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Catalytic Depolymerization of Chitin with Retention of $N$-Acetyl Group**

Mizuho Yabushita, Hirokazu Kobayashi,* Kyoichi Kuroki, Shogo Ito, and Atsushi Fukuoka* 

Abstract: Chitin, a polymer of $N$-acetylglucosamine units with $\beta$-1,4-glycosidic linkages, is the most abundant marine biomass. Chitin monomers containing $N$-acetyl groups are useful precursors to various fine chemicals and medicines. However, the selective conversion of robust chitin to $N$-acytlated monomers currently requires a large excess of acid or a long reaction time, which has limited its application. We demonstrate a fast catalytic transformation of chitin to monomers with retention of $N$-acetyl group by combining mechanocatalysis and homogeneous catalysis. Mechanical-force-assisted depolymerization of chitin with a catalytic amount of $\text{H}_2\text{SO}_4$ gave soluble short-chain oligomers. Subsequent hydrolysis of the ball-milled sample provided $N$-acytleyglucosamine in 53% yield, and methanolysis afforded 1-O-methyl-$N$-acytleyglucosamine in yields of up to 70%. Our process can greatly reduce the use of acid compared to the conventional process.

The transformation of biomass to value-added chemicals with particular functional groups is an attractive alternative to multistep functionalization of hydrocarbons from fossil fuels.[1] The most abundant terrestrial and marine biomasses are cellulose and chitin, respectively. Notably, chitin is a characteristic biopolymer consisting of nitrogen-containing monomer units, $N$-acytleyglucosamine ($\text{GlcNAC}$), with $\beta$-1,4-glycosidic linkages (Scheme 1). It is therefore an attractive renewable feedstock for organic nitrogen compounds.[2] However, unlike cellulose conversion, in which thermal and mechanocatalytic depolymerization can both yield monomers,[3] chitin requires selective depolymerization of glycosidic bond to prevent the loss of $N$-acytley group for the wide application. Hydrolysis of the glycosidic bonds of chitin gives $\text{GlcNAC}$, which is useful as a biologically active agent, an ingredient for cosmetics, and a raw material for $N$-containing organic compounds.[4] Meanwhile, methanolysis of chitin can afford 1-O-methyl-$N$-acytleyglucosamine (denoted as MeGlcNAC), which inhibits hemagglutination, and thus can be used for suppression of influenza and cancer.[5] MeGlcNAC is also applicable to the synthesis of biodegradable polyesters and polyamides,[6] organocatalysts,[7] ligands,[8] and gels.[9] Additionally, MeGlcNAC is a useful molecule for further transformations in glycoscience, since the hemiacetal group is protected by a methyl group.[10]

Despite a great deal of effort in various fields, efficient conversion of chitin to $N$-acytleyated monomers remains a challenge. In biochemistry, chitin is depolymerized by enzymes to form GlcNAC in yields of 64–77% under ambient conditions; however, it takes 10 days for the reaction to be completed due to low catalytic activity.[11] Conventional chemical/mechanical pretreatments slightly improved the reactivity of chitin by the amorphization, but a harsh pretreatment (supercritical water with converge milling) was necessary to accelerate the reaction.[12] In synthetic organic chemistry, $\text{HCl}$ provides GlcNAC in ca. 65% yield, but only when $\text{HCl}$ is used in large excess (i.e., the molar ratio of $\text{HCl}$ to GlcNAC units in chitin $= 100$) and at high concentrations of 15–36%.[13] Therefore, the serious corrosiveness of $\text{HCl}$ becomes an issue, and a huge amount of acidic waste is produced. Dilution of $\text{HCl}$ in the depolymerization reaction of chitin results in low selectivity for GlcNAC.[13] Moreover, the use of acid results in deacetylation of chitin,[13] but preservation of the $N$-acetyl group is essential for high physiological activity and wide application of the resulting chitin monomers.[14] The synthesis of MeGlcNAC has not been studied to the same extent as that of GlcNAC; in fact, its production from chitin has not yet been reported. Fischer glycosidation of GlcNAC has been necessary to obtain MeGlcNAC.[15] Herein, we report catalytic depolymerization of chitin to GlcNAC and MeGlcNAC (Scheme 1) in a process which reduces the use of acid by 99.8% and is completed within one day (including all handling).

