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Author(s)	Tsuchida, T.; 土田, 猛; Kakuta, T.
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Fabrication of SPS compacts from NbC-NbB₂ powder mixtures synthesized by the MA-SHS in air process

Takeshi Tsuchida* and Tsuyoshi Kakuta

Division of Materials Chemistry, Graduate School of Engineering, Hokkaido University, N13 W8, Kita-ku, Sapporo, 060-8628, Japan

Abstract

When the powder mixture of Nb/B/C=2/2/1 in molar ratio was mechanically activated (MA) by ball milling for 105 min and then exposed to air, it self-ignited spontaneously and the self-propagating high-temperature synthesis (SHS) was induced. The powder mixture of NbC and NbB₂ was obtained by this MA-SHS in air process and then sintered by spark plasma sintering (SPS) at a temperature as low as 1800°C. The sintered NbC/NbB₂ composite compact showed the fine, homogeneous microstructure consisting of the a few μm-sized grains and the Vickers hardness of 19.8 GPa, which was higher than 14.3 GPa and 16.6 GPa for the NbC/NbB₂ composite compacts sintered from the powder mixture of Nb/B/C=2/2/1 without MA-SHS and from the commercial powder mixture of NbC/NbB₂=1/1.

Keywords:

Mechanical activation (MA); Self-propagating High-temperature Synthesis (SHS); Spark plasma sintering (SPS); Ball milling; NbC/NbB₂ composite; Nb/B/C powder mixtures; Vickers hardness

*Corresponding author: Takeshi Tsuchida

Tel/Fax: +81 11 706 6578

e-mail address: tsuchida@eng.hokudai.ac.jp

1. Introduction

Carbides and borides of niobium have excellent properties such as high melting temperature, high strength, high thermal and electrical conductivity, and chemical stability. Therefore, the use of these ceramics in composites can be expected to offer potential candidates for a variety of high-temperature structural applications. Actually, the mechanical properties of the composites, such as mechanical strength, hardness and fracture toughness are controlled by densification and microstructure of the sintered bodies. In particular, because it is known that ceramic materials with fine microstructures, especially nanocomposites, exhibit improved mechanical properties [1], the preparation of the fine particles of NbC and NbB₂ and the homogeneous mixing between them are definitely important. We have so far developed a mechanical activation assisted self-propagating high-temperature synthesis (MA-SHS in air process), and successfully applied to the synthesis of carbides and nitrides of Al, Zr and Nb [2-9], and borocarbides of Al₃BC [10]. This technique is based on SHS induced by self-ignition reaction of disordered carbon when the mechanically activated metal-graphite powder mixtures were exposed to air [2, 4]. Thus, this process does not need any expensive reaction equipments and is a very simple, energy-saving technique. In the latest studies [9,11], it is expected that the fine powder mixtures of ZrC-ZrB₂ and NbC-NbB₂ synthesized by the MA-SHS in air process from the powder mixtures of Zr/B/C and Nb/B/C are a promising candidate as precursor of the ZrC/ZrB₂ and NbC/NbB₂ composites.

On the other hand, there is another problem that the sintering of NbC and NbB₂ powders to form the composite is a difficult process because of the high melting point and high vapor pressure of the constituents. Spark plasma sintering (SPS) is one of the most powerful candidates to densify poorly sinterable materials such as NbC and NbB₂. SPS is a noble sintering process that allows compaction of ceramic powders at very lower temperature with shorter holding time than in more conventional sintering process, and then depresses grain growth. The sample is loaded in a graphite die, uniaxial pressure is applied during the sintering by the top and bottom punches, and a pulsed direct current is allowed to pass through the die and the sample. The pulses generate spark discharge and/or plasma between powder particles. Because the die acts also as a heating source, the sample is effectively heated from both outside and inside. Although the mechanism for densification and grain growth in the SPS process have not been well elucidated yet, Tokita [12] pointed out that the generated spark discharge and/or plasma cleans the surfaces from adsorbed species, such as

CO₂, H₂O, OH⁻, so that enhances the grain-boundary diffusion process for densification. A systematic study of various spark plasma sintering (SPS) parameters such as temperature, holding time, heating rate, pressure, and pulse sequence was conducted by Shen et al. [13] in order to investigate their effect on the densification, grain-growth kinetics, hardness, and fracture toughness of a commercially available submicrometer-sized Al₂O₃ powder.

In the present study, the powder mixture of NbC and NbB₂ synthesized by the MA-SHS in air process from the powder mixture of Nb/B/C=2/2/1 was sintered by SPS to obtain the NbC/NbB₂ composite compact. Then, the Vickers hardness was compared with those of the NbC/NbB₂ composite compacts obtained from the powder mixture of Nb/B/C=2/2/1 without MA-SHS and the commercial powder mixture of NbC/NbB₂=1/1, and discussed on the basis of the microstructures of the sintered bodies.

