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Correlation between Kinetics and Chemical Bonding State of Catalyst Surface in Catalyzed Magnesium Hydride

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Abstract

We reported on the hydrogen desorption properties, microstructure, kinetics, and chemical bonding state of catalyst surface for composites of MgH₂ and 1 mol% Nb₂O₅ ball-milled for 0.02 h, 0.2 h, 2 h, 20 h under 1 MPa H₂ atmosphere, as well as hand-mixed (HM) one. Hydrogen desorption properties were significantly improved by ball-milling with Nb₂O₅. Then, we estimated by Kissinger Method the activation energy (E_a) of hydrogen desorption reaction that decreased with the increase of ball-milling time. Especially, E_a of the sample ball-milled for 0.2 h was drastically

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decreased, compared with that of the sample ball-milled for 0.02 h. TEM observations revealed that the distribution of Nb_2O_5 in MgH_2 was gradually improved during ball-milling. On the other hand, we confirmed by XPS that in the sample ball-milled for 0.2 h, $\text{Nb}_2\text{O}_{5-x}$ phase(s) existed at least on the surface. It can be suggested that these deoxidized $\text{Nb}_2\text{O}_{5-x}$ phases eventually decrease E_a as substantial catalyst rather than Nb_2O_5 itself.

Keywords: Hydrogen Storage Materials; MgH_2 ; Oxide Catalyst; XPS; TEM

1. Introduction

Many kinds of hydrogen absorbing materials have been studied in order to store hydrogen, such as sodium alanates [1],[2], metal nitrides [3],[4], as well as MgH_2 [5],[6], which is one candidate because of its high hydrogen capacity (7.6 mass%). However, the absorption and desorption reaction of Mg/MgH_2 itself is too slow. Recently, scientists found out that the composite of MgH_2 and transition metal compounds prepared by ball-milling showed improved kinetics for both hydrogen absorption and desorption [7]-[11], which can be seen as a good indicator for practical application. Barkhordarian et al. reported that the Nb_2O_5 had a superior catalytic

effect for absorption and desorption, compared to other metal oxide catalysts. The magnesium hydride catalyzed with 0.5 mol% Nb₂O₅, which was ball-milled for 100 h, finished full desorption within 90 seconds under vacuum conditions at 300 °C. The desorbed products could also absorb roughly 7 mass% of hydrogen within 60 seconds under 8.4 bar of hydrogen at 300 °C [10]. Hanada et al. reported that the desorbed composite, MgH₂ with 1 mol% Nb₂O₅ milled for 20 h, was able to absorb ~4.5 mass% of hydrogen under lower pressure than 1.0 MPa within 15 seconds even at room temperature [11]. However, the mechanism of this significant catalytic effect has not been clarified yet. Hanada et al. observed the ball-milled MgH₂-Nb₂O₅ mixture by the Transmission Electron Microscope (TEM) and found that the additive particles were homogeneously distributed in MgH₂ particles within nanometer scale, thus leading to the fast kinetics [12]. Additionally, they performed X-ray absorption spectroscopy for investigating the valence of metal oxide additives such as Nb₂O₅ and V₂O₅. The results showed that the metal oxide had been reduced by ball-milling with MgH₂ for 20 h. Therefore, they suggested that the improvement of hydrogen sorption was caused by the catalytic effect of reduced metal oxides [13]. Meanwhile, Porcu et al. confirmed via High Resolution TEM that in the MgH₂-Nb₂O₅ ball-milled composite MgNb₂O_{3.67} was present, which may elaborate on the improved kinetics [14]. With respect to the chemical bonding state of niobium, both results indicated that the

reduced niobium oxide existed in the sample after milling with MgH_2 . However, few works have been done to directly clarify the correlation between the kinetics and the reduced catalyst. The exact mechanism of how the reduced oxides affect the kinetics of reaction should be investigated further in order to develop catalytic effect in hydrogen storage materials, not only in MgH_2 . In this work, we have investigated the kinetics of desorption, the morphology and the chemical state on the surface of the catalyst in $\text{MgH}_2\text{-Nb}_2\text{O}_5$ composites ball-milled for varied time. Especially, we have tried to trace the change of chemical bonding states of catalyst surface by X-ray photoelectron spectroscopy (XPS) under the consideration that the state of catalyst surface is the most important factor affecting the kinetics of desorption.

