



Title	Copper-Catalyzed Enantioselective Coupling between Allylboronates and Phosphates Using a Phenol-Carbene Chiral Ligand: Asymmetric Synthesis of Chiral Branched 1,5-Dienes
Author(s)	Yasuda, Yuto; Ohmiya, Hirohisa; Sawamura, Masaya
Citation	Synthesis-stuttgart, 50(11), 2235-2246 https://doi.org/10.1055/s-0036-1591548
Issue Date	2018-06
Doc URL	http://hdl.handle.net/2115/74534
Rights	(c)2018 Georg Thieme Verlag KG
Type	article (author version)
File Information	Synthesis_revision.10.pdf



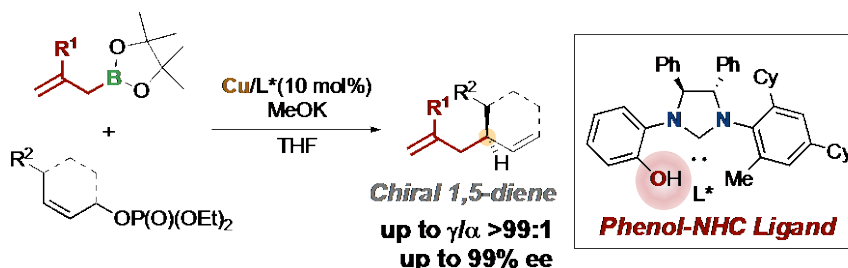
[Instructions for use](#)

Copper-Catalyzed Enantioselective Coupling between Allylboronates and Phosphates with a Phenol-Carbene Chiral Ligand: Asymmetric Synthesis of Chiral Branched 1,5-Dienes

Yuto Yasuda^a
Hirohisa Ohmiya^{*b}
Masaya Sawamura^{*a}

^a Department of Chemistry, Faculty of Science, Hokkaido University, Sapporo 060-0810, Japan
sawamura@sci.hokudai.ac.jp

^b Division of Pharmaceutical Sciences, Graduate School of Medical Sciences, Kanazawa University, Kakuma-machi, Kanazawa 920-1192, Japan
ohmiya@p.kanazawa-u.ac.jp



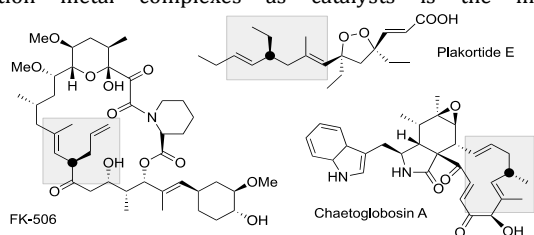
Received:
Accepted:
Published online:
DOI:

Abstract This report presents details of the Cu-catalyzed enantioselective allyl–allyl coupling reaction between allylboronates and (*Z*)-allylic phosphates using a new chiral *N*-heterocyclic carbene (NHC) ligand containing a phenolic hydroxyl group. The copper catalysis delivered enantioenriched chiral 1,5-dienes with a tertiary stereogenic center. Various functional group compatibility and the use of earth-abundant and relatively low-toxicity Cu as a metal were attractive features of this protocol. The utility of the chiral phenol-NHC ligand for enantioselective copper catalysis with organoboron compounds was demonstrated. Enantio-discrimination models were discussed.

Key words asymmetric catalysis · allylic substitution · synthetic method · copper catalysis · organoboron

INTRODUCTION AND BACKGROUND

Chiral 1,5-dienes with a stereogenic center at the allylic/homoallylic position are found in many important biologically active molecules such as FK-506, Plakortide E, and Chaetoglobosin A (Figure 1)^[1], and also serve as useful building blocks in organic synthesis due to the versatility of the two alkene functionalities for further transformations. While several methods have been developed to produce chiral 1,5-dienes^[2], enantioselective γ -substitution of allyl alcohol derivatives with organometal reagents (allyl–allyl coupling) using the chiral transition metal complexes as catalysts is the most



straightforward because it uses readily available substrates.^[3]

In 2010, Morken and co-workers reported the pioneering example of catalytic asymmetric allyl–allyl coupling; Pd-catalyzed allyl–allyl coupling between substituted allylboronates^[4] and (*E*)-allylic carbonates occurred with high regio- and enantioselectivity.^[5] This protocol enabled access to chiral 1,5-dienes containing contiguous stereogenic centers with high diastereo- and enantioselectivity. More recently, Feringa's group reported asymmetric cross-coupling between allyl Grignard reagents and (*E*)-allyl bromides using the copper-chiral monodentate phosphoramidite catalyst, but only moderate S_N2' regioselectivity was obtained.^[6] In 2014, Carreira *et al.* developed Ir-catalyzed regio- and enantioselective allylic cross-coupling between allylsilanes and secondary aromatic allylic alcohols.^[7] Despite these efforts, the allyl–allyl coupling was limited to use of acyclic (*E*)-allylic electrophiles.

Recently, the Cu-catalyzed enantioselective allyl–allyl coupling between allylboronates and allylic phosphates^[8–11] using a new chiral *N*-heterocyclic carbene (NHC) ligand bearing a phenolic hydroxyl group^[12,13] was reported. Reaction occurred with exceptional γ -regioselectivity and high enantioselectivity. Various functional groups were tolerated. The ability to use (*Z*)-aliphatic allylic substrates in the copper catalysis was complementary to Morken's Pd system (in their reports, only primary (*E*)-allylic electrophiles were used).^[5] The present report describes the details of further studies of this Cu-catalyzed enantioselective allyl–allyl coupling.^[8]

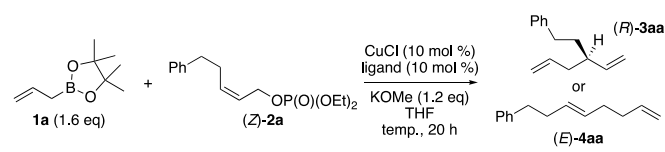
RESULTS AND DISCUSSION

Optimization. Enantioselective S_N2' allylic alkylation between non-allylic allylboranes (alkyl-9-BBN) and primary allylic substrates using a catalytic amount of a chiral bisphosphine/Cu(I) complex and a stoichiometric amount of potassium alkoxide base has been reported.^[14] The results prompted the initiation of a program to develop a Cu-catalyzed enantioselective allylic substitution with allylboron reagents.

For screening of the reaction conditions, commercially available allylboronic acid pinacolate esters were used instead of the allyl-9-BBN reagents (Table 1).^[15] While investigating an effective achiral ligand that could selectively produce the racemic, branched γ -substitution product (**3aa**), a ring-saturated NHC/Cu complex prepared *in situ* from 1,3-bis(2,4,6-trimethylphenyl)imidazolium chloride (SImes·HCl), CuCl, and KOMe was produced in high yield (93%) with exclusive γ -regioselectivity ($\gamma/\alpha > 99:1$) to form the branched coupling product **3aa** from reaction between 2-allyl-4,4,5,5-tetramethyl-1,3,2-dioxaborolane (**1a**) and γ -monosubstituted primary (Z)-allylic phosphate **2a** in THF at -20°C (entry 1). In contrast to the excellent performance of the ring-saturated NHC ligand SImes, the corresponding unsaturated NHC ligand IMes, derived from 1,3-bis(2,4,6-trimethylphenyl)imidazolium chloride (IMes·HCl), gave a mixture of branched and linear products with a low γ/α regioselectivity (62:38) and moderate total product yield (entry 2). Without a ligand, or with 1,10-phenanthroline (Phen) or 1,2-bis(diphenylphosphino)ethane (DPPE) ligands, no reaction occurred (entries 3–5). The use of triphenylphosphine (Ph_3P) as a monodentate phosphine ligand gave only the linear α -substitution product (*E*)-**4aa** (entry 6).

