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Author(s)	Sakurai, Toshimitsu; Ohno, Hiroshi; Genceli, F. Elif et al.
Citation	Journal of Glaciology, 56(199), 837-842 https://doi.org/10.3189/002214310794457335
Issue Date	2010-12
Doc URL	https://hdl.handle.net/2115/45541
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Type	journal article
File Information	JG56-199_837-842.pdf



Magnesium methanesulfonate salt found in the Dome Fuji (Antarctica) ice core

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ABSTRACT. Using micro-Raman spectroscopy, we identified the chemical forms of methanesulfonate salt particles in reference samples of the Dome Fuji (Antarctica) ice core. We found only $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$ among methanesulfonate salts, and this salt particle is most prevalent in the Last Glacial Maximum (LGM) ice. We suggest that during the LGM, $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$ may have formed in the atmosphere through the chemical reaction of $\text{CH}_3\text{SO}_3\text{H}$ with sea salts, but probably not in the firn and ice due to the neutralization of acid in LGM ice of inland Antarctica.

1. INTRODUCTION

Methanesulfonic acid ($\text{CH}_3\text{SO}_3\text{H}$) is an important compound formed by the oxidation of dimethyl sulfide (DMS) in the atmosphere (Hynes and others, 1986; Saigne and Legrand, 1987; Saltelli and Hjorth, 1995). $\text{CH}_3\text{SO}_3\text{H}$ is specific to marine biogenic activity, so its concentration in the remote troposphere has been a subject of great interest over the past decade. In ice cores, $\text{CH}_3\text{SO}_3\text{H}$ can be detected through ion chromatography (CH_3SO_3^-) by melting an ice sample. Its anion is of primary importance as a tool for reconstructing several activities related to climate change: marine productivity (Legrand and others, 1991, 1992), sea-ice extent (Curran and others, 2003) and major El Niño events (Legrand and Feniet-Saigne, 1991).

In the present time period, ice-core CH_3SO_3^- appears to derive mainly from post-depositional relocation, a process that occurs while the acid is in the gas and/or liquid state (Mulvaney and others, 1992; Jaffrezo and others, 1994; Pasteur and Mulvaney, 1999, 2000; McMorrow and others, 2004; Smith and others, 2004; Ruth and others, 2005). On the other hand, a recent study using Raman spectroscopy has provided new information on the chemical forms of certain water-soluble salts contained in the microparticles of ice cores. This evidence suggests that methanesulfonate may exist as a salt in glacial period ice (Ohno and others, 2005, 2006; Barletta and others, 2009). However, very little is known about the eutectic temperature and Raman spectra of methanesulfonate salts. This study identifies methanesulfonate salt particles in the Dome Fuji (Antarctica) ice core using micro-Raman spectroscopy, and proposes a formation process based on phase diagrams measured by differential scanning calorimetry (DSC) and Raman spectroscopy.

2. METHODS

2.1. Chemical form of methanesulfonate salts

Our micro-Raman spectrometer is a triple monochromator (Jobin-Yvon, T64000) equipped with a charge-coupled device (CCD) detector (Jobin-Yvon, Spectraview-2D). The device has good quantum efficiency from 500 to 900 nm.

The spectral resolution of the Raman spectrum is 0.6 cm^{-1} . The absolute frequency of the monochromator was calibrated using a neon emission line and silicon wafers. The cross slit of this measurement is $3\text{ }\mu\text{m}$. The cryostat is posed on an x-y-z translation stage, and fed by a N_2 gas flow from liquid nitrogen heated inside a dewar. More details of the method can be found in Genceli and others (2009) and Sakurai and others (in press).

