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Semi-open and closed system pyrolysis of Paleogene coal
for evaluating the timing of hydrocarbon gas expulsion

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Abstract

Hydrocarbon gas expulsion from coal source rocks takes place under conditions intermediate between an open and closed system. The compositional and isotopic changes of hydrocarbon gases generated from Paleogene coal with increasing maturation were investigated under semi-open and closed system laboratory pyrolysis to evaluate the maturity level at the time of hydrocarbon gas expulsion. Lower yields of saturated hydrocarbon gases and their higher $\delta^{13}$C values were observed in the semi-open system pyrolysis compared to those at the same maturity level in the closed system pyrolysis, since gaseous hydrocarbons depleted in $^{13}$C were removed during semi-open system pyrolysis. The $\delta^2$H values of hydrocarbon gases generated in the closed system pyrolysis were initially lower than those in the semi-open system pyrolysis, and then became markedly higher at the higher maturity level (vitrinite reflectance $[\text{VR},] > 1.2\%$), indicating significant hydrogenation of isotopically heavier inorganic hydrogen derived from water moisture in the coal matrix. The semi-open and closed system pyrolysis showed that the relationship between the isotopic composition of hydrocarbon gases expelled from coal and the maturity of coal depends on the state of openness or closedness of the hydrocarbon generation environment. Diagrams showing the relationship between the $\delta^{13}$C values of hydrocarbon gases expelled from
Paleogene coal and the VR, under conditions intermediate between semi-open and semi-closed systems were proposed for evaluating the timing of hydrocarbon gas expulsion in sedimentary basins.

Keywords:

Paleogene coal; carbon and hydrogen isotopic compositions; hydrocarbon gas expulsion, open and closed system pyrolysis
1. Introduction

The enrichment of $^{13}$C in thermogenic hydrocarbon gases, such as methane ($C_1$), ethane ($C_2$) and propane ($C_3$), with increasing maturation has been well studied in the laboratory and applied to the evaluation of the maturity level of natural gas (e.g., Clayton, 1991; Berner et al., 1992, 1995; Andresen et al., 1995; Berner and Faber, 1996; Michels et al., 2002). In the case of accumulated hydrocarbon gases in a reservoir, in many cases, the maturity level of the gases is related to the timing of gas expulsion from the source rocks if there is no thermal alteration after accumulation. The burial depth and temperature at the time of the expulsion in deep sedimentary basins are often recorded for natural gas according to its maturity level. The stable carbon and hydrogen isotopic composition ($\delta^{13}$C and $\delta^2$H) of gaseous hydrocarbons represents one of only a few indicators for evaluating the maturity level of natural gas. The $\delta^{13}$C and $\delta^2$H values of hydrocarbon gases in reservoir rocks can therefore indicate the timing of the expulsion of the gas from coal seams.

In previous studies, laboratory pyrolysis experiments on coal have been performed to elucidate the progressive change in the $\delta^{13}$C values of coal-derived hydrocarbon gases with increasing maturation. These pyrolysis experiments can be classified into open system pyrolysis (e.g., Berner et al., 1995; Cramer et al., 1998,
Gaschinski et al., 2001), anhydrous closed system pyrolysis (e.g., Liu et al., 2007; Duan et al., 2012; Li et al., 2013, 2016; Takahashi et al., 2014), and hydrous closed system pyrolysis (e.g., Chung and Sackett, 1979; Andresen et al., 1995; Kotarba and Lewan, 2004). Under open system pyrolysis, secondary cracking does not occur, or is very limited, since volatile pyrolyzates generated by the primary cracking of organic matter are immediately removed from the reaction system. In contrast, secondary and tertiary cracking of pyrolyzates proceed further under closed system conditions. If hydrocarbon generation in the source rocks proceeds under open system thermal maturation, the expulsion of hydrocarbon gas from the source rocks takes place efficiently. However, if hydrocarbon generation proceeds under closed system conditions, hydrocarbon gas expulsion does not take place due to the retention of thermogenic products in the source rocks. In the natural environment, it is more likely that the source rocks are restricted during their evolution by neither fully open nor fully closed systems, and the generation and expulsion of hydrocarbon gases in the sedimentary basin proceed under conditions intermediate between an open and closed system. If the changes in the \( \delta^{13}C \) and \( \delta^2H \) values of hydrocarbon gases with thermal maturation under both open and closed system conditions are known, the timing of hydrocarbon gas expulsion from the coal source rocks can be estimated more accurately.
based on the carbon and hydrogen isotopic compositions of natural hydrocarbon gas.

Open system pyrolysis is generally performed under inert gas flow. In the present study, semi-open system pyrolysis was used to estimate the progressive changes in the $\delta^{13}$C and $\delta^2$H values of hydrocarbon gases and the vitrinite reflectance ($VR_r$) of the source coal with thermal maturation. We performed the semi-open system pyrolysis according to Michels et al. (2002), as follows. The coal sample was pyrolyzed in the reaction chamber under closed system conditions, and the free gases generated from the coal were recovered after completion of the pyrolysis. The remaining coal sample was then re-pyrolyzed at a higher temperature than in the previous pyrolysis experiment. The experimental process was repeated until the final pyrolysis experiment at the highest temperature. This series of pyrolysis experiments is known as semi-open system pyrolysis. Experiments of semi-open and closed system pyrolysis of coal can reveal the influence of the openness or closedness of the reaction environment on the relationship between the $\delta^{13}$C and $\delta^2$H values of hydrocarbon gases released from the coal, as well as the maturity level of the source coal. Changes in the isotopic composition of hydrocarbon gases from Paleogene coal with increasing thermal maturation under semi-open and closed system conditions provide useful datasets for evaluating the timing of hydrocarbon gas expulsion from coal source rocks.
2. Materials and methods