2. Experimental procedures

The powders of niobium metal (particle size of 50~130 μm, 98% purity, Kojundo Chemical Laboratory), amorphous boron (practical grade, Sigma Chemical Company) and natural graphite (mean flake size 5 μm, 97% carbon, 2% ash and 1% volatile component, Nippon Kokuen Industry) were used as starting materials. The powders were mixed in a molar ratio of Nb/B/C=2/2/1 in an agate mortar, loaded in a p-7 planetary ball mill (Fritsch, Idar-Oberstein, Germany) in air and then ground. The mixing ratio was supposed to be a corresponding ratio of NbC/NbB₂=1/1 as the product phases. After grinding for 105 min, the ground sample was transferred into a graphite crucible and exposed to air. Just then, it self-ignited and the exothermic reactions propagated into the reactant powders. As soon as the reactions started, the graphite crucible was covered with another one to prevent the sample from further oxidizing. The particle size and morphology of starting materials, and the detail of the MA-SHS in air process have been previously described [9]. The obtained powder mixtures (which is called MA-SHS sample hereafter) were loaded in a graphite die (10 mm in diameter) and spark plasma sintered at 1600-1800°C for 10-30 min under a vacuum of 6-10 Pa with SPS apparatus (SPS-510L, Sumitomo Coal Mining Co., Tokyo, Japan). The temperature was measured with an optical pyrometer focused on the surface of the graphite die and automatically heated from room temperature to 600°C at heating rate of 120°C min⁻¹ and then up to 1800°C at 60°C min⁻¹. A pressure of 40 MPa and an on-off pulse sequence of

12:2 were applied to the sample during entire sintering time. At the end of the holding time, the pressure was released and the current was shut off. For the comparison, the following samples were also sintered by SPS: the powder mixture of Nb/B/C=2/2/1 which was merely mixed in an agate mortar without mechanical activation (which is called the mixing sample); the powder mixture of NbC/ NbB₂=1/1 which was obtained by mixing the commercially available samples in an agate mortar (which is called the commercial sample). The purity of the commercial NbC and NbB₂ powders was 99 % (Kojundo Chemical Laboratory) and their particle size was estimated to be < 2 μm from SEM microphotographs shown in Fig.1.

The density of the sintered samples was measured by Archimedes method. Before X-ray diffraction (XRD), scanning electron microscopy (SEM) and the Vickers hardness (HV) measurements, the surface of the sintered samples was mirror-like polished with 1.0 μm diamond slurries. The crystalline phases were identified by XRD (RINT-2000, Rigaku Denki) using Ni-filtered CuKα radiation (30 kV, 15 mA). The microstructures were observed on the polished surfaces and the fracture surfaces with scanning electron microscopy (JSM-5410-SEM-EPMA-WDX combined microanalyzer, JEOL). For the SEM observation, the powder sample was dispersed in ethanol by ultrasonication, and was placed as a drop on a glass plate or a brass stub and then dried. The Vickers hardness at room temperature was evaluated by Vickers indentation technique at a load of 4.9 MPa. More than five indents were made at the middle of each sample.

3. Results and discussion

The reaction mechanism of the MA-SHS in air process for the formation of NbC and NbB₂ from the powder mixture of Nb/B/C=2/2/1 has been already discussed in the previous paper [9]. In the MA-SHS in air process, it was found to be essential that graphite is converted into finely divided, disordered carbon by mechanical activation and mixed with the particles of Nb and B in an intimate contact state.

Fig.2 shows the XRD patterns of powder mixture of Nb/B/C=2/2/1 ground for various times. The disappearance of the graphite peaks, and the decreasing in intensity and broadening of the Nb peaks with increasing grinding time indicate the formation of disordered carbon, and the size reduction and the increase in the lattice strain in Nb particles. After the grinding of 105 min, the powder mixture self-ignited in air and the subsequent exothermic

reactions violently occurred evolving white heat. The evolution of the formation enthalpies of NbC and NbB₂, i.e., -140 kJ mol⁻¹ and -250 kJ mol⁻¹ in the temperature range of 300-1500°C, sustained completely the SHS process. After the reaction, as shown in Fig.2, NbC and NbB₂ were detected as product phases. In addition, a very small amount of Nb metal remained unreacted, suggesting that a small part of amorphous boron and disordered carbon remained unreacted as well. From the secondary electron image (SEI) and associated elemental maps of the MA-SHS product shown in Fig.3, it was found that the particles of NbC and NbB₂ obtained by the MA-SHS in air process are a few μm-sized grains and homogeneously mixed each other, therefore they are expected to be a promising candidate as precursor of the NbC/NbB₂ composites.