Based on these results, focus was placed on ring-saturated chiral NHC ligands (Table 1, entries 7–14). The C_2 -symmetric imidazolium chloride [(*S,S*)-**L1**·HCl],^[16] which has two stereogenic carbon centers in the imidazolidine ring with two *N*-mesityl groups, did not result in enantioselectivity; the branched, nearly racemic coupling product **3aa** was obtained with exclusive regioselectivity and moderate yield (entry 7). When similar chiral NHC ligands (**L2**, **L3**) bearing 2-methylphenyl (**L2**) or 2-methoxyphenyl (**L3**) groups were used instead of one of the mesityl groups in **L1**, enantioselectivity was moderate (50% and 54% ee), but yields were low (entries 8 and 9). Next, the imidazolium salt **L4**·HBF₄ bearing a 2-hydroxyphenyl group was used (entry 10). A previous report indicated that **L4** exhibited high ligand performance in enantioselective Cu-catalyzed allylic substitution with terminal alkyne as pronucleophiles.^[12a] Fortunately, the Cu-**L4** catalyst system produced better results; the coupling product was formed in greater yield (77%) and with greater enantioselectivity (85% ee) compared to the system using (*S,S*)-**L2** or **L3**, without decreasing regioselectivity. These results indicated that the phenolic hydroxy group in **L4** has a functional role. Changing the *N*-mesityl group of **L4**·HBF₄ to a *N*-2,4-dicyclohexyl-6-methylphenyl group to afford the new phenol-NHC chiral ligand precursor **L5**·HBF₄, resulted in greater enantioselectivity (92% ee), a high yield (83%), and exceptional regioselectivity ($\gamma/\alpha > 99:1$) (entry 11). Furthermore, the enantioselectivity was achieved (99% ee) by decreasing the reaction temperature to -40°C (entry 12). In contrast, the naphthol-NHC chiral ligands **L6** and **L7**^[12b] were not effective (entries 13 and 14).

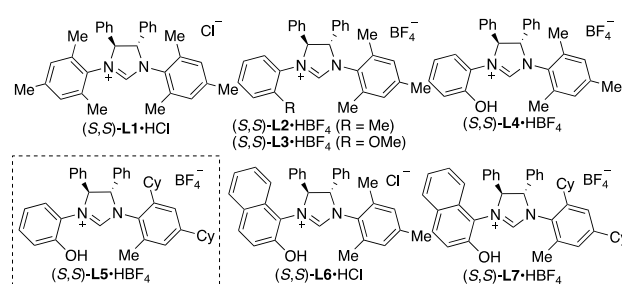
Table 1. Ligand effects for reaction between **1a** and (Z)-**2a**^[a]



Entry	Ligand	Temp ($^\circ\text{C}$)	Yield (%) ^[b]	γ/α (3aa : 4aa) ^[c]	Ee (%) ^[d]
1	SImes·HCl	-20	93	>99:1	–
2	IMes·HCl	-20	65	62:38	–
3	none	-20	0	–	–
4	Phen	-20	0	–	–

5	DPPE	-20	0	–	–
6	Ph_3P	-20	57	1:>99	–
7	(<i>S,S</i>)- L1 ·HCl	-20	49	>99:1	2
8	(<i>S,S</i>)- L2 ·HBF ₄	-20	13	>99:1	50
9	(<i>S,S</i>)- L3 ·HBF ₄	-20	10	>99:1	54
10	(<i>S,S</i>)- L4 ·HBF ₄	-20	77	>99:1	85
11	(<i>S,S</i>)- L5 ·HBF ₄	-20	83	>99:1	92
12 ^[e]	(<i>S,S</i>)- L5 ·HBF ₄	-40	80	>99:1	99
13	(<i>S,S</i>)- L6 ·HBF ₄	-20	34	92:8	67
14	(<i>S,S</i>)- L7 ·HBF ₄	-20	56	>99:1	82

[a] Data are taken from ref 8. Conditions: **1a** (0.24 mmol), (Z)-**2a** (0.15 mmol), CuCl/ligand (10 mol%), KOMe (0.18 mmol), THF (0.6 mL) for 20 h. [b] Yield of isolated product. [c] Determined by ¹H NMR analysis of the crude product. [d] The ee was determined by HPLC. [e] **1a** (1.9 mmol) and (Z)-**2a** (1.2 mmol) were used. Reaction was conducted for 48 h.



Next, the effect of copper salts (Table 2) was investigated. When CuOAc was used as the catalyst, regioselectivity (γ/α 91:9) and enantioselectivity (76% ee) decreased. The use of Cu(OAc)₂ instead of Cu(I) salt, resulted in a modest yield and low selectivities. Changing CuCl to CuOTf·toluene_{1/2} afforded similar selectivity but a lower product yield. Use of MesCu and cationic Cu(MeCN)₄PF₆ was as effective as CuCl, and gave comparable results.

The effect of solvent is shown in Table 2. Reaction with toluene gave only the racemic product in moderate yield (43%) and low regioselectivity (γ/α 64:36). Dichloromethane was ineffective in the reaction (26% yield, γ/α 71:29, 51% ee). No reaction occurred in MeCN or hexane.

Table 2. Effects of copper salts and solvents^[a]

Entry	Cu Salt	Solvent	Yield (%) ^[b]	γ/α □ [(<i>R</i>)- 3aa :(<i>E</i>)- 4aa] ^[c]	Ee (%) ^[d]
1 ^[e]	CuCl	THF	83	>99:1	92
2	CuOAc	THF	85	91:9	76
3	Cu(OAc) ₂	THF	68	73:27	52
4	CuOTf·toluene _{1/2}	THF	48	>99:1	90
5	MesCu	THF	85	>99:1	89
6	Cu(MeCN) ₄ PF ₆	THF	73	>99:1	88
7	CuCl	toluene	43	64:36	0
8	CuCl	DCM	26	71:29	51
9	CuCl	MeCN	0	–	–
10	CuCl	hexane	0	–	–

[a] Conditions: **1a** (0.24 mmol), (Z)-**2** (0.15 mmol), Cu salt/(*S,S*)-**L5**·HBF₄ (10 mol%), KOMe (0.18 mmol), solvent (0.6 mL), -20°C for 20 h. [b] Yield of isolated product. [c] Determined by ¹H NMR of crude product. [d] The ee was determined by HPLC. [e] Table 1, entry 11.

The effects of leaving groups and bases are summarized in Table 3. The use of allyl bromide or chloride as an allylic electrophile under the conditions described for Table 1, entry 11, decreased both regio- and enantioselectivity (entries 2 and 3).

The nature of the base had a significant impact on yield and selectivities (Table 3). Use of NaOMe instead of KOMe decreased regioselectivity (γ/α 89:11) and enantioselectivity (74% ee) (entry 4). No reaction occurred with LiOMe, due to the greater

solubility of LiCl or LiOP(O)(OEt)₂, which may form inactive Cu species through ionic interactions. (entry 5). The structure of the alkoxide moiety of the base also had a strong impact on yield, regioselectivity, and enantioselectivity. Thus, when KOMe was changed to a sterically more demanding base, KOtBu, product yield was moderate and both regioselectivity and enantioselectivity decreased significantly (entry 6). This result suggests that the trialkoxyboron ROBpin may participate in the reaction because Lewis acids activate the phosphate leaving group (see Figure 4). Coupling reaction with K₂CO₃ did not occur (entry 7).

Table 3. Effect of leaving group and base^[a]

Entry	Leaving group	Base	Yield (%) ^[b]	γ/α □ [(<i>R</i>)- 3aa :(<i>E</i>)- 4aa] ^[c]	Ee (%) ^[d]
1 ^[e]	OP(O)(OEt) ₂ (2a)	KOMe	83	>99:1	92
2	Cl ^[f]	KOMe	89	94:6	64
3	Br ^[e]	KOMe	85	82:18	20
4	OP(O)(OEt) ₂ (2a)	NaOMe	91	89:11	74
5	OP(O)(OEt) ₂ (2a)	LiOMe	0	–	–
6	OP(O)(OEt) ₂ (2a)	KOtBu	67	63:37	50
7	OP(O)(OEt) ₂ (2a)	K ₂ CO ₃	0	–	–

[a] Data are taken from ref 8. Conditions: **1a** (0.24 mmol), (*Z*)-**2** (0.15 mmol), CuCl/(*S,S*)-**L5**·HBF₄ (10 mol%), base (0.18 mmol), THF (0.6 mL), –20 °C for 20 h. [b] Yield of isolated product. [c] Determined by ¹H NMR of crude product. [d] The ee was determined by HPLC. [e] Table 1, entry 11. [f] (*Z*)-[5-Chloro-3-penten-1-yl]benzene. [g] (*Z*)-[5-Bromo-3-penten-1-yl]benzene.