2. Experimental Procedures

MgH_2 powder was purchased from Alfa Aesar with a purity of 98%. Nb_2O_5 powder with the purity of 99.99% was also purchased from the Kojundo chemical Laboratory, with a size of $\sim 1\ \mu\text{m}$. Ball-milling was performed using a Fritsch P7 device. The samples, mixtures of 300 mg MgH_2 and 1 mol% Nb_2O_5 , were ball-milled with 20 steel balls (7 mm in diameter) for 0.02 h, 0.2 h, 2 h and 20 h. The ball-to-powder ratio was 100:1. The hydrogen pressure inside the milling spot was 1 MPa and the rotating speed was 400 rpm during milling. Also, hand-mixed (HM)

sample with the same composition was prepared by agate mortar. All the samples were treated without exposure to air, because of the instability of MgH_2 . The hydrogen desorption properties were examined by thermal desorption spectroscopy (TDS) with several rates from 1 to 20 $^{\circ}\text{C}/\text{min}$ under a highly pure helium flow. X-ray diffraction (XRD) was performed by using *JEOL JDX-3500* powder diffractometer with $\text{Cu K}\alpha$ radiation. The samples were observed by 200 kV TEM (*JEOL JEM-2010*). XPS was carried out using a *JPS-90MX* photoelectron spectrometer with $\text{Mg K}\alpha$ radiation. The energy scale of the spectrometer is calibrated by setting oxygen (O1s) peak to 532 eV. The spectra were smoothed and subtracted the background according to the Shirley method. Peak separation was performed using Gaussian-Lorentz routines with the ratio of 0.8. The area ratio was fixed to 2/3 due to spin-orbit coupling and the distance 2.7 eV between $\text{Nb3d}_{3/2}$ and $\text{Nb3d}_{5/2}$ peaks.

3. Results and Discussion

Figure 1 shows the profiles of TDS of hydrogen for the composites with a heating rate of 5 $^{\circ}\text{C}/\text{min}$. It can be seen that as the ball-milling time increased, the peak temperatures of hydrogen desorption are decreasing, meaning the catalytic effect was gradually activated and improved during ball-milling. By measuring TDS at different heating rate, the activation energy (E_a) of desorption was calculated by Kissinger

Method [15], according to the following equation;

$$\ln(\beta/T_p^2) = -E_a/RT_p + \ln(k_0R/E_a),$$

where β is the heating rate, T_p indicates the peak temperatures of desorption, R is the gas constant and k_0 is the frequency factor. If plotting $\ln(\beta/T_p^2)$ as the function of the inverse of T_p , straight lines can be obtained, namely the Kissinger plot. From the slope of the straight lines E_a for the HM sample and those ball-milled for 0.02 h, 0.2 h, 2 h and 20 h was estimated to be 147 kJ/molH₂, 138 kJ/molH₂, 82 kJ/molH₂, 70 kJ/molH₂ and 63 kJ/molH₂, respectively. The correlation between E_a and ball-milling time was shown in Fig 2. It can be seen that E_a decreases with the increase of ball-milling time. Especially between the samples ball-milled for 0.02 h and 0.2 h, a large reduction of E_a (~ 56 kJ/molH₂) has been reached, which is worth discussion. This indicates that some changes were brought by ball-milling and hence improved the catalytic effect of Nb₂O₅. To understand the mechanism of the catalytic effect, these changes should be investigated.

XRD patterns of the samples were obtained to evaluate the crystalline information, as shown in Fig 3. The width of the peaks widened and the intensity drastically decreased when the samples were milled for a long time. It indicates that the grain size of both MgH₂ and Nb₂O₅ becomes smaller during ball-milling. This refinement of grains may suggest the decrease of size for both MgH₂ and Nb₂O₅ particles, leading to

