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Citation	Catalysis science and technology, 13(7), 2247-2254 https://doi.org/10.1039/d3cy00024a
Issue Date	2023-04-07
Doc URL	http://hdl.handle.net/2115/91243
Type	article (author version)
File Information	CST 13(7)2247 .pdf



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Catalysis of surface dispersed Cu^{2+} species on $t\text{-ZrO}_2$: Square-planar Cu catalyzed cross-coupling of arylboronic acid and imidazole

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Received 00th January 20xx,
Accepted 00th January 20xx

DOI: 10.1039/x0xx00000x

www.rsc.org/

We examined the Chan-Lam cross-coupling of imidazole and arylboronic acids under additive-free and mild conditions using heterogeneous CuO/MO_x catalysts such as metal oxide-supported CuO and Cu -doped metal oxides. Among them, $\text{CuO}/t\text{-ZrO}_2$ exhibited the best performance, which was more than ten times the higher turnover number in comparison with bulk CuO . By combining experimental and computational approaches, we elucidated the surface structure of $\text{CuO}/t\text{-ZrO}_2$. The catalyst contained abundant highly reducible Cu^{2+} species, in contrast to the other Cu^{2+} -containing catalysts. Notably, X-ray absorption spectroscopy and density functional theory calculations revealed the Cu^{2+} species to be square-planar [CuO_4]. Our results indicate that the highly active [CuO_4] species play a key role in the present coupling reaction.

Introduction

Heterogeneous copper catalysts are attracting increasing attentions in both academia and industry owing to their unique catalytic activity, easy product separation, and low cost.¹ In the field of organic synthesis, they are applied in a variety of catalytic reactions such as oxidation, reduction, cross-coupling, and click reaction.² However, insoluble heterogeneous catalysts exhibit much lower catalytic activities than homogeneous catalysts. Therefore, heterogeneous copper-mediated catalysis often requires elevated temperature, high catalyst loadings, and chemical additives. Metal oxide-supported CuO catalysts have been utilized to improve their catalytic activity because they can easily control structures and electronic properties of the copper species.^{2,3} Cu -doped metal oxide, which can modulate coordination modes, are also a robust catalyst.⁴ Hence, we supposed that a highly reactive Cu species on the metal oxide surface would promote organic transformation under mild conditions. Herein, we focus on surface dispersed Cu^{2+} species on tetragonal zirconia.

Cu/ZrO_2 catalysts have been applied to hydrogenation and dehydrogenation.⁵ Moreover, our group has reported the MeOH synthesis from carbon dioxide using a copper-doped zirconia catalyst

precursor.⁶ However, although crystal phases of ZrO_2 , mainly monoclinic and tetragonal phases, have rarely been studied in organic synthesis, they are thought to have a crucial effect on the catalytic activity. In 2021, Hu et al. demonstrated that monoclinic ZrO_2 -supported CuO ($\text{CuO}/m\text{-ZrO}_2$) exhibited superior performance in the hydrogenation of furfural and vanillin over Cu -doped tetragonal ZrO_2 ($\text{CuO}/t\text{-ZrO}_2$).^{5f} In contrast, $\text{CuO}/t\text{-ZrO}_2$ was found to be a more suitable catalyst for their acetalization. Cu -catalyzed Chan-Lam coupling, a carbon-heteroatom bond formation reaction using boronic acids, is one of the most important transformations for the preparation of biologically active compounds and functional materials.⁷ While a variety of homogeneous and heterogeneous Cu catalysts have been used for Chan-Lam coupling, CuO/ZrO_2 have never been explored. Thus, in this work, we demonstrate $\text{CuO}/t\text{-ZrO}_2$ catalyzed Chan-Lam coupling of arylboronic acids and imidazole under mild and additive-free conditions.⁸ Moreover, the highly reactive Cu species and catalytic cycle of the present reaction were investigated by using both experimental and computational approaches.

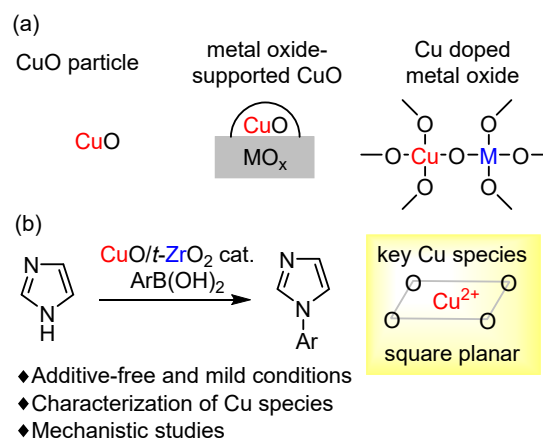


Fig. 1 (a) Heterogeneous Cu catalysts. (b) This work

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Electronic Supplementary Information (ESI) available: CCDC 1973631.