2.1. Samples

Eocene coal-bearing strata are distributed widely from north to south in the Ishikari sedimentary basin in central Hokkaido. The middle to late Eocene Ishikari Group comprises the major stratigraphic sequence in the basin. The Ishikari Group consists of mainly non-marine to littoral sedimentary rocks with abundant coal seams, which have been exploited as the Ishikari coal field since the late 19th century. The Ishikari Group is divided into nine lithostratigraphic formations (Imai, 1924). A coal sample was taken from outcrops of one of the formations exposed at Yubari, the Middle Eocene Yubari Formation, which is composed of alternating beds of sandstone, mudstone, and coal seams deposited in fluvial and floodplain to estuarine environments (Takano and Waseda, 2003). The Eocene coals in the Ishikari basin are important source rocks of the Yufutsu oil and gas field (Waseda et al., 1998; Yessalina et al., 2006).

The Eocene Yubari coal is composed mainly of vitrinite, with small amounts of degradinite (detrovitorinite) macerals. The VRr, and organic carbon, hydrogen, total nitrogen and total sulfur contents, of the Eocene Yubari coal used in the present study are shown in Table 1. Takahashi et al. (2014) reported the changes in the concentrations
and isotopic compositions of gas components generated from the same Eocene Yubari coal source under anhydrous closed system pyrolysis. The elemental composition and VR$_r$ of artificially matured coals obtained by anhydrous closed system pyrolysis of Eocene Yubari coal are shown in Table 1. The bulk carbon isotopic composition of the Yubari coal sample was determined by SI Science Co. (Japan) to be $-24.0\%$.

2.2. Semi-open system pyrolysis of coal

Semi-open system pyrolysis experiments were conducted using a 30-mL reactor made of stainless steel 316 (Nitto Koatsu, Co., Ltd, Japan). The Yubari coal sample was crushed to fragments of 5–7 mm and dried at 60$^\circ$C for 24 h. Approximately 5 g of coal fragment sample was used in each pyrolysis experiment. The reactor was evacuated for several minutes and subsequently purged with He gas for 5 min. The reactor was then heated rapidly to the pyrolysis temperature, and the temperature was held constant for 72 h in a GC-14A oven (Shimadzu Co., Japan). The pyrolysis temperature was carefully monitored using a thermometer to an accuracy of within ±0.5$^\circ$C. The gaseous products generated from the individual coal samples, each pyrolyzed at 330, 340, 350, 370, and 390$^\circ$C for 72 h, were sequentially collected in aluminum bags as samples that had undergone semi-open system pyrolysis. Higher
pyrolysis temperature and shorter reaction time in the artificial maturation of coal are assumed to be kinetically comparable to low temperature and long-term heating in a sedimentary basin. The oven was cooled below 100ºC before gas sampling. The pyrolysis gas expelled from the coal fragments was collected using a 1.0-L aluminum bag (GL Science Co.), and the total volume of expelled gas was measured at room temperature using a 100-mL gastight syringe (GL Science Co.).

2.3. Gas analysis

The composition of the gases from the semi-open system was measured by gas chromatography (GC) using an Agilent 7890A instrument (USA) equipped with a pulsed discharge helium ionization detector (PDHID) and a micropacked column containing ShinCarbon-ST 80/100 (2.0 m × 1.0 mm i.d.; Shinwa Co., Japan). The oven temperature for the GC was initially programmed at 40ºC for 3 min, increasing to 300ºC at 15ºC/min, and then held at 300ºC for 15 min. Ultra-high purity He was used as the carrier gas. A constant amount of gas was introduced into the GC column using a 10-µL sampling loop. Compounds were identified by comparing the retention times with those of reference standards in a gas mixture containing CH₄, C₂H₄, C₂H₆, C₃H₆, C₃H₈, i-C₄H₁₀, n-C₄H₁₀, H₂, CO, and CO₂. The detailed analytical method, as well as the
method for identification of the gases, are described in Saito et al. (2012).

2.4. Isotope analysis

Stable carbon isotopic analysis was performed by GC-combustion-isotope ratio mass spectrometry (GC-C-IRMS, Agilent 6890A GC and Finnigan [Germany] MAT 252) using the same micropacked column (ShinCarbon-ST 80/100, 2.0 m × 1.0 mm i.d.; Shinwa Co.) and He as the carrier gas. The GC oven temperature was programmed to 40ºC for 3 min, then to 300ºC at 15ºC/min, and finally held at 300ºC for 20 min. The combustion was performed in a micro-volume ceramic tube with CuO, NiO, and Pt wires at 900ºC. The stable hydrogen isotopic composition was measured using GC-high-temperature conversion–isotope ratio mass spectrometry (GC-HTC-IRMS; Finnigan Delta V Advantage) using the same micropacked column and He as the carrier gas. The oven temperature program was the same as that for the GC-C-IRMS. The thermal conversion furnace between the GC and IRMS was maintained at 1,450ºC. Stable carbon and hydrogen isotope data were recorded as $\delta^{13}C$ and $\delta^2H$ values relative to the VPDB and VSMOW scales, respectively, according to Coplen (2011). The $\delta^{13}C$ and $\delta^2H$ values were determined based on a single-point calibration by comparison with the MZ3-01 CH$_4$ standard ($\delta^{13}C$ of -31.1‰; SI Science Co., Japan) and the Indiana CH$_4$
#1 standard ($\delta^{13}$C of -38.85‰ and $\delta^2$H of -160.8‰; Indiana University, USA), respectively. The $\delta^{13}$C value of the MZ3-01 CH$_4$ standard has been calibrated against NBS 19. The $\delta^{13}$C and $\delta^2$H values of the Indiana CH$_4$ #1 standard are normalized to NBS 19/L-SVEC and VSMOW/SLAP, respectively. The H$_3^+$ factor of the mass spectrometer was determined once daily using external reference gas injections and was consistently < 10 ppm. Selected gas samples were measured repeatedly for $\delta^{13}$C and $\delta^2$H three to five times. The results showed that the analytical precisions of the $\delta^{13}$C and $\delta^2$H values were within ±0.5‰ and ±10‰, respectively.

