This was consistent with the seafloor observations of the gas bubbles, which consisted mostly of CH₄.

Except for CH₄, however, the major dissolved components (Na⁺, K⁺, Mg²⁺, Cl⁻, and SO₄²⁻) in the bottom fluid samples showed neither enrichment nor depletion in comparison with seawater. If the site would be an active seeping site of some liquid material from deep sub-seafloor, it is difficult to assume such liquid having little anomalies in the major chemical components compared to seawater. The seeping material seems to be limited to gas bubbles and liquid a few amount in the knoll. That is, simple upward migration of the gases from sub-seafloor gas reservoir and mixing into bottom seawater and/or shallow pore water might be more probable for the mechanism to explain the CH₄-enrichment in the bottom fluid samples.

**Effluent plume**

We found clear enrichment of CH₄ in the water column at depths between 536 m and 644 m, the potential densities (expressed as specific gravity anomalies) being almost uniform at 28–30 kg/m³ (Fig. 6). On the other hand, we did not find enrichment of CH₄ at depths less than 500 m. The maximum enrichment was 82.0 nmol/kg at a depth of 615 m for station 01-0 (bottom depth of 640 m), located just above the gas seeping site (Fig. 2), while the background level was around 1 nmol/kg at that depth (Fig. 6). In addition, we found smaller but clear anomalies at depths of 600–630 m for stations 01-1 and 01-2, both of which were about 3 km horizontally from the gas seeping site.

The depths of the seafloor at stations 01-1 and 01-2 were ca. 1050 and 870 m, respectively and thus it is unlikely that the vertical transport of CH₄ (or CH₄-enriched fluid) emitted from adjacent seafloor was the cause of CH₄ enrichment at stations 01-1 and 01-2. The lateral transport of the CH₄ plume from station 01-0 (gas seeping site) is more plausible.

The vertical transport of CH₄ in station 01-0 from the adjacent gas seeping site was limited to heights of less
than 140 m from the seafloor, probably owing to complete dissolution of the gas bubbles to that height. After dissolution, the plume spread horizontally along with the surface of equal density of the plume, while the concentrations reduced through either dilution or oxidation in the oxygenated, low-CH$_4$ water column.

**DISCUSSION**

*Initial generation of hydrocarbons in the bubbles*

Both chemical and isotopic compositions of gaseous alkanes (C$_1$-C$_4$) are useful for inferring the origin of hydrocarbons. Specifically, we can differentiate thermogenic hydrocarbons (generated by thermal degradation of organic matter) from biogenic hydrocarbons (generated by microbial degradation of organic matter) by plotting the stable carbon isotopic compositions of CH$_4$ ($\delta^{13}C_1$) vs. the light gas composition (C$_2$/(C$_2$+C$_3$)) on a widely used conventional diagram (Bernard et al., 1977). The application of this diagram in the identification of surface gas seepage, however, provides misleading information for gases that underwent later alteration, such as oxidation of biogenic gas or molecular separation during migration (Tsunogai et al., 2002a).

In Fig. 3, the C$_2$/(C$_2$+C$_3$) ratio of the Kuroshima bubble samples is plotted as a function of $\delta^{13}C$ for CH$_4$ ($\delta^{13}C_1$). Approximate source regions of both thermogenic and biogenic hydrocarbons are also plotted together with a hypothetical mixing line between them having respective representative compositions: $\delta^{13}C_1 = -45\%_{\text{VPDB}}$ and C$_2$/(C$_2$+C$_3$) = 10 for thermogenic hydrocarbons and $\delta^{13}C_1 = -80\%_{\text{VPDB}}$ or $-60\%_{\text{VPDB}}$ and C$_2$/(C$_2$+C$_3$) = 10$^4$ for biogenic hydrocarbons (Bernard et al., 1977). While the high C$_2$/(C$_2$+C$_3$) ratios suggest microbial generation, the $\delta^{13}C$ values of CH$_4$ (-40.1 ± 0.5$\%_{\text{VPDB}}$) suggest thermogenesis. It is difficult to explain this contradiction through mixing between microbial and thermogenic hydrocarbons, as schematically shown by the theoretical mixing line. Rather, there was significant secondary alteration subsequent to generation.

