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Li-C-H system, which can store about 5.0 mass% of rechargeable H2, has been reported as a promising hydrogen storage system by Ichikawa et al. [Appl. Phys. Lett. 86, 241914 (2005); Mater. Trans. 46, 1757 (2005)]. This system was investigated from the thermodynamic and structural viewpoints. However, hydrogen absorption/desorption mechanism and the state of hydrogen atoms absorbed in the composite have not been clarified yet. In order to find new or better hydrogen storage system, graphite powder and nano-structural graphite ball-milled under H2 and Ar atmosphere were prepared and milled with Li and Mg under Ar atmosphere in this study. Microstructural analysis for those samples by transmission electron microscope revealed that LiC6 and/or LiC12 were formed in Li-C-H system. On the other hand, MgC2 was found in Mg-C-H system ball-milled under H2 atmosphere, but not in the system ball-milled under Ar atmosphere. These results indicated that nano-structure in composites of nano-structural graphite is different from that of alkali (-earth) metal. For these reasons, metal-C-H system can be recognized to be a new family of hydrogen storage materials. © 2013 AIP Publishing LLC.

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I. INTRODUCTION

Hydrogen storage is one of the most important technologies to develop fuel cell vehicles. A lot of research on hydrogen storage materials has been reported over the past few decades.1–5 At present, carbon based materials have been paid much attention as a hydrogen storage materials, because carbon is light element and abundant materials. A lot of new carbon materials such as carbon nano tubes (CNTs), activated carbon, and hydrogenated nano-structural graphite (CnanoHx) have been researched to create more efficient hydrogen storage materials.4–12 Among these new carbon materials, we have paid attention to hydrogenated nano-structural graphite. CnanoHx is synthesized from graphite by the ball-milling under H2 gas atmosphere.13–22 So far, it has been reported that the hydrogenated nano-structural graphite (CnanoHx) can store about 7.4 mass% of hydrogen.13 CnanoHx material requires heating up to more than 700 °C to release the hydrogen and desorbs hydrogen with hydrocarbons (CH4 and C2H6). Thus, CnanoHx itself is difficult to use as hydrogen storage materials. So, Li-C-H system has recently been reported as a way to solve these problems.23,24 Li-C-H system was synthesized from LiH and nano-structural graphite by ball-milling under hydrogen atmosphere, and it can reversibly absorb and desorb 5 mass% of hydrogen at 350 °C under 3.0 MPa H2 gas atmosphere. Moreover, not only Li-C-H system but also Mg-C-H system has been reported as the superior hydrogen storage material than CnanoHx itself.25–30 However, hydrogen absorption/desorption mechanism and the state of hydrogen atoms absorbed in the composite have not been clarified yet. It has been difficult to discuss those by using X-ray diffraction (XRD), because the crystallite sizes of the compounds in Li-C-H and Mg-C-H systems were too small or amorphous. In this paper, by using transmission electron microscopy (TEM), we have investigated the compounds in both systems to understand the reaction mechanism. Especially, we have determined the reaction formula of hydrogenation and dehydrogenation in both systems.

II. EXPERIMENTAL PROCEDURES

Samples were prepared by ball-milling method. High purity graphite powder (99.99% purity) was milled under H2 and Ar gas atmosphere (1.0 MPa pressure) for 8 h, which was labeled as HG and AG, respectively (a planetary ball-mill equipment: Fritch, P7). No-milling graphite was labeled just as G. In this way, three kinds of graphite (G, HG, AG) were prepared. Then, these three kinds of graphite were milled with Li (Li:C = 1:2, 99.9%) under Ar gas atmosphere (1.0 MPa) for 3 h to synthesize the sample of Li-C-H system. We labeled three kinds of samples as Li-HG, Li-AG, and Li-G, respectively. In the same way, graphite were milled with Mg (Mg:C = 1:4), which were labeled as Mg-HG, Mg-AG, and Mg-G, respectively. For the hydrogenation and the dehydrogenation of all samples, the heat-treatments were carried out at 450 °C for 8 h under 3.0 MPa H2 gas atmospheres and at 450 °C for 8 h under vacuum condition, respectively. We observed all samples by TEM (JEM-2010, 200 keV) and high voltage electron microscope (HVE M, ARM-1300, 1250 keV). The samples were handled without any exposure
to air in whole process. The sample powder was placed on a micro-grid, and the grid was set into a heating specimen holder in the glove box under Ar gas atmosphere. The sample holder was put into a plastic bag under Ar gas atmosphere and loaded into TEM equipment in order to prevent the oxidation. The high resolution TEM (HRTEM) images were analyzed by fast Fourier transformation (FFT) and inverse FFT (IFFT).