Fig.4 and Fig.5 show the XRD patterns and the SEM microphotographs of the NbC/NbB₂ composite compacts sintered by SPS at 1800°C for 10 min under a vacuum of 6-10 Pa from three samples, (i) the MA-SHS sample, (ii) the mixing sample and (iii) the commercial sample. As can be seen from the XRD results, the peaks of NbC and NbB₂ obtained from the samples (ii) and (iii) were detected in the almost same intensity (Fig.4c, d), whereas the phases obtained from the sample (i) was identified in a following order of intensity, NbC > NbB₂ >> Nb₃B₄ > Nb (Fig.4b). From a decrease in intensity of unreacted Nb peaks after SPS runs, it is estimated that the following reactions among unreacted Nb, C and B elements were carried out, Nb + C → NbC and Nb + B → Nb₃B₄. Furthermore, according to EPMA observation, the light phase and the grey phase in Fig.5 are NbC and NbB₂, respectively. Among these composite compacts, the very fine, homogeneous microstructure can be observed in the MA-SHS and commercial sample compacts in Figs.5(a) and (c), and a lot of pores of ~ 1 μm are also seen in the former compact. The grain size of NbC and NbB₂ is < 5 μm and is nearly equivalent to the particle size of the samples before sintering as shown in Figs.1 and 3, indicating retardation of grain growth during SPS process. In contrast, the compact from the mixing sample of Nb/B/C=2/2/1 in Fig.5(b) showed the relatively heterogeneous microstructure containing the grain > 10 μm. The relative density and the Vickers hardness, HV of these compacts are summarized in Table 1: 91.0 % and 13.3 GPa for the MA-SHS sample compact; 96.7 % and 14.3 GPa for the mixing sample compact; 97.7 % and 16.6 GPa for the commercial sample compact. Apparently, the lowest value of relative density and HV for the MA-SHS sample compact seems to be attributable to its porous structure shown in Fig.5(a). Therefore, the elimination of pores from the MA-SHS

sample compact having a fine grain-sized, homogeneous microstructure can be expected to improve the mechanical property such as Vickers hardness. Then, the effect of the unreacted C and B elements that remained in the MA-SHS sample on the pore formation was investigated.

Fig.6 shows the SEM microphotographs of the SPS compacts obtained from the commercial sample which the C or B element was added to in a molar ratio of NbC/NbB₂/C (or B)=45/45/10. The XRD patterns of these compacts showed only the NbC and NbB₂ phases. Compared with the SEM microphotographs shown in Fig.5(c), it was found that the addition of C and B elements to the commercial NbC/NbB₂ sample led to an increase in the pore size, particularly the C-added sample showed the development of about 10 μm-sized pore. Such pore formation may be considered to have resulted from the oxidation and vaporization of the C and B elements although the SPS runs were carried out under vacuum. The values of relative density and Vickers hardness of these compacts were decreased by the formation of porous microstructure as shown in Table 1: 94.9 % and 14.5 GPa for the sample of NbC/NbB₂/C=45/45/10; 95.6 % and 15.3 GPa for the sample of NbC/NbB₂/B=45/45/10.

Then, we have tried to separate and remove the unreacted C (density, d=2.25) and B (d=2.34) elements from the MA-SHS sample (d=8.56 for Nb, 7.80 for NbC, 6.93 for NbB₂) by heavy liquid separation using tetrabromoethan (d=2.96). The sample of several grams was taken into a separatory funnel with tetrabromoethan of 100 ml and shaken for enough time and allowed to stand for three days. Then, only the sample settled in the bottom was separated, washed with acetone and dried. Such MA-SHS sample separated from the C and B elements was compacted by SPS at the same sintering conditions. The phases of the obtained compact were identified by XRD in the following order of intensity, NbC >> NbB₂ = Nb₃B₄ > Nb. In contrast to the increase in intensity of NbC peaks, the decrease in intensity of NbB₂ peaks and the increase in intensity of Nb₃B₄ peaks probably indicate the progress of the reaction, NbB₂ + Nb → Nb₃B₄. Fig.7 shows the SEM microphotograph of the compact sintered at 1800°C. The pores in the sintered compact are remarkably eliminated and the homogeneous microstructure consisting of a few μm-sized grains can be observed. And, its Vickers hardness increased to 19.8 GPa as shown in Table 1. Regardless of the low relative density of 94.0 %, this value was higher than 16.6 GPa of the commercial sample compact of the relative density of 97.7 %. It can be easily considered that the sintering behavior and the densification rate strongly depend on the reactivity of the starting powder, such as particle size, specific surface

area, surface defect structure and so on. Therefore, the better value of Vickers hardness of the MA-SHS sample compact obtained after elimination of the unreacted C and B elements seems to primarily depend on its smaller grain-sized, homogeneous microstructure. A similar hardness-grain size dependence has been obtained for the MA-SHS sample compact of the Zr/B/C system [14]. In addition, it was found from the SEM observation of the fracture surfaces that many cracks were deflected at the grain boundaries, which would lead the increase in the fracture toughness because of the limited mobility of dislocation in small grains. In order to discuss on the hardness and/or fracture toughness-grain size dependence in detail, the high resolution TEM study about the interface structure of ZrC/ZrB₂ and NbC/NbB₂ and the dislocation behavior would be required.