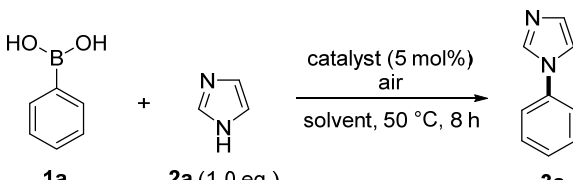
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Result and discussion

Catalyst performance

We commenced our study by screening catalysts for the Chan-Lam coupling of phenylboronic acid **1a** (0.2 mmol) and imidazole **2a** (0.2 mmol) in methanol and air at 50 °C for 8 h (entries 1–5, Table 1). When bulk CuO catalyst was used as the catalyst, the desired product **3a** was not formed (entry 1). When the CuO/TiO₂ catalyst was tested, **3a** was obtained in 23% yield (entry 2). The use of CuO/Al₂O₃ slightly improved the yield to 32% (entry 3), while Chan-Lam coupling catalyzed by CuO/*m*-ZrO₂ afforded **3a** in 42% yield (entry 4). Finally, CuO/*t*-ZrO₂ was determined to be the optimal catalyst for this transformation, affording **3a** in 81% yield (76% isolated yield, turnover number (TON) = 15, entry 5). This reaction did not proceed in THF and CH₃CN probably due to blocking of the copper coordination site (entries 6 and 7) and was significantly suppressed under argon atmosphere (20% yield, entry 8).

Table 1. Screening of catalysts and solvents^a



entry	catalyst	solvent	yield (%) ^b	TON
1	CuO	MeOH	<5	<1
2	CuO/TiO ₂	MeOH	23	4.6
3	CuO/Al ₂ O ₃	MeOH	32	6.4
4	CuO/ <i>m</i> -ZrO ₂	MeOH	42	8.2
5	CuO/ <i>t</i> -ZrO ₂	MeOH	81 (76)	16 (15)
6	CuO/ <i>t</i> -ZrO ₂	CH ₃ CN	<5	<1
7	CuO/ <i>t</i> -ZrO ₂	THF	<5	<1
8 ^c	CuO/ <i>t</i> -ZrO ₂	MeOH	20	4

^aReaction conditions: **1a** (0.2 mmol), **2a** (0.2 mmol), Cu catalyst (0.01 mmol, 5 mol%) in solvent (1.0 mL, 0.2 M) at 50 °C in air.

^b1,3,5-Trimethoxybenzene was used as an internal standard. Isolated yield is shown in parentheses. ^cUnder argon atmosphere.

Structural analysis by experiments

To consider the different activities of each Cu catalyst in Table 1, we analyzed surface Cu species. The CuO loading of the prepared catalysts was approximately 1 mmol g_{cat}⁻¹, as measured by XRF (Table S1, ESI). First, we examined the crystal structure of the as-prepared catalysts using XRD (Fig. 2). The XRD pattern of each catalyst presented peaks corresponding to its support. On the other hand, CuO peaks were either weak or absent, indicating that the CuO nanoparticles are smaller than the XRD detection limit (<10 nm) and/or that the Cu species are different from those in bulk CuO. In the latter case, they can be incorporated into the lattice of metal oxide supports. Fig. 3a shows the Cu-K-edge X-ray absorption near edge structure (XANES) spectra of the Cu-based catalysts. As references, data for CuO, Cu₂O, and Cu foil were also collected. The peak at ~8980 eV was attributed to the 1s → 4pπ transition, whereas that at ~8992 eV was attributed to the 1s → 4pσ transition.⁹ For all the Cu-based catalysts except for CuO/*t*-ZrO₂, the XANES spectra were similar to those of CuO. For CuO/*t*-ZrO₂, the XANES spectrum displayed a white line characterized by double peaks at 8988 and