Substrate Scopes. (*Z*)-Allylic phosphates with various aliphatic substituents were reacted with **1a** in the Cu-**L5** catalyst system (Table 4). When the 2-phenylethyl group of **2a** was replaced with a benzyl or octyl group, coupling proceeded with excellent γ -selectivity and preservation of enantioselectivity (entries 1 and 3). A sterically more demanding γ -substituent, such as a cyclohexyl group, was also tolerated and produced a high level of enantioselectivity (80% ee) (entry 4).^[17] Notably, the enantioselective reaction with 2-butene-1,4-diol derivatives, which have two potential leaving groups at different allylic positions, occurred with the allylic C–O bonds in the ether or carboxylic ester leaving groups untouched (entries 5–12 and 15).

The reaction has a great functional group compatibility (entries 2 and 5–12, 15–17). For example, an allylic phosphate (**2**) bearing a 1,3-benzodioxole (**2c**), THP ether (**2f**), benzyl ether (**2g**), silyl ether (**2h, q**),^[17] pivalate (**2i**), carbamate (**2p**) or *p*-toluenesulfonate (**2r**) group as the aliphatic γ -substituent, reacted with **1a** to produce the corresponding 1,5-diene derivatives in good yields with high enantioselectivities (85–97% ee) (entries 2, 5–8, 15–17). Methoxy, trifluoromethyl, bromo, or dimethylamino substituents were tolerated in the aromatic ring of the benzoate groups (entries 9–12). However, no reaction occurred with the allylic phosphate bearing a nitro (**2n**) or cyano (**2o**) group (entries 13 and 14).

The potential for scaling up the enantioselective allyl–allyl coupling was examined on a preparative scale (Eq. 1). Reaction between allylboronate **1a** (9.4 mmol, 1.5 g) and allylic phosphate **2h** (5.9 mmol, 2.0 g) afforded only the branched coupling product in 73% yield (0.98 g, 4.3 mmol) with 92% ee.

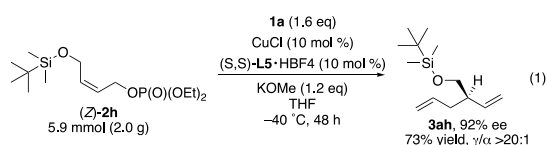


Table 4. Scope of allylic phosphates^[a]

Entry	Phosphate OP = OP(O)(OEt) ₂	Product	Temp (°C)	Yield (%) ^[b,c]	Ee (%) ^[d]
1			–50	62	86
2			–30	97	86
3			–40	85	84
4			–50	59	80
5	R = THP (2f)	3af	–30	83	90
6	R = Bn (2g)	3ag	–50	75	85
7	R = ^t BuMe ₂ Si (2h)	3ah	–50	78	97
8	R = Piv (2i)	3ai	–30	70	91
9	R = MeO (2j)	3aj	–30	85	92
10	R = F ₃ C (2k)	3ak	–30	61	90
11	R = Br (2l)	3al	–30	65	90
12	R = Me ₂ N (2m)	3am	–30	87	91
13	R = NO ₂ (2n)	3an	–10	0	–
14	R = CN (2o)	3ao	–10	0	–
15			–30	80	90
16			–40	78	96
17			–30	88	92

[a] Data are taken from ref 8 for entries 1–12 and 15–17. Conditions: **1a** (0.24 mmol), (*Z*)-**2** (0.15 mmol), CuCl/(*S,S*)-**L5**·HBF₄ (10 mol%), KOMe (0.18 mmol), THF (0.6 mL) for 48 h. [b] Yield of isolated product. [c] Constitutional isomer ratio γ/α > 20:1 (determined by ¹H NMR of crude product). [d] The ee was determined by HPLC.

The copper-catalyzed allyl–allyl coupling reaction between **1a** and enantioenriched 2-butene-1,4-diol derivative (*S*)-**2s** (99% ee) with (*R,R*)-**L5**·HBF₄ (*i.e.*, an enantiomeric isomer of (*S,S*)-**L5**·HBF₄) gave the corresponding 1,5-diene **3as** with two adjacent stereogenic centers with a 2*S*,3*S*-configuration in modest yield (51%) and high diastereoselectivity (d.r. 90:10) (Eq. 2). Without a ligand, or with (*S,S*)-**L5**·HBF₄, the reaction gave no or only a trace of the coupling product. Thus, the use of **L5** is mandatory, and (*R,R*)-**L5**-Cu system and the (*S*)-**2s** substrate is a matched pair, while (*S,S*)-**L5**/Cu complex and (*S*)-**2s** is mismatched. This result is in accord with a consideration of the Felkin-Anh model as depicted in Figure 2, which predicts preferred approach of organocopper nucleophile to *Si*-face of the allylic plain with a minimum A^{1,3}-strain.

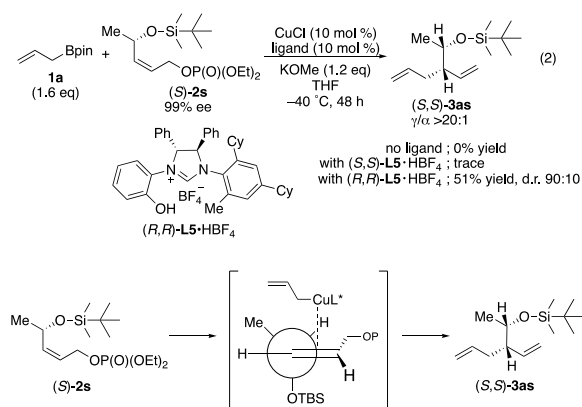


Figure 2. A Felkin-Anh-type stereoselection model

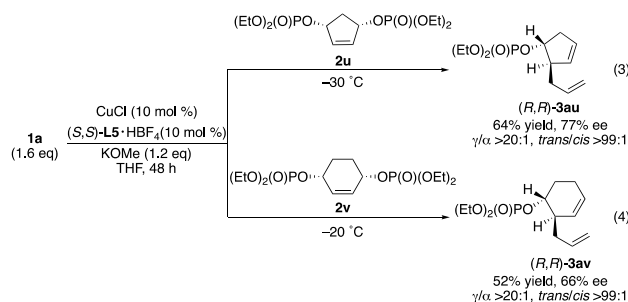
Reactions of β -substituted allylboronate derivatives using the Cu-L5 catalyst system were investigated (Table 5). Methallylboronate **1b** and 2-ethyl-2-propen-1-ylboronate **1c** reacted with (*Z*)-**2a** with excellent γ -selectivity ($\gamma/\alpha > 20:1$) and high enantioselectivity (95% and 80% ee, respectively) (entries 1 and 2). An allylboronate with a hexyl, benzyl, or phenyl group at the β -position (**1d-f**) underwent reaction to afford the coupling products with enantiomeric excesses greater than 80%, while the regioselectivities were moderate (entries 3–5).^[18] The allylboronate with a chloro group at the β -position did not react (entry 6). In addition, no reaction occurred with γ -substituted allylboronates such as *trans*- or *cis*-crotylboronates.

Table 5. Scope of allylboronates^[a]

Entry	Allylboronate	Phosphate OP = OP(O)(OEt) ₂	Product	Yield (%) ^[b]	γ/α ^[c]	Ee (%) ^[d]
1	1b	2a	3ba	77	>20:1	95
2	1c	2a	3ca	50	>20:1	80
3	1d	2a	3da	87	84:16	80
4	1e	2a	3ea	59	87:13	82
5	1f	2t	3ft	89	78:22	83
6	1g	2a	3ga	0	–	–

[a] Data are taken from ref 8 for entries 1–5. Conditions: **1** (0.24 mmol), (*Z*)-**2** (0.15 mmol), CuCl/(*S,S*)-L5·HBF₄ (10 mol%), KOMe (0.18 mmol), THF (0.6 mL), –50 °C (entry 1) or –30 °C (entries 2–6) for 48 h. [b] Yield of isolated product. [c] Determined by ¹H NMR of crude product. [d] The ee was determined by HPLC.

