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Development of A Novel Photocatalytic Reaction System for Oxidative Decomposition of Volatile Organic Compounds in Water with Enhanced Aeration

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Abstract

A novel photocatalytic reaction system, composed of solution and gas spaces which are divided by a thin Teflon film and a titanium(IV) oxide (TiO₂) coated mesh or cloth, for the treatment of contaminated aqueous solutions was developed to be operated with enhanced aeration without bubbling of air in the solution. First, photocatalytic activity of TiO₂ particles immobilized on two kinds of support materials, stainless steel mesh (SSM) and fiberglass cloth (FGC), was investigated for photocatalytic oxidation of 2-propanol, as a model volatile organic compound, dissolved in aerated aqueous solution. Both immobilized TiO₂ exhibited photocatalytic activity of 2-propanol oxidation to acetone and carbon dioxide (CO₂) and the activities of two kinds of immobilized TiO₂ were comparable. Presumably due to the presence of small amount of metal species originated from SSM which might work as reduction catalysts, molecular hydrogen (H₂) was also liberated on the TiO₂-immobilized SSM. Analysis of the weight loss after photoirradiation suggested that stability of the TiO₂-immobilized FGC was rather better than the TiO₂-immobilized SSM. On the basis of these results, FGC was employed in construction of a photocatalytic reactor equipped with an

oxygen (O₂) permeable Teflon membrane in order to make oxygen pass from a gas space to a solution space and to keep the surface of immobilized TiO₂ photocatalyst, facing to an aqueous solution containing volatile organic compounds, saturated with dissolved O₂. From the results of photocatalytic oxidative decomposition of 2-propanol, it was clarified that the TiO₂ surface could be supplied with O₂ effectively from the gas space through the membrane to accelerate the oxidation.

Keywords: titanium(IV) oxide; supported photocatalysts; oxygen permeable membrane; photocatalytic oxidation; volatile organic compounds; aqueous solution

1. Introduction

Volatile chlorinated organic compounds (VCOCs) such as trichloroethylene (TCE) and tetrachloroethylene (PCE) have been widely used as solvents in industrial processes. As a result of leak of them, e.g., from underground storage tanks and disposed plastics, soils and/or ground water supplies have often become contaminated by VCOCs. This is one of the recent major issues because these chemicals are toxic, often carcinogenic, and extremely persistent in the environment. A conventional method for the VCOCs removal from water using adsorbents, e.g., granular activated carbon (GAC) is effective for the decrease in the VCOCs concentration, but further treatment is required for recovery of GAC adsorbing toxic organic compounds.

Photocatalytic oxidation is an alternative method to purify water through decomposition of organic contaminants into carbon dioxide (CO₂) and mineral acid such as hydrogen chloride (HCl). There have been many papers dealing with photoinduced decomposition of a wide range of organic compounds including VCOCs, nonionic surfactants, and aromatic compounds in aqueous solutions using titanium(IV) oxide (TiO₂) as a photocatalyst [1-8]. The photocatalytic decomposition by TiO₂ is usually thought to be initiated by either direct [9,10] or hydroxyl radical (•OH) mediated [4-6] positive hole (h⁺) transfer to organic molecules and as a counterpart of redox reaction, conduction band electrons (e⁻) are consumed by reaction with molecular oxygen (O₂) dissolved in the reaction

solutions [7]. To accelerate the oxidation, it is necessary to scavenge e^- efficiently by O_2 . However, since the concentration of dissolved O_2 in aqueous solutions is limited (ca. 1 mmol dm^{-3}), recombination of e^- and h^+ , leading to low catalytic efficiency, is one of the practical problems in TiO_2 photocatalytic reactions operated in aqueous solutions. Bubbling of air through aqueous solutions is the simplest way to supply sufficient amount of O_2 and thereby to suppress the e^- - h^+ recombination, but this method cannot be adopted in the case of VCOCs because of their volatilization to open air. An alternative method to remediate contaminated water by TiO_2 photocatalyst is to decompose gaseous VCOCs, which is expelled away from the aqueous solution by air. However, it has been already claimed that the VCOC degradation in the gas-phase photocatalytic reaction produces unfavorable byproducts such as dichloroacetyl chloride and phosgene which are well-known highly toxic compounds [12,13]. It has been also reported that accumulation of chloroacetic acid on the TiO_2 surface decreases the photocatalytic activity of the catalyst when used in air [14,15].

