Dehydrogenation properties of AB–MH (Ammonia Borane–Metal Hydride, M = K, Na, Li, Ca, Mg, Al) composites were systematically investigated by thermal and mass analyses. The results suggest that the Pauling electronegativity of M, χp, is a good indicator to predict the phases of composites, the dehydrogenation temperature and the amount of by-product gases (NH3 and B2H6). The phases of composites were classified by χp as follows. MBH4 was formed for M = K, Na (χp ≤ 0.9), MNH2BH3 was formed for M = Na, Li (0.9 ≤ χp ≤ 1.0) and no new compounds were formed for M = Ca, Mg, Al (1.0 ≤ χp). The 1st dehydrogenation temperatures of the samples (M = Na, Li, Ca, Mg) were 10–20 °C lower than that of AB itself (χp ≤ 1.2). The amount of NH3 was decreased as χp increased. On the other hand, the amount of B2H6 was decreased as χp decreased. The emission of B3H6N3 could occur by the reaction of NH3 and B2H6. Finally, AB–MAlH4 (M = Na, Li) composites, which were prepared based on the indicator, showed superior potential as hydrogen storage materials because they did not desorb any by-products NH3, B2H6 and B3H6N3.
AB–MH (Ammonia Borane–Metal Hydride) composites: systematic understanding of dehydrogenation properties

Yuki Nakagawa,*a Shigehito Isobe,*ab Yudai Ikarashi*a and Somei Ohnuki3

Dehydrogenation properties of AB–MH (Ammonia Borane–Metal Hydride, M = K, Na, Li, Ca, Mg, Al) composites were systematically investigated by thermal and mass analyses. The results suggest that the Pauling electronegativity of M, \( \chi_p \), is a good indicator to predict the phases of composites, the dehydrogenation temperature and the amount of by-product gases (NH\(_3\) and B\(_2\)H\(_6\)). The phases of composites were classified by \( \chi_p \) as follows. MBH\(_4\) was formed for M = K, Na (\( \chi_p \leq 0.9 \)), MNH\(_2\)BH\(_3\) was formed for M = Na, Li (0.9 \( \leq \chi_p \leq 1.0 \)) and no new compounds were formed for M = Ca, Mg, Al (1.0 \( \leq \chi_p \)). The 1st dehydrogenation temperatures of the samples (M = Na, Li, Ca, Mg) were 10–20 °C lower than that of AB itself (\( \chi_p \approx 1.2 \)). The amount of NH\(_3\) was decreased as \( \chi_p \) increased. On the other hand, the amount of B\(_2\)H\(_6\) was decreased as \( \chi_p \) decreased. The emission of B\(_3\)H\(_6\)N\(_3\) could occur by the reaction of NH\(_3\) and B\(_2\)H\(_6\). Finally, AB–MAIH\(_4\) (M = Na, Li) composites, which were prepared based on the indicator, showed superior potential as hydrogen storage materials because they did not desorb any by-products NH\(_3\), B\(_2\)H\(_6\) and B\(_3\)H\(_6\)N\(_3\).

1. Introduction

Ammonia borane (NH\(_3\)BH\(_3\), AB) is an attractive candidate for hydrogen storage media because of its high hydrogen content (19.6 wt%, 0.145 kg L\(^{-1}\)). It can desorb \(~13\) wt% of hydrogen below 200 °C. The dehydrogenation of AB takes place in three steps within a single equivalent of hydrogen evolved in each step at ca. 70–110, 110–200, and 400–900 °C in temperature ramping experiments. One of the disadvantages for practical application is the emission of by-product gases such as ammonia (NH\(_3\)), diborane (B\(_2\)H\(_6\)), and borazine (B\(_3\)H\(_6\)N\(_3\)). For instance, release of ammonia causes damage to the fuel cell performance even at trace levels. Worse yet, NH\(_3\) and B\(_2\)H\(_6\) are harmful for living creatures. The dehydrogenation properties of AB–MH composites have been extensively studied, such as AB–LiH, AB–NaH, AB–LiH–NaH, AB–KH, AB–MgH\(_2\), AB–CaH\(_2\), AB–LiNH\(_2\), AB–LiBH\(_4\), AB–LiAlH\(_4\), and AB–LiNH\(_2\)–LiBH\(_4\), in order to improve the dehydrogenation properties of AB. For example, the AB–NaH composite desorbed hydrogen at 80–100 °C without the by-product diborane and borazine. The AB–LiNH\(_2\) composite desorbed hydrogen as low as 60 °C without the by-products diborane and borazine. However, systematic investigation on AB–MH composites has not been explored. On the other hand, thermodynamical stabilities of M(BH\(_4\))\(_n\) and M(AlH\(_4\))\(_n\) systems have been systematically investigated by using the Pauling electronegativity of M as an indicator. The correlation between the dehydrogenation temperature and the Pauling electronegativity of M was found computationally and experimentally. It is of great importance to understand the dehydrogenation properties of AB–MH composites systematically in order to decrease the dehydrogenation temperature and suppress by-product gas emission.

