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Tetrahedron Letters, 39(16), 2357-2360

https://doi.org/10.1016/S0040-4039(98)00199-3

1998-04-16

http://hdl.handle.net/2115/56475

article (author version)

(16) B-B + Allene (Com).pdf

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Platinum(0)-Catalyzed Diboration of Allenes with Bis(pinacolato)diboron

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Abstract: Bis(pinacolato)diboron [(Me₄C₂O₂)B-B(O₂C₂Me₄)] added to allenenes in the presence of Pt(PPh₃)₄ at 80 °C or Pt(dba)₂/((c-Hex)₃P at 50 °C to afford 2,3-bis(boryl)-1-propenes in excellent yields. The one-pot synthesis of substituted homoallyl alcohols was also examined by the allylboration of aldehydes with diborated products followed by the coupling with organic halides.

Transition-metal-catalyzed addition reaction of metal-metal reagents to allenenes provides a straightforward route to 2,3-bis(metal)-1-propene and its analogs, which include synthetically useful allylmetal part and vinylmetal part in the same molecule. Although disilanes,¹ distannanes,² silylstannanes,³ and germystannanes⁴ are known to add to allenenes in the presence of palladium(0) catalyst, the corresponding reaction of diborons has not been reported so far presumably due to difficulties in the boron-boron bond activation by its oxidative addition to the palladium(0) complexes.⁵ Recently, we have found that platinum(0) complex effectively activates the boron-boron bond by the oxidative addition to give bis(boryl)platinum(II) complex.⁶ This activation protocol is readily extended to the catalytic diboration of unsaturated hydrocarbons, such as alkynes,⁶ alkenes,⁷ and 1,3-dienes.⁸ In the course of our study, we wish to disclose herein the platinum(0)-catalyzed addition of bis(pinacolato)diboron (1) to allenenes (2) to afford 2,3-bis(boryl)-1-propenes (3) (Eq. 1).

\[
(RO)₂B-B(OR)₂ + \underset{\text{Pt catalyst}}{\text{toluene}} \rightarrow (RO)₂B-B(OR)₂ + (RO)₂B-B(OR)₂ \quad (\text{1})
\]

1,2-Propadiene (1.5 mmol) was allowed to react with 1 (1.0 mmol) for 16 h to optimize the reaction conditions. The addition completed at 80 °C with 3 mol% of Pt(PPh₃)₄ in toluene, producing the corresponding 3 in 99% yield. Ligand less platinum(0) complex such as Pt(dba)₂ also gave 3 even at room temperature, but the yield of the adduct is rather low (50%) because of catalyst decomposition. The catalytic activity of platinum(0) complex is markedly influenced by the nature of ligand. Comparison of the reaction rate at room temperature with Pd(dba)₂/PR₃ (1:1) elucidates the effectiveness of more sterically demanding ligand: e.g., (c-Hex)₃P (85%) > (4-MeOC₆H₄)₃P.
(66%) > (4-ClC₆H₄)₃P (46%) > (C₆H₅)₃P (25%) > Me₃P (7%). As for the solvents, less polar one such as toluene is more favorable than dioxane or DMF.

The representative results of the reaction between 1 and 2 are summarized in Table 1. A variety 2 having alkyl and aryl substituents were smoothly converted into the corresponding 3 (82-99%) in toluene by using of Pt(PPh₃)₄ at 80 °C (Conditions A) or Pt(dba)₂/(c-Hex)₃P at 50 °C (Conditions B) (Entries 1-6), while strong electron-donating substituents on 2 such as MeO or MeS groups sufficiently slowed down the addition rate (Entries 7 and 8). The regioselectivity of the addition depends upon both the substituents of 2 and the bulkiness of ligands. The reaction of monosubstituted 2 (Entries 2-4) with less bulky ligands preferentially afforded the internal adducts (3b), whereas that of 1,1-disubstituted (Entries 5 and 6) or heteroatom-substituted 2 (Entries 7 and 8) with more bulky ligands tend to give the terminal adducts (3a). The (Z)-configuration of 3a (Entries 4, 7, 8) was immediately established by their conversion into known homoallyl alcohols and the comparison of ¹H NMR spectra with authentic alcohols (Eq. 2).