III. RESULTS AND DISCUSSION

A. Li-C-H system

Figures 1(a)–1(c) show the bright field (BF) images and diffraction patterns of as-synthesized samples of Li-G, Li-HG, and Li-AG, respectively. From the BF images, it seems that the particle sizes of Li-G, Li-HG, and Li-AG are from several tens nm to several hundred nm. Additionally, the crystallite size of Li2C2 in Li-HG was around 30 nm, as shown in the insetted dark field (DF) image of (b). From the diffraction patterns, the diffraction spots from LiC6 (003), LiC6 (001), LiC6 (002), C (002) and the diffraction rings from Li2O (111), Li2O (220), Li2O (311) were indexed in Li-G. In cases of Li-HG and Li-AG, the diffraction spots from LiC6 (001), and no spots, the diffraction rings from LiC6 (110), LiC6 (121), LiC6 (231), Li2O (220), and Li2C2 (020), Li2C2 (121), Li2C2 (211), were identified, respectively, as listed in Table I. Two kinds of compounds between Li and C are identified. One is a graphite intercalation compound (GIC) of LiC6, the other is Li2C2, which has been reported in Refs. 27 and 29. The crystallinity of C in Li-G and Li-AG is the highest and the lowest. Therefore, it could be considered that LiC6 was generated from Li and C with high crystallinity (graphite), meanwhile Li2C2 was from Li and C with low crystallinity. Then, the samples were hydrogenated at 450°C for 8 h under 3.0 MPa H2 gas atmospheres. Figure 2 shows the BF and DF images and diffraction patterns of hydrogenated samples: (a) Li-G, (b) Li-HG, and (c) Li-AG. The compounds identified from the diffraction patterns are listed in Table I. LiH was observed in the all hydrogenated samples; on the other hand, C was observed in the hydrogenated samples of Li-G and Li-HG. The DF images show the crystallites of LiH with the size of around 200 nm, 50 nm, and 30 nm in the hydrogenated samples of Li-G, Li-HG, and Li-AG. From those results, the hydrogenated reaction of Li-G, Li-HG, and Li-AG can be considered as follows:

\[
\begin{align*}
\text{Li-G:} & \quad \text{LiC}_6 + \text{H}_2 \rightarrow \text{LiH} + \text{C}, \\
\text{Li-HG:} & \quad \text{LiC}_6 + \text{H}_2 \rightarrow \text{LiH} + \text{C}, \\
& \quad \text{Li}_2\text{C}_2 + \text{H}_2 \rightarrow \text{LiH} + \text{CH}_x, \\
\text{Li-AG:} & \quad \text{Li}_2\text{C}_2 + \text{H}_2 \rightarrow \text{LiH} + \text{CH}_x.
\end{align*}
\]

Here, CHx is hydrocarbon such as CH4, C2H6, and C2H4. As reported in Ref. 29, the hydrogenation reactions of Li-G, Li-HG, and Li-AG occur around 200°C, 130°C and 200°C, and 130°C. Considering this result, it can be indicated that LiC6 and Li2C2 are hydrogenated at around 200°C and 130°C, respectively.