8993 eV and no shoulder peak at 8980 eV. Thus, we supposed that the Cu species in CuO/*t*-ZrO₂ were different from those in CuO. Fig. 3b shows the Cu K-edge radial structure functions (RSFs). A strong backscattering peak at 1.5 Å was assigned to the nearest-neighbor Cu–O distance of 1.96 Å (see Table S2). The peak at 2.5 Å was assigned to the second-neighbor CuO, and first- and second-neighbor Cu–Cu coordinations. The fitting results revealed that CuO comprised distorted octahedral clusters [CuO₆]. For CuO/TiO₂ and CuO/*m*-ZrO₂, the RSF was similar to that of CuO. Therefore, we concluded that these samples contained CuO nanoparticles supported by the corresponding metal oxides. Both the XRD (Fig. 2) and RSF (Fig. 3b) indicate that the CuO nanoparticle size was <10 nm. In the case of CuO/Al₂O₃, the second peak at 2.5 Å was weak. Following fitting, we could only recognize the peak of the second-Cu–O coordination but not those of the first- and second-neighbor Cu–Cu coordinations (Table S2). This indicated the sample mainly contained isolated octahedral clusters [CuO₆]. Notably, for CuO/*t*-ZrO₂, the peak at 2.5 Å was not observed. The coordination number of Cu–O was 3.9±0.4, indicating that the Cu species in CuO/*t*-ZrO₂ was a four-fold coordinated Cu cluster [CuO₄]. As shown in Figs. S1 and S2, the RSF and XRD of the spent CuO/*t*-ZrO₂ were similar to that of the as-prepared CuO/*t*-ZrO₂, which confirms that the [CuO₄] species was present even after the reaction.

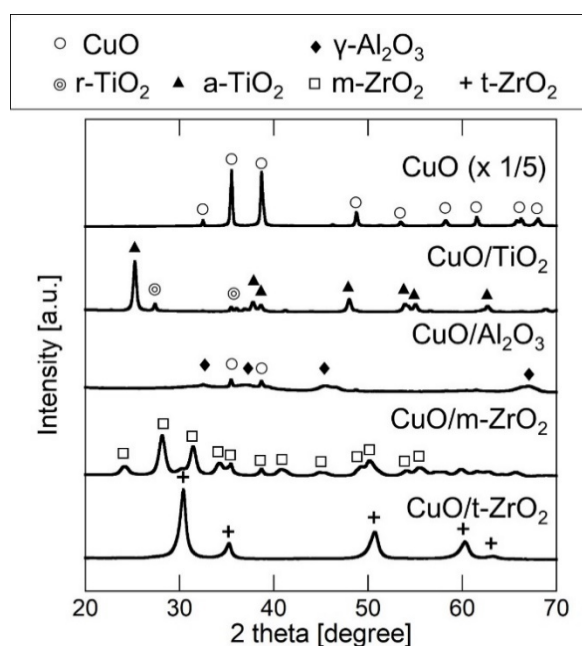


Fig. 2 XRD patterns of the prepared catalysts and CuO.

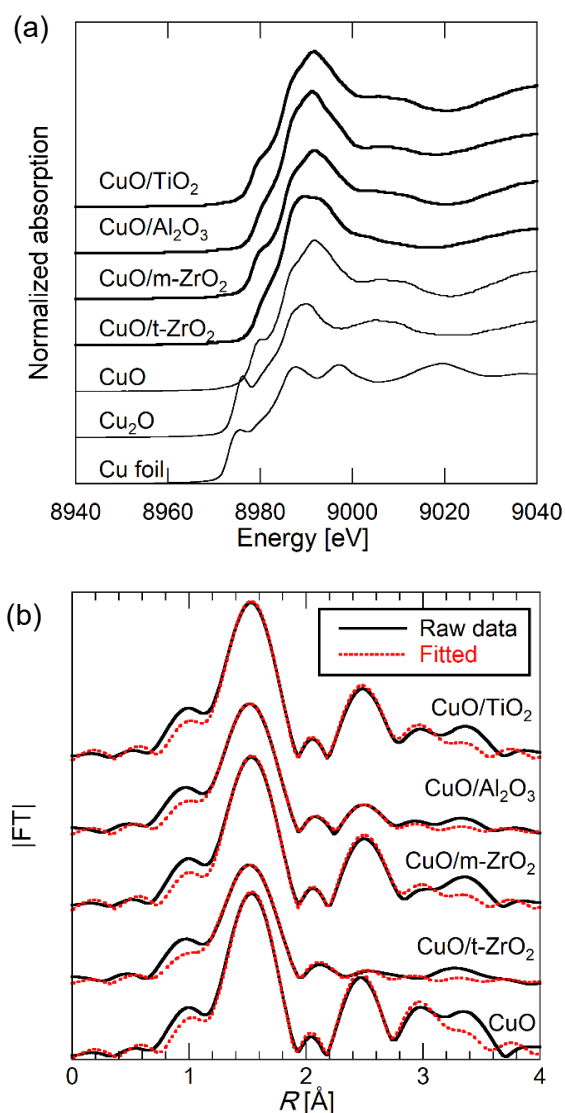


Fig. 3 (a) Cu-K edge XANES spectra of the Cu-based catalysts, CuO, Cu₂O, and Cu foil. (b) Fourier transforms of the k^3 -weighted extended X-ray absorption fine structure (EXAFS) oscillations measured at room temperature near the Cu K-edge of the as-prepared catalysts, Cu foil, Cu₂O, and CuO; k range = 30–110 nm⁻¹ and R range = 1.0–3.4 Å.