2.5. Measurement of vitrinite reflectance

The VR$_r$ was measured using a reflecting microscope (Eclipse LV100ND; Nikon, Ltd., Japan) equipped with a stabilized halogen light source and photonic multichannel analyzer (PMA12; Hamamatsu Photonics Co., Japan). The VR$_r$ value was measured for a spot diameter of 20 µm at a wavelength of 542.8 nm by comparison with the standard values of polished glasses with VR$_r$ values of 0.55, 0.79, 1.08, and 1.53%. The VR$_r$ value of the coal samples was determined based on a VR$_r$ dataset obtained from 40 to 50 spots. The accuracy of the VR$_r$ measurements was within ±0.03% for a VR$_r$ range of 0.3–2.0%.
3. Results and discussion

3.1. Changes in gas yields with maturation

The gaseous hydrocarbons and non-hydrocarbons generated from the Eocene Yubari coal during semi-open system pyrolysis were CH$_4$ to $n$-C$_4$H$_{10}$ (C$_1$ to C$_4$), C$_2$H$_4$, C$_3$H$_6$, $i$-C$_4$H$_{10}$, H$_2$, CO, and CO$_2$. Typical PDHID gas chromatograms for these gas components are shown in Fig. 1. The O$_2$ and N$_2$ in the chromatograms were contaminants from the air. C$_1$ was generally the most abundant hydrocarbon gas in the pyrolysis product. The relative abundance of C$_2$ to C$_3$ increased with increasing heating temperature, showing the progress of thermal cracking. The significant decrease in CO$_2$ with heating temperature was due to the rapid removal of oxygen-containing functional groups. The independent yields and cumulative yields (the sum of independent yields in each pyrolysis experiment) of the gas components in the series of semi-open system pyrolysis experiments are shown in Table 2. The VR$_r$ values for coal samples pyrolyzed in the semi-open system (Table 2) were obtained from the measured VR$_r$ values for coal samples pyrolyzed in the closed system (Table 1) under the same time-temperature conditions as those for the corresponding individual semi-open system pyrolysis. Taking into account the longer reaction time due to the multiple and sequential pyrolysis
processes, the VR\textsubscript{r} values in the semi-open system pyrolysis were slightly higher than those in the corresponding closed system pyrolysis. However, this difference was almost negligible in the context discussed below. The independent and differential yields (the difference in the yield of the gas component between two successive pyrolysis experiments) of gas components in the series of closed system pyrolysis experiments are shown in Table 3. Changes in the cumulative yield of the gas components of the semi-open system pyrolysis and in the independent yield of the gas components of the closed system pyrolysis with increasing VR\textsubscript{r} are shown in Fig. 2.

The yields of saturated hydrocarbon gases (C\textsubscript{1}, C\textsubscript{2} and C\textsubscript{3}) in the closed system pyrolysis were similar to the cumulative yields of hydrocarbon gases in the semi-open system pyrolysis at the lower maturity level (VR\textsubscript{r} < 1.2\%) (Fig. 2). However, they were clearly higher than the cumulative yields of hydrocarbon gases in the semi-open system pyrolysis at the higher stage of maturation (VR\textsubscript{r} > 1.2\%). This was due to the removal of gaseous hydrocarbons, which can be a source of C\textsubscript{1}, C\textsubscript{2} and C\textsubscript{3}, from the reaction chamber during the series of semi-open system pyrolysis. Unsaturated hydrocarbons, such as C\textsubscript{2}H\textsubscript{4} and C\textsubscript{3}H\textsubscript{6}, were comparatively minor compounds. These unsaturated hydrocarbons were more abundant in pyrolyzates from the semi-open system pyrolysis compared to those from the closed system pyrolysis. The abundance of unsaturated...
hydrocarbons was related to the degree of saturation by hydrogenation during the pyrolysis experiment. Lower abundance of unsaturated hydrocarbons in the pyrolyzates from the closed system pyrolysis suggested that the process of hydrogenation to alkenes proceeded more efficiently in the closed system pyrolysis. The abundance of H₂ in the reaction chamber of the closed system pyrolysis was higher than that of the semi-open system pyrolysis at the same maturation level, since H₂ was removed from the reaction chamber during the series of semi-open system pyrolysis. The lower concentration of unsaturated hydrocarbons in the closed system pyrolysis can be interpreted as being due to efficient hydrogenation, owing to the higher H₂ concentration in the reaction chamber of the closed system pyrolysis. However, the cumulative yields of CO₂ from the semi-open system pyrolysis increased significantly at the lower maturity level (VRₑ < 0.7%), while the yields of CO₂ from the closed system pyrolysis increased significantly at the relatively higher maturity level (VRₑ > 1.2%). The CO₂ in the thermogenic gas was derived from oxygen-containing functional groups, such as carbonyl and carboxyl groups. The fairly scattered plot for the yield of CO₂ from the closed system pyrolysis could be due to the adsorption of CO₂, on the surface of the coal sample and on the inner surface of the reaction chamber, during gas sampling collection into the aluminum bag. The CO₂ yield of the first semi-open system pyrolysis at 330°C for 72 h was
significantly higher than that of the closed system pyrolysis under the same conditions, indicating the uncertainty in determining the CO₂ yield. Lewan (1997) reported that a trace amount of water vapor in the anhydrous pyrolysis experiments could oxidize carbonyl groups to form carboxyl and hydrogen. Decarboxylation is one of the major sources of CO₂. The higher CO₂ yield in the closed system pyrolysis at the higher maturity level was attributable to a higher water vapor concentration in the closed system pyrolysis, since water vapor was removed from the reaction chamber during the semi-open system pyrolysis.