a better distribution of catalyst in the samples. TEM observation provides evidence for the inference above. In the bright field images shown in Fig 4a-4e (all of them were in the same scale), two kinds of particles could be distinguished by different contrast. The selected area diffraction was applied for both of them (see area A and B in Fig 4b). In Fig 4g the spots with streak (obtained from area B) are recognized as Nb₂O₅, while in Fig 4f the Debye rings (obtained from area A) are confirmed as Mg and MgO. Therefore, in the bright field images the particles with very dark contrast should be Nb₂O₅ while the others are Mg related phases. MgH₂ was not found because of the fast decomposition under the electron beam, as discussed in other papers [12],[14]. When comparing the images, it can be seen that the Nb₂O₅ particles were gradually attenuated by ball-milling and a homogeneous distribution was reached in the samples milled for long time. In Fig 4e, e.g., the Nb₂O₅ particles imbed homogeneously in the Mg related phases. The size of the Nb₂O₅ particles was estimated to be less than 100 nm. This gradually improved distribution of the catalyst with the refined size should be one factor responsible for the decreased E_a and improved kinetics.

However, we noticed that E_a decreased not gradually but in fact drastically when the sample was milled for 0.2 h. Thus there should be other factors that affect the kinetics more essentially. Under this consideration, XPS measurements were carried out for further investigation. The niobium (Nb3d) spectra were drawn in Fig 5 after

calibration. A decrease of the Nb3d signal with ball-milling can be seen, reasserting the results from elsewhere [16],[17]. The decrease could be explained as the milled MgH₂ particles covered the surface of Nb₂O₅, preventing the detector from getting enough signal from inside. It can be seen that the position of the Nb3d peaks detected almost remain the same, even after the sample was milled for a substantial period. However, the obviously widened peaks were obtained in the sample milled for 0.2 h. Further separation of those peaks shown in the inserted graph resulted in a coexistence of 4 peaks. There are two peaks at 210.2 eV and 207.5 eV, which are corresponding to Nb3d_{3/2} and Nb3d_{5/2} (210.1 eV and 207.4 eV respective in reference to [18]), and two new peaks appearing to the right. Though the signal is not strong enough to identify the exact phase for the new peaks, the shift to lower energy infers that the chemical state of Nb has changed at least on the surface of the sample. A certain or a complication of Nb₂O_{5-x} phase(s) appeared in the sample milled for 0.2 h. The deoxidized Nb₂O_{5-x} on the surface with the valence of Nb less than +5 may act as a more effective catalyst that decreases E_a and also improve the kinetics of desorption. Here it is also noticed that there was a small shoulder at higher energy side (~211 eV). We considered it as noise because fully oxidized niobium (Nb₂O₅) should have the Nb3d peak with the highest binding energy.

4. Conclusion

In conclusion, we were able to perform a study on the HM $\text{MgH}_2\text{-Nb}_2\text{O}_5$ composite as well as those ball-milled for 0.02 h, 0.2 h, 2 h and 20 h under 1 MPa H_2 atmosphere. An improved desorption kinetic was reached by ball-milling with Nb_2O_5 . E_a of HM, 0.02 h, 0.2 h, 2 h and 20 h samples was estimated to be 147 kJ/mol H_2 , 138 kJ/mol H_2 , 82 kJ/mol H_2 , 70 kJ/mol H_2 and 63 kJ/mol H_2 , respectively, thus decreasing in accordance with the improved kinetics. XRD and TEM results showed a decrease of particle size for both MgH_2 and Nb_2O_5 occurred with ball-milling. In the sample milled for 20 h, the Nb_2O_5 particles were found imbedding homogeneously in Mg related phases, with a size of less than 100 nm. The better distribution and the refined size of the Nb_2O_5 may be responsible for the decreased E_a and improved kinetics for desorption. The chemical state of Nb was found changed at least on the surface, as evident by XPS results. The peak separation was performed on the XPS spectrum of the sample milled for 0.2 h, indicating that a certain or a complication of deoxidized $\text{Nb}_2\text{O}_{5-x}$ phase(s) appeared, which may play an more important role in decreasing E_a and improving kinetics for desorption.