Fig. 4 shows high-angle annular dark field scanning transmission electron microscopy (HAADF-STEM) images of as-prepared CuO/*t*-ZrO₂. We can observe several particles with the size = ca. 10 nm. Each particle contained Zr and Cu species, according to the elemental mapping. Additionally, we checked the morphology of the spent CuO/*t*-ZrO₂. The catalyst was used for the above Chan-Lam coupling (Entry 5). Of note, only the nanoparticles containing Zr and Cu were seen in Fig. 4b, which means that Cu species are well-dispersed even after the reaction. Hence, the HAADF-STEM observation is in line with the results of Cu K-edge XAS (Figs. 3 and S1).

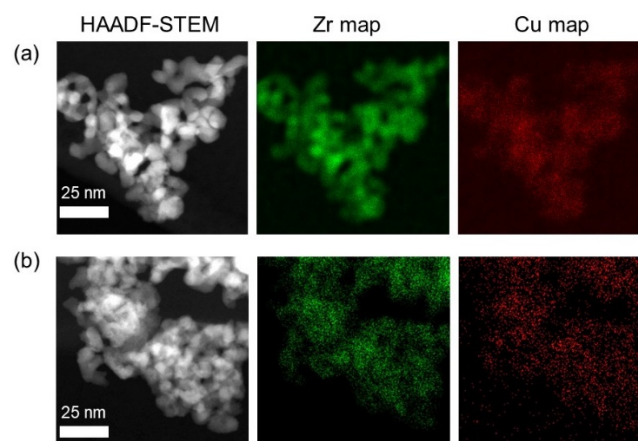


Fig. 4 HAADF-STEM images and elemental mapping of (a) as-prepared and (b) spent CuO/*t*-ZrO₂.

Thus, we determined that the types of MO_x determines the Cu coordination structure. We next proceeded to investigate whether reducibility of the Cu²⁺ species is related to the coordination structure. Such knowledge can be informative because the proposed reaction pathway of Cu-catalyzed Chan-Lam coupling is considered to include a redox cycle of the Cu species.⁷ Thus, we performed H₂-TPR to analyze the reducibility of the Cu²⁺ species in the as-synthesized Cu-based catalysts. Fig. 5 shows the H₂-TPR profiles of CuO and the prepared catalysts. For CuO, the broad reduction peak at 250–400 °C indicated that the Cu²⁺ species in CuO was reduced by H₂ above 250 °C. For the Cu-based catalysts, the reduction peaks were observed below 250 °C. Hence, we supposed that the interaction between Cu and the metal oxides improves the reducibility of the Cu²⁺ species. Two peaks, namely a low-temperature and a high-temperature peak were observed for

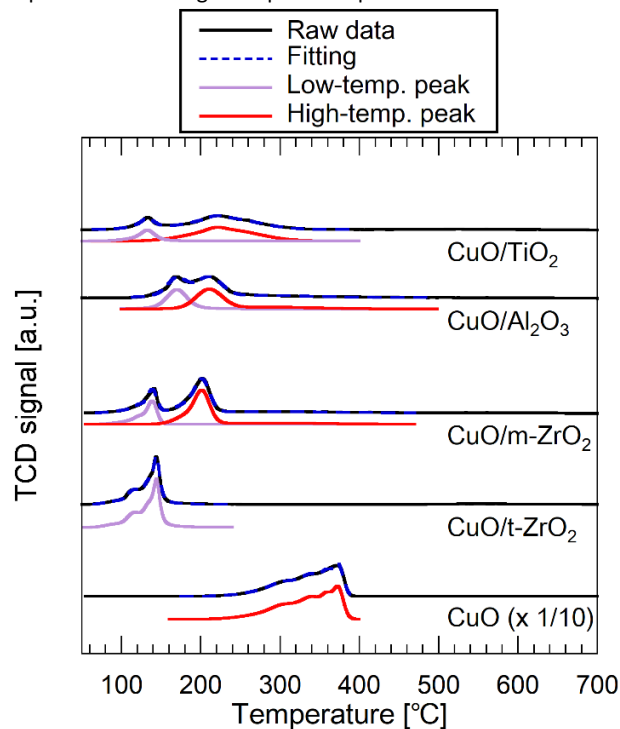


Fig. 5 H₂-TPR profiles of the Cu-based catalysts and CuO (reference).

Table 2. H₂ consumption in the H₂-TPR experiment.

sample	H ₂ consumption [mmol g _{cat} ⁻¹]		
	Low temp. ^a	High temp. ^b	Total
CuO/TiO ₂	0.21	0.70	0.91
CuO/Al ₂ O ₃	0.42	0.60	1.02
CuO/ <i>m</i> -ZrO ₂	0.30	0.68	0.98
CuO/ <i>t</i> -ZrO ₂	0.71	0	0.71

^a Based on the low-temperature peak in Fig. 5. ^b Based on the high-temperature peak in Fig. 5.