The first key step in our reaction is mechanocatalytic conversion of chitin to soluble oligomers. It is known that mechanical force enhances chemical reactions under mild conditions.[16] Chitin was impregnated with a small amount of $\text{H}_2\text{SO}_4$ [substrate to catalyst ratio (S/C) = 8.1], and the resulting mixed solid was ball-milled at 500 rpm for 6 h to achieve mechanocatalytic hydrolysis (Scheme 1, product name: Chitin-$\text{H}_2\text{SO}_4$-BM). The water used for hydrolysis should be physisorbed water (1.5 wt%) in chitin. The process quantitatively converted chitin to soluble compounds such as GlcNAC (4.7% yield), ($\text{GlcNAC}_2$) (7.8% yield), ($\text{GlcNAC}_3$) (11% yield), ($\text{GlcNAC}_4$) (9.7% yield), and ($\text{GlcNAC}_5$) (8.6% yield), where ($\text{GlcNAC}_n$) denotes an oligomer consisting of $n$ GlcNAC units (Figure 1, Chitin-$\text{H}_2\text{SO}_4$-BM). Other products were longer and branched oligomers, as determined by liquid chromatography/mass spectroscopy (LC/MS) and nuclear magnetic resonance (NMR) spectroscopy (Figures S1 and S2). It is known that side chains improve the solubility of oligomers.[16] Chitin oligomers are useful as antitumor, immunoenhancing, and antimicrobial agents[17] as well as useful precursors of monomers (vide infra).

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Supporting information for this article is given via a link at the end of the document.
Two-step depolymerization of chitin to N-acetylated monomers, catalyzed by H$_2$SO$_4$. The first step is mechanocatalytic depolymerization of chitin to water-soluble oligomers, and the second is solvolysis of the oligomers to N-acetylated monomers.

In controlled experiments (Figure 1), pristine chitin contained no soluble compounds, while chitin that had been ball-milled in the absence of H$_2$SO$_4$ (Chitin-BM) afforded soluble compounds in 5.2% yield, with the chitin becoming amorphous (Figure S3), and H$_2$SO$_4$-impregnated chitin that had been aged at 298 K for 6 h (Chitin-H$_2$SO$_4$) had a soluble fraction of only 3.7%. Hence, both ball-milling and H$_2$SO$_4$ are necessary for depolymerization of chitin. To examine the synergy between mechanical force and H$_2$SO$_4$, chitin was ball-milled, impregnated with H$_2$SO$_4$, and aged at 298 K for 6 h (Chitin-BM-H$_2$SO$_4$); this sample had a soluble fraction of only 7.8%. Therefore, mechanical force was only effective for the depolymerization of chitin in the presence of H$_2$SO$_4$.

Notably, acetyl groups in the GlcNAc units were preserved after the mechanocatalytic reaction, since Chitin-H$_2$SO$_4$-BM consisted of N-acetylated oligomers and did not contain acetic acid. This result is specific to the mechanocatalysis, as the conventional hydrolysis using a small amount of acid results in significant removal of N-acetyl group. In the mechanocatalytic reaction (Scheme 2), we propose that amorphization of chitin first occurs by the ball-milling, which enhances the accessibility of acid to chitin due to lower steric hindrance. Then, mechanical force assists the cleavage of the connecting points of polymers, in which polymers are subjected to macro-scale power as tensile stress. Thus, the use of mechanical force resulted in selective activation of glycosidic bonds connecting the GlcNAc units rather than amide bonds in the acetamide.
These results clearly show that was obtained predominantly (24% yield) due to the harsh deacetylated product [i.e., 1 possible in the presence of H2SO4].

Recently, Yan et al. demonstrated that alcoholysis of chitin was the wide applications in various fields as described above.

Indeed, 94% of the MeGlcNAc was recovered after heat treatment at 463 K (Table S1, entry S2). It is notable that the product MeGlcNAc does not contain a hemiacetal group.

The second key point in our reaction is to overcome the stability issue by using methanolysis instead of hydrolysis, as the product MeGlcNAc does not contain a hemiacetal group. Indeed, 94% of the MeGlcNAc was recovered after heat treatment at 463 K (Table S1, entry S2). It is notable that the advantage of MeGlcNAc is not only the good stability but also the wide applications in various fields as described above.