In conclusion, it was verified that the MA-SHS in air process is a unique and energy-saving technique to synthesize concurrently the fine particles of NbC and NbB₂. The NbC/NbB₂ composite compact was sintered by SPS technique from this sample at a temperature as low as 1800°C, and showed a fine grain-sized, homogeneous microstructure and Vickers hardness, HV of 19.8 GPa, which was higher than 14.3 and 16.6 GPa for the NbC/NbB₂ composite compacts obtained from the powder mixture of Nb/B/C=2/2/1 without MA-SHS and from the commercial powder mixture of NbC/NbB₂=1/1. In addition, it would be possible to synthesize the nanocomposite of NbC/NbB₂, when the MA-SHS sample consisting of nanosized particles would be available as a starting powder.

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Figure captions

Fig.1 The SEM microphotographs of the commercial powders of NbC and NbB₂.

Fig.2 The XRD patterns of the sample of Nb/B/C=2/2/1 which were ground for 0-105 min. The weight ratio of sample to balls, R is 1/10. After grinding of 105 min, the formation of NbC and NbB₂ by the MA-SHS in air process was detected.

Fig.3 The secondary electron image (SEI) and associated elemental maps of the MA-SHS sample obtained by grinding of 105 min in Fig.2.

Fig.4 The XRD patterns of (a) the MA-SHS sample obtained from the powder mixture of Nb/B/C=2/2/1, and the NbC/NbB₂ composite compacts obtained by SPS at 1800°C from (b) the MA-SHS sample, (c) the mixing sample of Nb/B/C=2/2/1 without MA-SHS and (d) the commercial sample mixed in a ratio of NbC/NbB₂=1/1. *: The peak of carbon resulted from carbon paper that was used to wrapped the sample powders in SPS runs.

Fig.5 The SEM microphotographs of the NbC/NbB₂ composite compacts sintered by SPS at 1800°C from (a) the MA-SHS sample, (b) the mixing sample and (c) the commercial sample.

Fig.6 The SEM microphotographs of the SPS compacts obtained from the commercial sample which the C or B element was added to in a molar ratio of (a) NbC/NbB₂/B=45/45/10 and (b) NbC/NbB₂/C=45/45/10.

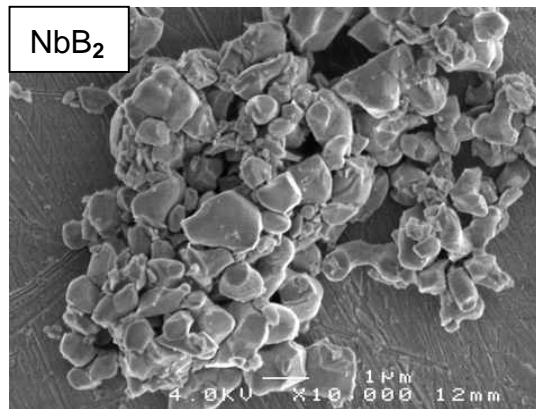
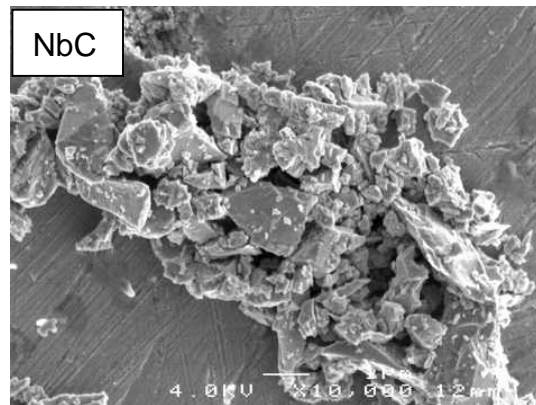
Fig.7 The SEM microphotograph of the NbC/NbB₂ composite compact sintered by SPS at 1800°C from the MA-SHS sample which the unreacted C and B elements were separated and removed from by heavy liquid separation using tetrabromoethan.

Table 1. The relative density and the Vickers hardness, HV of the NbC/NbB₂ composite compacts obtained by SPS at 1600-1800°C from the MA-SHS sample (○), the mixing sample and the commercial sample without the MA-SHS process (×), which are described in detail in text.