3.2. Changes in carbon isotopic composition with maturation

In the semi-open system pyrolysis, the volatile products generated from the individual coal samples pyrolyzed at 330, 340, 350, 370, and 390°C for 72 h were collected sequentially. The mean VR_r value marking the maturity level of the hydrocarbon gases from the independent semi-open system pyrolysis was the logarithmic mean determined using the two VR_r values of successive semi-open system pyrolysis. The δ¹³C values of the hydrocarbon gases from the semi-open system can be assumed to represent those of the hydrocarbon gases expelled from the coal source rocks under the open system at the mean VR_r value. The mean VR_r and δ¹³C values of
the hydrocarbon gases from the semi-open system are shown in Table 4.

Hydrocarbon gas expulsion does not take place under fully closed systems. The expulsion of hydrocarbon gases is thought to take place under conditions close to an open system in the early stage of diagenesis because of the comparatively high porosity and permeability of the coalbed, and significant water expulsion with increasing compaction. Hydrocarbon gas expulsion proceeds gradually under conditions closer to a closed system with the lower porosity and permeability of the coalbed in the later stages of diagenesis. Hydrocarbon gas expelled from the coal under the semi-closed system is therefore a mixed gas of newly-generated and pre-generated hydrocarbon gases, which is poor in the hydrocarbon gases generated in the early stage of diagenesis. The differential hydrocarbon gases of the successive closed system pyrolysis in the present study (Table 3) were composed of a hydrocarbon gas mixture without the gas generated in the early stage of pyrolysis, similar to the hydrocarbon gas expelled from the coal under the semi-closed system. The apparent $\delta^{13}$C value of the differential hydrocarbon gas in the series of closed system pyrolysis experiments was estimated based on the extrapolated difference in the $\delta^{13}$C value and volume of hydrocarbon gas between two successive pyrolysis experiments (Table 5). The measured VR$_r$ corresponding to each closed system pyrolysis was used as the maturity level of the hydrocarbon gas expelled.
from the coal under the semi-closed system (Table 5). The changes in the $\delta^{13}C$ values of C$_1$, C$_2$, and C$_3$ ($\delta^{13}C_1$, $\delta^{13}C_2$, and $\delta^{13}C_3$), which were assumed to represent those expelled from Eocene Yubari coal under the semi-open and semi-closed systems with increasing VR$_r$, are shown in Fig. 3.

The $\delta^{13}C_1$ value was generally lower than the $\delta^{13}C_2$ and $\delta^{13}C_3$ values. The observed relationship of $\delta^{13}C_1 < \delta^{13}C_2 < \delta^{13}C_3$ is characteristic for thermogenic hydrocarbon gases (e.g., McCarty and Felbeck, 1986; Chung et al., 1988; Liguo et al., 2009). The $\delta^{13}C_1$, $\delta^{13}C_2$, and $\delta^{13}C_3$ values of the expelled gas under the semi-open system increased with increasing maturation, although a decrease in the $\delta^{13}C_1$ value in the initial stage of coal maturation (VR$_r < 0.9\%$) was also observed (Fig. 3). The $\delta^{3}C$ value of CH$_4$ generated in the initial stage of the pyrolysis experiments is often high (e.g., Berner and Faber, 1996; Cramer et al., 1998; Takahashi et al., 2014). According to Chung and Sackett (1980) and Smith et al. (1985), the initial thermogenic CH$_4$ is thought to be generated from a pool of isotopically heavy methyl groups bound to relatively labile C–O or C–S linkages. $\delta^{3}C_1$ then becomes lighter in the initial stage of pyrolysis and subsequently becomes heavier again with increasing pyrolysis temperature in the higher maturation stage (VR$_r > 0.9\%$). Enrichment of $^{13}C$ in natural gas with increasing temperature has been well documented (e.g., Clayton, 1991;
Andresen et al., 1995; Cramer et al., 1998; Tang et al., 2000). The increase in $\delta^{13}C_1$ in the higher maturation stage is due to the mixing of CH$_4$ generated from isotopically heavier and more tightly bound aliphatic structures in kerogen and bitumen. A comparison of the $\delta^{13}C$ values of hydrocarbon gases expelled under the semi-open and semi-closed systems at the same VR$_r$ showed that the former were higher than the latter at the maturity level of VR$_r > 0.7\%$ (Fig. 3). The difference in the $\delta^{13}C$ value between the semi-open and semi-closed systems increased with increasing maturation. Hydrocarbon gases rich in $^{12}C$ pre-generated under the closed system were mixed with the newly-generated ones to form the expelled gas rich in $^{12}C$, while in the case of the semi-open system, pre-generated hydrocarbon gases rich in $^{12}C$ were removed from the semi-open system to form the expelled gas composed mainly of newly-generated hydrocarbon gases depleted in $^{12}C$. The higher $\delta^{13}C$ values of hydrocarbon gases expelled from the semi-open system were due to the removal of pre-generated hydrocarbon gases rich in $^{12}C$ from the reaction system.