For the fundamental research to clarify the reaction mechanism and to achieve complete mineralization of organics in aqueous solution into carbon dioxide (CO_2), most works used reactors with suspensions of TiO_2 particles [1-6]. However, the suspension system cannot be applied as a practical technique because it needs filtration to recover catalysts from reaction mixtures. In this regard, two different methods have been proposed for the continuous use without removal from cycle to cycle; one is using TiO_2 pellets [7] and the other is immobilization of TiO_2 particles on various supports [8], though these methods were often confronted with problems, e.g., decrease in effective surface area, appreciable loss of TiO_2 from supports, decreased adsorption of organic substances on the TiO_2 surface, and mass transfer limitations.

Our research interest is, therefore, focused on development of a system for efficient photodegradation of aqueous VCOCs using an active immobilized TiO_2 photocatalyst. In this paper, we have investigated photocatalytic activity of TiO_2 particles immobilized on two kinds of supports, and have developed a novel photocatalytic reaction system, where O_2 molecule could be supplied to the surface of TiO_2 photocatalysts from the gas phase through an O_2 permeable membrane to accelerate the oxidation of organic pollutants without air

bubbling to induce volatilization of organic compounds into open air.

2. Experimental

2.1. Materials

Degussa P25 TiO₂ powder was supplied from Nippon Aerosil. From the analysis of percent proportion of the most intense XRD peaks of anatase ((101), 2 theta = 25.3.) and rutile ((110), 2 theta = 27.5.).[16], the TiO₂ powder is composed from 83% of anatase and 17% of rutile. Its BET surface was measured to be 50 m² g⁻¹ by nitrogen adsorption at 77 K. A stainless steel mesh (SSM, 144 holes cm⁻², 0.2 mm in thickness) and fiberglass cloth (FGC, Central Glass EGW 110TH-153 (104), 0.24 mm in thickness) were used as supports of TiO₂ particles. Laboratory grade water was prepared with a Milli-Q pure water system (Yamato-Millipore). Titanium(IV) tetra-2-propoxide (Ti(OⁱPr)₄, Wako Pure Chemical), 2-propanol (Wako), and acetic acid (Wako) were used without further purification.

2.2. Immobilization of TiO₂ on stainless steel mesh and fiberglass cloth

To a 50-cm³ portion of 2-propanol, Ti(OⁱPr)₄ (0.94 g) was added dropwise with vigorous stirring in open air. After agitation for overnight at ambient temperature, a slightly turbid solution thus obtained was diluted with 60 cm³ of 2-propanol. Then, 3.88 g of TiO₂ powder was added and the resulting suspension was stirred for several hours to obtain homogeneous dispersion. The dispersion was applied onto SSM and FGC heated at ca. 350 K using a spray bottle followed by heat treatment at 473 K for 1 h in air. For all experiments reported in this paper, amount of TiO₂ was in the range of 7-10 mg cm⁻² of support materials, SSM or FGC.

2.3. Photoirradiation under static conditions

The TiO₂-immobilized SSM or FGC (0.7 cm × 2.0 cm) and an aqueous 2-propanol (0.25 vol%) solution (5 cm³) were placed in a glass tube (transparent at > 300 nm, 18 mm in diameter and 180 mm in length), and they were photoirradiated by a high-pressure mercury arc (Eiko-sha, 400 W) in air at 298 K. The TiO₂-coated face of SSM and FGC was kept perpendicular to the light beam. After irradiation for 1 h, a portion (0.2 cm³) of the gas phase of the sample was withdrawn with an air-tight syringe and subjected to gas

chromatographic analyses (GC) of molecular hydrogen (H_2) and carbon dioxide (CO_2) using a Shimadzu GC-8A gas chromatograph equipped with MS-5A (for H_2) and Porapak-Q (for CO_2) columns, and a TCD detector. The yield of acetone was also measured by GC using a Shimadzu GC-14B gas chromatograph equipped with a PEG-20M column and an FID detector.