The coupling reaction between **1a** and *cis*-4-cyclopentene-1,3-diol diphosphate **2u** catalyzed by the Cu-L5 system occurred with adequate enantioselectivity, giving the *trans*-1,2-isomer **3au** (Eq. 3).^[17,19] Reaction with *cis*-4-cyclohexene-1,3-diol derivative **2v** gave the *trans*-1,2-isomer **3av** with moderate enantiocontrol (66% ee) (Eq. 4). These stereochemical results indicate that the present Cu-catalyzed reaction proceeded through an *anti*-S_N2' type reaction pathway. Morcken's Pd-catalyzed protocol has not been applied to this type of cyclic allylic electrophile involving a *Z* alkene moiety.^[5]



Mechanistic Study and Proposed Reaction Pathway. The alkene geometry of the allylic phosphates influenced both regio- and enantioselectivity. Thus, reaction between **1a** and (*E*)-**2a** under the conditions described in Table 1, entry 11 gave the linear α -substitution product (*E*)-**4aa** preferentially with minor formation of (*S*)-**3aa** (71% ee), the antipode of the product derived from (*Z*)-**2a** (Eq. 5).^[19] This α -selectivity appeared to occur *via* allylic 1,3-migration of Cu in the allylcopper(III) species. To gain insight into the nature of the postulated allylcopper(III) species, secondary allylic phosphate **2a'**, a constitutional isomer of (*Z*)-**2a**, was reacted under the same conditions (Eq. 6).^[19] Interestingly, this reaction gave results similar to reaction with (*E*)-**2a**. The convergence observed in the regioselectivity and stereochemistry suggests that (*E*)-**2a** and **2a'** led to a common equilibrium mixture of the allylcopper(III) species prior to reductive elimination to form the product mixture (Figure 3).^[20]

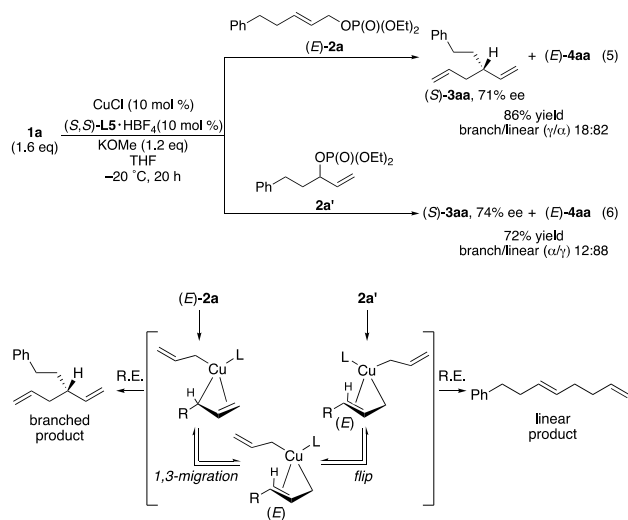


Figure 3. Allylcopper(III) species

Based on the assumption that γ -selective reaction of (*Z*)-**2** also occurs *via* allylcopper(III) intermediates, a catalytic cycle was proposed for the enantioselective allyl-allyl coupling catalyzed by the Cu-L5 system, as shown in Figure 4.^[20] An alkoxy copper(I) complex (**A**) was formed from reaction between CuCl, L5·HBF₄, and KOMe. For this complex, the chiral NHC ligand coordinated to Cu as an anionic C,O-bidentate ligand. Then, transmetalation between **A** and an allylborate (**B**)^[21] afforded the potassium phenoxo(allyl)cuprate (**C**). Compound **C** formed a π -complex (**D**) with the allylic phosphate **2**, in which the Cu was *anti* to the phosphate leaving group. The MeOBpin may activate the phosphate group as a Lewis acid.^[14d] Subsequently, oxidative addition produced (π -en- σ -yl)copper(III) complex **E1** with a secondary sp³-carbon atom bound to Cu. Facile reductive elimination of **E1**, which was faster than allylic 1,3-Cu-migration to form **E2**, produced the

branched γ -substitution product **3** and regenerated alkoxy copper(I) complex **A** (Figs. 5 and 6 for enantio-discrimination models).

Thus, the crucial dependence of regioselectivity on the *E/Z* geometry of the allylic substrate (**2**) can be explained by the greater instability of the allylcopper(III) intermediate **E2** compared to the corresponding Cu(III) species (**E1**) produced from the *E* substrate, due to larger steric repulsion in the allyl moiety derived from the *Z* substrate (Fig. 3). Similar consideration may be applicable to the corresponding transition states that determine the *E/Z*-isomeric product distribution.

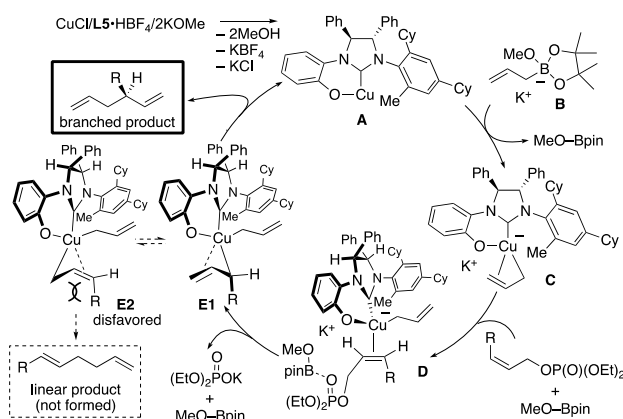


Figure 4. Postulated catalytic reaction pathway.

Reactions of *E*- and *Z*-isomers of **2a** gave the antipodes of **3aa**, (Table 1, entry 11 vs. Eq. 5), and the products from (*Z*)-**2a** were produced with greater enantioselectivity. This led to the proposal of the enantioselection models shown in Figs. 5 and 6. In the π -complex **D** (Fig. 4), the chiral copper center adopted a tetrahedral coordination geometry including the C,O-bidentate chelation. The K⁺ ion bridged the phenoxide oxygen and MeO-Bpin, which interacted with the phosphate leaving group as a Lewis acid. These assumptions suggest four possible π -complexes (**D1** and **D2** for the major enantiomer; **D3** and **D4** for the minor enantiomer) (Fig. 5). Complex **D1** is the most favorable, because it results in the fewest steric repulsions between catalyst and allylic phosphate. Complexes **D2**, **D3**, and **D4** appear to be destabilized by the steric repulsions between the γ -substituent (R) and nonhydroxylated *N*-aryl group or one of the phenyl groups on the imidazolidine ring. These conclusions are consistent with the experimental observations that the enantioselectivity is influenced by the steric nature of the γ -substituent (R) (Table 4, entries 3 and 4) or the nonhydroxylated *N*-aryl group in the phenolic NHC ligands (Table 1, entries 10 and 11).

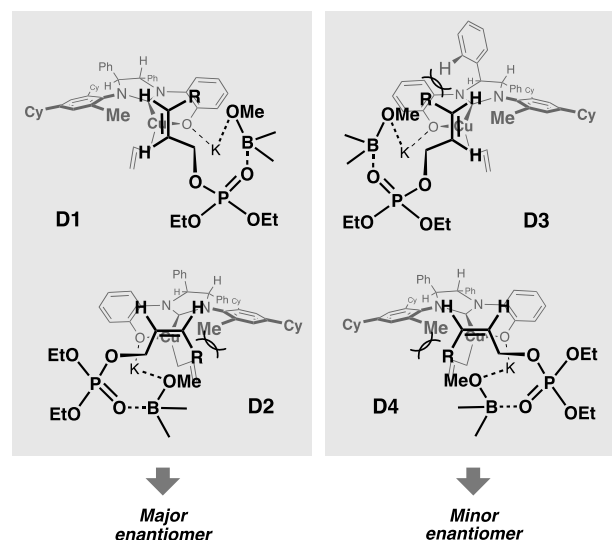


Figure 5. Models for enantioselection with (*Z*)-**2**

For reaction of (*E*)-**2** (Fig. 6), the nonhydroxylated *N*-aryl group and/or the phenyl group in the imidazolidine ring caused steric repulsion toward the γ -substituent (R) for all postulated complexes **D1'**–**D4'**. These considerations explain the more efficient enantioselection in the reaction with allylic substrates having the *Z*-configuration.