The microparticles in the Dome Fuji ice core were analyzed using micro-Raman spectroscopy. The Dome Fuji ice-core section at Termination I (the Holocene/Last Glacial Maximum (LGM) transition) was cut with a bandsaw at depths of 340.38 (Holocene), 382.86, 414.36, 463.85, 502.35 (Termination I) and 576.50 m. The final depth corresponds to the LGM. All sections have dimensions of $10 \times 10 \times 3\text{ mm}^3$. The Raman spectra of microparticles in the ice core were compared with Raman spectra taken from reference specimens of simple salts such as reagent-grade chemicals 98 wt% $\text{CH}_3\text{SO}_3\text{Na}$ (Alfa Aesar), 99 wt% $(\text{CH}_3\text{SO}_3)_2\text{Mg}$ (Daniels Fine Chemicals), 98 wt% $\text{CH}_3\text{SO}_3\text{K}$ (Tokyo Kasei Kogyo) and 98 wt% $(\text{CH}_3\text{SO}_3)_2\text{Ca}$ (Tokyo Kasei Kogyo) at room temperature. $\text{CH}_3\text{SO}_3\text{Na}\cdot\text{H}_2\text{O}$, $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot\text{H}_2\text{O}$ and $\text{CH}_3\text{SO}_3\text{K}\cdot\text{H}_2\text{O}$ were measured by Raman spectroscopy at -30°C , and $(\text{CH}_3\text{SO}_3)_2\text{Ca}\cdot\text{H}_2\text{O}$ was measured at -50°C , temperatures below the eutectic temperature, at which solid hydrated crystal salts formed. All four solutions were prepared with ultrapure water ($18.3\text{ m}\Omega$) by weighing out the required amounts of the aforementioned reagent-grade chemicals. Due to their high percentage purities, we assume that these chemicals are anhydrous. Note that we distinguish the spectra on the basis of peak shapes as well as peak positions.

2.2. The eutectic temperatures of methanesulfonate salts

The phase diagrams of the methanesulfonate salts, including their eutectic temperatures, are determined using a differential scanning calorimeter (DSC; Rigaku Co., DSC8230) with liquid nitrogen cooling. Further details on the analytical technique of DSC measurement are provided by, for example, Wunderlich (1973). The temperature reproducibility of this

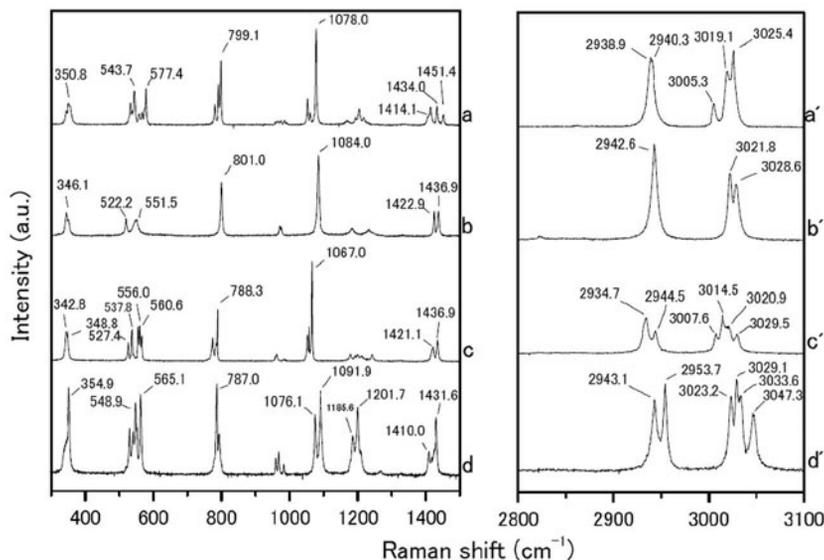


Fig. 1. Raman spectra from (a–d) 300–1500 cm^{-1} and (a'–d') 2800–3100 cm^{-1} . (a–d) are reagent-grade crystals of $\text{CH}_3\text{SO}_3\text{Na}$, $(\text{CH}_3\text{SO}_3)_2\text{Mg}$, $\text{CH}_3\text{SO}_3\text{K}$ and $(\text{CH}_3\text{SO}_3)_2\text{Ca}$ respectively, which were measured at room temperature.

instrument is better than $\pm 0.2^\circ\text{C}$, based on a three-point temperature calibration using indium (156.6°C), water (0.0°C) and cyclohexane (-83.0°C) (Charsley, 1993).