2.4. Photooxidation using a reactor equipped with an oxygen permeable membrane

A schematic view of the reaction apparatus for the photocatalytic oxidation of organic contaminants in water using a TiO_2 -immobilized FGC and an oxygen (O_2) permeable membrane is shown in Figure 1. The reactor consists of two chambers separated by a Teflon membrane (0.025 mm in thickness, ca. 9 cm in diameter); an upper chamber made from Teflon was used for circulation of reaction solution (a solution space) and a lower chamber made from stainless steel was used for gas supply, O_2 , Ar, or air (a gas space). TiO_2 (ca. 350 mg)-immobilized FGC (ca. 8 cm in diameter) was placed on the Teflon membrane so that the loaded TiO_2 photocatalyst could be supplied with O_2 from the gas space through the Teflon membrane. A peristaltic pump (Cole Palmer Masterflex 7553-70) was used to supply the solution space of the reactor with an aqueous 2-propanol solution from a Teflon reservoir (2500 cm^3). Photoirradiation was performed through a Pyrex window (8 cm in diameter) fixed on the top of the liquid space. Two sets of experiments were conducted using this reactor. In one set of experiments (a circulation system), aqueous 2-propanol (18 $mmol\ dm^{-3}$) deaerated by bubbling with Ar was circulated at the flow rate of 140 $cm^3\ min^{-1}$ and photoirradiation was performed by a black light (Toshiba, 3.0 $mW\ cm^{-2}$ at 360 nm). A portion (0.2 cm^3) of a reaction solution in the reservoir was withdrawn in every one hour and subjected to GC analysis to quantify the produced acetone. During the experiment, the solution in the reservoir was magnetically stirred (ca. 500 rpm) to retain homogeneity. Supplied gases in the gas space were changed from Ar to O_2 or from O_2 to Ar. In another set of experiments (a single pass system), an aerated aqueous 2-propanol (2.5 $mmol\ dm^{-3}$) was supplied to the reactor at the flow rate of 55 $cm^3\ min^{-1}$ and drain from the outlet of the reactor was collected in a glass bottle and was analyzed at every one hour. A xenon (Xe) arc (300 W, Ushio) was employed as a light source. During the experiment, air or Ar was

continuously supplied to the gas space at ambient pressure. Concentrations of dissolved O₂ in the solution after passing the reactor were also monitored by a TOA-DKK DO55G dissolved oxygen meter.

3. Results and Discussion

3.1. Photocatalytic activity of TiO₂ loaded on SSM and FGC

Photoinduced oxidation of 2-propanol in aerated aqueous solution was conducted as a test reaction of photoactivity of the TiO₂-coated SSM and FGC. Table 1 summarizes amounts of products and percent proportion of TiO₂ remaining on the supports after photoirradiation for 1 h. Both immobilized TiO₂ particles showed activity for photooxidation of 2-propanol and acetone was produced as a main product along with liberation of appreciable amount of CO₂ which might be produced by further oxidation of acetone [17]. The formation of these products was hardly observed without photoirradiation of TiO₂-coated support or with irradiation in the absence of TiO₂ photocatalyst; a negligible amount of acetone could be detected by irradiation in the absence of TiO₂. Due to the difficulty to control the amount of TiO₂ loading precisely, there was an obvious difference in the amount of immobilized TiO₂ between SSM and FGC, and exact comparison of the activity cannot be discussed. However, in a semi-quantitative sense, there was no significant difference in the amount of acetone and CO₂ between them, suggesting that the microenvironment of the loaded TiO₂ was almost the same. It was noted that small amount of molecular hydrogen (H₂) was detected on the TiO₂-immobilized SSM while no H₂ liberation was observed when the TiO₂-immobilized FGC was used as a photocatalyst. Since the TiO₂ surface has poor ability to reduce protons (H⁺) to liberate H₂ [18,19], certain metal species in SSM might participate in the observed H₂ liberation, i.e., metal species may work as active sites for reduction of H⁺ with photoexcited electrons in TiO₂. Thus, in order to evaluate the photocatalytic ability of oxidation without any other side reactions, FGC seems to be more preferable compared to the SSM. Another advantage of FGC as a support is stability of loaded TiO₂ on FGC higher than that on SSM; the FGC support could keep ca. 90% of TiO₂ while less than 80% of TiO₂ remained on SSM under the present reaction

conditions (Table 1). On the basis of these results, the TiO₂-immobilized FGC was employed in further experiments.