In this study, AB–MH (M = K, Na, Li, Ca, Mg, Al) composites were synthesized by a ball milling method. Phases of AB–MH composites were determined by X-ray diffraction (XRD) analysis and dehydrogenation properties were discussed in terms of dehydrogenation temperature and amounts of hydrogen and by-product gas emissions. From the results, we proposed an indicator to predict the phases of composites, the dehydrogenation temperature and the amount of by-product gas emission (NH\(_3\) and B\(_2\)H\(_6\)). Finally, we created superior composites according to the indicator and evaluated their dehydrogenation properties.

2. Experimental

The starting materials NH\(_3\)BH\(_3\), NaH, LiH, CaH\(_2\), NaAlH\(_4\), and LiAlH\(_4\) (purity 97%, 55–65% [moistened with oil], 95%, 99.99%, 90%, and 95%, respectively) were purchased from Sigma Aldrich Co. Ltd. MgH\(_2\) (purity 98%) was purchased from Alfa Aesar. These materials were used as-received.
without any purification. AlH₃ was prepared by the chemical reaction between LiAlH₄ and AlCl₃ in ether solution.²⁴ KH was prepared according to ref. 35. All samples were handled in an argon glovebox without exposing samples to air. AB–MH (M = K, Na, Li, Ca, Mg, Al) and AB–MAH₄ (M = Na, Li) composites were prepared by ball-milling under 0.1 MPa Ar with 400 rpm for 30 min, 1.0 MPa H₂ with 300 rpm for 5 min, respectively. Ball-milling processes were performed by using a planetary ball-mill apparatus (Fritsch Pulverisette 7) with 20 stainless steel balls (7 mm in diameter) and 300 mg samples (ball : powder ratio = 70 : 1, by mass). The phase analysis was conducted by X-ray powder diffraction (XRD, PANalytical, X’Pert-Pro with Cu Kα radiation). The dehydrogenation properties were examined by thermal desorption mass spectrometry measurements (TDMS, ULVAC, BGM-102) combined with thermogravimetry and differential thermal analysis (TG-DTA, Bruker, 2000SA). The heating rate was 2 or 5 °C min⁻¹ and the helium gas flow rate was 300 mL min⁻¹. TDMS measurements combined with TG-DTA were performed twice for each sample. The amounts of hydrogen and by-product gases were defined by integrating the peaks of mass spectra. The value of integral for by-product gases was normalized by that for hydrogen. Then, the weight loss of each gas was estimated by the total weight loss obtained from TG results. All measurements were immediately performed after synthesizing the AB–MH composite in order to prevent the composite from decomposing.