<table>
<thead>
<tr>
<th>Sample</th>
<th>Temperature for hydrogenation (°C)</th>
<th>Compounds in as-synthesized</th>
<th>Compounds in hydrogenated</th>
<th>Compounds in dehydrogenated</th>
</tr>
</thead>
<tbody>
<tr>
<td>Li-G</td>
<td>130</td>
<td>C, LiC₆, Li₂O</td>
<td>LiC₆, Li₂C₂, Li₂O</td>
<td>LiC₆, LiH</td>
</tr>
<tr>
<td>Li-HG</td>
<td>130, 200</td>
<td></td>
<td>C, LiH</td>
<td>LiC₆, Li₂C₂</td>
</tr>
<tr>
<td>Li-AG</td>
<td>200</td>
<td></td>
<td>C, LiH, LiC₄</td>
<td>LiC₆, Li₂C₂</td>
</tr>
<tr>
<td></td>
<td></td>
<td></td>
<td>LiH, Li₂O, LiOH</td>
<td></td>
</tr>
</tbody>
</table>

FIG. 1. BF and DF images and diffraction patterns of as-synthesized samples: (a) Li-G, (b) Li-HG, and (c) Li-AG.

**TABLE I.** Temperature for hydrogenation, compounds identified by electron diffraction patterns in as-synthesized samples, compounds identified by electron diffraction patterns in hydrogenated samples, and compounds identified by electron diffraction patterns in dehydrogenated samples of Li-G, Li-HG, and Li-AG.
The BF and DF images and diffraction patterns of dehydrogenated samples are shown in Figure 3(g) Li-G, (h) Li-HG, and (i) Li-AG. After dehydrogenation, the graphite intercalation compound of LiC_{12} appeared in Li-G and Li-HG. And Li_{2}C_{2} appeared in Li-HG and Li-AG, as listed in Table I. The DF images show the crystallites of LiC_{12} and Li_{2}C_{2} in the dehydrogenated samples. The crystallite size of LiC_{12} in Li-G was relatively larger than that in Li-AG, and the crystallite size of Li_{2}C_{2} was comparable to that in Li-AG. This result would indicate that carbon (graphite) with relatively large crystallite and carbon (amorphous like) with small crystallite changed to LiC_{12} and Li_{2}C_{2}, respectively. From those results, the hydrogenated reaction of Li-G, Li-HG, and Li-AG can be considered as follows:

\[
\begin{align*}
\text{Li-G:} & \quad \text{LiH} + C \text{(graphite)} \rightarrow \text{LiC}_{12} \text{(or LiC}_{6} \text{)} + \text{H}_{2}, \\
\text{Li-HG:} & \quad \text{LiH} + C \text{(graphite)} \rightarrow \text{LiC}_{12} \text{(or LiC}_{6} \text{)} + \text{H}_{2}, \\
\text{Li-IH:} & \quad \text{LiH} + C \text{(amorphous like)} \rightarrow \text{Li}_{2}C_{2} + \text{H}_{2}, \\
\text{Li-AG:} & \quad \text{LiH} + C \text{(amorphous like)} \rightarrow \text{Li}_{2}C_{2} + \text{H}_{2}.
\end{align*}
\]

Figure 4 shows (a) BF, (b) DF images, and (c) diffraction pattern of the dehydrogenated sample of Li-G. In the diffraction pattern, a couple of diffraction spots were assigned to LiC_{12} (002). In the DF image, layer structure due to the crystallite of LiC_{12} was observed. Here, the Li atoms were intercalated into graphite structure, however, we found LiH particle in the sample after hydrogenation (Fig. 2(d)). Li atoms should move to outside of graphite from intercalated sites, when LiH forms. To clarify the formation site of LiH, we obtained the HRTEM image of the hydrogenated sample of Li-G, as shown in Figure 5(b). FFT image was analyzed from (a) HRTEM image, and (c) IFFFT image was analyzed from (b). Finally, (a') image shows the location of graphite and LiH. From this result, it was indicated that Li atoms could move along in-plane direction of graphite and could form LiH at the edge of graphite. Now stage, although we are not sure the driving force of Li atoms moving, it was confirmed that hydrogenation of GIC proceeded with generating LiH at the edge site of graphite.