3.2. Photooxidation of 2-propanol using a reactor equipped with an oxygen permeable membrane in a closed circulation mode

In order to confirm the effect of O₂ permeation through a Teflon membrane, photooxidation of 2-propanol in deaerated aqueous solution using the reactor equipped with the membrane was performed with keeping the gas space filled with Ar or O₂. The reaction solution was circulated in this system with a flow rate of 140 cm³ min⁻¹ during the experiment (see Figure 1). Figure 2 shows time courses of acetone production when the atmosphere in the gas space was switched from Ar to O₂ (upper) or from O₂ to Ar during photoirradiation. In the dark, without photoirradiation, negligible acetone production was observed, as expected, when the gas space was filled with Ar. On the other hand, a detectable amount of acetone was produced under ultraviolet light irradiation even though O₂ content in the system was negligible, but the total yield of acetone was saturated after ca. 2-h irradiation. The production of small amount of acetone under Ar atmosphere is attributable to the oxidation of 2-propanol by positive holes (h⁺) along with the formation of Ti³⁺ species by trapping of electrons (e⁻) at defective sites in TiO₂ [20-22]. As discussed in previous report [22], such accumulation of e⁻ was also observed by photoirradiation to a deaerated aqueous TiO₂ suspension containing methanol and triethanolamine in the absence of reduction catalyst such as platinum. The formation of Ti³⁺ is also supported by the change in color from white to pale blue during the photoirradiation. It has been reported that P25 TiO₂ contains ca. 50 mol g⁻¹ of sites to become Ti³⁺ by photoirradiation and this could interpret ca. 10 mol of acetone production assuming two Ti³⁺ is produced along with a molecule of acetone. Additional Ti³⁺ in TiO₂ originated from Ti(OⁱPr)₄ may account for the rest of acetone production.

The most striking feature in these plots is that just switching the atmosphere in the gas space changes the rates of acetone productions; rates of acetone production in the presence of O₂ in the gas space were about three times higher than those in the presence of Ar. These results clearly suggest that the photocatalyst in the solution space can be supplied with

O₂ from the gas space through the Teflon membrane to enhance photooxidation of aqueous 2-propanol.

Table 2 summarizes the results of amounts of acetone production and 2-propanol consumption from the deaerated aqueous 2-propanol solution in the presence of air, Ar, or O₂ in the gas space after photoirradiation for 1 h. Photoirradiation induced production of acetone when any kinds of gases were supplied in the gas space. In the presence of O₂ (air or O₂) in the gas space, amounts of produced acetone were considerably higher than those in the absence of O₂ (Ar atmosphere). Moreover, the amount of acetone and that of consumed 2-propanol became more than twice when air was substituted by pure O₂ in the gas space, i.e., the presence of O₂ in the gas phase is responsible for the increase of photooxidation of reactants in aqueous solution and the photoactivity depends on the concentration of O₂ in the gas space. Another point worth consideration is that the ratio of 2-propanol consumption to acetone production increased with increase in concentration of O₂ in gas phase. This might be due to the enhancement of further oxidation of acetone to yield CO₂ (see Table 1) with high concentration of O₂ in the gas space.

When similar experiments were conducted using 20-times thicker Teflon membrane (0.5 mm), there was no appreciable dependence of photocatalytic activity on the atmosphere of the gas space, and the change of the color of FGC from white to blue was observed even when O₂ was supplied in the gas space. This indicates that mass transfer of O₂ from the gas space should be dominated by the thickness of the Teflon membrane. Permeability coefficient of O₂ through Teflon, polytetrafluoroethylene, is $3.7 \times 10^{-7} \text{ cm}^3 \text{ (STP) cm}^{-2} \text{ cm s}^{-1} \text{ MPa}^{-1}$ [23]. Assuming 0.1-MPa (ca. 1 atm) pressure difference between solution and gas spaces, amount of O₂ permeating through a 0.025 and 0.5 mm (50 cm²) membrane is ca. 120 and 6 mol h⁻¹. Thus, it is reasonable to find no effect on the photocatalytic oxidation when the thicker membrane was used. Details of the dependence of thickness of are now investigated and will be reported elsewhere.