3. Results and discussion

3.1 Phase analysis of AB–MH composites

Fig. 1 shows the XRD patterns of each AB–MH composite and AB after milling and heating to 200 °C. Broad diffraction peaks around 20° and 27° in all profiles originate from the film and grease to prevent the sample oxidation. A small amount of (NH₂BH₃)₄ was observed in the milled AB. This phase was considered as impurity in as-received AB. After heating AB, there were no peaks in the profile, indicating that AB became amorphous. For the AB–NaH composite, peaks were consistent with sodium amidoborane (NaNH₂BH₃), which was confirmed by Xiong et al.¹⁸ After heating, NaBH₄ was confirmed by XRD, which is different from the experimental results of previous investigators.¹⁹–²¹ One of the possibilities of NaBH₄ formation would come from the different milling condition. There are some cases that non-equilibrium phases appear by mechanical ball milling.²² As described in ref. 20, the formation of BH₄⁻ ions may occur under the milling condition. For the AB–LiH composite, peaks were consistent with lithium amidoborane–ammonia borane (LiNH₂BH₃·NH₂BH₃). This was the intermediate phase during the synthesis of LiNH₂BH₃, which was confirmed by Wu et al.²⁷ Thus, metal amidoborane (MNH₂BH₃ (M = Na, Li)) was confirmed after milling for M = Na, Li. For the AB–KH composite, KBH₄ was observed instead of potassium amidoborane (KNH₂BH₃) after milling. The wet chemical synthesis of KNH₂BH₃ by stirring KH and AB in benzene for 2 days suggests that a lower rotation speed and longer time of ball milling enables generation of KNH₂BH₃.²²

In cases of AB–MH (M = Ca, Mg, Al) composites, no new compounds were observed after milling. Only AB and each MH phase were observed. It was reported that more than 6 hours milling by using a shaker mill can generate Ca(NH₂BH₃)₂.²³ However, synthesis of Mg(NH₂BH₃)₂ by ball milling was unsuccessful.²⁴ After heating, only each MH phase was observed for M = Mg, Ca. These results are in good agreement with previous reports.²³–²⁴ In the case of M = Al, Al was observed after heating, which indicated that AlH₃ desorbed hydrogen below 200 °C. From Fig. 1, we described the crystalline phases of AB–MH composites according to the Pauling electronegativity of M, Xₚ, in Table 1. The previous study about the stability of M(NH₂BH₃)ₙ by the computational method indicates that electronegativity is a good indicator to understand the stability generally.²³ As shown in Table 1, MBH₄ was formed for M = K, Na (Xₚ ≤ 0.9), MNH₂BH₃ was formed for M = Na, Li (0.9 ≤ Xₚ ≤ 1.0) and no new compounds were formed for M = Ca, Mg, Al (1.0 ≤ Xₚ). This suggests that the trend of the phases of AB–MH composites follows the trend of Xₚ. MH tends to cause the reaction with AB faster when Xₚ decreases.
3.2 Dehydrogenation properties of AB–MH composites

Fig. 2 shows the dehydrogenation temperatures of AB–MH (M = Na, Ca, Li, Mg, Al) composites below 200 °C. The results of the AB–KH composite were not plotted because it did not desorb any gases during heating below 200 °C. The 1st and 2nd dehydrogenation peak temperatures in mass spectra are plotted in Fig. 2(a) and (b), respectively. The dashed lines show the temperature of milled AB. As shown in Fig. 2(a), 1st temperatures for M = Na, Li, Ca, Mg (\(c_p \approx 1.2\)) were decreased by 10–20 °C as compared with milled AB. These results are in good agreement with other results.\(^{10,23,24}\) The 1st temperature for M = Al was not changed as milled AB. It is interesting to note that this trend is opposite to the trend in the dehydrogenation temperatures of M(BH\(_4\))\(_n\) and M(AlH\(_4\))\(_n\).\(^{30–33}\) The 2nd temperatures for M = Ca, Mg, Al were correlated with the electronegativity of M as shown in Fig. 2(b). It is difficult to compare the whole results of temperatures because phases are different between those for M = Na, Li and those for M = Ca, Mg, Al. The AB–CaH\(_2\) composite showed the lowest temperature of all the 2nd dehydrogenation temperatures. The dehydrogenation from AB–MH composites could be promoted by the solid-phase interaction between AB and MH. This interaction would affect the intramolecular N–H, B–H and B–N chemical bonds and intermolecular dihydrogen bond.

The amounts of hydrogen and by-product gases of AB–MH composites were investigated by TG-MASS. Fig. 3(a) shows the estimated weight losses of hydrogen and by-product gases below 200 °C. Milled AB desorbed by-product gases more than 60%, indicating the substantial amount of by-product gas emission. All the AB–MH composites showed by-product gas emission of 6–18%, suggesting that most of the by-product gases were suppressed. Fig. 3(b) shows the hydrogen purity of AB–MH composites below 200 °C. The heating rate was 2 °C min\(^{-1}\).