3.3. Changes in hydrogen isotopic composition with maturation

Tables 4 and 5 show the $\delta^2H$ values of C$_1$, C$_2$ and C$_3$ ($\delta^2H_{C1}$, $\delta^2H_{C2}$ and $\delta^2H_{C3}$) expelled from the Eocene Yubari coal under the semi-open and semi-closed systems,
respectively. The $\delta^2\text{H}$ values were measured in the hydrocarbon gases obtained from each pyrolysis experiment for the semi-open system. In the semi-closed system, the $\delta^2\text{H}$ values of the hydrocarbon gases were the apparent values calculated for the differential hydrocarbon gases, as described above for the apparent $\delta^1\text{C}$ values of the differential hydrocarbon gases. The $\delta^2\text{H}_{C1}$, $\delta^2\text{H}_{C2}$ and $\delta^2\text{H}_{C3}$ values tended to increase with increasing $\text{VR}_r$, although they decreased, or did not change significantly, in the initial stage of maturation (Fig. 4).

The $\delta^2\text{H}$ values of thermogenic hydrocarbon gases can be controlled by the hydrogenation of radicals and double bonds. The $\delta^2\text{H}$ value of hydrogen released by cleavage of the C–H bond increases with increasing thermal maturation, since the dissociation energy of the C–$^2\text{H}$ bond is higher than that of the C–H bond. Therefore, the $\delta^2\text{H}$ values of thermogenic hydrogen increase with increasing thermal maturation. However, the $\delta^2\text{H}$ values of the hydrocarbon gases in both the semi-open and semi-closed systems did not change significantly in the early stage of maturation ($\text{VR}_r < 1.0\%$), indicating that hydrogenation did not affect the $\delta^2\text{H}$ value of the hydrocarbon gases in this stage. The $\delta^2\text{H}$ value of the hydrocarbon gases subsequently increased with increasing maturation, suggesting that hydrogenation of isotopically heavier hydrogen was occurring at the higher maturation stage. The rate of increase in the $\delta^2\text{H}$ value of the
hydrocarbon gases in the semi-closed system was remarkably high compared to that in the semi-open system. The different rate of increase in the $\delta^2$H value was attributable to the different concentrations of hydrogen between the semi-open and semi-closed systems. The hydrogenation of radicals and double bonds proceeded more efficiently in the semi-closed system pyrolysis than in the semi-open system pyrolysis, since the hydrogen concentration in the reaction system of the semi-closed system pyrolysis was always greater than that of the semi-open system. This is evidenced by lower concentrations of unsaturated C$_2$H$_4$ and C$_3$H$_6$ in the closed system pyrolysis compared to those in the semi-open system pyrolysis (Fig. 2).

Another possible reason for the higher $\delta^2$H value of the hydrocarbon gases in the semi-closed system is that the hydrogen in the reaction chamber of the closed system pyrolysis was partly composed of isotopically heavier hydrogen derived from water moisture in the coal matrix. Hydrogen derived from water can contribute directly to the hydrogenation of radicals and double bonds simultaneous with hydrocarbon generation (e.g., Seewald, 1994; Lewan, 1997; Schimmelmann et al., 1999, 2001). The $\delta^2$H value of inorganic hydrogen (water-derived hydrogen) is generally higher than that of organic hydrogen. Water vapor was removed from the reaction chamber with gas samples during the sampling stage of the semi-open system pyrolysis, while isotopically
heavier hydrogen derived from water moisture remained in the reaction chamber. The remarkable increase in the $\delta^2$H values of the hydrocarbon gases at the later stage of maturation ($VR_r > 1.2\%$) in the semi-closed system could be due to the significant incorporation of hydrogen derived from water moisture.

The relationship of $\delta^2$H$_{C1} <\delta^2$H$_{C2} <\delta^2$H$_{C3}$ was observed, similar to the relationship among the $\delta^{13}$C$_1$, $\delta^{13}$C$_2$, and $\delta^{13}$C$_3$ values (Figs. 3 and 4). The dissociation energy of the C–C bond with $^2$H is larger than that without $^2$H (e.g., Schoell, 1980). The lowest $\delta^2$H$_{C1}$ value is thought to be because the C–C bond without $^2$H is more easily broken compared to that with $^2$H. Ni et al. (2011) proposed a kinetic model of hydrogen isotope fractionation during the generation of hydrocarbon gases based on quantum chemistry calculations for $n$-octane homolytic cracking. This kinetic model could be applicable to evaluation of the maturity level of natural gas. However, hydrogen exchange between interstitial water and sedimentary organic matter in the early stage of diagenesis has also been suggested as a possible mechanism (Schimmelmann et al., 1999, 2001; Lis et al., 2006; Kikuchi et al., 2010). The $\delta^2$H of higher land plants is also related to the climate humidity or aridity (e.g., Tipple and Pagani, 2010; Duan et al., 2011b; Yang et al., 2011). Estimation of the kerogen type of source rocks, and the maturity level of hydrocarbon gases at the time of hydrocarbon gas expulsion from the
source rocks based on the hydrogen isotopic composition, requires caution.