Samples	MA-SHS	Sintering Temp. (°C)	Holding Time (min)	HV (GPa)	Relative Density (%)
Nb/B/C=2/2/1	×	1600	30	8.8	95.3
	×	1800	10	14.3	96.7
	○	1700	10	14.1	90.2
	○	1800	10	13.3	91.0
	○*	1610	10	15.9	94.6
	○*	1800	10	19.8	94.0
NbC/NbB ₂ =1/1	×	1800	10	16.6	97.7
NbC/NbB ₂ /B=45/45/10	×	1800	10	15.3	95.6
NbC/NbB ₂ /C=45/45/10	×	1800	10	14.5	94.9

* : The MA-SHS samples from which the unreacted C and B elements were removed by heavy liquid separation.

Fig.1



— 1 μ m

Fig.2

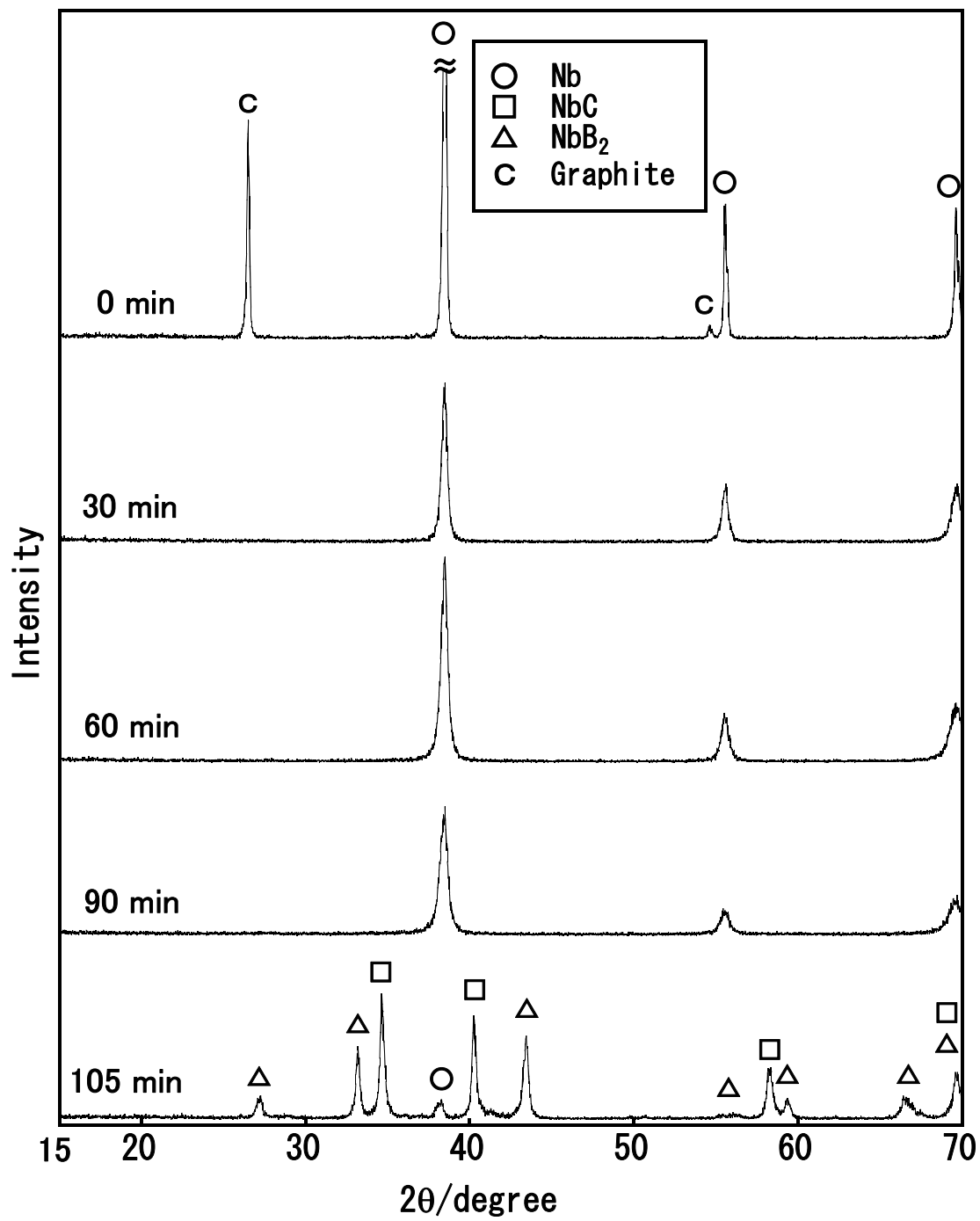


Fig.3

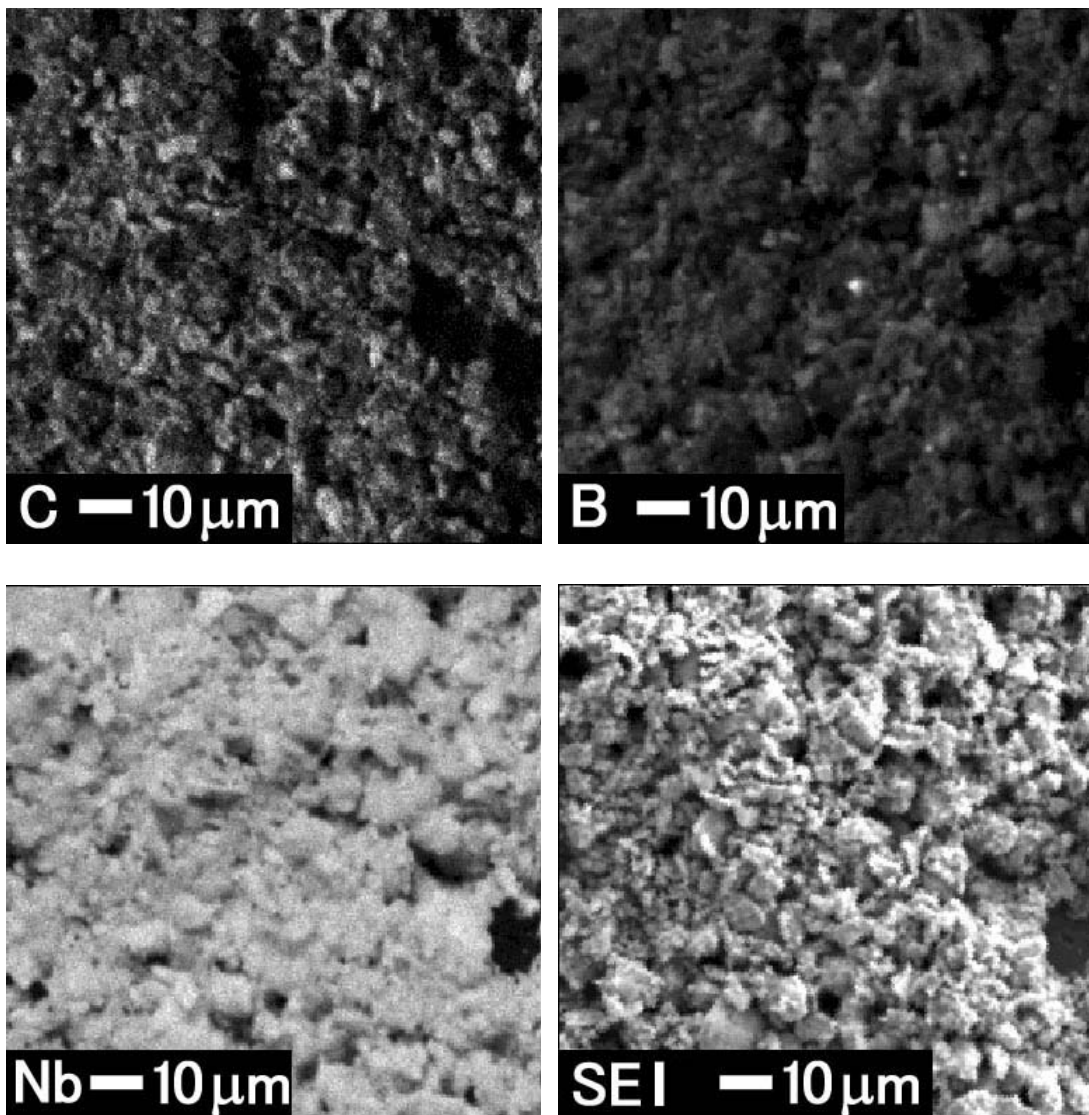


Fig.4

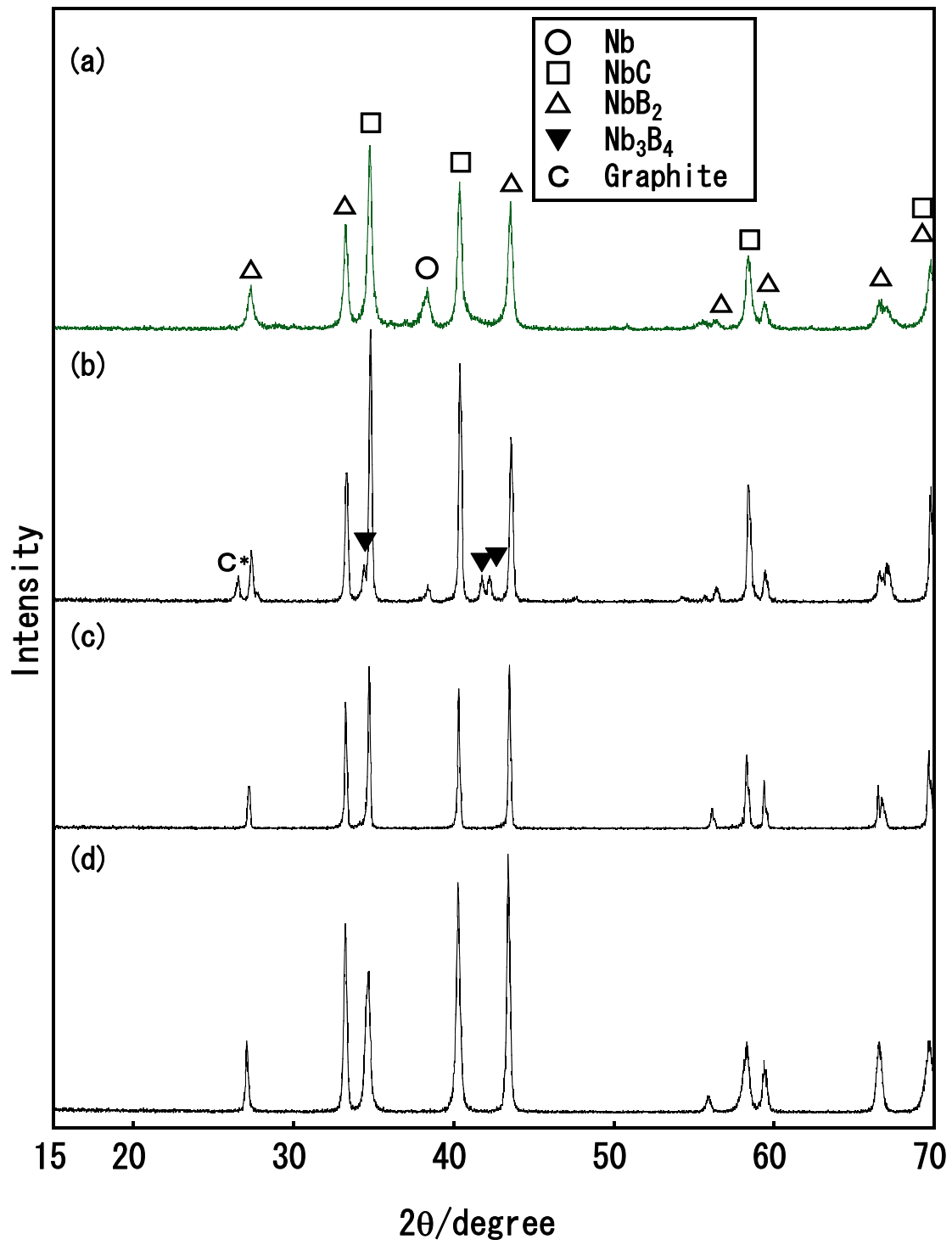


Fig.5

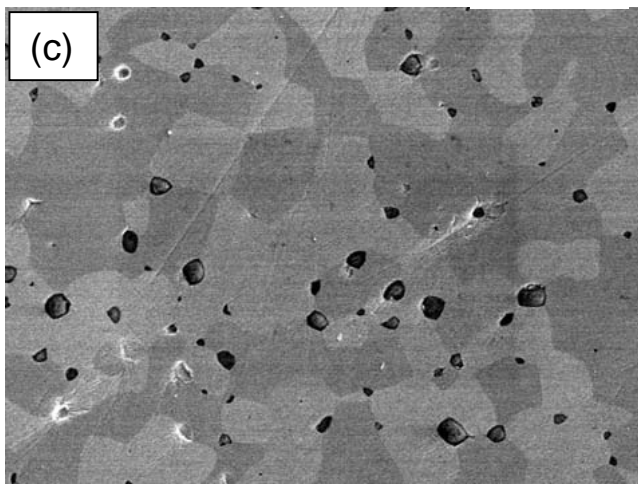