The DSC measurements are carried out on specimens of $\text{CH}_3\text{SO}_3\text{Na}\cdot\text{H}_2\text{O}$, $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot\text{H}_2\text{O}$, $\text{CH}_3\text{SO}_3\text{K}\cdot\text{H}_2\text{O}$ and $(\text{CH}_3\text{SO}_3)_2\text{Ca}\cdot\text{H}_2\text{O}$. Solutions are prepared by combining ultrapure water ($18.3\text{ m}\Omega$) with reagents. The DSC cells are made of aluminum and have a capacity of $30\ \mu\text{L}$. In cases where the samples sublimated, the cells were sealed before use. During the measurements for $\text{CH}_3\text{SO}_3\text{Na}\cdot\text{H}_2\text{O}$, $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot\text{H}_2\text{O}$ and $\text{CH}_3\text{SO}_3\text{K}\cdot\text{H}_2\text{O}$, the cells were cooled to -70°C at a rate of $20^\circ\text{C min}^{-1}$ and then heated to room temperature at $10^\circ\text{C min}^{-1}$. These three types of samples typically had a mass of about 2 mg. The $(\text{CH}_3\text{SO}_3)_2\text{Ca}\cdot\text{H}_2\text{O}$ samples typically had a mass of about 15 mg. The $(\text{CH}_3\text{SO}_3)_2\text{Ca}\cdot\text{H}_2\text{O}$ cells were cooled to -80°C at

a rate of $20^\circ\text{C min}^{-1}$ and then heated to room temperature at a rate of 1°C min^{-1} . These conditions were chosen because the latent heat of $(\text{CH}_3\text{SO}_3)_2\text{Ca}\cdot\text{H}_2\text{O}$ is very low.

3. RESULTS

3.1. Raman spectra of methanesulfonate salts

We measured the Raman spectra of several reference specimens: the simple salts $\text{CH}_3\text{SO}_3\text{Na}$, $(\text{CH}_3\text{SO}_3)_2\text{Mg}$, $\text{CH}_3\text{SO}_3\text{K}$ and $(\text{CH}_3\text{SO}_3)_2\text{Ca}$ (Fig. 1), and their solid hydrated crystal salts with ice under eutectic conditions (Fig. 2). The methanesulfonate salts show three major peaks around 1050–1100, 2930–2950 and 3010–3030 cm^{-1} . The peak around 1050–1100 cm^{-1} can be assigned to the SO_3^{2-} stretching frequencies (Socrates, 2001; Barletta and others,

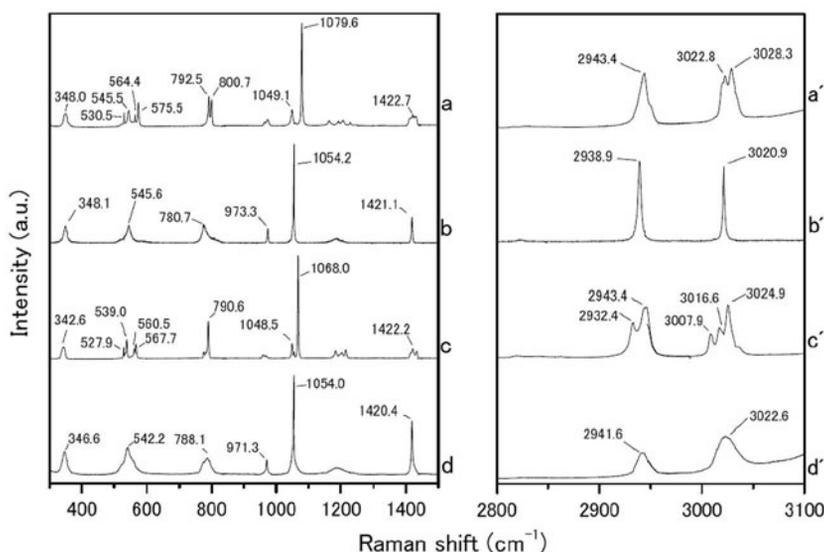


Fig. 2. Raman spectra from (a–d) 300–1500 cm^{-1} and (a'–d') 2800–3100 cm^{-1} . (a–d) are hydrated crystals of $\text{CH}_3\text{SO}_3\text{Na}\cdot n\text{H}_2\text{O}$, $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$, $\text{CH}_3\text{SO}_3\text{K}\cdot n\text{H}_2\text{O}$ and $(\text{CH}_3\text{SO}_3)_2\text{Ca}\cdot n\text{H}_2\text{O}$ respectively, which formed solid hydrated crystal salt under eutectic conditions. Measurement temperatures were -30°C for $\text{CH}_3\text{SO}_3\text{Na}\cdot n\text{H}_2\text{O}$, $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$ and $\text{CH}_3\text{SO}_3\text{K}\cdot n\text{H}_2\text{O}$, and -50°C for $(\text{CH}_3\text{SO}_3)_2\text{Ca}\cdot n\text{H}_2\text{O}$.

