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A transparent electrochromic metal-insulator switching device with three-terminal transistor geometry

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Proton and hydroxyl ion play an essential role for tuning functionality of oxides because their electronic state can be controlled by modifying oxygen off-stoichiometry and/or protonation. Tungsten trioxide (WO_3), a well-known electrochromic (EC) material for smart window, is a wide bandgap insulator, whereas it becomes a metallic conductor H_xWO_3 by protonation. Although one can utilize electrochromism together with metal-insulator (MI) switching for one device, such EC-MI switching cannot be utilized in current EC devices because of their two-terminal structure with parallel-plate configuration. Here we demonstrate a transparent EC-MI switchable device with three-terminal TFT-type structure using amorphous (a-) WO_3 channel layer, which was fabricated on glass substrate at room temperature. We used water-infiltrated nano-porous glass, CAN (calcium aluminate with nanopores), as a liquid-leakage-free solid gate insulator. At virgin state, the device was fully transparent in the visible-light region. For positive gate voltage, the active channel became dark blue, and electrical resistivity of the a- WO_3 layer drastically decreased with protonation. For negative gate voltage, deprotonation occurred and the active channel returned to transparent insulator. Good cycleability of the present transparent EC-MI switching device would have potential for the development of advanced smart windows.

A transparent electrochromic metal-insulator (EC-MI) switching device that can be electrically switched from a colorless transparent insulator to a colored metallic conductor would be ideal for use in future energy-saving technologies, such as advanced smart-windows. In the OFF state such transparent EC-MI switching device fully transmits visible light, whereas in the ON state it does not transmit light. Furthermore, the device is switched from insulator to conductor at the same time, which can function as an ON/OFF power switch for other electronic devices.

Among many potential electrochromic materials, tungsten trioxide (WO_3)¹ shows the greatest suitability for the EC-MI switching devices. Stoichiometric WO_3 is a transparent insulator with a bandgap (E_g) of 2.6–3.0 eV² and has a defect perovskite-type structure with space group $P2_1/n$, in which A-sites in the ABO_3 lattices are vacant³. If the vacant A-sites become occupied by protons (H^+), i.e. the formation of tungsten bronze, it becomes an electrical conductor and opaque to visible light following the valence-state change of W ion from W^{6+} to W^{5+4} . Thus, the protonation/deprotonation of WO_3 is promising for the realization of simultaneous electrical switching between colorless/colored and insulating/conducting states, as shown in Fig. 1.

Various types of WO_3 -based EC devices have been actively developed for applications in energy-saving smart windows^{5,6}; however, their MI switching behavior has not been utilized with EC switching simultaneously, because of their two-terminal structure, which involves a parallel-plate electrode configuration. Electrostatic charge modulation using three-terminal thin-film transistors (TFTs)^{7,8} on a WO_3 thin film could be used to realize simultaneous EC-MI switching; however, it remains difficult to fully switch their coloring state, because of

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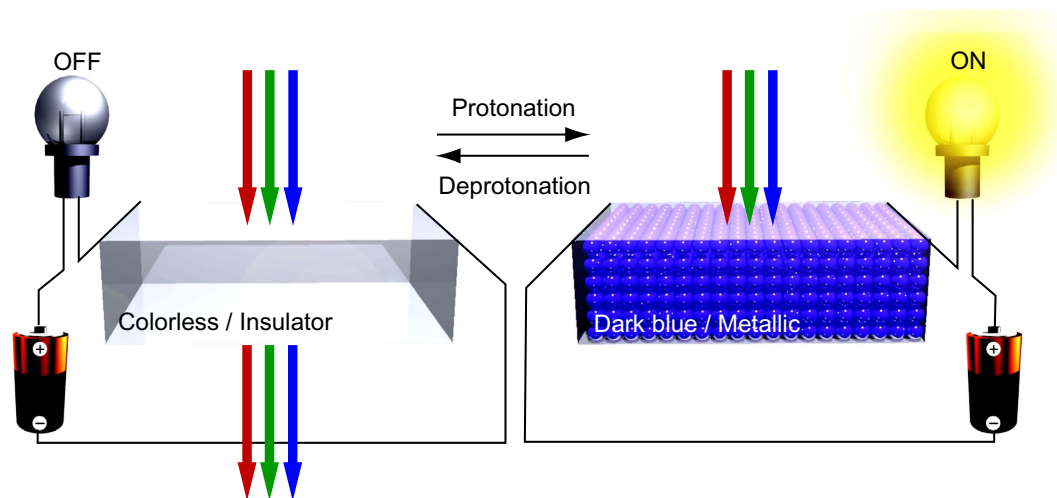


Figure 1. Concept of an electrochromic metal-insulator switching device. The device can be switched from a colorless transparent/insulator state to a dark blue/metallic state simultaneously by electrochemical protonation/deprotonation at RT in air. In the ON state, the visible light cannot be transmitted through the device, whereas it can be fully transmitted in the OFF state. Further, the device can function as an ON/OFF power switch for other electronic devices at the same time. Such a device would be useful for advanced smart window application.

limitations associated with the carrier-doping range and modifiable thickness⁹. Several researchers have developed liquid-electrolyte gated transistors^{10,11} and successfully demonstrated EC-MI switching of WO_3 . However, because these devices require liquid electrolytes, they are less practical for application, which depends on effective sealing.