![Chemical structure](image)

(2)

The catalytic cycle for the present reaction may involve the oxidative addition of 1 to platinum(0) complex to give bis(boryl)platinum(II) intermediate (4), the insertion of 2 into B-Pt bond to provide vinyl- or p-allylplatinum(II) species (5 or 6), and the reductive elimination of 3 (Fig. 1). The stoichiometric reaction between 1,2-heptadiene and cis-Pt(BO₂C₆Me₄)₂(PPh₃)₆ resulted in an 83% yield of 3 (3a:3b = 17:83), strongly supporting the above catalytic cycle (Eq. 3). The slow reaction of 2 having strong electron-donating groups (Entries 7 and 8) may reflect the rate of insertion step, which was consistent with the relative reactivity observed on the insertion of diarylacetylenes to the B-Pt bond. The selective formation of 3a in the reaction of 1,1-disubstituted 2 may be due to the steric requirement; however, the observed regio- and stereoselectivity for monosubstituted and heteroatom-substituted 2 remains controversial.
Figure 1. The Catalytic Cycle for the Diboration

\[
\text{cis-Pt[B(OR)\textsubscript{2}](PPh\textsubscript{3})\textsubscript{2}} + \text{BuH} \xrightarrow{C\textsubscript{6}D\textsubscript{5}} \text{50 °C/6 h} \xrightarrow{\text{BuH}} \text{PhCHO} \xrightarrow{\text{PdCl\textsubscript{2}}(dppf)/aq KOH dioxane/90 °C/16 h} \text{75% PhOH}
\]

83 % (3a:3b = 17:83)

The ready availability of various 3 from 2 now offers a simple route to substituted homoallyl alcohols. For example, the sequential reaction involving the allylboration of benzaldehyde (1.1 mmol) with 2,3-bis(boryl)-1-propene (1.0 mmol) in dioxane and the cross-coupling with iodobenzene (1.1 mmol) in the presence of PdCl\textsubscript{2}(dppf) (3 mol\%) and aqueous KOH (3 mmol) gave the corresponding homoallyl alcohol in 75% yield (Eq. 4).

A representative procedure for 3: To Pt(PPh\textsubscript{3})\textsubscript{4} (0.03 mmol) and 1 (1.0 mmol) were successively added toluene (6 ml) and 1,2-propadiene (1.5 mmol), and the resulting solution was then stirred at 80 °C for 16 h in a sealed reaction tube. Concentration of the reaction mixture and Kugelrohr distillation gave 2,3-bis(boryl)-1-propene: bp 130 °C/0.1 mmHg (oven temperature); \textsuperscript{1}H NMR (400 MHz, CDCl\textsubscript{3}) \textsuperscript{d} 1.24 (s, 12 H), 1.26 (s, 12 H), 1.82 (s, 2 H), 5.58 (br s, 1 H), and 5.71 (d, 1 H, \textit{J} = 3.4 Hz); \textsuperscript{13}C NMR (100 MHz, CDCl\textsubscript{3}) \textsuperscript{d} 24.74, 83.05, 83.35, and 128.44; \textsuperscript{11}B NMR (128 MHz, CDCl\textsubscript{3}) \textsuperscript{d} 30.03 and 33.46.

References


**Table 1.** The Synthesis of 3 (Eq. 1)\(^a\)

<table>
<thead>
<tr>
<th>Entry</th>
<th>Allene (2)</th>
<th>Product (3a, 3b)</th>
<th>Conditions A</th>
<th>Conditions B</th>
</tr>
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<tbody>
<tr>
<td></td>
<td></td>
<td></td>
<td>Yield/%(^b) 3a:3b(^c)</td>
<td>Yield/%(^b) 3a:3b(^c)</td>
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<td><img src="image" alt="Product" /></td>
<td>90 7:93 82 8:92</td>
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<tr>
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<tr>
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<td>96 50:50 84 85:15</td>
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<tr>
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<td>98 76:24 99 98:2</td>
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<tr>
<td>7</td>
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<td><img src="image" alt="Product" /></td>
<td>81(^d) 100:0 85(^d) 100:0</td>
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<tr>
<td>8</td>
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<td><img src="image" alt="Product" /></td>
<td>48(^d) 50:50 82(^d) 82:18</td>
<td></td>
</tr>
</tbody>
</table>

\(^a\)All reactions were carried out in toluene (6 ml) using 1 (1.0 mmol), 2 (1.5 mmol), and platinum catalyst (0.03 mmol).

Conditions A: Pt(PPh\(_3\))\(_2\)/80 °C/16 h. Conditions B: Pt(db)\(_2\)/(c-Hex)\(_3\)P/50 °C/18 h. The exact procedure, see the text.

\(^b\)Isolated yields based on 1. \(^c\)Regioisomeric purity was determined by GLC and \(^1\)H NMR analyses.

\(^d\)0.1 mmol of catalyst was used.
PLATINUM(0)-CATALYZED ADDITION REACTION OF DIBORON REAGENT TO ALLENES

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\[
(\text{RO})_2\text{B}\text{-B(OR)}_2 + \text{H} \xrightarrow{\text{Pt catalyst}} (\text{RO})_2\text{B}\text{(OR)}_2 + \text{Me}_4\text{C}_2\text{O}_2
\]

\( (\text{RO})_2 = \text{Me}_4\text{C}_2\text{O}_2 \)