3.3. Photooxidation of 2-propanol using a reactor equipped with an oxygen permeable membrane in a single-pass mode

In view of the practical application in the decomposition of VCOCs dissolved in

water, a single-pass (flow) system, where the reaction solution was supplied without circulation, is preferable compared to the circulation (batch) system as conducted in the experiments shown above. Moreover, in order to prove the effect of O₂-permeable membrane employed in the present reactor on the purification of actual contaminated water, it is necessary to confirm the advantage of the use of an O₂-permeable membrane in the photocatalytic decomposition of organic compounds in an aerated (not deaerated) solution. Thus, the photooxidation of 2-propanol in the aerated, i.e., air saturated, aqueous solution was performed using the single-pass system. Figure 3 shows variation of amount of acetone production and concentration of dissolved O₂ in the reaction solution after passing the reactor versus photoirradiation time. As clearly seen in the time courses of acetone production, photooxidation activity in the presence of air in the gas space showed more than 1.5 times higher than that in the presence of Ar. It was also noted that no appreciable decrease in the concentration of dissolved O₂ was observed when the gas space was filled with air while the O₂ concentration was decreased when air in the gas space was substituted by Ar owing to the consumption by the reaction (Probably due to the further oxidation to give CO₂ as discussed in the preceding section, molar amount of consumed O₂ exceeded that of produced acetone. However, we did not analyze amount of CO₂ in this system and do not discuss here). These results indicate that the O₂ permeable membrane effectively works for enhancement of aeration of reaction solution, leading to increment of photocatalytic activity even when the reaction solution contains dissolved O₂ before passage in the reactor. In the present experiments, the pass length was only 8 cm and it is expected that the longer the pass, the more significant the effect of the O₂-permeable membrane.

4. Conclusions

The present study has shown that enhancement of aeration by an O₂ permeable membrane can increase photocatalytic activity for the decomposition of organic compounds in aqueous solution for the first time [24]. The system has advantage for the application of photodegradation of aqueous VCOCs not to require air purging which causes unfavorable volatilization into open air. In this study, we used only 2-propanol as a model compound

and reaction conditions such as thickness of the membrane, flow rate of air (or O₂) were not optimized. Further studies using VCOCs and a O₂-permeable membrane, having large permeability coefficient, other than Teflon, e.g., silicone are now under way.

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References

- [1] D. F. Ollis, E. Pelizzetti, N. Serpone, *Photocatalysis —Fundamentals and Applications*, Chap. 18, John Wiley & Sons, Chichester, 1997.
- [2] D. F. Ollis, E. Pelizzetti, N. Serpone, *Environ. Sci. Technol.* 25 (1991) 1522.
- [3] M. R. Hoffmann, S. T. Martin, W. Choi, *Chem. Rev.* 95 (1995) 69.
- [4] C.-Y. Hisao, C.-L. Lee, D. F. Ollis, *J. Catal.* 82 (1983) 418.
- [5] R. W. Matthews, *J. Phys. Chem.* 91 (1987) 3328.
- [6] H. Al-Ekabi, N. Serpone, *J. Phys. Chem.* 92 (1988) 5726.
- [7] S. Yamazaki, S. Matsunaga, K. Hori, *Wat. Res.* 35 (2001) 1022.
- [8] S. Horikoshi, N. Watanabe, H. Onishi, H. Hidaka, N. Serpone, *Appl. Catal. B: Environ.* 37 (2002) 117.
- [9] O. I. Mimic, Y. Zhang, K. R. Cromack, A. D. Trifunac, M. C. Thurnauer, *J. Phys. Chem.* 97 (1993) 13284.
- [10] U. Stafford, K. A. Gray, P. V. Kamat, *J. Phys. Chem.* 98 (1994) 6343.
- [11] J. Fun, J. T. Yates Jr., *J. Am. Chem. Soc.* 118 (1996) 4686.