Herein, we demonstrate a *liquid-leakage-free* transparent EC-MI switching device. Figure 2a schematically illustrates the device structure, which has a typical three-terminal TFT- geometry composed of an active channel, a gate insulator, and source-drain-gate electrodes. We used an amorphous (a-) WO_3 thin film as the active channel layer, because EC switching of a- WO_3 film prepared on glass substrate at room temperature (RT) has been previously reported¹². The gate insulator consists of an a- $12\text{CaO}\cdot 7\text{Al}_2\text{O}_3$ (a-C12A7) thin film with nanoporous structure (calcium aluminate with nanopore, CAN)^{13,14}. It should be noted that the film porosity of CAN film can be controlled by oxygen pressure (P_{O_2}) during thin-film deposition at RT (porous structure can be observed at $P_{\text{O}_2} > 1$ Pa), and the nanopores with average diameters of 10–20 nm connect with each other, when the porosity reaches ~30% of fully dense film, leading to the percolation conduction of water in the CAN film. Since C12A7 is a hygroscopic material, water vapor in air is automatically absorbed into the CAN film, like a solid sponge, via the capillary action in the interconnected nanopores. Therefore, water electrolysis can be used in the solid gate insulator. A NiO/ITO (indium tin oxide) bilayer film was used as the gate transparent electrode, and ITO thin films were used as the transparent source and drain electrodes.

The device with leakage-free water can be considered as the nanosized electrochemical cell with a nanogap parallel plate structure, which enables the high electric-field application for ion migration and effective protonation/deprotonation of the a- WO_3 layer (Fig. 2b). Thus, a gate voltage (V_g) application induces water electrolysis in the CAN film, and produced H^+ and OH^- ions move to protonate the a- WO_3 layer ($\text{WO}_3 + x\text{H}^+ + xe^- \rightarrow \text{H}_x\text{WO}_3$)¹⁵ and hydroxylate the NiO layer ($\text{NiO} + \text{OH}^- \rightarrow \text{NiOOH} + e^-$)¹⁶, respectively. The NiO counter layer is expected to work as the OH^- absorber, which maintains a better electrochemical balance and should improve the reversibility and reproducibility of device operation. Alternative positive and negative V_g applications induce the reversible protonation/deprotonation of a- WO_3 layer, switching it from a transparent insulator to a dark blue conductor. The present EC-MI switching device with the two combined functions of color changing as a display and electrical switching as a transistor in one device can be reversibly operated at RT without sealing; thus, it may be suitable for a wide-range application in future energy-saving technologies, such as advanced smart-windows.

Results

Device fabrication. The EC-MI switching device was fabricated on an alkaline-free glass substrate (Corning® EAGLE XG®) by using stencil masks. All the thin-film fabrication processes were conducted at RT using pulsed laser deposition (PLD) with KrF excimer laser ($\lambda = 248$ nm). Details of the device fabrication process are provided in experimental section. First, 20-nm-thick transparent conducting ITO films (resistivity, $\rho = 1.0$ m Ω cm at RT) were deposited as source and drain electrodes. An 80-nm-thick a- WO_3 channel film, a 300-nm-thick CAN gate insulator film^{13,14}, and a NiO (20 nm)/ITO (20 nm) gate electrode film were then deposited in turn. The channel size was 800 $\mu\text{m} \times$ 400 μm . The density of a- WO_3 film was 5.96 g cm⁻³, evaluated by grazing incidence X-ray reflectivity, which corresponds to 82% of ideal density of a WO_3 crystal (7.29 g cm⁻³)¹⁷. The AC conductivity of the CAN film was 3.7×10^{-8} S cm⁻¹ at RT, slightly less than the 5.6×10^{-8} S cm⁻¹ of ultrapure water¹⁸.