CuO/*m*-ZrO₂, CuO/TiO₂, and CuO/Al₂O₃. Approximately 30% of the Cu²⁺ species were reduced below 150 °C, whereas the rest of the species were reduced at ~200 °C. Most of the Cu²⁺ species were expected to be reduced to metallic Cu by H₂ molecules below 500 °C because the total H₂ consumption was almost the same as the Cu loading (1 mmol g_{cat}⁻¹). For CuO/*t*-ZrO₂, only the peak below 150 °C was observed. Notably, according to Table 2, the H₂ consumption based on the corresponding peak area (0.71 mmol g_{cat}⁻¹) was smaller than that of the Cu loading (1 mmol g_{cat}⁻¹). This phenomenon was also predicted in our previous work.^{6b} We then carried out XANES analysis of the H₂-reduced CuO/*t*-ZrO₂ based on the linear combination fitting. While 68% of the Cu²⁺ species were reduced to Cu⁰, the remainder unchanged. These results indicated that ~70% of the Cu²⁺ species in CuO/*t*-ZrO₂ are highly reducible, while the others exist as ultra-stable Cu²⁺ species. Our previous hypothesis was in line with the present H₂-TPR results.

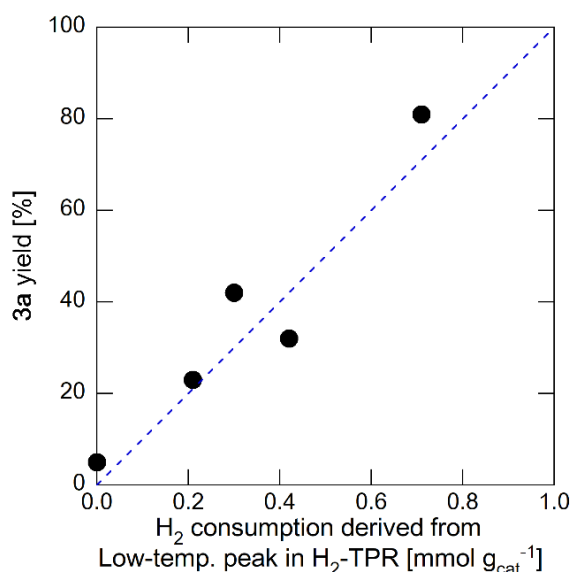


Fig. 6 Plots of the product (**3a**) yield against H₂ consumption calculated from the low-temperature H₂-TPR peaks.

Of note, we found the relationship between the product yield and Cu²⁺ reducibility. Fig. 6 shows the plots of the yield of **3a** against the H₂ consumption calculated from the low-temperature H₂-TPR peaks. These two values are positively correlated. As described above, the low-temperature peak is assignable to the reduction of highly reducible Cu²⁺ species. The peak area is related to the number of the

Cu²⁺ species. Hence, we concluded that the highly reducible Cu²⁺ species are essential to Chan-Lam coupling.

Structural analysis using DFT calculations

We have experimentally revealed that the Cu²⁺ species in CuO/*t*-ZrO₂ are unique and essential for Cu-catalyzed Chan-Lam coupling. To gain molecular-level insight into the surface of CuO/*t*-ZrO₂, we next performed DFT calculations.

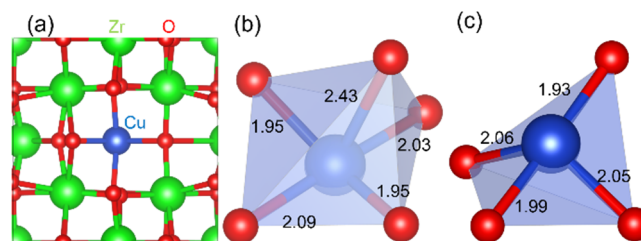


Fig. 7 Structures of (a) bulk CuO/*t*-ZrO₂, (b) CuO₅ and (c) CuO₄ clusters drawn by VESTA.¹⁰ Black lines are the boundaries of a supercell in (a) and bond lengths (Å) are shown in (b) and (c). Color code: green, Zr; blue, Cu; and red, O.