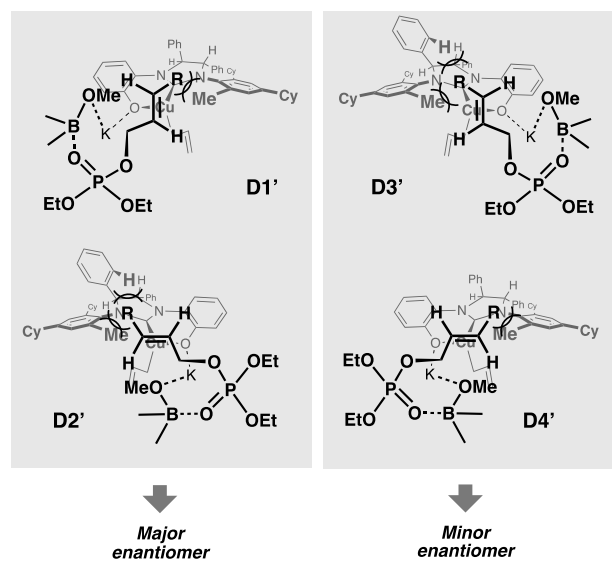


Figure 6. Models for enantioselection with (*E*)-**2**

CONCLUSIONS

In conclusion, a versatile method was developed for Cu-catalyzed enantioselective allyl-allyl coupling between allylboronates and (*Z*)-acyclic and cyclic allylic phosphates to form various chiral 1,5-diene derivatives. Catalysis of a Cu(I) complex with a new phenol-NHC chiral ligand enabled the reaction, demonstrating the utility of this class of chiral ligands for enantioselective copper catalysis with organoboron compounds.^[12] This Cu-catalyzed protocol provides efficient access to functionalized, enantio-enriched chiral 1,5-dienes with a stereogenic carbon center at the allylic/homoallylic position. The broad functional group compatibility and the use of earth-abundant and relatively low-toxic Cu as a metal are attractive features of this protocol. A functional role for the phenolic

hydroxy group in the chiral NHC ligand was suggested by the results.

EXPERIMENTAL SECTION

Instrumentation and Chemicals

NMR spectra were recorded on a JEOL ECX-400, operating at 400 MHz for ^1H NMR, 100.5 MHz for ^{13}C NMR and 128 MHz for ^{15}N NMR. Chemical shift values for ^1H and ^{13}C are referenced to Me_4Si and the residual solvent resonances, respectively. Chemical shifts are reported in δ ppm. Mass spectra were obtained with Thermo Fisher Scientific Exactive, JEOL JMS-T100LP or JEOL JMS-700TZ at the Instrumental Analysis Division, Global Facility Center, Creative Research Institution, Hokkaido University. HPLC analyses were conducted on a HITACHI ELITE LaChrom system with a HITACHI L-2455 diode array detector or a HITACHI Chromaster with a HITACHI 5430 diode array detector. Optical rotations were measured on a JASCO P-2200. TLC analyses were performed on commercial glass plates bearing 0.25-mm layer of Merck Silica gel 60F₂₅₄. Silica gel (Kanto Chemical Co., Silica gel 60 N, spherical, neutral) was used for column chromatography. IR spectra were measured with a Perkin-Elmer Spectrum One. Melting points were measured on a Yanaco MP-500D apparatus. Gel permeation chromatography (GPC) was performed by LC-908 (Japan Analytical Industry Ltd., two in-line JAIGEL-2H, CHCl_3 , 3.5 mL/min, UV and RI detectors).

All reactions were carried out under nitrogen or argon atmosphere. Materials were obtained from commercial suppliers or prepared according to standard procedures unless otherwise noted. CuCl , and KOMe were purchased from Aldrich Chemical Co., stored under nitrogen, and used as received. THF was purchased from Kanto Chemical Co., and stored under argon. 2-Allyl-4,4,5,5-tetramethyl-1,3,2-dioxaborolane (**1a**) was obtained from commercial suppliers. Allylboronates **1g** were prepared according to the reported procedures.^[22] Allylboronates **1b-f**, allylic phosphates **2a-r**, **2t-v**, **2a'** and NHC ligands **L2-7** were reported in the literature.^[8, 12a] (*E*)-**4aa** are found in the literature.^[3f]

Preparation of Allylic Phosphate (*S*)-**2s**

(*S*)-3-Butyn-2-ol (99% ee) (1.2 mL, 15 mmol) and imidazole (2.0 g, 30 mmol) were dissolved in DCM (30 mL) at 0 °C. Then, TBDMSCl (3.4 g, 22 mol) was added to the mixture, and the solution was stirred at rt for 8 h. The reaction was quenched with H_2O and extracted with DCM (15 mL \times 3). The combined organic layer was dried over MgSO_4 . Then, the drying agent was removed by filtration, and a filtrate was evaporated under reduced pressure. A residue was purified by flash chromatography on silica gel (0–5% EtOAc/hexane) to give TBDMS-protected alcohol derivative (2.6 g, 14 mmol) in 96% yield.

$^n\text{BuLi}$ (9.6 mL, 15.8 mmol) was dropped to a solution of TBDMS-protected alcohol (2.6 g, 14 mmol) in Et_2O (29 mL) at –78 °C. The reaction mixture was stirred for 30 min before formaldehyde (877 mg, 29.2 mmol) was added to the mixture. The reaction vessel was removed from a cooling bath. After stirring at rt for 12 hours, the reaction was quenched with H_2O . The mixture was diluted with Et_2O and the organic layer was separated. The aqueous phase was extracted with Et_2O (15 mL \times 2) and the combined organic layer was washed with brine, dried over MgSO_4 , filtered, and concentrated under reduced pressure by rotary evaporation. Purification by flash chromatography on silica gel (0–10% EtOAc/hexane) afforded the corresponding propargyl alcohol (2.3 g, 10 mmol) in 73% yield.

A solution of propargylic alcohol (1.1 g, 5 mmol) in hexane (4.3 mL)/acetone (790 μL) was added to the mixture of Pd/CaCO_3 (40.0 mg) and quinoline (397 μL , 3.6 mmol), and then the reaction mixture was filled with hydrogen gas. After confirming the completion of the reaction by ^1H NMR, Pd/CaCO_3 was removed by filtration, and the resulting solution was evaporated under reduced pressure. The residue was purified by flash chromatography on silica gel (0–10% EtOAc/hexane) to

give (*S,Z*)-4-[(*tert*-butyldimethylsilyloxy)-2-penten-1-ol] in 80% yield (874 mg, 4.0 mmol).

To a solution of (*S,Z*)-4-[(*tert*-butyldimethylsilyloxy)-2-penten-1-ol] (874 mg, 4.0 mmol) in pyridine (6.1 mL), $(\text{EtO})_2\text{P}(\text{O})\text{Cl}$ (697 μL , 4.8 mmol) and DMAP (24 mg, 0.2 mmol) were sequentially added at 0 °C. After being stirred at rt for 3 h, the reaction mixture was diluted with EtOAc (20 mL) and was treated with H_2O (2 mL). The resulting mixture was washed with saturated CuSO_4 aq. (10 mL \times 3) and brine, dried over anhydrous MgSO_4 , filtered, and evaporated under reduced pressure. The residue was purified through flash column chromatography on silica gel (10–45% EtOAc/hexane) to provide (*S,Z*)-4-[(*tert*-butyldimethylsilyloxy)-2-penten-1-yl] diethyl phosphate (**2s**) in 94% yield (1.3 g, 3.8 mmol). The ee value of (*S*)-**2s** (99% ee) was determined by chiral HPLC analysis of the *p*-OMe-benzoate derivative of (*S,Z*)-4-[(*tert*-butyldimethylsilyloxy)-2-penten-1-ol] (CHIRALCEL® OD-3 column, 4.6 mm \times 250 mm, Daicel Chemical Industries, hexane/2-propanol 99:1, 0.5 mL/min, 40°C, 220 nm UV detector, retention time = 13.7 min, for *R* isomer and 17.7 min for *S* isomer).