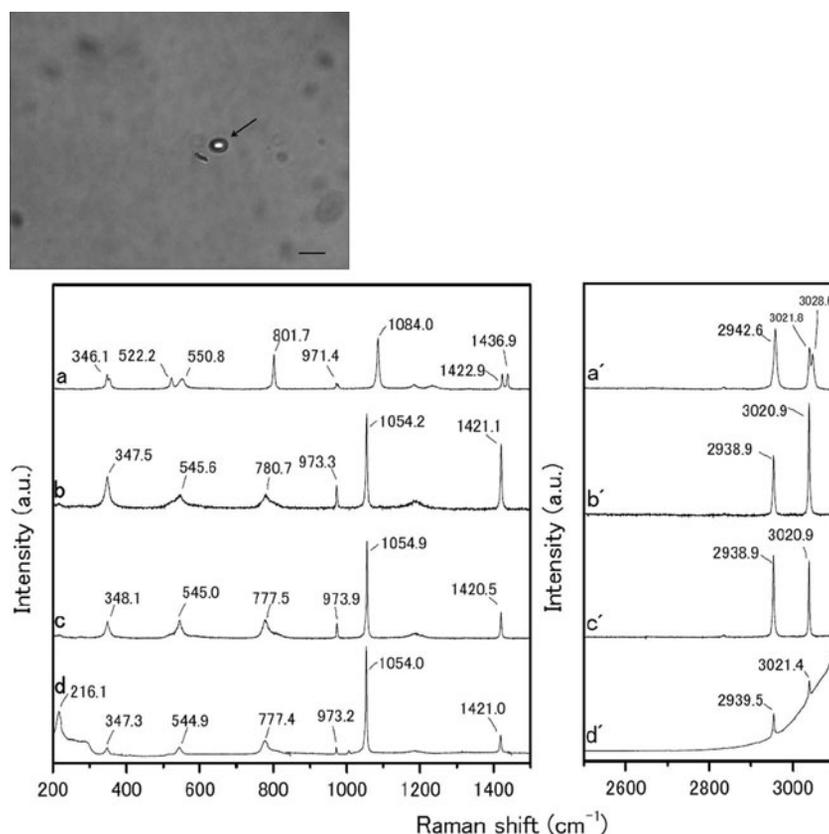


Fig. 3. Optical microscopy of a methanesulfonate salt particle (arrow) found in the ice sample at 576.5 m depth (24.5 ka BP). The scale bar is 10 μm . The Raman spectra from (a–d) 200–1500 cm^{-1} and (a'–d') 2500–3100 cm^{-1} shown below the photograph come from (a) $(\text{CH}_3\text{SO}_3)_2\text{Mg}$, (b) $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$ (after standing at room temperature for several days and measured at -30°C), (c) $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$ and (d) the methanesulfonate salt inclusion from the Dome Fuji ice core (the particle in the photograph).

2009). The peaks around 2930–2950 and 3010–3030 cm^{-1} can be assigned to CH_3 symmetric stretching and CH_3 asymmetric stretching respectively (Durig and others, 2000; Socrates, 2001; Barletta and others, 2009).

The Raman spectra of the simple salts are significantly different from those of the corresponding hydrated crystals. The differences are especially evident in the SO_3^{2-} , OH and CH_3 stretching frequencies. The SO_3^{2-} stretching mode generates a peak at 1084.0 cm^{-1} for the $(\text{CH}_3\text{SO}_3)_2\text{Mg}$ salt. The $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot\text{H}_2\text{O}$ solution, on the other hand, has a peak at 1054.2 cm^{-1} . The OH stretching mode does not appear for any simple salt. In contrast, all the solutions have a large peak around 3000–3500 cm^{-1} from OH stretching modes. A more interesting result from these measurements relates to the CH_3 symmetric and asymmetric stretching modes from 3010 to 3030 cm^{-1} . In Raman spectra from simple salts, we detect mode splitting. For instance, the spectrum of $(\text{CH}_3\text{SO}_3)_2\text{Mg}$ has peaks at 3021.8 and 3028.6 cm^{-1} from CH_3 asymmetric stretching (Fig. 3a). The $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot\text{H}_2\text{O}$ sample, on the other hand, produces a single sharp peak at 3020.9 cm^{-1} . These disagreements must be caused by the crystal structure, suggesting that hydrate is forming in the methanesulfonate salt, i.e. the structure in solution is $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$.