3.4. Evaluation of the timing of gas expulsion from coal

Oil and gas fields distributed in Southeast Asia are frequently associated with Cenozoic coal fields, where coal and coaly strata are important source rocks of oil and gas (e.g., Hoffmann et al., 1984; Waseda and Nishita, 1998; Hossain et al., 2014). Cenozoic coals have the potential to generate oil, since they are relatively rich in hydrogen and aliphatic structures compared with older Carboniferous coals (e.g., Aihara, 1979; Takahashi and Aihara, 1989; MacGregor, 1994; Petersen and Nytoft, 2006; Petersen et al., 2009). The maturity level of hydrocarbon gases in reservoirs generally indicates the maturity level at the time of their expulsion from the source rocks. Therefore, the relationship between the $\delta^{13}$C values of hydrocarbon gases and $VR_r$ can be applicable to evaluation of the timing of hydrocarbon gas expulsion from the coaly source rocks. Changes in the $\delta^{13}$C values of hydrocarbon gases from older pre-Cenozoic coals (Type III kerogen) with increasing maturation have been comparatively well studied in the laboratory (e.g., Arneth and Matzigkeit, 1986; Liu et al., 2007; Duan et al., 2011a; Li et al., 2013, 2016). However, the relationship between the $\delta^{13}$C values of $C_1$, $C_2$ and $C_3$ expelled from hydrogen-rich Cenozoic coal and the $VR_r$ of the source coal is
still poorly understood. In addition, to date, laboratory experiments relating to changes in the $\delta^{13}$C values of C$_1$, C$_2$ and C$_3$ with maturation have not examined the conditions under which the expulsion of hydrocarbon gases from the coal occurs. The relationship between the $\delta^{13}$C values of the hydrocarbon gases and VR$_r$, which is often applied in oil and gas exploration, is based on open system pyrolysis (e.g., Berner et al., 1995; Berner and Faber, 1996; Cramer et al., 2001). However, it is reasonable to assume that hydrocarbon gas expulsion in sedimentary basins occurs under conditions intermediate between those of open and closed systems.

Figure 5 shows a diagram of the relationship between the $\delta^{13}$C values of C$_1$, C$_2$ and C$_3$ expelled from Cenozoic coal and the VR$_r$ of the source coal at the time of gas expulsion, based on the experimental results (Fig. 3) of the semi-open and closed system pyrolysis in this study. In the sedimentary basin, the expulsion of hydrocarbon gases in the early stage of diagenesis is thought to take place under conditions close to an open system (semi-open system). The expulsion of hydrocarbon gas proceeds gradually under conditions closer to a closed system (semi-closed system), with decreasing porosity and permeability of the coalbed in the later stage of diagenesis. Therefore, in the major stage of gas expulsion from coal seams, newly generated hydrocarbon gases are mixed with pre-generated coalbed gas, and some of the mixed
gas is expelled from the coal seams under the semi-closed system. The $\delta^{13}$C value of the hydrocarbon gases expelled from the coal under the semi-open system was generally higher than that in the gases expelled under the semi-closed system at the same maturity level (Fig. 3), since the hydrocarbon gases expelled under the semi-open system did not contain pre-generated hydrocarbon gases depleted in $^{13}$C. The expulsion of hydrocarbons from the coal source rocks in sedimentary basins generally proceeds under conditions intermediate between semi-open and closed system. Therefore, the evolutionary path of the $\delta^{13}$C values of hydrocarbon gases expelled from the coal source rocks should be located between the paths of evolution under semi-open and semi-closed systems shown in Fig. 5. Natural gas from the Yufutsu oil and gas field in Hokkaido is derived from Paleogene coals (Waseda and Nishita, 1998). Waseda et al. (2002) reported $\delta^{13}$C values of hydrocarbon gases of $\delta^{13}$C$_1$ = -35.3‰, $\delta^{13}$C$_2$ = -25.6‰, and $\delta^{13}$C$_3$ = -23.6‰. These values are plotted along evolutionary paths of $\delta^{13}$C in Fig. 5, and were consistent with the results of our laboratory pyrolysis experiments. The carbon isotopic composition of natural gas from the Yufutsu oil and gas field lies close to the evolutionary paths of the $\delta^{13}$C values of the hydrocarbon gases expelled under the semi-open system, suggesting that hydrocarbon gas expulsion from the coal source rocks in the sedimentary basin is occurring under conditions close to a semi-open
In addition, the isotopic compositions of hydrocarbon gases reflect post-genetic phenomena, such as expulsion, migration, biodegradation, and mixing. The carbon isotopic fractionation that occurs during expulsion of CH₄ in coals and total organic carbon (TOC)-rich shales can affect the isotopic fractionation of CH₄ (e.g., Prinzhofer and Huc, 1995; Prinzhofer and Pernaton, 1997; Zhang and Krooss, 2001; Takahashi et al., 2014), although the carbon isotopic fractionation of C₂ and C₃ during their expulsion from coal has been reported as almost negligible (Takahashi et al., 2014). The δ¹³C value of C₁ in the reservoir decreases due to the biodegradation of oil and the mixing of microbial C₁, which is significantly depleted in ¹³C. Although the δ¹³C value of biogenic C₂ is also much lower than that of thermogenic C₂ (Mattavelli et al., 1992; Waseda and Didyk, 1995), the amount of biogenic C₂ is very limited (Oremland et al., 1988). Evaluation of the maturity level of natural gas based only on the δ¹³C value of C₁ requires caution. The δ³⁴C value of the thermogenic gas is also related to the δ³⁴C value of the source organic matter. The δ³⁴C value of the bulk Eocene Yubari coal used for our pyrolysis experiments was ~24‰ (VRᵣ = 0.46%) (Table 1). The δ¹³C value of the source organic matter should be taken into account at the time of application of the relationship between the δ¹³C values of C₁, C₂ and C₃, and VRᵣ shown in Fig. 5. The diagrams
shown in Fig. 5 should be verified by empirical studies. However, the conceptual view
in Fig. 5 is useful for assessment of the maturity of natural gas, and is applicable to the
evaluation of the timing of gas expulsion from coal in the sedimentary basin.