B. Mg-C-H system

Figures 6(a)–6(c) show the BF images and the diffraction patterns of as-synthesized samples of Mg-G, Mg-HG, and Mg-AG, respectively. From the BF images, it seems that the particle sizes of Mg-G, Mg-HG, and Mg-AG are from several tens nm to several hundred nm. From the diffraction patterns, the diffraction spots from Mg (101), MgC_{2} (210), and the diffraction rings from C (002), MgO (200) were indexed in Mg-G. In Mg-HG, and Mg-AG, the diffraction spots from Mg (101), Mg (102), MgC_{2} (210), MgC_{2} (311), MgC_{2} (400), and MgC_{2} (200), the diffraction rings from C (002), MgO (200), and no spots, the diffraction rings from C (002), MgO (200), were identified, respectively, as listed in Table II. The result of diffraction patterns of Mg-G was almost similar with the result of Mg-HG. MgC_{2} was generated from the solid-solid reaction between C and Mg during ball-milling. Parts of Mg and graphite remained without any reactions. MgO, which was composed of much small
FIG. 3. BF and DF images and diffraction patterns of dehydrogenated samples: (g) Li-G, (h) Li-HG, and (i) Li-AG.

FIG. 4. (a) BF images, (b) DF images, and (c) diffraction pattern of the dehydrogenated sample of Li-G.

FIG. 5. (a) HRTEM image of hydrogenated sample of Li-G, (b) FFT image from (a), (c) IFFT image from (b), (a'), and (a) with indication of crystallites of graphite and LiH.
crystallites, could be considered as an impurity in the starting material of Mg. In the previous report, Mg and C has never been found in XRD results. However, in this study, MgC₂ has been identified in the samples of Mg-G and Mg-HG. Figures 7(a)–7(c) show the BF images and the diffraction patterns of as-hydrogenated samples of Mg-G, Mg-HG, and Mg-AG, respectively. From the BF images, it seems that the particle sizes of Mg-G, Mg-HG, and Mg-AG are the almost same as the as-synthesized samples. From the diffraction patterns, the diffraction spots from MgH₂ (110) and the diffraction rings from C (002), MgO (200), MgO (220) are indexed in Mg-G. In Mg-HG and Mg-AG, the diffraction spots from MgH₂ (110), MgH₂ (101), MgH₂ (202), Mg (101), the diffraction rings from C (002), MgO (200), and the diffraction spots from MgH₂ (110), MgH₂ (202), MgH₂ (220), the diffraction rings from MgO (220), MgO (200), were identified, respectively, as listed in Table II. From Table II, we can consider that MgH₂ is hydrogenated from Mg and/or MgC₂ in Mg-G and Mg-HG. In case of Mg-AG, no crystalline was found in as-synthesized sample, and MgH₂ was found in as-hydrogenated sample. Therefore, Mg-AG could be totally different from Mg-G and Mg-HG. Meanwhile, the crystallite size of graphite in Mg-G is slightly different from that in Mg-HG, as shown in Figure 8. Figure 8 shows HRTEM images of graphite in as-synthesized Mg-G, Mg-HG, and Mg-AG with their average size. Compared with Mg-G and Mg-HG, there was no crystalline compounds in Mg-AG. Those results about the crystallinity of each sample obtained by TEM experiment were agreeable with the previous results obtained by XRD, except for identifying the Mg-C compound, that is, this is the first report on MgC₂ existing in Mg-C-H system.

With respect to the hydrogenation reaction of Mg-C-H system, it has been reported that the hydrogenation reactions of Mg-G, Mg-HG, and Mg-AG happened at about 220 °C, 220 °C, and 400 °C, respectively. As is listed in Table II, we

![FIG. 6. BF images and diffraction pattern of the as-synthesized samples: (a) Mg-G, (b) Mg-HG, and (c) Mg-AG.](image1)

![FIG. 7. BF images and diffraction pattern of hydrogenated samples: (a) Mg-G, (b) Mg-HG, and (c) Mg-AG.](image2)

![FIG. 8. HRTEM images of graphite: (a) Mg-G, (b) Mg-HG and (c) Mg-AG.](image3)
TABLE II. Temperature for hydrogenation, compounds identified by electron diffraction patterns in as-synthesized samples, and compounds identified by electron diffraction patterns in hydrogenated samples of Mg-G, Mg-HG, and Mg-AG.