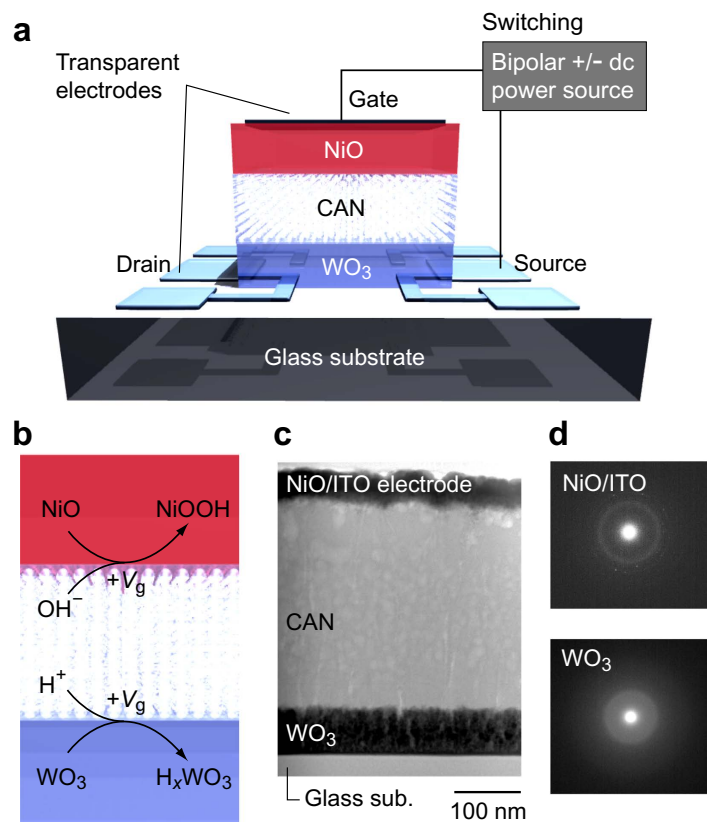


Figure 2. A transparent EC-MI switching device. (a) Schematic device structure of three-terminal TFT-type device, composed of a- WO_3 (80 nm), CAN (300 nm), and NiO (20 nm)/ITO (20 nm) layers. Transparent ITO thin films are used for all the electrodes. (b) Device operation mechanism. During the positive V_g application, protonation of a- WO_3 layer and hydroxylation of NiO layer occur simultaneously. Conversely, a- H_xWO_3 and NiOOH return to a- WO_3 and NiO during the negative V_g application. (c) Cross-sectional BF-STEM image of the device. Trilayer structure is clearly seen. Many light spots in the CAN layer indicate nanopores, which is fully occupied with water. (d) Selected area electron diffraction patterns of NiO/ITO layer (upper) and a- WO_3 layer (lower).

Figure 2c shows a bright-field scanning transmission electron microscopy (BF-STEM, in which heavier atoms appear darker) image of the cross-section of the resultant device, which reveals the multi-layer structure of ITO (20 nm)/NiO (20 nm)/CAN (300 nm)/a- WO_3 (80 nm) on a glass substrate. Numerous light spots with diameters of 10–20 nm are clearly seen in the CAN film, indicating the presence of high-density nanopores in the CAN film. Figure 2d summarizes the selected-area electron diffraction patterns of the NiO/ITO gate electrode layer (upper) and the a- WO_3 channel layer (bottom). A broad halo was observed for the a- WO_3 film, confirming the amorphous structure. Meanwhile, a ring diffraction pattern was seen for NiO/ITO film, originating from polycrystalline nature of NiO film, which was also confirmed by the grazing incidence X-ray diffraction measurements on each thin film.

Metal-insulator switching. We first evaluated the MI switching of the device by measuring sheet resistance (R_s) and thermopower (S) at RT, after applying and subsequently switching off V_g . Figure 3a plots R_s as a function of applied $\pm V_g$ at RT. First, a positive V_g up to +10 V was applied to protonate the a- WO_3 channel (left panel, each V_g application time of 20 s), and then negative V_g up to -10 V was applied to deprotonate the a- H_xWO_3 channel (right panel, each V_g application time of 10 s). The initial a- WO_3 channel was highly insulating (R_s was not in measurable range), but the reduction in R_s of more than six orders of magnitude was observed by applying positive V_g ; the R_s exponentially decreased from $17 \text{ M}\Omega \text{ sq.}^{-1}$ at +2 V and their saturation was observed at $\geq +8 \text{ V}$, where the R_s reached to $30 \Omega \text{ sq.}^{-1}$ at +10 V. It should be noted that the protonated a- H_xWO_3 channel was stable under ambient and vacuum conditions at RT after the + V_g application, confirming the non-volatility of device operation due to the electrochemical reaction. Subsequently, by applying negative V_g up to -10 V, R_s clearly recovered, reaching an insulating state ($6.8 \text{ M}\Omega \text{ sq.}^{-1}$).

We then measured the S for a- WO_3 channel protonated and deprotonated at each $\pm V_g$, because S is a good measure to evaluate the electronic-structure change resulting from carrier doping (protonation)^{19,20}. Figure 3b shows the relationship between S and $1/R_s$ for the device at RT. S -values were always negative, indicating that the channel layer is an n-type conductor. The $|S|$ linearly decreased from $47 \mu\text{V K}^{-1}$ to $11 \mu\text{V K}^{-1}$ with logarithmic increase in $1/R_s$ and the linear relation was reversibly observed with the application of $\pm V_g$, suggesting that