- [12] M. D. Driessen, A. L. Goodman, T. M. Miller, G. A. Zaharias, V. H. Grassian, *J. Phys. Chem. B* 102 (1998) 549.
- [13] S.-J. Hwang, C. Petucci, D. Raftery, *J. Am. Chem. Soc.* 120 (1998) 4388.
- [14] S. Yamazaki-Nishida, S. Cervera-March, K. J. Nagano, M. A. Anderson, K. Hori, *J. Phys. Chem.* 99 (1995) 15814.
- [15] S. Yamazaki-Nishida, X. Fu, M. A. Anderson, K. Hori, *J. Photochem. Photobiol. A: Chem.* 97 (1996) 175.
- [16] T. Torimoto, N. Nakamura, S. Ikeda, B. Ohtani, *Phys. Chem. Chem. Phys.* 4 (2002) 5910.
- [17] H. Yamashita, M. Harada, J. Misaka, M. Takeuchi, K. Ikeue, M. Anpo, *J. Photochem. Photobiol. A: Chem.* 148 (2002) 257.
- [18] S. Sato, J. M. White, *J. Phys. Chem.* 85 (1981) 592.
- [19] B. Ohtani, K. Iwai, S.-i. Nishimoto, S. Sato, *J. Phys. Chem. B* 101 (1997) 3349.
- [20] G. Rothenberger, D. Fitzmaurice, M. Gr tzel, *J. Phys. Chem.* 96 (1992) 5983.
- [21] S. A. Haque, Y. Tachibana, D. R. Klug, J. R. Durrant, *J. Phys. Chem. B* 102 (1998) 1745.
- [22] S. Ikeda, N. Sugiyama, S.-y. Murakami, H. Kominami, Y. Kera, H. Noguchi, K. Uosaki, T. Torimoto, B. Ohtani, *Phys. Chem. Chem. Phys.* 4 (2003) 778.
- [23] Chemical Society of Japan, *Kagaku-Binran Ohyo-hen II Materials*, Maruzen (Tokyo) (1986) 1177.
- [24] Japanese patent application, 2002-228445.

Table 1

Photocatalytic activity for oxidation of aqueous 2-propanol by immobilized TiO₂ powders on various supports.

support	W _{TiO₂} ^a / mg	R _{ac} ^b / mol h ⁻¹	R _{CO₂} ^c / mol h ⁻¹	R _{H₂} ^d / mol h ⁻¹	W ^e / %
blank ^f	-	0.02	0	0	-
SSM ^g	14	0.53	0.13	0.07	21
FGC ^h	10	0.48	0.10	0	11

^aTotal amount of TiO₂ immobilized on the support. ^bRate of acetone production.

^cRate of CO₂ liberation. ^dRate of H₂ liberation. ^ePercent weight loss of immobilized TiO₂ after photoirradiation for 1 h. ^fPhotoirradiation without catalyst.

^gStainless steel mesh. ^hFiberglass cloth. Catalyst; Degussa P25 immobilized on 1.4 cm² of SSM or FGC.

Table 2

Variation in amounts of acetone production and 2-propanol consumption in the presence of various gases in the gas space^a

gas phase ^b	P _{ac} ^c / mol	C _{pr} ^d / mol	C _{pr} /P _{ac}
Ar	3.2	- ^e	- ^e
air	8.5	10.2	1.2
O ₂	18.4	55.6	3.0

^aAll reactions were carried out by the reactor equipped with an O₂ permeable membrane (Fig. 1) for 1 h. Catalyst, TiO₂ (Degussa P25, ca. 350 mg) immobilized FC (ca. 50 cm²). Reaction solution was deaerated with Ar before photoirradiation.

^bSupplied gases in the gas space. ^cAmount of acetone production. ^dAmount of 2-propanol consumption. ^eNot measured.

Figure Captions

Figure 1. Schematic illustration of the photocatalytic oxidation system with a reactor equipped with an O₂ permeable membrane.

Figure 2. Photooxidation of 2-propanol dissolved in a deaerated aqueous solution using the reactor shown in Figure 1. The reaction solution was circulated with a flow rate of 140 cm³ min⁻¹.

Figure 3. Variation of amount of acetone and concentration of dissolved O₂ after passing through the reactor with a flow rate of 55 cm³ min⁻¹ as a function of photoirradiation time. Ar (open symbols) or O₂ (filled symbols) was supplied in the gas space during the experiment. Squares and diamonds denote amount of produced acetone and concentration of dissolved O₂, respectively.