Recently, Yan et al. demonstrated that alcoholysis of chitin was possible in the presence of H2SO4 (S/C = 0.9), although the deacetylated product [i.e., 1-O-(2-hydroxyethyl)-glucosamine] was obtained predominantly (24% yield) due to the harsh conditions required for direct depolymerization of the robust chitin. Therefore, we can expect selective synthesis of MeGlcNAc from the reactive chitin oligomers contained in Chitin-H2SO4-BM.

Chitin samples were subjected to methanolysis at 463 K to demonstrate selective synthesis of MeGlcNAc. The reaction of Chitin-H2SO4-BM provided MeGlcNAc in 68% overall yield, even when only a small amount of H2SO4 was used (S/C = 8.1) (entry 7). The ratio of α- to β-anomers in MeGlcNAc was 5.0, and the high stereoselectivity is favorable for further applications (the formation mechanism of the α-anomer is discussed in Figure S5). The turnover number (TON) of H2SO4 for MeGlcNAc production was 5.6, showing that the acid acted as a catalyst.

This result was in sharp contrast to those obtained from conventional systems, which require a large excess of HCl (TON < 0.01) to obtain the N-acetylated monomer. With regard to by-products, only a trace amount of methyl acetate (0.90%) was detected by gas chromatography. Almost no deacetylation occurred during methanolysis under our conditions. The yield of methanol-insoluble humins was 4.7 wt% based on the weight of chitin used. The yield of MeGlcNAc was increased to 70% under optimum conditions (S/C = 4.1) (entry 8). In contrast, methanolysis of other samples—pristine chitin, Chitin-H2SO4, Chitin-BM, and Chitin-BM-H2SO4—gave low yields of MeGlcNAc (<16%; entries 3–6). These results clearly show that mechanocatalytic depolymerization of chitin is different from that of cellulose; the methanolysis of cellulose gives a good yield of methylglucoside with no pretreatment.

In summary, we demonstrated the catalytic conversion of chitin to N-acetylated monomers, notably MeGlcNAc. The process consists of mechanocatalytic depolymerization of chitin to soluble oligomers, followed by solvolysis to monomers. Mechanical force was found to be essential to achieving selective catalytic depolymerization with retention of N-acetyl groups hanging from the units. This is the remarkable merit of mechanocatalysis in the selective depolymerization of chitin. The merit is different from that in the conversion of simple sugar polymers like cellulose that have only one functional group to be hydrolyzed (glycosidic bond).

<table>
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<tr>
<th>Entry</th>
<th>Sample</th>
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<td>MeGlcNAc</td>
</tr>
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<td>463</td>
<td>Water</td>
<td>32</td>
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<td>Chitin-H2SO4-BM</td>
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<td>Water</td>
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<tr>
<td>3</td>
<td>Chitin</td>
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<td>463</td>
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<td>463</td>
<td>Methanol</td>
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[a] Common reaction conditions: chitin weight 406 mg; solvent 40 mL; rapid heating-cooling conditions (time course is shown in Figure S4). [b] Based on moles of GlcNAc units. [c] No H2SO4 was impregnated on chitin.
group, demonstrating the remarkable merit of mechanochemical catalysis. The second thermocatalytic methanolysis converted oligomers to the N-acetylated monomer MeGlcNAc in 68–70% yield, owing to the improved reactivity of the substrate and the good stability of the product. In the overall process, the amount of acid used for the reaction was drastically reduced (99.8%) compared with the conventional process. Our method overcomes critical issues in conventional systems for chitin depolymerization, and opens the door to various applications of chitin-based compounds.

**Keywords:** biomass • catalysis • chitin • hydrolysis • mechanochemistry


Preserving N-acetyl groups: Chitin is depolymerized to N-acetylated monomers in good yields by combining mechanochemistry and homogeneous catalysis. The first mechanocatalytic hydrolysis of chitin provides water-soluble oligomers, which are more reactive than chitin. The subsequent solvolysis of the oligomers easily produces monomers without deacetylation.

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Page No. – Page No.

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