One possible secondary process is the $\delta^{13}C$ isotopic fractionation due to microbial oxidation of CH$_4$ under an aerobic/anaerobic condition. Tsunogai et al. (2002a) reported up to 50$%_{\text{VPDB}}$ $\delta^{13}C$ enrichment in residual CH$_4$ through anaerobic microbial oxidation within surface marine sediments. The carbonate crusts including chimney-like structures found in the knoll also suggest anaerobic oxidation of CH$_4$ and subsequent carbonate precipitation following the reactions.

\[
\begin{align*}
\text{CH}_4 + \text{SO}_4^{2-} & \rightarrow \text{HCO}_3^- + \text{HS}^- + \text{H}_2\text{O} \\
2\text{HCO}_3^- & \rightarrow \text{CO}_2 + \text{CO}_3^{2-} + \text{H}_2\text{O}
\end{align*}
\]

Ca + CO$_3^{2-}$ $\rightarrow$ CaCO$_3$. (6)

An aerobic oxidation of CH$_4$, however, results in the reduction of the C$_2$/(C$_2$+C$_3$) ratio in most cases (Whiticar et al., 1986), and thus it is difficult to explain the observed values. Furthermore, the $\delta^{13}C$-enrichment up to 20$\%_{\text{VPDB}}$ requires more than 90$\%_{\text{VPDB}}$ oxidation of CH$_4$. In case of anaerobic oxidation of CH$_4$, while half of oxidized carbon precipitate as CaCO$_3$, the rest of oxidized carbon should remain in bubbles as CO$_2$ through the reaction (5), under steady state condition. Because the bubbles contain CO$_2$ relative to CH$_4$ (Table 1), such enormous oxidation in the bubbles is unlikely.

An alternative possible process is the molecular fractionation of hydrocarbons generated originally through the thermogenic process. That is, the removal of heavier hydrocarbons from thermogenic hydrocarbons could explain the observed contradiction. Post-genesis CH$_4$ enrichment through molecular fractionation is ubiquitous for natural hydrocarbons (Nagy, 1960; Price and Schoell, 1995; Mango, 2001; Snowdon, 2001). There is almost one order of magnitude less C$_3$ than C$_2$ (C$_2$/C$_3 > 9$) in the bubbles which suggests some molecular fractionation is responsible for the C$_1$-enrichment in the hydrocarbons. Furthermore, the $\delta^{13}C$ value of C$_3$ (-19.1$\%_{\text{VPDB}}$), iso-C$_3$ (-22.4$\%_{\text{VPDB}}$), and n-C$_4$ (-19.9$\%_{\text{VPDB}}$) in comparison with C$_2$ (C$_2$/C$_3 > 9$) and C$_4$ (n-C$_4$ + iso-C$_4$ (-28.0$\%_{\text{VPDB}}$) suggests significant secondary alteration for the hydrocarbons. Regarding thermogenic hydrocarbons, the $\delta^{13}C$ value for each hydrocarbon reflects the integrated kinetic isotope effect during the cleavage of chemical bonds from parental material. This begins with long-chain organic matter and thus smaller carbon number molecules tend to exhibit lower $\delta^{13}C$, with there being a reverse correlation with the reciprocal of the carbon number (Chung et al., 1988; Mango and Elrod, 1999; Du et al., 2003) as shown by the dotted arrow in Fig. 5. The $\delta^{13}C$ values of C$_1$ and n-C$_4$ in the bubbles are larger than the values expected for C$_3$ and n-C$_4$ produced from sedimentary organic matter (~20 to ~30$\%_{\text{VPDB}}$) and also the value for coexisting C$_3$ (~28.0$\%_{\text{VPDB}}$) through thermogenic decomposition (Chung et al., 1988; Mango and Elrod, 1999; Du et al., 2003) or even microbial processes (Oremland et al., 1988; Taylor et al., 2000; Nakagawa et al., 2003). Hence, removal processes for the heavier hydrocarbons accompanying substantial isotopic fractionation must be highly probable for the C$_1$-enrichment in the hydrocarbons in the bubbles.

