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Low temperature electronic transport and electron transfer through organic macromolecules
Study of electron transport in hydrocarbon gases

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The drift velocity and the effective ionization coefficient of electrons in the organic gases, C_2H_2, C_2H_4, C_2H_6, CH_3OH, C_2H_5OH, C_6H_6, and C_6H_5CH_3, have been measured over relatively wide ranges of density-reduced electric fields (E/N) at room temperature (around 300 K). The drift velocity was measured, based on the arrival-time spectra of electrons by using a double-shutter drift tube over the E/N range from 300 to 2800 Td, and the effective ionization coefficient (\(\bar{\alpha} - \eta\)) was determined by the steady-state Townsend method from 150 to 3000 Td. Whenever possible, these parameters were compared with those available in the literature. It has been shown that the swarm parameters for these gases have specific tendencies, depending on their molecular configurations.

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I. INTRODUCTION

The elementary molecules containing carbon and hydrogen atoms (organic molecules) are of importance in manufacturing processes of graphene, carbon nanotubes (CNT), carbon nanowalls (CNW), diamond-like carbon (DLC) layers, and so on. In such processes, the molecules in the gas phase are often made available in the formation and growth of these structures through intermediate plasma states. For the application of a variety of molecules, however, there still remain some data associated with electron collisions to be measured in order to supply fundamental parameters needed in designing with simulation models. Unfortunately, quantities such as electron swarm parameters have seldom been measured in some of these gases over the past few decades.

In this article, we present the experimental results on the electron transport parameters (swarm parameters) for the organic molecules, C_2H_2 (acetylene), C_2H_4 (ethylene), C_2H_6 (ethane), CH_3OH (methanol), C_2H_5OH (ethanol), C_6H_6 (benzene), and C_6H_5CH_3 (toluene), obtained from the arrival-time spectra (ATS) method using a double-shutter drift tube and a steady-state Townsend (SST) method. The experimental swarm parameters for these gases have been scarcely reported so far. Using the ATS method, we have measured the drift velocity, \(\bar{W}_m\), which is defined from the ATS of electrons as

\[
\bar{W}_m = \frac{d\langle t \rangle}{dz}.
\]

Here, \(\langle t \rangle\) represents the average of the arrival-time of electrons at a position \(z\) in the drift tube. The experiment is conducted using a double-shutter drift tube, in which the arriving electrons are collected by the electrode (with shutter-2) at a certain distance from the cathode (with shutter-1) to release initial electrons. The drift velocity, \(\bar{W}_m\), is deduced by dividing the mean time of arrival of electrons by the distance between the electrodes. The experimental apparatus and procedure for the ATS method were the same as those described by Hasegawa et al. In the present study, the position of the collector plate was set at four locations, and the final value of drift velocity was determined by averaging the values determined at the four cases for each value of E/N.

B. Steady-state Townsend method

The effective ionization coefficient \(\bar{\alpha}\) is given by

\[
\bar{\alpha} = \frac{1}{N(z)} \frac{dN(z)}{dz}
\]

with

\[
N(z) = \int_0^\infty n(z,t)dt,
\]
where \( n(z,t) \) is the electron number density at a position \( z \) and time \( t \). Under the SST conditions, the electron current growth, \( I \), is given by

\[
I = I_0 \frac{\alpha}{\alpha - \eta} \exp \left\{ (\alpha - \eta) d \right\} \frac{\eta}{\alpha - \eta} \frac{1}{1 - \gamma_T \frac{\alpha}{\alpha - \eta} \exp \left\{ (\alpha - \eta) d \right\} - 1}
\]

(4)

where \( d \) is the gap spacing between the electrodes, \( I_0 \) is the initial current value, and \( \gamma_T \) is the secondary coefficient, respectively. To obtain the effective ionization coefficient, we used the SST method since it provides a more precise value of this coefficient than that obtained by the ATS method.