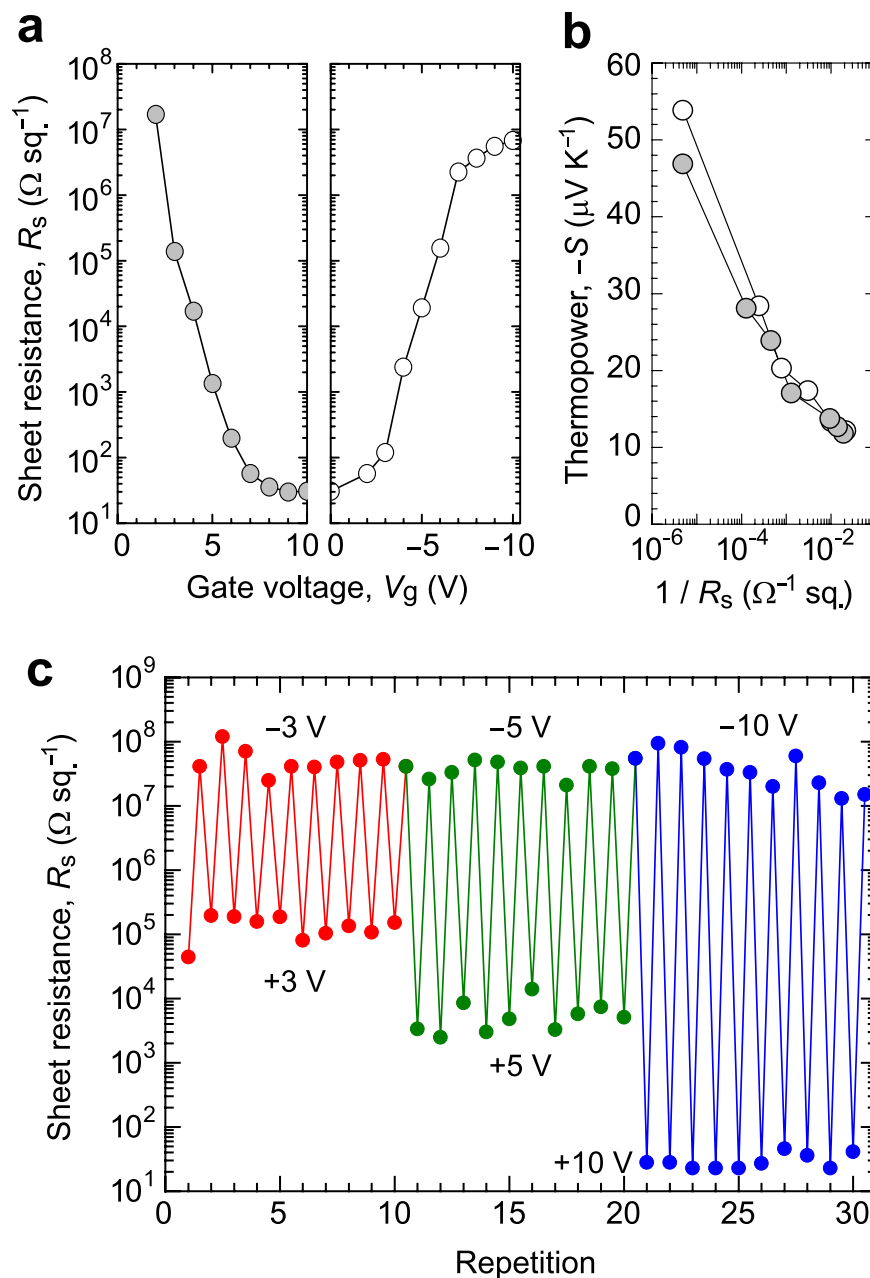


Figure 3. Metal-insulator switching of the transparent EC-MI switching device. (a) Sheet resistance (R_s) as a function of applied V_g for a- WO_3 layer. R_s values were measured after V_g application, where positive V_g up to +10 V was applied for protonation of a- WO_3 film (left panel, each V_g application time of 20 s), and then negative V_g up to -10 V was applied for deprotonation of a- H_xWO_3 film (right panel, each V_g application time of 10 s). (b) Thermopower (S) as a function of $1/R_s$ at RT. The linear relation between $-S$ and logarithmic $1/R_s$ was reversibly observed by $\pm V_g$ application. (c) Repetitive switching property by applying various $V_g = \pm 3$, ± 5 , and ± 10 V. The R_s modulation ratio can be controlled to be $\sim 10^3$ for ± 3 V, $\sim 10^4$ for ± 5 V, and $\sim 10^6$ for ± 10 V, respectively.

protonation of the a- WO_3 channel provides electrons to the conduction band, and the energy derivative of the electronic density of states (DOS) near the Fermi energy (E_F), $\left[\frac{\partial \text{DOS}(E)}{\partial E} \right]_{E=E_F}$, becomes moderate, resulting in the consequent reduction of $|S|$ -values.

Figure 3c shows the results of repeated R_s switching of the a- WO_3 channel at various V_g , e.g. ± 3 , ± 5 , and ± 10 V, where the V_g application time was 20 s for protonation ($+V_g$) and 10 s for deprotonation ($-V_g$). Clear cyclability of R_s switching was observed at each V_g , and the R_s modulation ratio was largely dependent on the applied V_g : the ON/OFF ratio was $\sim 10^3$ at ± 3 V, $\sim 10^4$ at ± 5 V, and $\sim 10^6$ at ± 10 V. Reversible and reproducible R_s switching with large ON/OFF ratios occurred rapidly on the second-time scale.