One possible molecular and isotopic fractionation process is fractional crystallization during hydrate formation. Gas hydrate is an ice-like crystalline mineral in which hydrocarbons are held within rigid cages of water molecules. While the structure I is usually pure CH$_4$...
(Kennicutt, 1995), structure II and structure H include C3–C4 hydrocarbons as significant components (Sloan, 1998). The molecular distribution of the bubbles could be consistent with those subducting the hydrocarbons forming the structure II/H hydrate from the original thermogenic hydrocarbons enriched in C2–C4 hydrocarbons, as a consequence of active gas hydrate crystallization during migration in the sub-seafloor. Such hydrate crystallization is significant in the knoll and accompanied by significant carbon isotopic fractionation, the residual bubbles would be CH4-enriched relative to the C2–C4 hydrocarbons and the C2–C4 hydrocarbons would be 13C-enriched relative to the initial hydrocarbons. In previous field observations, however, little carbon isotopic fractionation was found during crystallization of gas hydrates, including C2–C4 hydrocarbons (Brooks et al., 1986; Kennicutt et al., 1988). It is difficult to explain the observed 13C-enrichment for C3 and n-C4 in the bubbles as a consequence of the fractional crystallization during hydrate formation, and thus we should assume alternative processes.

An alternative process for explaining both molecular and isotopic fractionations is the fractional decomposition of hydrocarbons through some sub-seafloor microbial activity. James and Burns (1984) found that C2 and C3 hydrocarbons were significantly 13C-enriched compared with other hydrocarbons in deep subsurface hydrocarbon reservoirs, which they attributed to selective microbial consumption of C1 and C2 alkanes. In Fig. 5, δ13C values for the hydrocarbons in the bubbles are plotted as a function of the reciprocal of each carbon number, together with values for the Sirus-I Exmouth Plateau, offshore Australia, for which significant secondary alteration has been suggested (James and Burns, 1984). The alteration would probably be due to anaerobic microorganisms (Perry, 1980; Stephens and Dalton, 1986; Ashraf et al., 1994; Head et al., 2003). The two sets of results coincide strikingly well. Regarding the microbial processes for these deep subsurface hydrocarbons, preferential degradation of C3 followed by that of C4 and then that of C2 has also been noted (Boreham et al., 2001) with particularly high levels of biodegradation in lower-temperature (<60°C) reservoirs (Wenger et al., 2002). In addition, Strapoc et al. (2007) found kinetic isotope fractionation in propane presumably produced through anaerobic microbial consumption. The kinetic isotope fractionation during aerobic microbial consumption of C2–C4 hydrocarbons has been quantified in laboratory incubation experiments as well (Kinnaman et al., 2007; Bouchard et al., 2008). Thus, microbial consumption of C3 (significantly), C4 (moderately), and C2 (partly) in thermogenic hydrocarbons could explain the observed molecular and isotopic fractionations of hydrocarbons.

To verify this hypothesis quantitatively, we estimated the initial C4/(C2+C3) (the C4/(C2+C3) prior to the microbial alteration) by correcting the microbial alteration in the concentrations and δ13C values of C2–C4 hydrocarbons, assuming (1) the microbial alteration was limited to C2–C4 hydrocarbons, so that the observed concentrations and the δ13C values in the bubbles corresponds to those of the initial for methane (103 ± 3 vol% and −40.1‰) and C4 (0.4 ppmv and −28.0‰), (2) the initial δ13C values of C3, C4, and C2 were −33.0‰, −30.5‰, and −29.0‰, respectively, showing a reverse correlation with the reciprocal of carbon number in the initial C1–C4 hydrocarbons, (3) the KIE (−1000(1/α − 1)) in carbon isotopes of C2 during the microbial alteration was 3 ± 1, which corresponds to the KIE during the atmospheric oxidation of C2 (Tsunogai et al., 1999), (4) the KIEs in carbon isotopes of C1 and C3 during the microbial alteration were 2 ± 0.7 and 1.5 ± 0.5, respectively, which corresponds to 2/3 and 2/4 of the KIE of C2, due to the dilution effect of KIEs by unreacted carbon in a molecule (Tsunogai et al., 1999). The assumed KIE for C3 corresponds well to that determined by the field observation (Strapoc et al., 2007).