In the experiment for the organic gases in this study, electron attachment is significant, while the secondary coefficient \( \gamma_T \) was found to be negligible. Thus, Eq. (4) readily reduces to

\[
I = I_0 \left[ \frac{\alpha}{\alpha - \eta} \exp \left\{ (\alpha - \eta) d \right\} \frac{\eta}{\alpha - \eta} \right].
\]

(5)

We determined \( \alpha \) and \( \eta \) by fitting this equation to the relation between the current and the gap length using a linearized least-squares method.

The experimental apparatus and analytical procedures for the SST method are basically same as those given by Hasegawa et al.5,7 The current measurement was repeated 12–26 times for each gap length and \( E/N \), and the effective ionization coefficient was determined by averaging the values from the corresponding current measurements.

**C. Purity of the gases and experimental conditions**

The purity of the gases used and the experimental conditions for the field strength and gas pressure are listed in Table I. The drift length between the cathode and shutter-2 to collect electrons in the ATS experiment was 10–40 mm, and the gap length of the electrodes in the SST experiment (using the Harrison type electrodes) was 2–15 mm. Both experiments were performed at room temperature (295–300 K), except for \( \text{C}_6\text{H}_6 \) and \( \text{C}_6\text{H}_5\text{CH}_3 \), which were performed at about 60°C.

**III. RESULTS AND DISCUSSION**

Figure 1 shows a comparison of the drift velocities \( W_m \) obtained for all gases. Two general tendencies are apparent:

**TABLE I. Experimental conditions in this study.**

<table>
<thead>
<tr>
<th>Gas</th>
<th>Purity (%)</th>
<th>Reduced electric field, E/N (Td)</th>
<th>Pressure (Torr)</th>
</tr>
</thead>
<tbody>
<tr>
<td></td>
<td></td>
<td>ATS</td>
<td>SST</td>
</tr>
<tr>
<td>( \text{C}_2\text{H}_2 )</td>
<td>99.99</td>
<td>300–2500</td>
<td>140–3000</td>
</tr>
<tr>
<td>( \text{C}_2\text{H}_4 )</td>
<td>99.99</td>
<td>300–2500</td>
<td>140–3000</td>
</tr>
<tr>
<td>( \text{C}_2\text{H}_6 )</td>
<td>99.95</td>
<td>300–2500</td>
<td>130–3000</td>
</tr>
<tr>
<td>( \text{CH}_3\text{OH} )</td>
<td>99.8</td>
<td>300–2300</td>
<td>150–3000</td>
</tr>
<tr>
<td>( \text{C}_2\text{H}_5\text{OH} )</td>
<td>99.5</td>
<td>300–2500</td>
<td>100–3000</td>
</tr>
<tr>
<td>( \text{C}_2\text{H}_6 )</td>
<td>99.7</td>
<td>300–2800</td>
<td>100–3000</td>
</tr>
<tr>
<td>( \text{C}_6\text{H}_5\text{CH}_3 )</td>
<td>99.7</td>
<td>300–2300</td>
<td>100–3000</td>
</tr>
</tbody>
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The experimental apparatus and analytical procedures for the SST method are basically same as those given by Hasegawa et al.5,7 The current measurement was repeated 12–26 times for each gap length and \( E/N \), and the effective ionization coefficient was determined by averaging the values from the corresponding current measurements.

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<td>SST</td>
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<tr>
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<td>99.99</td>
<td>300–2500</td>
<td>140–3000</td>
</tr>
<tr>
<td>( \text{C}_2\text{H}_4 )</td>
<td>99.99</td>
<td>300–2500</td>
<td>140–3000</td>
</tr>
<tr>
<td>( \text{C}_2\text{H}_6 )</td>
<td>99.95</td>
<td>300–2500</td>
<td>130–3000</td>
</tr>
<tr>
<td>( \text{CH}_3\text{OH} )</td>
<td>99.8</td>
<td>300–2300</td>
<td>150–3000</td>
</tr>
<tr>
<td>( \text{C}_2\text{H}_5\text{OH} )</td>
<td>99.5</td>
<td>300–2500</td>
<td>100–3000</td>
</tr>
<tr>
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<td>99.7</td>
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</tr>
<tr>
<td>( \text{C}_6\text{H}_5\text{CH}_3 )</td>
<td>99.7</td>
<td>300–2300</td>
<td>100–3000</td>
</tr>
</tbody>
</table>