<table>
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<th>MH</th>
<th>KH</th>
<th>NaH</th>
<th>LiH</th>
<th>CaH(_2)</th>
<th>MgH(_2)</th>
<th>AlH(_3)</th>
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<td>Pauling electronegativity (x_p) of M</td>
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<td>0.9</td>
<td>1.0</td>
<td>1.0</td>
<td>1.2</td>
<td>1.5</td>
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<td>Crystalline phases after milling</td>
<td>KBH(_4)</td>
<td>NaNH(_2)BH(_3)</td>
<td>LiNH(_2)BH(_3)·NH(_2)BH(_3)</td>
<td>CaH(_3)</td>
<td>MgH(_2)</td>
<td>AlH(_3)</td>
</tr>
<tr>
<td>Crystalline phases after heating (200 °C)</td>
<td>KBH(_4)</td>
<td>NaNH(_2)</td>
<td>LiH</td>
<td>CaH(_3)</td>
<td>MgH(_2)</td>
<td>Al</td>
</tr>
<tr>
<td>Decomposition temperature/°C</td>
<td>425</td>
<td>400</td>
<td>555</td>
<td>627</td>
<td>446</td>
<td>n/a</td>
</tr>
</tbody>
</table>

Fig. 2 Dehydrogenation temperatures of AB–MH composites below 200 °C classified by the Pauling electronegativity of M: (a) 1st peak temperatures and (b) 2nd peak temperatures. The heating rate was 2 °C min\(^{-1}\). The black square symbols show the results of the 1st measurement and the white square symbols show those of the 2nd measurement.

Fig. 3 (a) Weight losses of hydrogen and by-product gases and (b) hydrogen purity of AB–MH composites below 200 °C. The heating rate was 2 °C min\(^{-1}\).
AB–MH composites. A purity of about 90% was obtained for M = Li, Mg, Al. On the other hand, the purity for M = Na, Ca was only less than 80%. In order to investigate the content of by-product gases, we estimated the amounts of by-product gases of NH₃, B₂H₆ and B₃H₆N₃. Fig. 4 shows the amounts of by-product gases of AB–MH composites. NH₃, B₂H₆ and B₃H₆N₃ were analysed by mass spectrometry. The dashed line shows the amount of milled AB. As shown in Fig. 4(a), the amount of NH₃ was decreased as \( \chi_p \) increased. The AB–AlH₃ composite almost suppressed the emission of NH₃. On the other hand, the amount of B₂H₆ was decreased as \( \chi_p \) decreased as shown in Fig. 4(b). AB–MH (M = Na, Li, Ca (\( \chi_p \leq 1.0 \))) composites completely suppressed B₂H₆. Comparing the results of Fig. 3(b), it was found that the low hydrogen purity for M = Na, Ca was ascribed to a large amount of NH₃ emission. The emission process of NH₃ in the NaNH₂BH₃ system was reported by Fijałkowski et al., which explained the fact that the formation and decomposition of the ionic salt caused the emission. The ionic radius of M⁺ seems to be correlated with the emission of NH₃. The ionic radii of Na⁺ (102 pm) and Ca²⁺ (100 pm) are different from those of Li⁺ (76 pm), Mg²⁺ (72 pm), and Al³⁺ (54 pm). The trend in the amount of B₂H₆ was similar to that in the case of M(BH₄)ₙ. In the case of M(BH₄)ₙ, those for \( \chi_p \leq 1.5 \) suppressed the emission of B₂H₆. The previous study showed diammoniate of diborane (DADB), \[ \text{[(NH₃)₂BH₂]+[BH₄]-} \], an ionic isomer of AB, is formed during the induction period before dehydrogenation of AB. If DADB is regarded as a kind of borohydride, the emission process of B₂H₆ in AB is considered to be similar to that in M(BH₄)ₙ. Further investigations are needed to clarify the suppression mechanisms of NH₃ and B₂H₆ emission in AB–MH composites. Fig. 4(c) shows that only the AB–MgH₂ composite desorbed B₃H₆N₃. The emission of B₃H₆N₃ may correlate with the emission of NH₃ and B₂H₆. B₂H₆N₃ can be generated by the reaction between NH₃ and B₂H₆ with a molar ratio of 2 : 1. The AB–MgH₂ composite desorbed NH₃ and B₂H₆, then they would react to form B₃H₆N₃. In the other composites, emission of either NH₃ or B₂H₆ was suppressed, which would result in the suppression of B₃H₆N₃ emission. As a result, \( \chi_p \) would be a good indicator to predict the dehydrogenation temperatures and the amounts of by-product gases. We summarized the dehydrogenation properties of composites in Fig. 5. Fig. 5 shows the amounts of hydrogen desorbed below 200 °C versus dehydrogenation temperatures. The AB–CaH₂ composite showed the lowest dehydrogenation temperature and the AB–AlH₃ composite showed the largest amount of hydrogen of all the composites. However, it was indicated that there were no materials fulfilling large amounts of hydrogen desorbed at low temperatures. In order to accomplish these requirements, further investigations, e.g. combining AB with more than two metal hydrides, will be needed.
3.3 Creation of AB–MH composites based on the indicator