4. Conclusions

In this study, we investigated the compositional and isotopic changes of
hydrocarbon gases generated from Paleogene coal with increasing maturation under
semi-open and closed system laboratory pyrolysis. The different experimental
conditions, i.e., semi-open or closed system, were well represented in the compositional
and isotopic changes of the hydrocarbon gases with thermal maturation. The increases
in the $\delta^{13}C$ values of C$_1$, C$_2$ and C$_3$ under the semi-open system pyrolysis were greater
than those under the closed system pyrolysis, since volatile hydrocarbon gases depleted
in $^{13}C$ were intermittently removed from the reaction chamber during the semi-open
system pyrolysis. The relationship between the $\delta^{13}C$ values of the hydrocarbon gases
expelled from the coal and the VR$_r$ of the source coal depended on the state of openness
or closedness of the reaction environment. Therefore, evaluation of the maturity level of
hydrocarbon gases, based on their isotopic composition, should take into consideration
the state of openness or closedness of the coal source rocks during burial diagenesis.
Diagrams showing the relationship between the δ¹³C values of C₁, C₂ and C₃ and VRᵣ of the source coal were drawn based on the experimental results from the semi-open and closed system pyrolysis. The δ¹³C values of C₁, C₂ and C₃ with increasing VRᵣ should fall between the two paths of evolution of δ¹³C versus VRᵣ obtained by semi-open and closed system pyrolysis, since the expulsion of hydrocarbon gas from the coal takes place under conditions intermediate between a semi-open and closed system. Interpretation of the diagrams is useful for evaluating the maturity level of hydrocarbon gases and the timing of their expulsion from the coal source rocks in sedimentary basins.

The δ²H values of hydrocarbon gases generated under the closed system pyrolysis increased significantly at the higher maturity level (VRᵣ > 1.2%). We suggest that this was due to significant hydrogenation of isotopically heavier inorganic hydrogen derived from water moisture in the coal matrix. The δ²H values of thermogenic hydrocarbons are also controlled by the paleohydrology of vegetation and hydrogen exchange with pore water during maturation in the sedimentary basin. The evaluation of the maturity level of hydrocarbon gases based on their hydrogen isotopic composition demands caution.
Acknowledgments

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References


Mattavelli, L., Ricchito, T., Martinenghi, C., 1992: Deep isotopic light methane in


Takahashi, K.U., Suzuki, N., Saito, H., 2014. Compositional and isotopic changes in


Waseda, A., Nishita, H., 1998. Geochemical characteristics of terrigenous- and marine-

Waseda, A., Iwano, H., and Takeda, N., 2002. Geochemical study on origin and
maturity of natural gases. Journal of The Japanese Association for Petroleum
Technology 67, 3-15. (in Japanese with English abstract)

Yang, H., Liu, W., Leng, Q., Hren, M.T., Pagani, M., 2011. Variation in n-alkane $\delta^D$
values from terrestrial plants at high latitude: Implications for Paleoclimate

Paleogene crude oils from the Yufutsu oil- and gasfield and offshore wildcats, Japan.

fractionation of methane during gas migration by diffusion through sedimentary
rocks at elevated temperature and pressure. Geochimica et Cosmochimica Acta 65
2723–2742.
Table 1. Vitrinite reflectance, elemental compositions, and H/C and O/C atomic ratios of Eocene Yubari coal and artificially matured coal (AMC #1 to #5) pyrolyzed at 330–390°C for 72 h under an anhydrous closed system.

Table 2. Independent and cumulative yields of gases generated from Eocene Yubari coal pyrolyzed under a semi-open system.

Table 3. Yields of gases generated from Eocene Yubari coal pyrolyzed under an anhydrous closed system.

Table 4. Stable carbon and hydrogen isotopic compositions of gases generated from Eocene Yubari coal pyrolyzed under a semi-open system.

Table 5. Stable carbon and hydrogen isotopic compositions of gases generated from Eocene Yubari coal pyrolyzed under a closed system. The apparent isotopic composition for the increased differential gas fraction was determined from the
mass balance of the gases.

[Figure Captions]

Fig. 1. Typical pulsed discharge helium ionization detector (PDHID) chromatograms of gas components generated from Eocene Yubari coal samples pyrolyzed at (a) 330°C, (b) 340°C, and (c) 370°C for 72 h. VR_r is the vitrinite reflectance of pyrolyzed coal.

Fig. 2. Changes in the yield of gas components with increasing VR_r under the semi-open and closed system pyrolysis of Eocene Yubari coal. The cumulative yield of the gas components in the semi-open system pyrolysis and the independent yield of gas components in the closed system pyrolysis are plotted.

Fig. 3. Changes in the δ^{13}C values of methane, ethane, and propane expelled from Eocene Yubari coal with increasing VR_r under semi-open and semi-closed systems. The δ^{13}C values of the hydrocarbon gases in the semi-open system are measured values, while those in the semi-closed system are the calculated values for the increase in the differential gas fraction between two successive pyrolysis
experiments. The VR<sub>r</sub> in the semi-open system is the logarithmic mean value of two successive pyrolysis experiments, while that in the semi-closed system is the measured value.

Fig. 4. Changes in the $\delta^2$H values of methane, ethane, and propane expelled from Eocene Yubari coal with increasing VR<sub>r</sub> under semi-open and semi-closed systems. The $\delta^2$H values of the hydrocarbon gases in the semi-open system are measured values, while those in the semi-closed system are the calculated values for the increase in the differential gas fraction between two successive pyrolysis experiments. The VR<sub>r</sub> in the semi-open system is the logarithmic mean value of two successive pyrolysis experiments, while that in the semi-closed system is the measured value.