Under these assumptions, we estimated the initial C4/(C2+C3) from the concentrations and δ13C values of C1–C4 hydrocarbons in the bubbles, by correcting the microbial alteration applying the Rayleigh equation (7) (Coleman et al., 1982) to each hydrocarbon.

$$\delta^{13}C_{C2} - \delta^{13}C_{C4} = 1000(1/\alpha - 1)\ln(C_f/C_0),$$

(7)

where C0 and Cf were the initial and final concentrations, respectively, and δ13C0 and δ13 Cf were the initial and final δ13C values, respectively. The estimated initial concentrations of C1–C4 hydrocarbons corresponds to the initial C1/(C2+C3) of 235 (median), 19 (minimum) and 612 (maximum). All the initial C1/(C2+C3) ratios correspond to those of the initial for methane (103 ± 3 vol% and −40.1‰) and C4 (0.4 ppmv and −28.0‰), so that the observed concentrations and the δ13C values in the bubbles corresponds to those of the initial for methane (103 ± 3 vol% and −40.1‰) and C4 (0.4 ppmv and −28.0‰), so that the observed concentrations and the δ13C values in the bubbles corresponds to those of the initial for methane (103 ± 3 vol% and −40.1‰) and C4 (0.4 ppmv and −28.0‰), so that the observed concentrations and the δ13C values in the bubbles corresponds to those of the initial for methane (103 ± 3 vol% and −40.1‰) and C4 (0.4 ppmv and −28.0‰).
ratios (solid circles) together with those for the mantle (8R*) (where R* denotes the 4He/He ratio of air of 1.4 × 10^-5) (Lupton, 1983), continental crust (about 0.02R*) (Lupton, 1983; Mamyrin and Tolstikhin, 1984) and sediments (the pore gases in the sedimentary layer) around Japanese islands (Deep Sea Drilling Project sites 583D, 583F, and 584; Sano and Wakita, 1987) (open triangles). To exclude samples contaminated by air to a significant extent, the data for sediments are limited to sediments having 4He/20Ne ratios exceeding 0.5. In addition, bold lines for the theoretical mixing with seawater (air saturated seawater; 1Ra) are presented for the mantle and crust. Similarly, the mixing with seawater is presented for both the sedimentary pore gases and Kuroshima bubbles by dotted lines. By extrapolating the dotted mixing lines to the region with higher 4He/20Ne ratios, we can obtain the endmember 4He/He ratios for the bubbles and sediments as 0.44Ra and 0.15Ra respectively, through removing seawater contaminants from the samples, assuming all 20Ne in the samples was derived from the seawater.

While the endmember 4He/He ratio for the bubbles is lower than the ratios for the mantle or seawater, it is much higher than the ratio for the crust or sediments. It is difficult to explain the composition of Kuroshima bubbles merely by the in situ contribution of 4He-enriched radiogenic helium in the crust/sediments. Although the mixing of the seawater components with the crust or sediments could produce 4He/He ratios similar to those for the bubbles, it is impossible to explain the high 4He/20Ne ratio of the bubbles through the mixing of seawater with sediments or crust as schematically shown in Fig. 4. Thus, there must be a contribution of mantle helium to the hydrocarbons in the bubbles, and thus thermogenic generation due to the contribution of geothermal heat is highly probable for the initial generation of hydrocarbons in the bubbles.

Comparison with the nearby natural gas seepages

The gas bubbles exhibit chemical and isotopic compositions that are similar to those of shallow-seafloor gas bubbles at the Taketomi spring. The Taketomi submarine spring (10–30 m depth) is 20 km northwest of the knoll and on the same forearc seafloor of the southern Ryukyu arc (Fig. 1). Gas bubbles composed mostly of CH4 (70–80%) are discharged together with hot water having a temperature of up to 40–60°C (Kaneshima et al., 1983; Oomori, 1987; Oomori et al., 1993; Hirayama et al., 2007). The δ13C values of CH4 (30 to 107 times those in the bottom fluid samples) show a substantial enrichment in 13C; +4% enrichment on average and +8% enrichment at maximum (Table 2). In addition, the lowermost δ13C value (−40.4‰) of CH4 in the bottom fluid samples corresponds to values for the bubbles. It is unlikely there is a contribution of CH4 from ambient water as for the cause of the δ13C variation because the concentrations of CH4 in the bottom fluid (Table 2) were 107 times those in ambient bottom water (Fig. 6). We conclude that kinetic microbial oxidation of CH4 must be responsible for the deviation: that is, CH4 in the bottom fluid samples is the residue of CH4 after the oxidation.