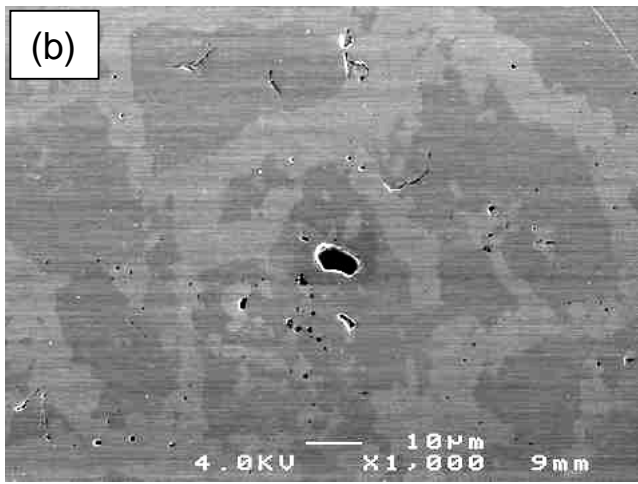
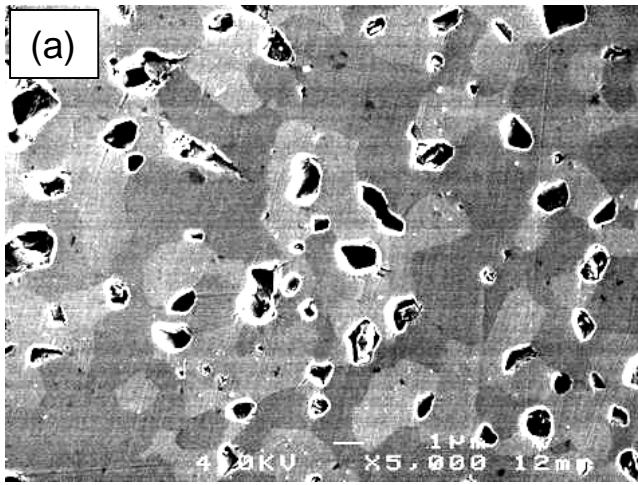
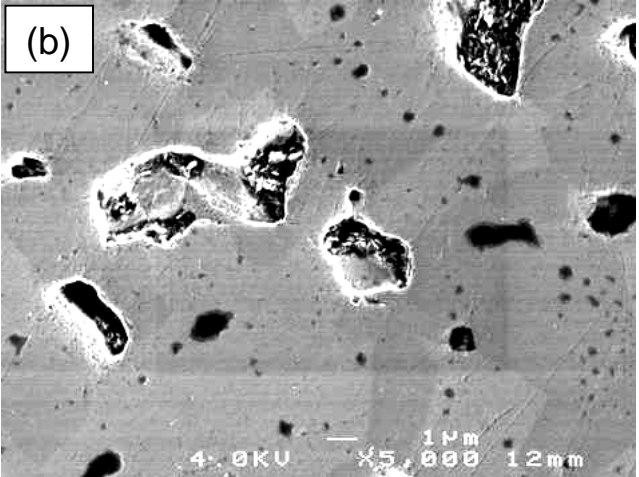
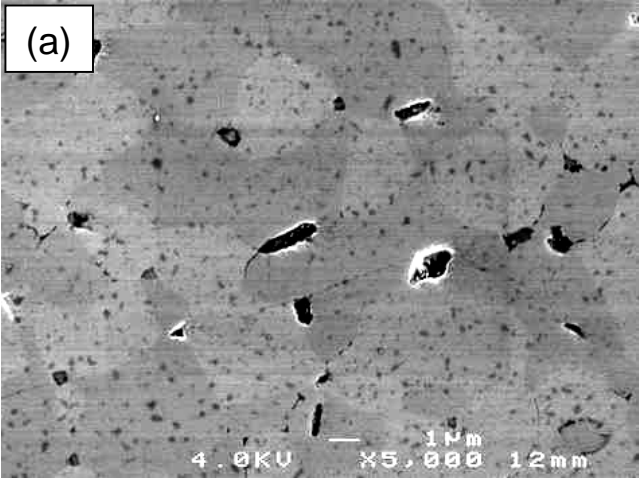
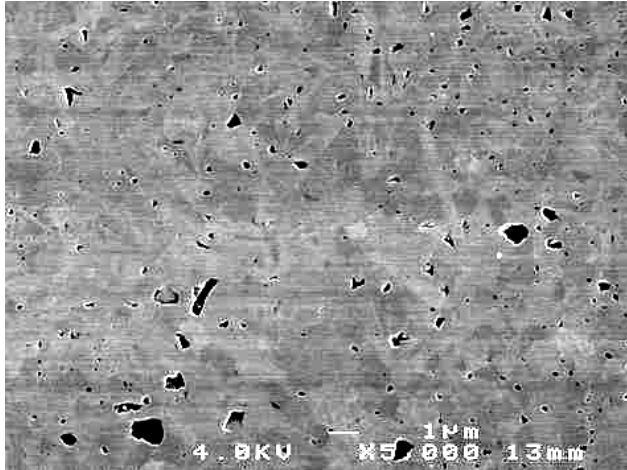


Fig.6



— 1 μm

Fig.7



— 1 μ m