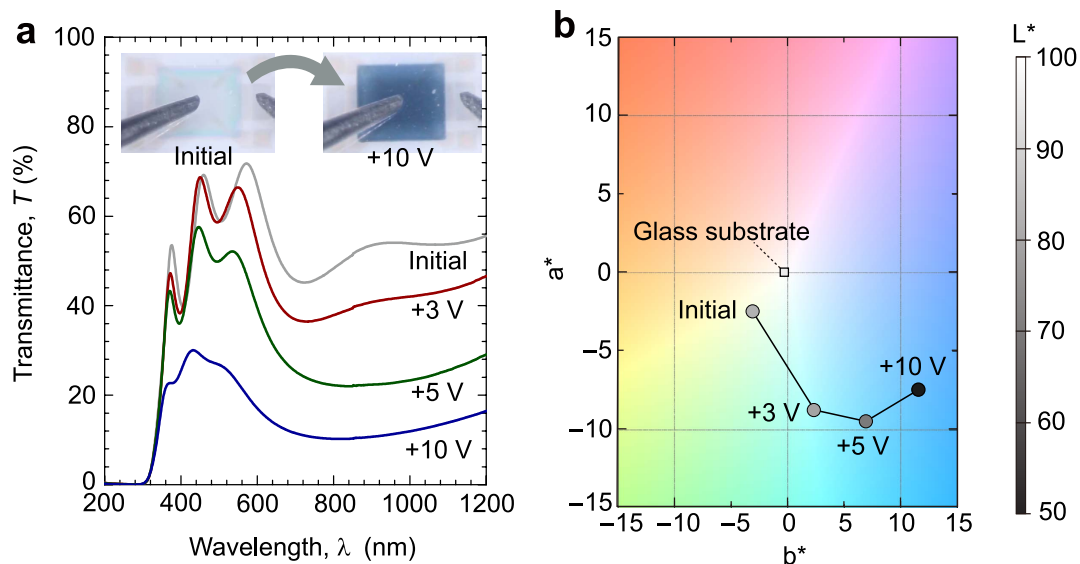


Figure 4. Electrochromic switching of the transparent EC-MI switching device. (a) Optical transmission spectra of the device. Significant decrease of the transmission is seen with increasing V_g . The inset shows the photographs of the device; after the application of $V_g = +10$ V, the device became opaque/dark blue (right), whereas the device was colorless transparent at the initial state (left). (b) $L^*a^*b^*$ color space plot of the device at each state in (a).

The device operation is considered to be largely dependent on the time-scaled processes consisted of ionic polarization in water, electric double layer formation, electrochemical reaction, and H^+ diffusion in a- WO_3 channel under the V_g applications. The rate-determining step among them for the protonation of WO_3 has been reported to be the surface reaction process^{21,22}. Therefore, the difference of the application time of $+V_g$ (protonation) and $-V_g$ (deprotonation) should originate from the activation barrier for the surface reaction, i.e. the in-diffusion and out-diffusion of H^+ transport have different interfacial resistances²³, which can be seen in the different gate current flowing in the device even at the same $\pm V_g$ (it will be shown later). In addition, the V_g -dependent ON/OFF ratio of the R_s modulation (Fig. 3c) suggests that the H^+ diffuses from the surface and the penetration depth along the out-of-plane direction can be controlled by the V_g , where the entire channel region is protonated by applying $V_g \geq 8$ V.

Electrochromic switching. We next evaluated the EC switching of the device. Figure 4a shows the optical transmission spectra of the device in the initial state (black line) and protonated states at +3 V (red line), +5 V (green line), and +10 V (blue line). The transmission (T) of the device was largely changed by the applied $+V_g$. The initial device was, to some extent, transparent in the visible light region. After protonation, T at $\lambda = 700$ nm reduced to 24% that of the initial state (transmission modulation of 35%). The inset shows the device picture at the initial and protonated states (+10 V); the color clearly changes from transparent colorless to opaque dark blue, as can be seen in $L^*a^*b^*$ color space (Fig. 4b), reflecting the optical modulation of the a- WO_3 channel by RT protonation. Thus, simultaneous EC-MI switching was realized in the device. It should be noted that there is no direct evidence for the hydroxylation of NiO to make NiOOH in this device, which should also affect the optical transmission spectra; NiO changes its color from transparent to deep brown by hydroxylation²⁴. However, considering that the no gas generation was confirmed in this device, which is indirect evidence to suggest the OH^- ion absorption in NiO counter layer, and there are many reports on the hydroxylation of NiO with aqueous solutions and their effect on electrochromic devices²⁴, NiO should play as OH^- absorption layer but further confirmation is necessary.