Fig. 7 shows the optimized structures of the bulk CuO/*t*-ZrO₂. The Cu atom is penta-coordinated (Fig. 7b) in the lowest-energy structure, in contrast to the four-fold coordination of Cu in CuO and the eight-fold coordination of Zr in *t*-ZrO₂. However, we also observed a four-fold-coordinated Cu cluster (Fig. 7c) as a metastable structure (energetically slightly higher than 0.04 eV per formula unit). The examined bulk CuO/*t*-ZrO₂ structures displayed a 1:1 five- to four-fold coordination ratio, indicating that the two species were almost mixed equally in the bulk. Notably, the four-fold-coordinated Cu cluster shown in Fig. 7c is not completely square planar but rather a distorted tetrahedral.

Fig. 8 shows the surface structure of the solid-solution CuO/*t*-ZrO₂ surface. The coordination number of the surface Cu atom is four, which is less than the five observed in the most stable bulk structure. The Cu–O bond lengths of the [CuO₄] cluster shown in Fig. 8b range from 1.94 to 2.00 Å. The shape of the [CuO₄] cluster is almost square planar, which is different from that of bulk CuO/*t*-ZrO₂.

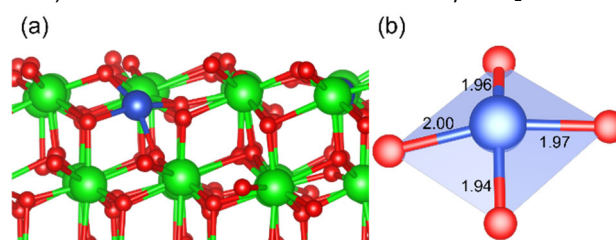


Fig. 8 Side view of the CuO/*t*-ZrO₂ surface and [CuO₄] cluster. The bond lengths (Å) are shown in (b).

Reaction mechanism.

We next considered the reaction mechanism of Cu-catalyzed Chan-Lam coupling over CuO/*t*-ZrO₂. One of the important factors here is the positive correlation between the product yields and Cu²⁺ reducibility (Fig. 6). This suggests that the redox cycle of the Cu²⁺ species is essential for Cu-catalyzed Chan-Lam coupling to occur. We therefore adopted a computational approach to further understand the role of the Cu²⁺ species at the molecular level. Fig. 9 shows that the adsorption of both the imidazole and benzene molecules on the surface Cu sites of the CuO/*t*-ZrO₂ surface is exothermic, with an

adsorption energy of -0.37 eV. When the adsorbates adsorb to the surface Cu atom, the adsorbates transfer protons to the O atoms at the catalyst surface. The adsorbed structure in Fig. 9 exhibits characteristic Cu coordination: The surface Cu atom forms four bonds with C (in benzene), N (in imidazole), and two surface O atoms (in the $[\text{CuO}_4]$ cluster) atoms by breaking two Cu–O bonds in the $[\text{CuO}_4]$ cluster at the CuO/*t*-ZrO₂ surface. Notably, the surface Cu atom moves away from the surface by 1.54 Å along the surface normal direction upon adsorption. The Cu–O bond lengths were 1.84 Å and 2.16 Å, and their strength was almost the same as that at the CuO/*t*-ZrO₂ surface (Fig. 8). These findings indicate that the imidazole and benzene molecules adsorb onto the Cu active site that detaches from the $[\text{CuO}_4]$ cluster. After the product (**3a**) is formed, the Cu atom returns to the $[\text{CuO}_4]$ cluster, as shown in Fig. 8.

$$E_{\text{ads}} = -0.37 \text{ eV}$$

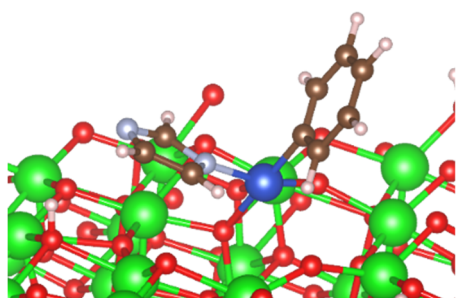


Fig. 9 Adsorption structure of imidazole and benzene molecules at the CuO/*t*-ZrO₂ surface. The adsorption energy (E_{ads}) is also displayed. Color code: green, Zr; blue, Cu; red, O; brown, C; light blue, N; and white, H.

Next, we assumed a reaction mechanism for the Chan-Lam coupling of phenylboronic acid **1a** and imidazole **2a** using the CuO/*t*-ZrO₂ catalyst (Fig. 10).^{7b} Initially, transmetalation of **2a** to form the Ph-imidazole-CuO₂ intermediate from $[\text{CuO}_4]$. The oxidation of Cu²⁺ to Cu³⁺ occurs by O₂ in air.¹¹ Subsequent reductive elimination provides the desired product **3a** and Cu¹⁺ species. Finally, the initial $[\text{CuO}_4]$ species is regenerated through the aerobic oxidation of Cu¹⁺.¹²

Fig. 10 Proposed reaction mechanism

Based on the proposed reaction mechanism, analytical results, and DFT calculations, we assumed that the highly reducible square planar $[\text{CuO}_4]$ in CuO/*t*-ZrO₂ exhibited superior performance because of smooth transmetalation from Cu–O bondings to Cu–C and Cu–N

bondings. In contrast, the octahedral $[\text{CuO}_6]$ in CuO/Al₂O₃ exhibited lower reactivity for Chan-Lam coupling, probably because of the less reducible Cu species. The reactions using CuO/TiO₂ and CuO/*m*-ZrO₂ showed lower reactivities because the Cu species were not doped, but supported on metal oxides.