If the oxidation occurred only in the aerobic water column after venting into the water column, there should be 13C-enrichment in accordance with the depletion in the concentration. However, we could not find a clear correlation between the concentrations and δ13C values.
of CH$_4$ in the bottom fluid samples (Table 2). Furthermore, while the samples D1356W2, D1356W3, and D1357W3 had concentrations 10$^7$ times that of ambient seawater (almost to saturation), the $\delta^{13}$C values indicated a clear $^{13}$C-enrichment in comparison with the bubbles. It is unlikely such rapid oxidation in the water column had progressed in the fluid before dilution with the surrounding ambient water. We conclude that the majority of the oxidation progressed under the seafloor. This is in agreement with the observation by Inagaki et al. (2004), who observed $^{13}$C-enrichment of CH$_4$ up to 20% within pore water of the seeping site. The CH$_4$ in the liquid samples contained the residual CH$_4$ oxidized anaerobically during diffusive migration in the sedimentary layer through the reaction (4) (Kulm et al., 1986; Tsunogai et al., 2002a).

Because the reaction requires the mixing of seawater SO$_4^{2-}$ with CH$_4$, it might be limited to a thin surface layer, probably less than 1 meter thick. By reducing the alkalinity of the mixture, carbonates have precipitated in the knoll as observed at another cold seepage site (Kulm and Suess, 1990). On the other hand, the observed homogenous $^{13}$C values of CH$_4$ in the bubbles suggest that they had not undergone anaerobic oxidation, probably owing to the rapid, advective migration of the bubbles in the sedimentary layer. The trace concentration of CO$_2$ in the bubbles also supports that there was minimum oxidation of the bubbles prior to seepage.

Contribution of gas hydrate dissociation

As already discussed, the contribution of gases derived from the dissociation of sub-seafloor methane hydrate is not required to explain the chemical and isotopic compositions of hydrocarbons in the knoll as observed at another cold seepage site (Kulm and Suess, 1990). In addition, the chemical and isotopic compositions of noble gases suggest that the dissociation of methane hydrate under the seafloor is not a realistic source of the bubbles. Because the $^4$He/$^3$He ratios and $^4$He/$^{20}$Ne ratios of the bubbles highly deviate from the ratios for seawater, it is unlikely that there is a helium source either in the seawater column or in the shallow sedimentary layer where the dissociation of hydrate has been assumed. If we assume the dissociation of methane hydrate is the only source of the bubbles in the knoll, the helium in the bubbles must have been trapped in the methane hydrate as well, together with CH$_4$ in the bubbles. However, it is difficult to explain the helium enrichment in the bubbles derived from hydrate because helium is rarely incorporated into the structure of natural methane hydrate (Dickens and Kennedy, 2000; Winckler et al., 2002). In the case of the pure natural gas hydrates taken from Hydrate Ridge, Cascadia Margin, for instance, the $^4$He/CH$_4$ ratios were from 0.0003 to 0.009 ppmv (Winckler et al., 2002), while the ratios for the Kuroshima bubbles were more than 10 ppmv. The dissociation of methane hydrate cannot be the only source of the bubbles in the knoll. Thus, if we persist with the contribution of methane hydrate, we must assume additional He-enriched gases for the source of the bubbles, which seems unrealistic for the area. Because the $^4$He/CH$_4$ ratios in the Kuroshima bubbles can be classified into the representative $^4$He/CH$_4$ ratios of thermogenic natural gases in and around Japan islands (Wakita and Sano, 1983; Urabe et al., 1985; Poreda et al., 1988), direct leakage of hydrocarbons from sub-seafloor gas reservoirs to seafloor is the most probable process to explain the origin of the bubbles in Kuroshima Knoll.