The density-normalized effective ionization coefficient for all gases under study is shown in Fig. 3. In general, the ionization coefficient for the cyclic molecules, \( \text{C}_6\text{H}_6 \) and \( \text{C}_6\text{H}_5\text{CH}_3 \), is smaller than that for the others, particularly in the low \( E/N \) region below 500 Td. The results are separately shown in Figs. 4–6 for three groups: Fig. 4 for \( \text{C}_2\text{H}_2, \text{C}_2\text{H}_4, \) and \( \text{C}_2\text{H}_6 \), Fig. 5 for \( \text{CH}_3\text{OH} \) and \( \text{C}_2\text{H}_5\text{OH} \), and Fig. 6 for \( \text{C}_2\text{H}_6 \) and \( \text{C}_6\text{H}_5\text{CH}_3 \), and compared with other previously reported data.13–16 As in Fig. 4, although there exists a slight difference between our results and other investigations for each gas, the ionization coefficient in \( \text{C}_6\text{H}_6 \) is larger than those for the other two molecules over the \( E/N \) range in this study. As regards \( \text{CH}_3\text{OH} \), our result is in good agreement with the value of \( W_m \) is smaller as the molecular weight of the gas molecules becomes smaller for the normal chain structure cases of \( \text{C}_2\text{H}_2, \text{C}_2\text{H}_4, \) and \( \text{C}_2\text{H}_6 \), while \( W_m \) is larger for the gas molecules containing the hydroxyl radical (-OH) than for those without it. Figure 2 shows the drift velocities for \( \text{C}_2\text{H}_2, \text{C}_2\text{H}_4, \) and \( \text{C}_2\text{H}_6 \) in comparison with the data found in the literatures.8–12 Our present results are in good agreement with those from other investigations from 300 to 1000 Td. To the best of our knowledge, the drift velocities in these gases for \( E/N \) above 1000 Td have been measured for the first time in this study.
However, the ionization coefficient for C₆H₆ is up to a factor of 2 larger than that of Ref. 18 for E/N from 400 to 800 Td as shown in Fig. 6. The ionization coefficients for C₂H₂ and C₂H₄ for E/N above 2000 Td, CH₃OH and C₂H₅OH for E/N above 200 Td, and C₆H₆ for E/N above 2000 Td have not been reported previously, and the coefficients in C₆H₅CH₃ for E/N from 200 to 3000 Td have been measured for the first time. In Figs. 7 and 8, the ionization coefficient (a/N) and the electron attachment coefficient (g/N) are plotted separately over a relatively low range of E/N (150–400 Td). As shown in Fig. 8, the resultant values for the attachment coefficient g/N for C₂H₂ and C₂H₄ are shown.
coefficients scatter widely, presumably due to the presence of impurities and for also a statistical reason, and are generally much smaller than the corresponding ionization coefficients over this range of E/N.

The uncertainty of the values obtained for these coefficients was evaluated by the ratio of $\sigma/m$ ($\sigma$: the square of unbiased variance and $m$: the average) for the cases of drift distances in the ATS observations and for the iteration of current measurements in the SST method. The maximum and minimum values of $\sigma/m$ are shown in Table II, where $m$ and $\sigma$ are defined by

$$m = \frac{1}{n} \sum_{i=1}^{n} x_i$$  \hspace{1cm} (6)

and

$$\sigma = \sqrt{\frac{1}{n-1} \sum_{i=1}^{n} (x_i - m)^2}.$$  \hspace{1cm} (7)

Here, $x_i$ represents a measured value and $n$ is the number of values.