3.2. Chemical form of methanesulfonate salts in the ice core

We found microparticles with Raman spectra very similar to those of the methanesulfonate salt $(\text{CH}_3\text{SO}_3)_2\text{Mg}$. Among the six depths considered, particles of this kind are only

present in LGM (576.50 m depth) ice. Furthermore, no other types of methanesulfonate salt were found among the 758 microparticles examined (Fig. 4). The microparticles of $(\text{CH}_3\text{SO}_3)_2\text{Mg}$ were typically several microns in diameter (Fig. 3 photograph).

To identify the chemical form of the methanesulfonate salts present in the ice core, we compared the Raman spectra of microparticles (Fig. 3) with the reference spectra. The Raman spectra of the methanesulfonate salt particles found in the Dome Fuji ice correspond with the $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$ reference specimen, *not* with the anhydrous $(\text{CH}_3\text{SO}_3)_2\text{Mg}$ specimen or any other chemical form.

The Raman spectra of $(\text{CH}_3\text{SO}_3)_2\text{Ca}\cdot n\text{H}_2\text{O}$ and $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$ are very similar in many respects. However, they can be distinguished by their symmetric and asymmetric CH_3 stretching frequencies: 2938.9 and 3020.9 cm^{-1} for $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$, or 2941.6 and 3022.6 cm^{-1} for $(\text{CH}_3\text{SO}_3)_2\text{Ca}\cdot n\text{H}_2\text{O}$. We used Gaussian fitting to estimate the positions and half-bandwidths of the peaks. The corresponding peaks in Raman spectra obtained from methanesulfonate salt particles in the ice core are located at 2938.9 and 3020.9 cm^{-1} . The half-bandwidths of the peaks are 3.9 ± 0.1 and 2.9 ± 0.9 cm^{-1} respectively. The peak locations and half-bandwidths of the $(\text{CH}_3\text{SO}_3)_2\text{Ca}\cdot n\text{H}_2\text{O}$ reference hydrated crystal are 2941.6 cm^{-1} , 3023.1 cm^{-1} and 15.8 ± 0.2 cm^{-1} , 23.7 ± 3.4 cm^{-1} respectively. This evidence strongly suggests that the methanesulfonate salt found in the ice core is $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$ and not $(\text{CH}_3\text{SO}_3)_2\text{Ca}\cdot n\text{H}_2\text{O}$.

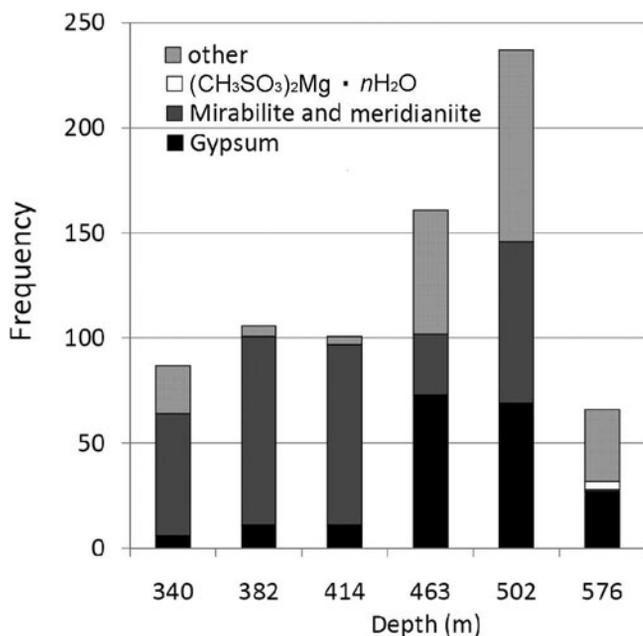


Fig. 4. The distribution of microparticle types found in ice samples taken from Termination I of the Dome Fuji ice core. The climate ranges are 340 m (Holocene), 382–502 m (Termination I) and 576 m (LGM). The numbers of microparticles measured at each depth are 87, 106, 101, 161, 237 and 66 (a total of 758 particles). The numbers of microparticles in the ‘gypsum’ category are 6, 11, 11, 73, 69 and 27 respectively. The numbers in the ‘mirabilite and meridianiite’ category are 58, 90, 86, 29, 77 and 1 respectively. Only four (CH₃SO₃)₂Mg·nH₂O microparticles were found, and these only in the LGM.