Submarine massive hydrate is ubiquitous in the deep ocean floor around the world (Brooks et al., 1986; Suess et al., 1999). Because hydrate decomposes when the temperature rises, decomposition of sub-seafloor methane hydrate has been assumed as the source of seafloor discharges of CH$_4$ bubbles at many sites of the world ocean, as presented for the case of the Kuroshima Knoll in this study. Both the present progression of global climate change and the findings of the massive deposits of methane hydrate on the seafloor at active CH$_4$ seeping sites seem to support the assumption. Most of the submarine massive hydrates, however, seems to be stable at present (Sassen et al., 2001). In addition, methane hydrate can be produced as a result of natural CH$_4$ seeping activity on the seafloor. Regarding the origin of the seafloor seeping bubbles, we also need to take into account other pos-
Distribution and fate of methane in the water column

We observed a clear enrichment of CH₄ in the water column at depths between 536 and 694 m (Fig. 6). On the other hand, we did not find any CH₄ enrichment shallower than 500 m at the stations. All the rising bubbles in Kuroshima Knoll have dissolved within 140 m of the seafloor. This is consistent with observations of other bubbles seeping from seafloor (Suess et al., 1999; Valentine et al., 2001). Because the depth is substantially deeper than the maximum mixed layer depth (ca. 100 m) in the area, direct input into atmosphere is negligible for the gas seeping from Kuroshima Knoll.

To clarify whether the CH₄ decrease in the plume can be explained by simple dilution by the background seawater, the measured δ¹³C values of CH₄ in the plume are plotted against the reciprocal of the CH₄ concentration for the samples taken within the plumes (column water samples taken from 500 to 700 m depths) (Fig. 7). The least-squares fitted line for the plume samples is also shown in the figure. The linear correlation suggests that simple mixing between the seeping CH₄ and CH₄ in background seawater is the main process determining both the concentration and δ¹³C of CH₄ in the plume, and the δ¹³C value of seeping CH₄ is uniform during the mixing (Tsunogai et al., 1998, 2000, 2005). That is, aerobic microbial oxidation in the oxic water column is negligible for CH₄ in the plume, and thus CH₄ is diluted rapidly by eddy diffusion rather than oxidation in the water column. In conclusion, all the rising bubbles at Kuroshima Knoll have dissolved within 140 m of the seafloor. After dissolution, the plume spreads horizontally along with the surface of equal density of the water column, while the concentration is reduced through dilution by eddy diffusion, rather than by oxidation.

Quantification of sub-seafloor anaerobic oxidation

The intercept of the least-squares fitted line for the plume samples in Fig. 7 corresponds to the end-member δ¹³C value of CH₄ seeping from the knoll (Tsunogai et al., 1998), which can be considered the weighted average δ¹³C value of CH₄ emitted from the knoll to the water column. Assuming the δ¹³C values of the bubbles (−40.1‰VPDB) represent the initial δ¹³C of CH₄ prior to sub-seafloor anaerobic oxidation, we can evaluate the extent of the sub-seafloor anaerobic microbial oxidation quantitatively using the discrepancy between the weighted average δ¹³C value (−38.6‰VPDB) and the δ¹³C values of the bubbles because the discrepancy (±2.5‰VPDB) corresponds to the average extent of the progress of the sub-seafloor anaerobic oxidation of CH₄ at Kuroshima Knoll.

Assuming the kinetic isotope effect due to the anaerobic microbial CH₄ oxidation (k₁₂/kᵢ₃) to be 1.008 (Tsunogai et al., 2002a), the fraction of residual CH₄ can be calculated from the δ¹³C variation using the Rayleigh equation (7). Using this equation, we estimated the average Cₑ/C₀ to be 0.82. Thus, about 20% of CH₄ has been oxidized before seeping into the water column.

Conclusions

Most portions of the hydrocarbons in the Kuroshima bubbles were supplied directly from deep gas reservoirs produced by thermogenic processes in the Miocene. The hydrocarbons, especially C₃- and C₄-, had been fractionated both chemically and isotopically through microbial destruction during the storage in the gas reservoir. In addition, the anaerobic oxidation of CH₄ within shallow sediments just beneath the seafloor removes about 20% of CH₄ on average, until seepage into ocean water column. After the seepage, all rising bubbles dissolved into the ocean water column within 140 m of the seafloor. After the dissolution, the plume spreads horizontally along with the surface of equal density in the water column, while the concentrations decrease through dilution by eddy diffusion, rather than by oxidation.

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References