Figure 1

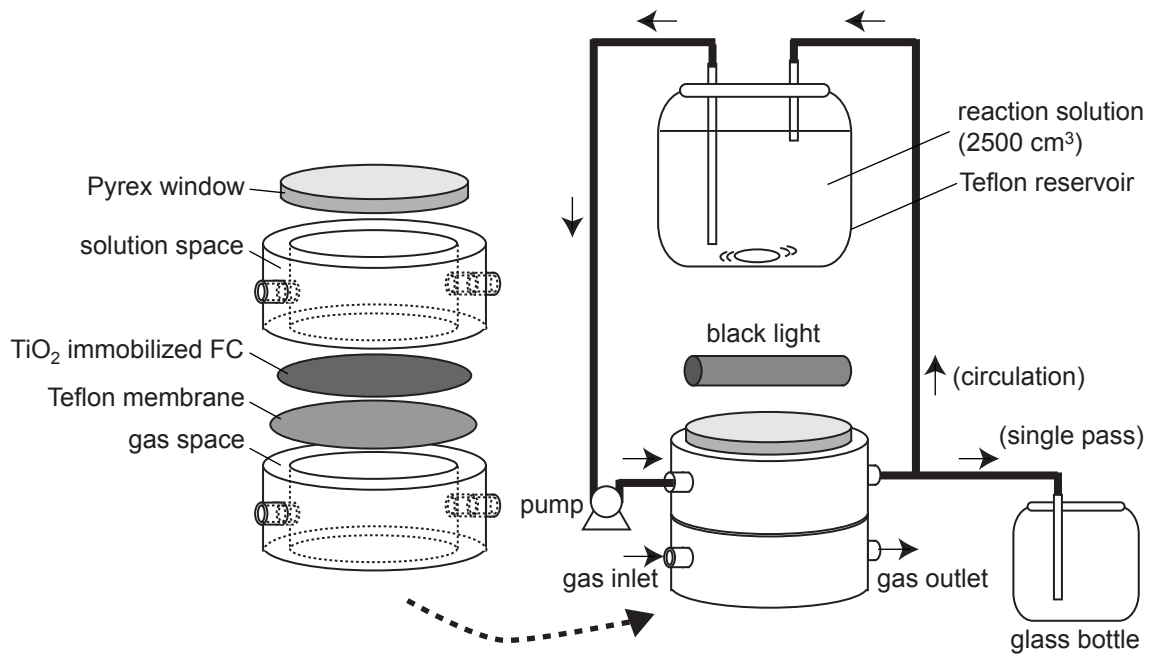


Figure 2

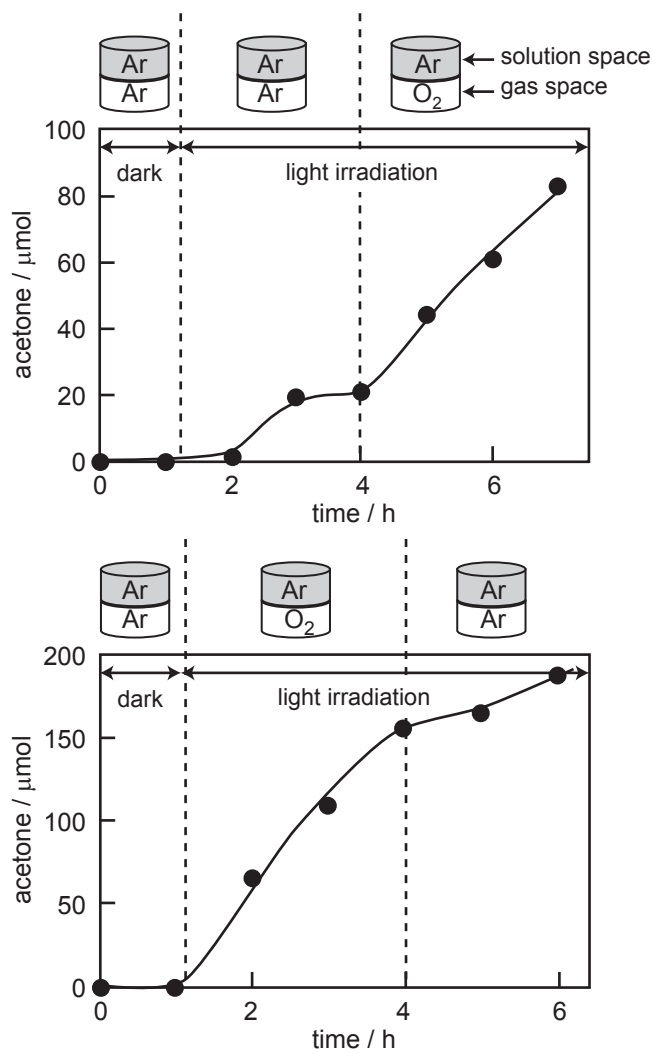


Figure 3