Substrate scope and recycling catalyst.

Under the optimal reaction conditions listed in Table 1, we next examined the substrate scope (Scheme 1). Chan-Lam coupling of 4- and 3-tolylboronic acids **1b** and **1c** with imidazoles **2** provided the corresponding products **3b** and **3c** in 93% and 99% yields, respectively. Sterically hindered 2-tolylboronic acid **1d** was also tolerated, affording the *N*-arylated imidazole **3d** in 76% yield. When 4-chloro- and 4-methoxyphenylboronic acids **1e** and **1f** were used as substrates, the desired products **3e** and **3f** were obtained in 71% and 77% yields, respectively. 1- and 2-Naphthylboronic acids **1g** and **1h** were successfully converted to the corresponding products **3g** and **3h** in good yields (**3g**: 77%, **3h**: 81%). The reaction using 3-thienylboronic acid as a heterocyclic substrate yielded the product **3i** in 37% yield. Although the cross coupling of **1b** with 2-methyl imidazole **2b** provided the desired product **3j** in 28% yield, benzimidazole **2c** was not a suitable substrate, probably due to steric bulkiness. When we reused CuO/*t*-ZrO₂ catalyst after Chan-Lam coupling of **1a** and **2a**, the yield of **3a** was decreased to 58%.

Scheme 1. Scope of substrates using CuO/*t*-ZrO₂ catalyst^{a,b}

^aReaction conditions: Reaction was conducted using **1** (0.2 mmol), **2** (0.2 mmol), and the catalyst (0.01 mmol, 5 mol%) in MeOH (1.0 mL, 0.2 M) at 50 °C for 8 h in air. ^bIsolated yield.

Conclusions

We demonstrated the efficient Chan-Lam coupling of imidazole and boronic acids when catalyzed by CuO/*t*-ZrO₂. The reaction proceeded smoothly under air atmosphere and mild conditions in the absence of external additives, affording the corresponding *N*-arylated compounds (TON = 15 for the synthesis of **3a**). Of note, the TON of CuO/*t*-ZrO₂ was more than ten times higher than that of CuO nanoparticles. XRD, XAS, and DFT structural analyses characterized the square planar Cu species, [CuO₄], dispersed on the CuO/*t*-ZrO₂ surface. H₂-TPR analyses indicated that the [CuO₄] species are highly reducible, and thus, that they may have good redox properties. DFT calculations also suggested the co-adsorption of imidazole and benzene on the surface of the Cu site, which can be an essential step in Chan-Lam coupling. Notably, similar to ligands for homogeneous catalysts, the types of support materials for heterogeneous catalysts determine the coordination structure of the Cu species: CuO nanoparticles are formed on the TiO₂ and *m*-ZrO₂ surfaces, octahedral [CuO₆] species are formed in Al₂O₃, and [CuO₄] species are selectively formed on the *t*-ZrO₂ surface. This study opens up the possibility that heterogeneous copper catalysts can be applied in organic synthesis reactions. Further catalytic applications of CuO/*t*-ZrO₂ and mechanistic studies are ongoing in our laboratory.

Experimental section

Catalyst preparation

In this study, four types of CuO/MO_x catalysts (MO_x = TiO₂, Al₂O₃, *m*-ZrO₂, and *t*-ZrO₂) were prepared using an incipient wetness impregnation method; Evonik P-25 was used as the TiO₂. Al₂O₃ (JRC-ALO-8) and *m*-ZrO₂ (JRC-ZRO-3) were supplied by the Catalysis Society of Japan, while amorphous ZrO₂ (NND) was provided by Daiichi Kigenso Kagaku Kogyo. First, MO_x was impregnated with an aqueous solution of Cu(NO₃)₂·3H₂O (Fujifilm Wako, Japan). The obtained powder was then dried at 110 °C and subsequently the powder was calcined at 500 °C for 3 h. Their CuO contents was 1 mmol g⁻¹ and, commercial CuO was used as the reference (Fujifilm Wako, Japan).