Characterization Data for Allylic Phosphate

(*S,Z*)-4-[(*tert*-Butyldimethylsilyloxy)-2-penten-1-yl] diethyl phosphate (**2s**) (99% ee)

Colorless Oil. IR (neat) 667, 775, 830, 975, 1026, 1254, 1369, 1393, 1473, 2858, 2930, 2957 cm^{-1} . ^1H NMR (400 MHz, CDCl_3) δ 0.04 (s, 3H), 0.06 (s, 3H), 0.89 (s, 9H), 1.20 (d, $J = 6.4$ Hz, 3H), 1.34 (t, $J = 7.2$ Hz, 6H), 4.11 (quintet, $J = 7.2$ Hz, 4H), 4.55–4.69 (m, 3H), 5.48 (dt, $J = 11.2, 6.0$ Hz, 1H), 5.64 (dd, $J = 11.2, 8.0$ Hz, 1H). ^{13}C NMR (100 MHz, CDCl_3) δ –4.8, –4.6, 16.1 (d, $J = 6.7$ Hz), 18.1, 24.6, 25.7, 63.0 (d, $J = 4.8$ Hz), 63.7 (d, $J = 5.7$ Hz), 65.1, 122.3 (d, $J = 6.6$ Hz), 139.2. HRMS–ESI (m/z): [$\text{M}+\text{Na}$]⁺ calcd for $\text{C}_{15}\text{H}_{33}\text{O}_5\text{NaPSi}$, 375.17271; found, 375.17227. $[\alpha]_{\text{D}}^{25} +150$ (c 1.17, CHCl_3).

Procedure for the Copper-Catalyzed Enantioselective Allyl-Allyl Coupling

The reaction in Table 4, entry 1 is representative. CuCl (1.5 mg, 0.015 mmol), $\text{L5}\cdot\text{HBF}_4$ (9.8 mg, 0.015 mmol) and KOMe (12.6 mg, 0.18 mmol) were placed in a vial containing a magnetic stirring bar. The vial was sealed with a Teflon®-coated silicon rubber septum, and then the vial was evacuated and filled with argon. THF (0.6 mL) was added to the vial, and then the mixture was stirred at room temperature for 30 min. Next, 2-allyl-4,4,5,5-tetramethyl-1,3,2-dioxaborolane (**1a**) (45.3 μL , 0.24 mmol) was added. Finally, allylic phosphate **2b** (42.6 mg, 0.15 mmol) was added at –50°C. After 48 h stirring at –50°C, the reaction was quenched with saturated NH_4Cl aq. and extracted with diethyl ether (1 mL \times 3). The combined organic layer was dried over MgSO_4 . Then, the drying agent was removed by filtration, and the resulting solution was evaporated under reduced pressure. The residue was purified by flash chromatography on silica gel (hexane) to give **3ab** (15.9 mg, 0.09 mmol) in 62% yield.

Characterization Data for Allyl-Allyl Coupling Products

Full details for compounds **3aa-am**, **3ap-ar**, **3ba-ea**, **3ft** and **3au-av** can be found in our previous report.^[8]

(*R*)-(3-Vinyl-5-hexen-1-yl)benzene (**3aa**)

The product **3aa** was purified by flash chromatography on silica gel (hexane) (178.9 mg, 0.96 mmol, 80% isolated yield from (*Z*)-**2a**).

(*S*)-(2-Vinyl-4-penten-1-yl)benzene (**3ab**)

The product **3ab** was purified by flash chromatography on silica gel (hexane) (15.9 mg, 0.09 mmol, 62% isolated yield from **2b**).

(*S*)-5-(2-Vinyl-4-penten-1-yl)benzo[*d*][1,3]dioxole (**3ac**)

The product **3ac** was purified by flash chromatography on silica gel (5% Et₂O/hexane) (31.5 mg, 0.14 mmol, 97% isolated yield from **2c**).

(R)-4-Vinyl-1-dodecene (3ad)

The product **3ad** was purified by flash chromatography on silica gel (hexane) (24.7 mg, 0.12 mmol, 85% isolated yield from **2d**).

(R)-Hexa-1,5-dien-3-ylcyclohexane (3ae)

The product **3ae** was purified by flash chromatography on silica gel (pentane) (14.6 mg, 0.09 mmol, 59% isolated yield from **2e**).

(R)-2-[(2-Vinyl-4-penten-1-yl)oxy]tetrahydro-2H-pyran (3af)
(Diastereomeric ratio 1:1)

The product **3af** was purified by flash chromatography on silica gel (0–2% EtOAc/hexane) (24.3 mg, 0.12 mmol, 83% isolated yield from **2f**).

(R)-{[(2-Vinyl-4-penten-1-yl)oxy]methyl}benzene (3ag)

The product **3ag** was purified by flash chromatography on silica gel (0–1% EtOAc/hexane) (22.8 mg, 0.11 mmol, 75% isolated yield from **2g**).

(R)-tert-Butyldimethyl[(2-vinylpent-4-en-1-yl)oxy]silane (3ah)

The product **3ah** was purified by flash chromatography on silica gel (0–1% EtOAc/hexane) (26.6 mg, 0.12 mmol, 78% isolated yield from **2h**).

(R)-2-Vinyl-4-penten-1-yl Pivalate (3ai)

The product **3ai** was purified by flash chromatography on silica gel (0–2% EtOAc/hexane) (20.6 mg, 0.11 mmol, 70% isolated yield from **2i**).

(R)-2-Vinyl-4-penten-1-yl 4-Methoxybenzoate (3aj)

The product **3aj** was purified by flash chromatography on silica gel (0–2% EtOAc/hexane) (31.2 mg, 0.13 mmol, 85% isolated yield from **2j**).

(R)-2-Vinyl-4-penten-1-yl 4-(Trifluoromethyl)benzoate (3ak)

The product **3ak** was purified by flash chromatography on silica gel (0–2% EtOAc/hexane) (25.9 mg, 0.09 mmol, 61% isolated yield from **2k**).

(R)-2-Vinyl-4-penten-1-yl 4-Bromobenzoate (3al)

The product **3al** was purified by flash chromatography on silica gel (0–2% EtOAc/hexane) (28.8 mg, 0.10 mmol, 65% isolated yield from **2l**).

(R)-2-Vinyl-4-penten-1-yl 4-(Dimethylamino)benzoate (3am)

The product **3am** was purified by flash chromatography on silica gel (0–5% EtOAc/hexane) (34.0 mg, 0.13 mmol, 87% isolated yield from **2m**).

(R)-1-Benzyl 4-(2-Vinyl-4-penten-1-yl)piperidine-1,4-dicarboxylate (3ap)

The product **3ap** was purified by flash chromatography on silica gel (5–10% EtOAc/hexane) (43.0 mg, 0.12 mmol, 80% isolated yield from **2p**).

(R)-tert-Butyldimethyl[(4-vinyl-6-hepten-1-yl)oxy]silane (3aq)

The product **3aq** was purified by flash chromatography on silica gel (0–1% EtOAc/hexane) (29.8 mg, 0.12 mmol, 78% isolated yield from **2q**).