On the basis of the indicator described in Section 3.2, we created superior AB–MH composites. As shown in Fig. 4, the emissions of NH₃ and B₂H₆ were suppressed by combining with AlH₃ and NaH (LiH), respectively. Therefore, we thought the idea of combining MAH₄ (M = Na, Li) with AB because MAH₄ (M = Na, Li) is the compound consisting of MH (M = Na, Li) and AlH₃. The milling conditions of AB–MAH₄ composites are described in the Experimental part. The milling time of 5 min would be appropriate because milling for 30 min resulted in the decomposition of composites during milling. The mass spectra of AB–NaAlH₄ and AB–LiAlH₄ composites are shown in Fig. 6. It is interesting that both composites did not desorb NH₃, B₂H₆, and B₃H₆N₃ at all within the accuracy of our apparatus. Only hydrogen gas was desorbed. About 4 wt% of hydrogen was desorbed below 240 °C in both composites. The suppression of by-product gas emission was also confirmed in AB–Li₃AlH₆ composites. Thus, AB–MAH₄ (M = Na, Li) composites showed superior potential as hydrogen storage materials. Investigations of the reaction processes of both composites are currently in progress.

4. Conclusion

In order to decrease the dehydrogenation temperature and suppress by-product gas emission, we investigated the phases and dehydrogenation properties of AB–MH composites prepared by ball-milling. MBH₄ was formed for M = K, Na (χₚ ≤ 0.9), MNH₂BH₃ was formed for M = Na, Li (0.9 ≤ χₚ ≤ 1.0) and no new compounds were formed for M = Ca, Mg, Al (1.0 ≤ χₚ). 1st dehydrogenation temperatures for M = Na, Li, Ca, Mg (χₚ ≤ 1.2) were decreased by 10–20 °C as compared with the milled one. The amount of NH₃ was decreased as electronegativity increased. The AB–AlH₃ composite almost suppressed the emission of NH₃. On the other hand, the amount of B₂H₆ was decreased as electronegativity decreased. AB–MH (M = Na, Li, Ca (χₚ ≤ 1.0)) composites completely suppressed B₂H₆. B₃H₆N₃ emission was observed for M = Mg. The emission could be occurred by the reaction between NH₃ and B₂H₆. These results suggested that the Pauling electronegativity of M, χₚ, is a good indicator to predict the phases of composites, dehydrogenation temperature and the amount of by-product gas emission (NH₃ and B₂H₆). AB–MAH₄ (M = Na, Li) composites, which were prepared based on the indicator, showed superior potential as hydrogen storage materials because they did not desorb any by-products NH₃, B₂H₆ and B₃H₆N₃. These results would be helpful for clarifying the improvement mechanism of dehydrogenation properties and designing new hydrogen storage materials.

Notes and references

8. International Chemical Safety Cards, ICSC number: 0414.


33 J. Graetz, *Chem. Soc. Rev.*, 2009, **38**, 73–82.