To evaluate the structural phase transition during preservation, we left the (CH₃SO₃)₂Mg-H₂O solution in a Petri dish at room temperature for several days to sublimate the water, then measured the Raman spectrum of the sample. The resulting spectrum remains clearly (CH₃SO₃)₂Mg·nH₂O, not (CH₃SO₃)₂Mg (Fig. 3b and c). This result suggests that once (CH₃SO₃)₂Mg·nH₂O has formed it can be preserved as a structure of hydrate. It further suggests that a structural phase transition to the anhydrous state or some other phase could not have occurred.

3.3. Phase diagrams of methanesulfonate salts

The phase transitions of CH₃SO₃Na-H₂O, (CH₃SO₃)₂Mg-H₂O, CH₃SO₃K-H₂O and (CH₃SO₃)₂Ca-H₂O generate typical DSC profiles during the heating process. The phase diagrams of the solutions are shown in Figure 5.

CH₃SO₃Na-H₂O

We performed measurements in the composition range 0–52 wt% (Fig. 5a). The eutectic temperature of this salt is $-29.3 \pm 0.2^\circ\text{C}$. Concentrations higher than 50 wt% CH₃SO₃Na-H₂O were deposited from the solution at room temperature. The enthalpy of fusion at the eutectic composition of 50 wt% is $\Delta H_{\text{Na}} = 12.6 \pm 0.9 \text{ kJ mol}^{-1}$.

(CH₃SO₃)₂Mg-H₂O

We performed measurements in the composition range 0–36 wt% (Fig. 5b). The average eutectic transition temperature over all samples is $-5.0 \pm 0.5^\circ\text{C}$. Above a solute concentration of 14 wt%, salt deposition occurred at room temperature. The enthalpy of fusion at the eutectic composition of 14 wt% is $\Delta H_{\text{Mg}} = 13.3 \pm 0.04 \text{ kJ mol}^{-1}$.

CH₃SO₃K-H₂O

We performed measurements in the composition range 0–60 wt% (Fig. 5c). The average eutectic transition temperature is $-17.8 \pm 0.1^\circ\text{C}$. The enthalpy of fusion at the eutectic composition of 44 wt% is $\Delta H_{\text{K}} = 17.9 \pm 0.08 \text{ kJ mol}^{-1}$.

(CH₃SO₃)₂Ca-H₂O

We performed measurements in the composition range 0–44 wt% (Fig. 5d). The average eutectic transition temperature is $-32.6 \pm 0.2^\circ\text{C}$. At compositions of 44 wt% or higher, our measurements of this compound were not easily reproducible (although the other three species were detectable at their eutectic compositions). We therefore do not have a good estimate for the enthalpy of fusion, but one profile gave $\Delta H_{\text{Ca}} = 0.46 \pm 0.04 \text{ kJ mol}^{-1}$.

4. DISCUSSION

The eutectic temperature of (CH₃SO₃)₂Mg·nH₂O ($-5.0 \pm 0.5^\circ\text{C}$) is much higher than any temperature within the Dome Fuji ice core, from the surface to the lowest depth. The average temperature of the core at 10 m depth is -58°C , and the deepest recorded temperature (from the first Dome Fuji ice-core project, at 2503 m) is $\sim -10^\circ\text{C}$ (Hondoh and others, 2002). This result indicates that (CH₃SO₃)₂Mg·nH₂O is thermodynamically stable in the Dome Fuji ice core between the surface and 2503 m. We shall now propose two formation mechanisms for (CH₃SO₃)₂Mg·nH₂O in Antarctica, one in the ice-sheet firn and the other in the polar atmosphere.

CH₃SO₃H is formed by the oxidation of DMS, which is emitted by biological activity (e.g. Saltelli and Hjorth, 1995) on the open sea. Sea salts are emitted by sea spray and probably from surface sea ice (e.g. Rankin and others, 2002). The chemical reaction of CH₃SO₃H with sea salt can form methanesulfonate salt. In previous studies using ion chromatography on coastal ice cores obtained at Dolleman Island and Siple Dome, Antarctica, (Mulvaney and others, 1992; Kreutz and others, 1998) the concentration of CH₃SO₃⁻ in the winter layer (which was relocated from the summer layer) determined whether sodium methanesulfonate or magnesium methanesulfonate formed at a depth of several meters. The temperatures of the cores at 10 m depth were -16.8°C and -25°C at Dolleman Island and Siple Dome respectively, suggesting that deeper ice might be at even higher temperatures. Taking into consideration our data on the eutectic temperatures of methanesulfonate salts, CH₃SO₃Na·nH₂O (-29.3°C) could not be present in the Dolleman Island ice core but (CH₃SO₃)₂Mg·nH₂O (-5.0°C) should be.