(R)-4-Vinyl-6-hepten-1-yl 4-Methylbenzenesulfonate (3ar)

The product **3ar** was purified by flash chromatography on silica gel (0–5% EtOAc/hexane) (38.7 mg, 0.13 mmol, 88% isolated yield from **2r**).

tert-Butyldimethyl[[[(2S,3S)-3-vinyl-5-hexen-2-yl]oxy]silane (3as)

The product **3as** was purified by flash chromatography on silica gel (0–2% EtOAc/hexane) (22.2 mg, 0.09 mmol, 61% isolated yield from **2s**). Colorless Oil. IR (neat) 664, 772, 830, 910, 994, 1072, 1254, 1361, 1374,

1463, 1472, 1641, 2858, 2887, 2930, 2957, 3078 cm⁻¹. ¹H NMR (400 MHz, CDCl₃) δ 0.05 (s, 6H), 0.89 (s, 9H), 1.08 (d, *J* = 6.4 Hz, 3H), 1.98–2.12 (m, 2H), 2.34 (m, 1H), 3.70 (quintet, *J* = 6.4 Hz, 1H), 4.95–5.08 (m, 4H), 5.59–5.88 (m, 2H). ¹³C NMR (100 MHz, CDCl₃) δ -4.8, -4.3, 17.9, 20.8, 25.7, 34.5, 51.7, 70.5, 115.2, 115.8, 137.4, 139.1. HRMS-APCI (*m/z*): [M+H]⁺ calcd for C₁₄H₂₉O₂Si, 241.19822; found, 241.19826. [α]_D²⁵ +34 (*c* 0.37, CHCl₃). The d.r. (90:10) was determined by ¹H NMR analysis of the crude product. The absolute configuration of **3as** was assigned by consideration of the stereochemical pathway.

(R)-{(5-Methyl-3-vinyl-5-hexen-1-yl)benzene (3ba)

The product **3ba** was purified by flash chromatography on silica gel (hexane) (23.1 mg, 0.12 mmol, 77% isolated yield from **(Z)-2a**).

(R)-{(5-Methylene-3-vinylheptyl)benzene (3ca)

The product **3ca** was purified by flash chromatography on silica gel (hexane) (16.1 mg, 0.07 mmol, 50% isolated yield from **(Z)-2a**).

(R)-{(5-Methylene-3-vinylundecyl)benzene (3da)

The product **3da** was purified by flash chromatography on silica gel (hexane) (35.3 mg, 0.13 mmol, 87% isolated yield from **(Z)-2a**).

(R)-{(2-Methylene-4-vinylhexane-1,6-diyl)dibenzene (3ea)

The product **3ea** was purified by flash chromatography on silica gel (0–1% Et₂O/hexane) (24.7 mg, 0.08 mmol, 59% isolated yield from **(Z)-2a**).

(R)-{(4-Methyl-1,5-hexadien-2-yl)benzene (3ft)

The product **3ft** was purified by flash chromatography on silica gel (hexane) (23.0 mg, 0.13 mmol, 89% isolated yield from **2t**). The isolated branched product (**3ft**) was contaminated with a trace amount of the linear product.

(1R,2R)-2-Allyl-3-cyclopenten-1-yl Diethyl Phosphate (3au)

The product **3au** was purified by flash chromatography on silica gel (20–50% EtOAc/hexane) (25.1 mg, 0.10 mmol, 64% isolated yield from **2u**).

(1R,2R)-2-Allyl-3-cyclohexen-1-yl Diethyl Phosphate (3av)

The product **3av** was purified by flash chromatography on silica gel (20–50% EtOAc/hexane) (21.3 mg, 0.08 mmol, 52% isolated yield from **2v**).

Acknowledgment

This work was supported by Grants-in-Aid for Scientific Research (B) (No. 15H03803), JSPS, to H.O. and by CREST and ACT-C, JST, to M.S.

Supporting Information




Is there **Supporting Information** to be published? Click here to indicate YES or NO (text and links will be updated prior to publication).

References

- [1] Breitmaier, E. *Terpenes: Flavors, Fragrances, Pharmaca, Pheromones*; Wiley-VCH: Weinheim, 2006. See also refs 5–7, 9 and references therein.
- [2] a) van Tamelen, E. E.; Schwartz, M. A. *J. Am. Chem. Soc.* **1965**, *87*, 3277.; b) Stork, G.; Grieco, P. A.; Gregson, M. *Tetrahedron Lett.*, **1969**, 1393.; c) Grieco, P. A.; Masaki, Y. *J. Org. Chem.*, **1974**, *39*, 2135.; d) Negishi, E.; Valente, L. F.; Kobayashi, M. *J. Am. Chem. Soc.* **1980**, *102*, 3298.; e) Negishi, E.; Liou, S.-Y.; Xu, C.; Huo, S. *Org. Lett.* **2002**, *4*, 261.
- [3] For selected papers, see: a) Trost, B. M.; Keinan, E. *Tetrahedron Lett.* **1980**, *21*, 2595.; b) Godschalx, J.; Stille, J. K. *Tetrahedron Lett.* **1980**, *21*, 2599.; c) Nakamura, H.; Bao, M.; Yamamoto, Y. *Angew.*