Discussion

To clarify the device operation mechanism, we investigated the relationship between the flowing current and R_s of the a- WO_3 channel under various V_g conditions. Figure 5a,b summarize the retention-time (t) dependence of the gate current (I_g) along with R_s for the a- WO_3 channel under $\pm V_g$ values of 3, 5, and 10 V, where the $+V_g$ was initially applied for protonation (Fig. 5a) and $-V_g$ was subsequently applied for deprotonation (Fig. 5b). I_g and R_s were measured during and after the V_g application, respectively. Upon $+V_g$ application (Fig. 5a), $+I_g$ increased with t for all applied $+V_g$ (the I_g at +10 V showed an irregular jump and exponentially increased and then exhibited the saturation at last), while R_s simultaneously decreased, indicating that electrochemical protonation of the a- WO_3 channel occurred. Meanwhile, upon $-V_g$ application (Fig. 5b), $-I_g$ decreased with t at each $-V_g$ while R_s simultaneously increased due to electrochemical deprotonation. This observation suggests the following device operation mechanism. The $+V_g$ application first accumulates charge carriers at film surface via electrostatic field effect with small $+I_g$, resulting in the formation of parallel-plate capacitor, i.e. the electric field starts to be well applied on the channel surface, and a further $+V_g$ application generates dissociated H^+ and OH^- ions and