Characterization of catalysts

X-ray diffraction (XRD), X-ray fluorescence (XRF), X-ray absorption spectroscopy (XAS), scanning transmission electron microscopy (STEM), and H₂ temperature-programmed reduction (H₂-TPR) were carried out to characterize the as-synthesized catalysts. The detailed procedures are described in the ESI.

Reaction test

Each test-tube was charged with phenylboronic acid **1a** (0.2 mmol, 24.4 mg), imidazole **2a** (0.2 mmol, 13.6 mg) and the catalyst (0.01 mmol Cu) in 1 mL MeOH. The resulting solution was stirred at 50 °C in air. After 8 h, the organic layer was separated with the Cu catalyst and concentrated under reduced pressure. The residue was purified by silica gel column chromatography (hexane:ethyl acetate ratio = 5:5) to obtain the pure product **3a**. TON values were determined by dividing the quantity of **3a** produced after 8 h by the moles of the Cu catalyst.

Density functional theory (DFT) calculations

DFT calculations were performed using the Perdew-Burke-Ernzerhof (PBE)¹³ exchange and correlation functional and D3 correction¹⁴ using the cp2k program package,¹⁵ based on the computational procedures reported in our previous work.¹⁶ The Goedecker-Teter-

Hutter pseudopotentials¹⁷ and a double- ζ valence plus polarization (DZVP-MOLOPT-SR-GTH)¹⁸ Gaussian basis set for the orbitals were employed. The cutoff associated with mapping Gaussians onto a multi-grid was 60 Ry and the planewave cutoff for the finest level of the multi-grid was 400 Ry. We employed 3 × 3 × 4 Monkhorst-Pack¹⁹ *k*-point meshes and Γ point for the *k*-point samplings of the bulk and surface, respectively. The target accuracy for the SCF (self-consistent field) convergence was 10⁻⁵.

DFT calculations of bulk CuO/*t*-ZrO₂

The initial bulk geometry of the supercell was generated using a periodic 2 × 2 × 1 supercell of *t*-ZrO₂.²⁰ Fermi-Dirac smearing was employed at an electronic temperature of 300 K. To model the solid solution, the Zr and O atoms were replaced with a Cu atom and oxygen vacancy (V_O), respectively, to maintain the neutrality of the system. Bulk CuO/*t*-ZrO₂ was prepared by ZrO₂ with a Cu-V_O pair to afford Cu_xZr_{1-x}O_{2-x} with a molar ratio of $x = 1/8 = 0.125$. We first replaced an O atom in the supercell (Zr₈O₁₆) with V_O. Using the optimized structure with the lowest energy, we next replaced the Cu atom with a Zr atom to form Zr₇CuO₁₅ that has the lowest energy.

DFT calculations of surface CuO/*t*-ZrO₂

The *t*-ZrO₂ (101) surface was generated from the optimized bulk structure. The rectangular slab consisted of Zr₄₀O₈₀ and the dimensions of the super cell were 12.702 × 14.472 Å² and 35.000 Å along the surface-parallel and surface-normal directions, respectively. The bottom ZrO₂ layer of the six layers was fixed to the bulk structure, and the geometry of the remaining atoms was optimized. To prepare CuO/*t*-ZrO₂, we replaced two Cu-V_O pairs in the top layer of the zirconia surfaces. The molar ratio of the top two ZrO₂ layers in Cu_xZr_{1-x}O_{2-x} was $x = 1/8 = 0.125$. The Supporting Information details the computational conditions and results of the formation free energy of the CuO/*t*-ZrO₂ surface,²¹ and the atomic charges around the Cu atoms by Hirshfeld analysis²² using a hybrid functional (PBE + 10.5% Hartree-Fock exchange).²³

DFT calculations of adsorption energy

The adsorption energy was calculated using the equation.

$$E_{\text{ads}} = E_{\text{substrate+adsorbate}} - E_{\text{substrate}} - E_{\text{adsorbate}}$$

where $E_{\text{substrate+adsorbate}}$, $E_{\text{substrate}}$, and $E_{\text{adsorbate}}$ are the energies of the substrate and adsorbate, substrate, and adsorbate, respectively. The adsorbates were benzene and imidazole molecules to model the reactants of the reaction considered in this study.

Conflicts of interest

There are no conflicts of interest to declare.

Acknowledgements

The authors gratefully acknowledge generous support by JSPS KAKENHI Grant Numbers 21K04988, Research Foundation for JFE 21st Century Foundation. The authors thank the Center for Instrumental Analysis, Ibaraki University, for its help with XRD and XPS. The synchrotron radiation experiments were performed at BL14B2 of SPring-8 with the approval of the Japan Synchrotron Radiation Research Institute (JASRI, Proposal no. 2022A1782). We used the supercomputers of the Institute for Solid State Physics, the University of Tokyo, and the Research Institute for Information Technology, Kyushu University.

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