The annual surface temperature (e.g. 1996/97) on the Dome Fuji ice sheet ranges from -79.7°C to -18.6°C , with a mean of -54.4°C (Yamanouchi and others, 2003). The annual average surface temperature during the LGM has been determined from stable-isotope studies as roughly -60°C (Watanabe and others, 2003). There should also have been a seasonal variation around this mean. From previous studies, Na⁺ is the most abundant ion, while Mg²⁺ and CH₃SO₃⁻ are minor ions in both periods, i.e. during the LGM and in current times (e.g. Watanabe and others, 2003; Iizuka and others, 2004, 2006). Thus, these previous studies suggest that CH₃SO₃Na·nH₂O can form at the surface of the Dome Fuji ice during both periods, and that

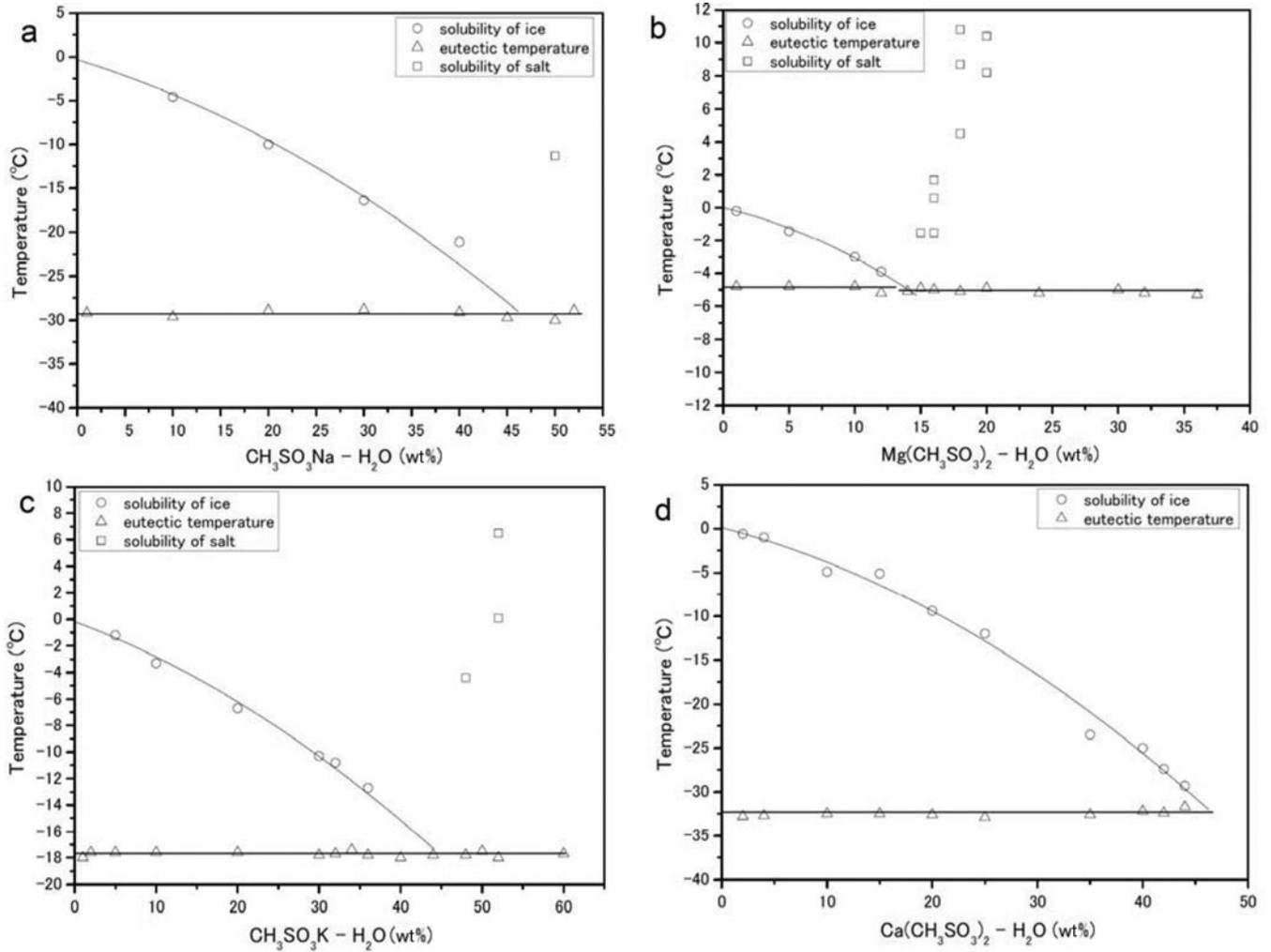
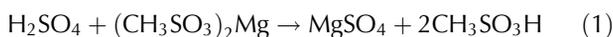