<table>
<thead>
<tr>
<th>Sample</th>
<th>Temperature for hydrogenation (°C)</th>
<th>Compounds in as-synthesized</th>
<th>Compounds in hydrogenated</th>
</tr>
</thead>
<tbody>
<tr>
<td>Mg-G</td>
<td>220</td>
<td>C, Mg, MgC2, MgO</td>
<td>C, MgH2, MgO</td>
</tr>
<tr>
<td>Mg-HG</td>
<td>220</td>
<td>C, Mg, MgC2, MgO</td>
<td>C, Mg, MgH2, MgO</td>
</tr>
<tr>
<td>Mg-AG</td>
<td>440</td>
<td>Amorphous, MgO</td>
<td>C, MgH2, MgO</td>
</tr>
</tbody>
</table>

found MgC2 in as-synthesized sample and MgH2 in as-hydrogenated sample of both Mg-G and Mg-HG. So, we expected that Mg and MgC2 absorbed hydrogen at 220 °C to be MgH2. The hydrogenation reactions of Mg and MgC2 were considered as follows:

\[
\text{Mg} + \text{H}_2 \rightarrow \text{MgH}_2, \\
\text{MgC}_2 + \text{H}_2 \rightarrow \text{MgH}_2 + 2\text{C}.
\]

In general, Mg can absorb hydrogen over 300 °C. It would be indicated that MgC2 somehow helped Mg absorbing hydrogen. Actually, it was confirmed that MgC2 phase was located nearby Mg phase. (Figure 9 shows (a) HRTEM image of Mg-G, (b) FFT pattern of a boxed part of (a), and (c) IFFT pattern obtained from spots of Mg and MgC2 in (b).) Meanwhile, it was reported that Mg-AG can absorb hydrogen at 440 °C. In Mg-AG, amorphous materials or composite with poor crystallinity, which called (Mg-C)nano in Ref. 30, absorbed hydrogen to be MgH2. However, the reaction mechanism of Mg-AG has not clarified yet in this study, however, we confirmed that metal-C-H system has high potentials for promising hydrogen storage materials.

IV. CONCLUSION

In this study, we have investigated the compounds in the Li-C-H and Mg-C-H systems to understand the reaction mechanism. Conclusively, we have determined the reaction formula of the hydrogenation and dehydrogenation in both systems. In the L-C-H system, we identified two kinds of compounds between Li and C. One was LiC6 (LiC12) of GIC, the other was LiC2. The hydrogenation reaction formula of each compound were determined as

\[
\text{LiH} + \text{C} \rightarrow \text{LiC12 (or LiC6)} + \text{H}_2, \\
\text{LiH} + \text{C} \rightarrow \text{Li}_2\text{C}_2 + \text{H}_2.
\]

And the dehydrogenation reaction formula of each compound were determined as

\[
\text{LiC12 (or LiC6)} + \text{H}_2 \rightarrow \text{LiH} + \text{C}, \\
\text{Li}_2\text{C}_2 + \text{H}_2 \rightarrow \text{LiH} + \text{CH}_x.
\]

Here, it should be noticed that de/hydrogenation Li2C2 is irreversible because of emission of CHx. Additionally, we have discussed the hydrogenation mechanism of GIC, namely, it has been indicated that the hydrogenation of GIC proceeded with generating LiH at the edge site of graphite.

In the Mg-C-H system, we identified the MgC2 as the compound between Mg and C. MgC2 absorbed hydrogen at 220 °C to be MgH2

\[
\text{MgC}_2 + \text{H}_2 \rightarrow \text{MgH}_2 + 2\text{C}.
\]

According to this reaction equation, MgC2 can theoretically absorb and desorb 4.1 mass% of hydrogen at around 200 °C. Additionally, Mg with MgC2 can absorb and desorb hydrogen at around 200 °C, though Mg itself absorb and desorb hydrogen at around 400 °C. The mechanism of this improvement has not clarified yet in this study, however, we confirmed that metal-C-H system has high potentials for promising hydrogen storage materials.

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