Fig. 5. The paths of evolution of the $\delta^{13}$C values of methane, ethane, and propane expelled from Eocene Yubari coal with increasing VR<sub>r</sub> under semi-open and semi-closed systems. (a) Relationship between the $\delta^{13}$C values of methane and ethane. (b) Relationship between the $\delta^{13}$C values of ethane and propane.
Table 1. Vitrinite reflectance, elemental compositions, and H/C and O/C atomic ratios of Eocene Yubari coal and artificially matured coal (AMC #1 to #5) pyrolyzed at 330–390°C for 72 h under an anhydrous closed system.

<table>
<thead>
<tr>
<th>Sample or Pyrolysis No.</th>
<th>Temp. (°C)</th>
<th>Time (h)</th>
<th>VRr (%)</th>
<th>C (%)</th>
<th>H (%)</th>
<th>O* (%)</th>
<th>N (%)</th>
<th>S (%)</th>
<th>Atomic H/C</th>
<th>Atomic O/C</th>
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*: by difference (dry ash free)
Table 2. Independent and cumulative yields of gases generated from Eocene Yubari coal pyrolyzed under a semi-open system.

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<tr>
<th>Pyrolysis No.</th>
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<th>CO</th>
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Yield of Gases (μL/gCoal)

*: VRr value from corresponding closed system pyrolysis shown in Table 1.
Table 3. Yields of gases generated from Eocene Yubari coal pyrolyzed under an anhydrous closed system.

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<th>Pyrolysis No.</th>
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<th>Time (h)</th>
<th>VR (%)</th>
<th>Yield of Gases (μL/gCoal)</th>
<th>H₂</th>
<th>CO</th>
<th>C₁</th>
<th>CO₂</th>
<th>C₂H₄</th>
<th>C₂</th>
<th>C₃H₆</th>
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<th>Yield of Gases (μL/gCoal)</th>
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Table 4. Stable carbon and hydrogen isotopic compositions of gases generated from Eocene Yubari coal pyrolyzed under a semi-open system.

<table>
<thead>
<tr>
<th>Sample or Pyrolysis No.</th>
<th>Temp. (ºC)</th>
<th>Time (h)</th>
<th>VRr (%)</th>
<th>VRmean * (%)</th>
<th>$\delta^{13}$C (%)</th>
<th>$\delta^2$H(‰)</th>
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<td>Original Coal</td>
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<td>1.70</td>
<td>1.58</td>
<td>-31.0</td>
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*: Logarithmic mean of VRr values from two successive pyrolysis experiments.
Table 5. Stable carbon and hydrogen isotopic compositions of gases generated from Eocene Yubari coal pyrolyzed under a closed system. The apparent isotopic composition for the increased differential gas fraction was determined from the mass balance of the gases.

<table>
<thead>
<tr>
<th>Pyrolysis No.</th>
<th>Temp. (ºC)</th>
<th>Time (h)</th>
<th>VR&lt;sub&gt;r&lt;/sub&gt; (%)</th>
<th>δ&lt;sup&gt;13&lt;/sup&gt;C&lt;sub&gt;1&lt;/sub&gt; (%o)</th>
<th>δ&lt;sup&gt;13&lt;/sup&gt;C&lt;sub&gt;2&lt;/sub&gt; (%o)</th>
<th>δ&lt;sup&gt;13&lt;/sup&gt;C&lt;sub&gt;3&lt;/sub&gt; (%o)</th>
<th>δ&lt;sup&gt;2&lt;/sup&gt;H&lt;sub&gt;C1&lt;/sub&gt; (%o)</th>
<th>δ&lt;sup&gt;2&lt;/sup&gt;H&lt;sub&gt;C2&lt;/sub&gt; (%o)</th>
<th>δ&lt;sup&gt;2&lt;/sup&gt;H&lt;sub&gt;C3&lt;/sub&gt; (%o)</th>
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</table>

*: Calculated value for differential gas fraction.
Fig. 1

(a) 330°C-72h (VR_t=0.60%)

(b) 340°C-72h (VR_t=0.97%)

(c) 370°C-72h (VR_t=1.46%)

Retention time (min)

Intensity

N2, O2, CO, CH4, CO2, C2H6, C3H8, i-C4H10, n-C4H10
Fig. 2

- **CH₄**
  - Semi-open system
  - Closed system

- **C₂H₆**
  - Semi-open system
  - Closed system

- **C₃H₈**
  - Semi-open system
  - Closed system

- **CO₂**
  - Semi-open system
  - Closed system

- **H₂**
  - Semi-open system
  - Closed system

- **C₂H₄+C₃H₆**
  - Semi-open system
  - Closed system

Yield (μL/g coal) vs. Vitrinite Reflectance (%)
Fig. 3

- **CH₄**
- **C₂H₆**
- **C₃H₈**

- **Semi-open system**
- **Semi-closed system**

Vitrinite Reflectance (%) vs. $\delta^{13}$C (‰ vs. VPDB)
Fig. 4

- **CH$_4$**
  - Semi-open system
  - Semi-closed system

- **C$_2$H$_6$**

- **C$_3$H$_8$**

Vitrinite Reflectance (%)
Fig. 5

(a) 

\[ \delta^{13}C_2 \text{ (%o vs. VPDB)} \]

\[ \delta^{13}C_1 \text{ (%o vs. VPDB)} \]

- Semi-open system
- Semi-closed system

(b) 

\[ \delta^{13}C_3 \text{ (%o vs. VPDB)} \]

\[ \delta^{13}C_2 \text{ (%o vs. VPDB)} \]

- Semi-open system
- Semi-closed system