Fig. 5. Phase diagrams (eutectic temperatures are given in parentheses) constructed from DSC measurements. (a–d) correspond to $\text{CH}_3\text{SO}_3\text{Na}-\text{H}_2\text{O}$ ($-29.3 \pm 0.2^\circ\text{C}$), $(\text{CH}_3\text{SO}_3)_2\text{Mg}-\text{H}_2\text{O}$ ($-5.0 \pm 0.5^\circ\text{C}$), $\text{CH}_3\text{SO}_3\text{K}-\text{H}_2\text{O}$ ($-17.8 \pm 0.1^\circ\text{C}$) and $(\text{CH}_3\text{SO}_3)_2\text{Ca}-\text{H}_2\text{O}$ ($-32.6 \pm 0.2^\circ\text{C}$) respectively. The circles, triangles and squares trace the solubility of ice, the eutectic temperature and the solubility of salt respectively.

$(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$ is probably a minor compound in the Dome Fuji ice core. However, among the 758 microparticles analyzed, we found no methanesulfonate salts except for $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$ in LGM ice. The discrepancy can be explained by the surface temperature and acidity of the ice at the time it formed.

Taking into consideration the eutectic temperatures of the Mg and Na salts, LGM summer temperatures at the surface could have been high enough to melt $\text{CH}_3\text{SO}_3\text{Na}\cdot n\text{H}_2\text{O}$ but not $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$. Furthermore, $\text{CH}_3\text{SO}_3\text{H}$ might sublimate from the firn after the $\text{CH}_3\text{SO}_3\text{Na}\cdot n\text{H}_2\text{O}$ melts, lowering the concentration of CH_3SO_3^- in the ice. This situation, dominant in the Holocene and in Termination I, suggests that $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$ could have formed in the Holocene and Termination I sections of the Dome Fuji ice core. According to previous studies of the firn, however, sulfuric acid and sea salt can combine to produce sulfate salts (Iizuka and others, 2006). The high acidity of Holocene and Termination I ice (Fujita and others, 2002) may have allowed the reaction



to occur during the post-depositional process, eliminating $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$ particles from the ice.

In the LGM ice, methanesulfonate salts form by fixation on alkaline particles of marine or continental origin during long-range aerosol transport to polar areas (Delmas and others, 2003). During this period, high levels of dust neutralized the acids, greatly reducing their abundance (Röthlisberger and others, 2003; Iizuka and others, 2008). This neutralization of the LGM ice permits the existence of $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$.

To summarize, we propose that $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$ may have formed in the air and been transported as an aerosol. The salt was then deposited on the ice sheet, where it had a chance to survive and be preserved in LGM ice because the acids were neutralized.

5. CONCLUSION

We found $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$ in the LGM ice from Dome Fuji, and this salt is restricted to the coldest period in the glacial cycle. Our study suggests that $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$ formed in the atmosphere of Antarctica. In acidic Holocene ice, however, $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$ was dissolved by acid after being deposited. Only during the LGM did $(\text{CH}_3\text{SO}_3)_2\text{Mg}\cdot n\text{H}_2\text{O}$ that formed in the atmosphere have a chance to survive on the ice sheet and be preserved in the Dome Fuji ice core.

ACKNOWLEDGEMENTS

We thank all participants in the Dome Fuji ice-core drilling project. We also thank A. Hori and A. Miyamoto for valuable discussions, and the scientific editor B. Kulesa. This study was supported by Grants-in-Aid for Creative Scientific Research (grant Nos. 14GS0202 and 21710002), by the Global COE program (Establishment of Center for Integrated Field Environmental Science) and by a Special Education Study Expense fund (for cooperation between universities) provided by the Ministry of Education, Culture, Sports, Science and Technology (MEXT) and the Japan Society for the Promotion of Science (JSPS), Japan.

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