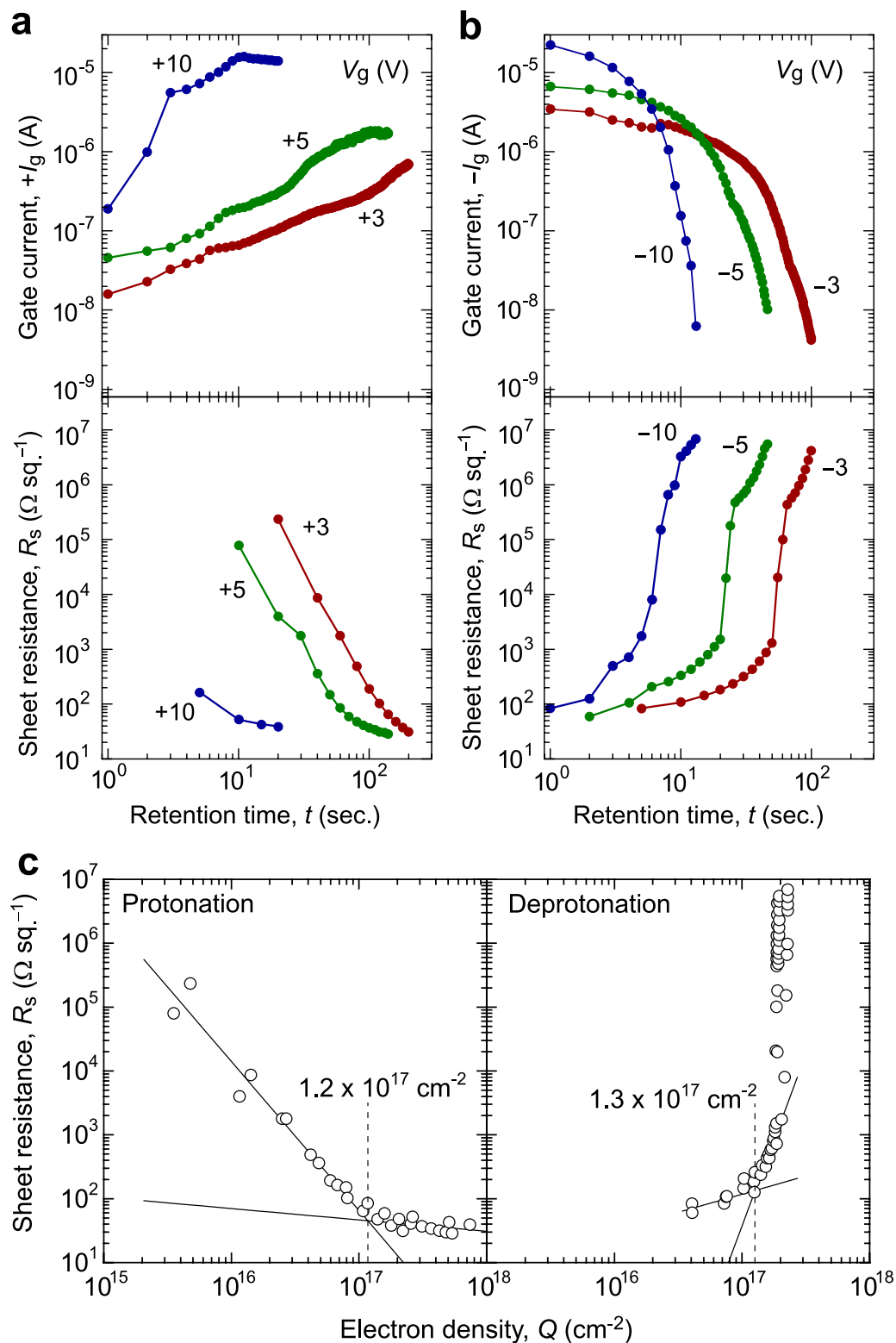


Figure 5. Switching mechanism of the transparent EC-MI switching device. (a,b) Retention time (t) dependence of gate current (I_g , upper) and sheet resistance (R_s , lower) under application of various $\pm V_g$'s from 3, 5, and 10 V, where (a) $+V_g$ was initially applied for protonation and (b) $-V_g$ was subsequently applied for deprotonation. (c) Electron-density (Q) dependence of R_s under application of the various V_g 's. The Q was calculated as the integrated value of the I_g - t plots in (a,b). The universal changes in R_s are presented by the black lines. The critical Q of protonation and deprotonation was 1.2 – $1.3 \times 10^{17} \text{ cm}^{-2}$, which corresponds well with the ideal Q ($1.3 \times 10^{17} \text{ cm}^{-2}$) required for the 100% protonation/deprotonation reaction of $\text{WO}_3 + \text{H}^+ + \text{e}^- \rightleftharpoons \text{HWO}_3$.

attracted them to each film surface, which should be primary origin of the significant increase of I_g at the initial stage. The H^+ and OH^- ions diffuse from the surface to bulk region of a- WO_3 film and NiO counter layer, where the improvement of conductivity of each layer exponentially increases the I_g .

We then compared the variation in R_s of the a- WO_3 channel with respect to applied electron density (Fig. 5c) estimated as $Q = \frac{C}{S \times q}$, where C is total coulomb amount calculated from the integral value of the I_g-t plots in Fig. 5a,b, S is the surface area of a- WO_3 channel, and q is elementary charge, respectively. For protonation ($+V_g$ application), R_s steeply decreased with increasing Q up to $1.2 \times 10^{17} \text{ cm}^{-2}$ ($\equiv 1.5 \times 10^{22} \text{ cm}^{-3}$) and then kept unchanged with further electron injection. In contrast, for deprotonation ($-V_g$ application), R_s increased moderately at Q up to $1.3 \times 10^{17} \text{ cm}^{-2}$ ($\equiv 1.6 \times 10^{22} \text{ cm}^{-3}$), after which it sharply switched to insulating state. The Q described here does not mean the carrier concentration in the film but electrochemically-used electron density ($\approx H^+$ concentration). Referring to the previous report on H^+ implantation²⁵, almost the same dose density ($\sim 10^{17} \text{ cm}^{-2}$) of H^+ was used for metallization of WO_3 , but the carrier concentration was measured to be $7.2 \times 10^{20} \text{ cm}^{-3}$, where the efficiency of carrier generation was only $\sim 60\%$; this situation should be equivalent for the present result. The change of R_s with Q seems to occur differently by protonation and deprotonation processes, which should come from the inhomogeneous H^+ distribution in a- H_xWO_3 channel along the out-of-plane direction and the H^+ -concentration dependent activity (chemical potential) of H^+ transport²⁶. The critical Q of MI switching of the a- WO_3 channel corresponds with the ideal Q value ($1.3 \times 10^{17} \text{ cm}^{-2}$) required for the 100% protonation/deprotonation of a- WO_3 , according to the following reaction: $WO_3 + H^+ + e^- \rightleftharpoons HWO_3$ (the Q is estimated by $Q = \frac{\rho \times T}{M} \times N_A$, where M is molar mass of WO_3 , ρ is the film density, T is the film thickness, and N_A is Avogadro constant, respectively).

The activation energy (E_a) of the electrical conductivity for the fully protonated a- H_xWO_3 channel that was determined by R_s-T measurements (Fig. S1) was $E_a = 4.3 \times 10^{-3} \text{ eV}$. This was an order of magnitude smaller than $E_a = 5.0 \times 10^{-2} \text{ eV}$ reported for the a- H_xWO_3 film ($x = 0.32$)²⁷. Since the E_a is a function of x due to the band filling, the much lower E_a in the a- WO_3 channel layer supports that it is effectively protonated by water-electrolysis with high electric-field application in the present TFT-type structure. This result also supports the above conclusion. In addition, R_s decreased and increased along the universal line under all values of V_g (Fig. 5c), indicating that all the provided electrons were used for electrochemical protonation/deprotonation of the a- WO_3 channel, obeying Faraday's laws of electrolysis, and that the device operation can be controlled by the current density.

The present device with leakage-free water can be reversibly switched from a colorless transparent insulator to a colored metallic conductor in a short amount of time (~ 10 s, see Video S1). In the ON state, visible light cannot transmit through the device, whereas it can fully transmit in the OFF state. The device can also function as an ON/OFF power switch for other electronic devices. The device is mainly composed of amorphous oxide films, which can be deposited at RT with no substrate heating required. This means there are no limitations on the type of substrate materials that can be used for the device. Moreover, the device can be operated without sealing thanks to the liquid-leakage-free CAN gate insulator. These features are suitable for the development of large-area devices and mass production; thus, the present device may find practical application in future energy-saving technologies, such as advanced smart windows.