The reason why the ionization coefficients in the ring type molecules (C6H6 and C6H5CH3) are smaller than those of the chain type molecules as in Fig. 3 may be attributable to the energy transfer from the incident electron to the molecule. With regard to the presence of radical -OH in the molecules, we can see a major difference between C2H5OH and C2H6 in Fig. 3, where the ionization coefficient for C2H5OH is much smaller than that for C2H6 for E/N above 1800 Td. This may indicate that -OH acts as a less electron-attracting effect in the molecule. As a matter of fact, the ionization probability and the behavior of electron swarms should be discussed based on the electron collision cross sections. Unfortunately, to our knowledge, the cross section data for C2H6 and C6H5CH3 and for CH3OH and C2H5OH are insufficient to explain electron transport parameters in terms of them. However, one feature in Fig. 4 can be interpreted from a view point of the momentum transfer cross section is the smallest among the three molecules.

Although the swarm parameters sometimes differ depending on their definitions, particularly at the high E/N region in which the ionization process can occur (e.g., Ref. 4), the exact definition of some parameters and their measurement methods have been established. Our study presents experimental parameters for the mean arrival-time drift velocity ($W_m$) and the Townsend effective ionization coefficient ($\tilde{z}$) (including the separate result for the ionization coefficient $z$ and the attachment coefficient $\eta$ over a low range of E/N). The resultant data for the swarm parameters in the organic gases treated here can be used to predict quantitative properties of electron transport in the gases under the action of electric fields.

IV. CONCLUSION

In this study, the drift velocity ($W_m$) of electrons and the effective ionization coefficient ($z - \eta$) were measured in the organic gases (C2H2, C2H4, C2H6, CH3OH, C2H5OH, C6H6, and C6H5CH3) for relatively wide ranges of E/N. The drift velocity was obtained by using a double-shutter drift tube and the effective ionization coefficient was determined by the SST method. Some of them were measured for the first time. We recognize that these parameters have specific tendencies, depending on the molecular structure of the gases, such as chain-like or ring-type, and also by the presence or absence of -OH in the molecule. Some of the results were compared with those in the literature, and good agreement was observed in general. The results in this study on the swarm parameters would be beneficial in a variety of applications of these gases in weakly ionized plasma states.

ACKNOWLEDGMENTS

The authors are indebted to Y. Okazaki, H. Yokoya, M. Ishizuka, Y. Togashi, and S. Narita for helping us to perform the experiments. This work was supported by JSPS KAKENHI Grant No. 19310033.


TABLE II. Ratio of the standard deviation to the mean value.

<table>
<thead>
<tr>
<th>Gas</th>
<th>Number of positions</th>
<th>Number of current measurements</th>
<th>$\sigma/m$ (max.)</th>
<th>$\sigma/m$ (min.)</th>
<th>$\sigma/m$ (max.)</th>
<th>$\sigma/m$ (min.)</th>
</tr>
</thead>
<tbody>
<tr>
<td>C2H2</td>
<td>4</td>
<td>12–23</td>
<td>0.349</td>
<td>0.349</td>
<td>4.01</td>
<td>2500</td>
</tr>
<tr>
<td>C2H4</td>
<td>4</td>
<td>13–23</td>
<td>0.225</td>
<td>0.225</td>
<td>3.90</td>
<td>2500</td>
</tr>
<tr>
<td>C2H6</td>
<td>4</td>
<td>13–23</td>
<td>0.182</td>
<td>0.182</td>
<td>3.71</td>
<td>1800</td>
</tr>
<tr>
<td>CH3OH</td>
<td>4</td>
<td>14–26</td>
<td>0.290</td>
<td>0.290</td>
<td>5.27</td>
<td>1000</td>
</tr>
<tr>
<td>CH2OH</td>
<td>4</td>
<td>14–26</td>
<td>0.395</td>
<td>0.395</td>
<td>4.79</td>
<td>700</td>
</tr>
<tr>
<td>C2H5</td>
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<td>13–23</td>
<td>0.192</td>
<td>0.192</td>
<td>5.51</td>
<td>2300</td>
</tr>
<tr>
<td>C2H5CH3</td>
<td>4</td>
<td>12–19</td>
<td>0.300</td>
<td>0.300</td>
<td>4.16</td>
<td>1500</td>
</tr>
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