- Chem.* **2001**, *113*, 3308.; *Angew. Chem., Int. Ed.* **2001**, *40*, 3208.; d) Karlström, A. S. E.; Bäckvall, J.-E. *Chem. Eur. J.* **2001**, *7*, 1981.; e) Jiménez-Aquino, A.; Flegeau, E. F.; Schneider, U.; Kobayashi, S. *Chem. Commun.* **2011**, *47*, 9456.; f) Yuan, Q.; Yao, K.; Liu, D.; Zhang, W. *Chem. Commun.* **2015**, *51*, 11834.
- [4] Diner, C.; Szabo, K. J. *J. Am. Chem. Soc.* **2017**, *139*, 2.
- [5] a) Zhang, P.; Brozek, L. A.; Morken, J. P. *J. Am. Chem. Soc.* **2010**, *132*, 10686.; b) Zhang, P.; Le, H.; Kyne, R. E.; Morken, J. P. *J. Am. Chem. Soc.* **2011**, *133*, 9716.; c) Brozek, L. A.; Ardolino, M. J.; Morken, J. P. *J. Am. Chem. Soc.* **2011**, *133*, 16778.; d) Ardolino, M. J.; Morken, J. P. *J. Am. Chem. Soc.* **2014**, *136*, 7092.; e) Ardolino, M. J.; Morken, J. P. *Tetrahedron* **2015**, *71*, 6409.; f) Le, H.; Kyne, R. E.; Brozek, L. A.; Morken, J. P. *Org. Lett.* **2013**, *15*, 1432.; g) Wang, X.; Wang, X.; Han, Z.; Wang, Z.; Ding, K. *Angew. Chem.* **2017**, *129*, 1136.; *Angew. Chem., Int. Ed.* **2017**, *56*, 1116.
- [6] Hornillos, V.; Pérez, M.; Fañanás-Mastral, M.; Feringa, B. L. *J. Am. Chem. Soc.* **2013**, *135*, 2140.
- [7] Hamilton, J. Y.; Hauser, N.; Sarlah, D.; Carreira, E. M. *Angew. Chem.* **2014**, *126*, 10935.; *Angew. Chem., Int. Ed.* **2014**, *53*, 10759.
- [8] Yasuda, Y.; Ohmiya, H.; Sawamura, M. *Angew. Chem.* **2016**, *128*, 10974.; *Angew. Chem., Int. Ed.* **2016**, *55*, 10816.
- [9] For synthesis of chiral 1,5-diene derivatives through Cu-catalyzed enantioselective coupling with diboron, allene and allylic phosphate, see: Meng, F.; McGrath, K. P.; Hoveyda, A. H. *Nature* **2014**, *513*, 367.
- [10] For reviews on Cu-catalyzed allylic substitutions, see: a) Alexakis, A.; Bäckvall, J. E.; Krause, N.; Pàmies, O.; Diéguez, M. *Chem. Rev.* **2008**, *108*, 2796.; b) Harutyunyan, S. R.; den Hartog, T.; Geurts, K.; Minnaard, A. J.; Feringa, B. L. *Chem. Rev.* **2008**, *108*, 2824. c) Shintani, R. *Synthesis* **2016**, *48*, 1087.
- [11] For Cu-catalyzed enantioselective allylic substitutions using organoboron reagents and oxygen-functionalized NHC chiral ligands, see: a) Shintani, R.; Takatsu, K.; Takeda, M.; Hayashi, T. *Angew. Chem.* **2011**, *123*, 8815.; *Angew. Chem., Int. Ed.* **2011**, *50*, 8656.; b) Gao, F.; Carr, J. L.; Hoveyda, A. H. *Angew. Chem.* **2012**, *124*, 6717.; *Angew. Chem., Int. Ed.* **2012**, *51*, 6613.; c) Jung, B.; Hoveyda, A. H. *J. Am. Chem. Soc.* **2012**, *134*, 1490.; d) Takeda, M.; Takatsu, K.; Shintani, R.; Hayashi, T. *J. Org. Chem.* **2014**, *79*, 2354.; e) Shi, Y.; Jung, B.; Torker, S.; Hoveyda, A. H. *J. Am. Chem. Soc.* **2015**, *137*, 8948. See also ref 10c.
- [12] For our previous work, see: a) Harada, A.; Makida, Y.; Sato, T.; Ohmiya, H.; Sawamura, M. *J. Am. Chem. Soc.* **2014**, *136*, 13932.; b) Ohmiya, H.; Zhang, H.; Shibata, S.; Harada, A.; Sawamura, M. *Angew. Chem.* **2016**, *128*, 4855.; *Angew. Chem., Int. Ed.* **2016**, *55*, 4777.; c) Hojoh, K.; Ohmiya, H.; Sawamura, M. *J. Am. Chem. Soc.* **2017**, *139*, 2184.
- [13] For functionalized NHC ligands in asymmetric catalysis, see: a) Hameury, S.; Frémont, P.; Braunstein, P. *Chem. Soc. Rev.* **2017**, *46*, 632.; b) Pape, F.; Teichert, J. F. *Eur. J. Org. Chem.* **2017**, 4206.; c) Peris, E. *Chem. Rev.* 10.1021/acs.chemrev.6b00695
- [14] a) Shido, Y.; Yoshida, M.; Tanabe, M.; Ohmiya, H.; Sawamura, M. *J. Am. Chem. Soc.* **2012**, *134*, 18573.; b) Hojoh, K.; Shido, Y.; Ohmiya, H.; Sawamura, M. *Angew. Chem.* **2014**, *126*, 5054.; *Angew. Chem., Int. Ed.* **2014**, *53*, 4954. See also: c) Ohmiya, H.; Yokobori, U.; Makida, Y.; Sawamura, M. *J. Am. Chem. Soc.* **2010**, *132*, 2895.; d) Nagao, K.; Yokobori, U.; Makida, Y.; Ohmiya, H.; Sawamura, M. *J. Am. Chem. Soc.* **2012**, *134*, 8982.
- [15] The use of allyl-9-BBN reagent instead of **1a** under the conditions for Table 1, entry 12 resulted in decreases in both enantioselectivity (87% ee) and product yield (30%) but with the exclusive regioselectivity ($\gamma/\alpha > 99:1$) unchanged.
- [16] Scholl, M.; Ding, S.; Lee, C. W.; Grubbs, R. H. *Org. Lett.* **1999**, *1*, 953.
- [17] The absolute configurations of **3ae** and **3ah** were determined by comparison of the specific rotations with the values reported previously. See ref 5a. The absolute configuration of **3au** was determined by the Mosher's NMR spectroscopic method. Absolute configurations of the other products were assigned by consideration of the stereochemical pathway. See ref 8.
- [18] The linear α -substitution product was (*E*)-isomer.
- [19] Data are taken from ref 8 for Eq. 3–6.
- [20] Nakamura conducted DFT calculations on the mechanism of the reaction between [MeCu(CN)Li] and allyl acetate to form a square planar four-coordinate (γ - σ -enyl)copper(III) species [(π -en- σ -yl)copper(III) complex]. Our mechanistic proposal is in accord with the Nakamura's mechanism, in which the (γ - σ -enyl)copper(III) species is not in equilibrium with the corresponding (α - σ -enyl)copper(III) species; the regioselectivity is determined at the oxidative addition step as a consequence of the asymmetric nature of MeCuCN⁻. Our proposed mechanism is in accord with the Nakamura's mechanism in that the reaction proceeds through oxidative addition of a cuprate to form the (γ - σ -enyl)copper(III) species followed by reductive elimination. However, the coordination number of Cu in the allylcopper(III) complex is different by virtue of bidentate coordination of the anionic phenol-NHC chiral ligand (L). The strongly electron-donating NHC coordination should render the π -en coordination weaker, making the allylic 1,3-Cu-migration in the allylcopper(III) complex more feasible. See: a) Yoshikai, N.; Zhang, S.-L.; Nakamura, E. *J. Am. Chem. Soc.* **2008**, *130*, 12862. For the effect of a σ -donor ligand, see: b) Yamanaka, M.; Kato, S.; Nakamura, E. *J. Am. Chem. Soc.* **2004**, *126*, 6287.
- [21] Rapid formation of a tetravalent borate (**B**) was confirmed by ¹¹B NMR spectroscopy. See ref 8.
- [22] Zhang, P.; Roundtree, I. A.; Morken, J. P. *Org. Lett.* **2012**, *14*, 1416.

Biosketches

	<p>Yuto Yasuda graduated from Hokkaido University in 2014. He received his M. Sc. in 2016 from Hokkaido University under the supervision of Professor Masaya Sawamura. Currently, he is working on his Ph. D. thesis on the development of copper-catalyzed asymmetric C–C bond formation reactions. He is a JSPS research fellow (DC1) since 2016. Awards: CSJ Student Presentation Award (2014), CSJ Student Presentation Award (2016), Poster Award in 1st Singapore Japan Germany Trilateral Symposium on Precision Synthesis & Catalysis (2017).</p>
	<p>Hirohisa Ohmiya is a Professor at Kanazawa University. He received his Ph.D. degree from Kyoto University (Japan) in 2007 under the supervision of Professor Koichiro Oshima. He spent one year as a JSPS postdoctoral fellow in the group of Professor Timothy F. Jamison at Massachusetts Institute of Technology (USA). In 2008, he became an Assistant Professor at Hokkaido University (Japan) working with Professor Masaya Sawamura. He was promoted to Associate Professor in 2010. Since 2017, he has been a Full Professor at Kanazawa University. He received the Chemical Society of Japan Award for Young Scientists (2014) and the Young Scientist's Prize for the Commendation of Science and Technology by MEXT (2015).</p>
	<p>Masaya Sawamura is a Professor at Hokkaido University. He received his Ph.D. degree from Kyoto University in 1989 under the supervision of Professor Yoshihiko Ito. In 1989, he joined the faculty of the same department as an Assistant Professor. He spent one year as a researcher at Harvard University (Professor Stuart L. Schreiber, 1993–1994). In 1995, he moved to Tokyo Institute of Technology and to the University of Tokyo to join the group of Professor Eiichi Nakamura as an Assistant Professor. He was promoted to Lecturer in 1996 and to Associate Professor in 1997. Since 2001, he has been a Full Professor at Hokkaido University. He received the Chemical Society of Japan Award for Young Scientists (1996), the Chemical Society of Japan Award for Creative Work (2012), the SSOCJ (The Society of Synthetic Organic Chemistry, Japan) Nissan Chemical Industries Award for Novel Reaction & Method (2014), and the Nagoya Medal (Silver Medal, 2017).</p>

Include photos and short biographical text for all authors. If the photo is of a group of people, specify who is shown where.

Photographs (.jpg or .tif format) must be a minimum of 300 dpi in order to reproduce well in print.

Checklist (have these on hand for manuscript submission in ScholarOne):

- cover letter, including a statement of the work's significance
- full mailing address, telephone and fax numbers, and e-mail address of the corresponding author
- email address for each author
- original Word file
- original graphics files zipped into one zip file
- eye-catching graphical abstract as an individual file
- 5–8 key words
- separate Supporting Information file
- separate zipped Primary Data files including cover sheet (optional)