In summary, we have demonstrated a liquid-leakage-free transparent EC-MI switching device, which has a three-terminal TFT-type structure consisting of transparent oxide thin films of a- WO_3 active channel, CAN gate insulator, NiO/ITO gate electrode, and ITO source-drain electrodes. At initial state, the device was fully transparent in the visible light region and the WO_3 channel was insulator. For $+V_g$ application, the device became dark-blue-colored state and the R_s of the WO_3 channel drastically decreased due to the protonation of a- WO_3 channel. For $-V_g$ application, deprotonation of a- WO_3 channel occurred and the device returned to colorless-transparent-insulator state. The reversible EC operation (transmission modulation of 35% at λ of 700 nm) and MI switching (R_s -modulation ratio $\sim 10^6$) were simultaneously demonstrated. The present transparent EC-MI switching device with leakage-free water is composed mostly of amorphous oxide thin films, which can be deposited at RT, and can be operated without sealing. Such low cost device will find the practical application for future energy saving technologies such as advanced smart-windows.

Methods

Device fabrication. The present liquid-leakage-free transparent EC-MI switching device (active channel area: $400 \mu\text{m}$ in width and $800 \mu\text{m}$ in length) was fabricated on an alkaline-free glass substrate (Corning® EAGLE XG®, substrate size: $10 \times 10 \times 0.7 \text{ mm}^3$) by pulsed laser deposition (PLD) using stencil masks. All the thin-film fabrication was conducted at RT, where a KrF excimer laser (wavelength of 248 nm, repetition rate of 10 Hz) was used to ablate ceramic target disks. First, 20-nm-thick metallic ITO films were deposited at O_2 pressure (P_{O_2}) of 4 Pa as the transparent source and drain electrodes. Then, a- WO_3 channel layer was deposited under P_{O_2} of 7 Pa, where the deposition rate of a- WO_3 film was 6 nm min^{-1} . 300-nm-thick CAN gate insulator was deposited under P_{O_2} of 5 Pa to make CAN film nanoporous structure^{13,14}. NiO (20 nm)/ITO (20 nm) bilayer film was deposited as the transparent gate electrode on the CAN film surface, where the nanopores in the CAN film is small enough to prevent the NiO/ITO film penetrate into the CAN film and reach a- WO_3 layer during the deposition.

Structural characterization. Crystallinity of the fabricated thin films were investigated by grazing incidence X-ray diffraction analyses (Cu $K\alpha_1$, ATX-G, Rigaku Co.), which revealed that all oxide layers were amorphous in nature expect for NiO polycrystalline film. Cross-sectional thin-film samples for TEM observations were prepared by focused-ion-beam (FIB) micro-sampling technique, in which the multilayer structure region of the TFTs was cutout and thinned by FIB (FB-2000A, HITACHI) to obtain samples for cross-sectional observation. The cross-sectional microstructure and electron diffraction pattern of the a- WO_3 devices were examined by high-resolution TEM and STEM (JEM-ARM200F, 200 kV, JEOL Ltd.).

Electrical and optical property measurements. I_g was measured between the gate and source electrodes during the V_g application using a source measurement unit (Keithley 2450). The electrical and optical properties were measured after switching the V_g off, because of the non-volatile device operation due to the electrochemical reaction. R_s were measured by the d.c. four-point probe method (van der Pauw configuration). For the retention time dependence of I_g and R_s (Fig. 5a,b), I_g was measured at 1 s intervals, and R_s was measured at each interval of 20 s for +3 V, 10 s for +5 V, 5 s for +10 V, 5 s for -3 V, 2 s for -5 V, 1 s for -10 V, respectively. Thermopower (S) was measured by giving a temperature difference (ΔT) of ~ 4 K in the film using two Peltier devices, where the actual temperatures of both sides of a- WO_3 channel layer were monitored by two tiny thermocouples. The thermo-electromotive force (ΔV) and ΔT were simultaneously measured, and the S -values were obtained from the linear slope of the ΔV - ΔT plots. Optical transmission spectra were measured by UV-Vis/NIR microscope with the light irradiation area of $100 \mu\text{m}$ in diameter (MSV-5200, JASCO). The relative humidity value, at which the device operation was tested, was $\sim 30\%$ at 25°C . Since the present device electrochemically operates in a cycled process between a- WO_3 cathodic layer ($\text{WO}_3 + x\text{H}^+ + xe^- \rightarrow \text{H}_x\text{WO}_3$) and NiO anodic layer ($\text{NiO} + \text{OH}^- \rightarrow \text{NiOOH} + e^-$), the water in CAN gate insulator should not be lost during the device operation and the CAN does not degrade after many cycles because of the no gas generation. However, it is considered that the humidity in air slightly affect the device operation, which should be tested in the near future.

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Author Contributions

T.K., T.O., M.H. and T.M. performed the device fabrication, measurements and data analyses. T.K. performed TEM/STEM analyses. All authors discussed the results and commented on the manuscript. H.O. planned and supervised the project.

Additional Information

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