Acknowledgement

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Figure captions

Fig. 1. Profiles of TDS of hydrogen for HM sample and those ball-milled for 0.02 h, 0.2 h, 2 h and 20 h

Fig. 2. Correlation between E_a and ball-milling time for HM sample (0 h) and those ball-milled for 0.02 h, 0.2 h, 2 h and 20 h

Fig. 3. X-ray diffraction profiles of HM sample and those ball-milled for 0.02 h, 0.2 h, 2 h and 20 h

Fig. 4. TEM micrographs of HM and ball-milled samples: bright field images of HM (a), 0.02 h (b), 0.2 h (c), 2 h (d), 20 h (e) and the selected area diffraction from area A (f) and B (g)

Fig. 5. XPS spectra of HM sample and those ball-milled for 0.02 h, 0.2 h, 2 h, and 20

h. The insert graph shows the Nb3d peak separation of the sample milled for 0.2 h

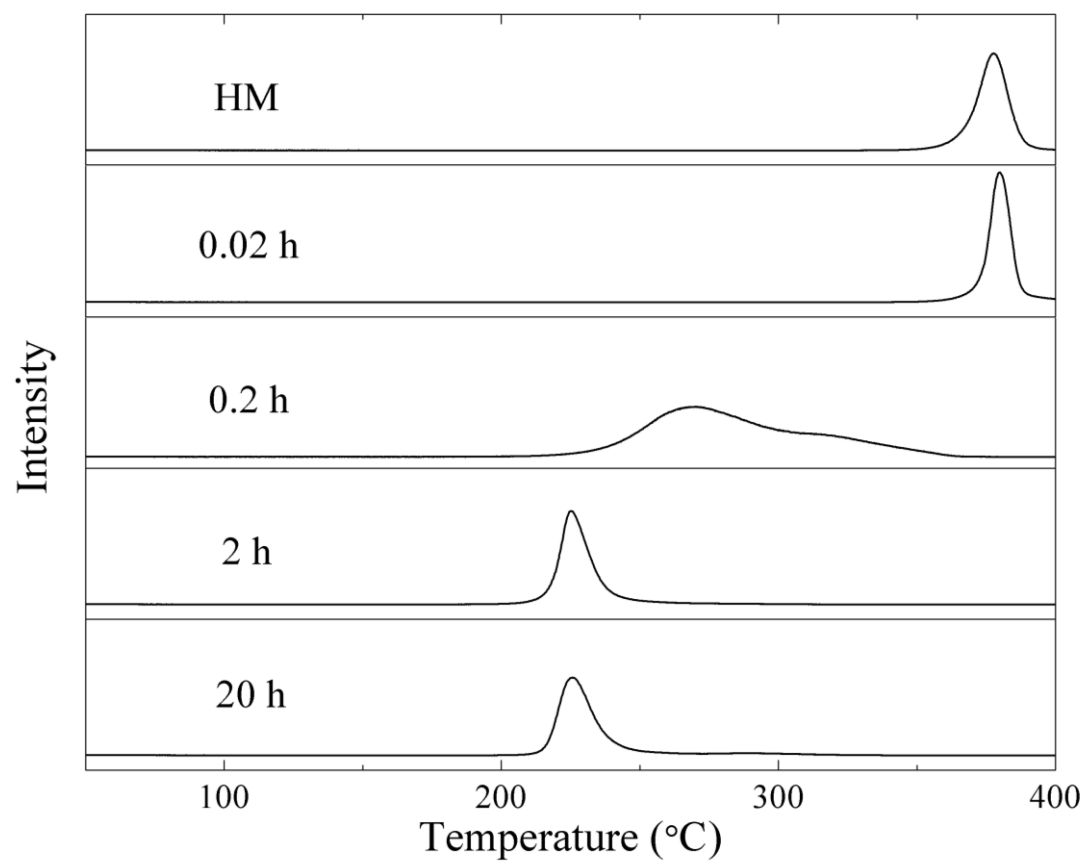


Figure 1

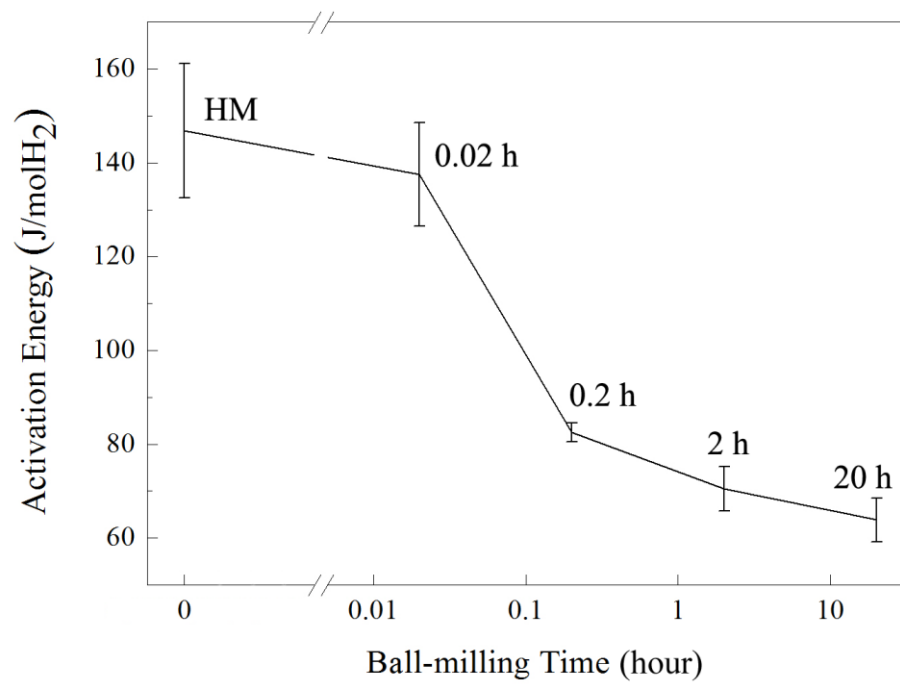


Figure 2

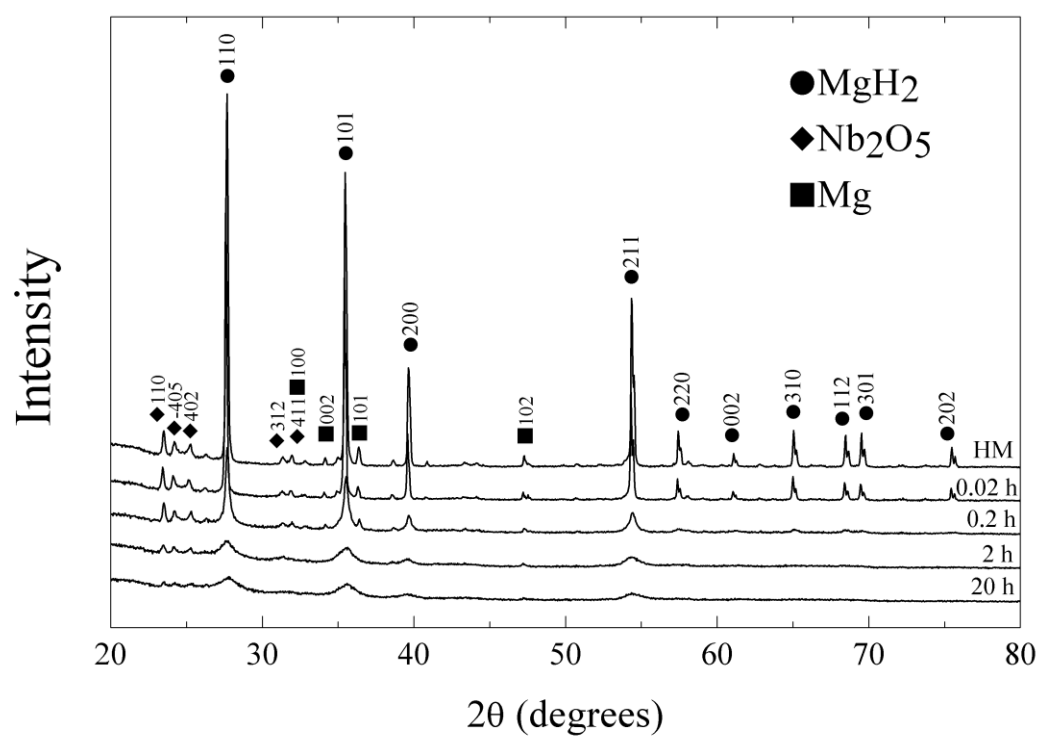


Figure 3

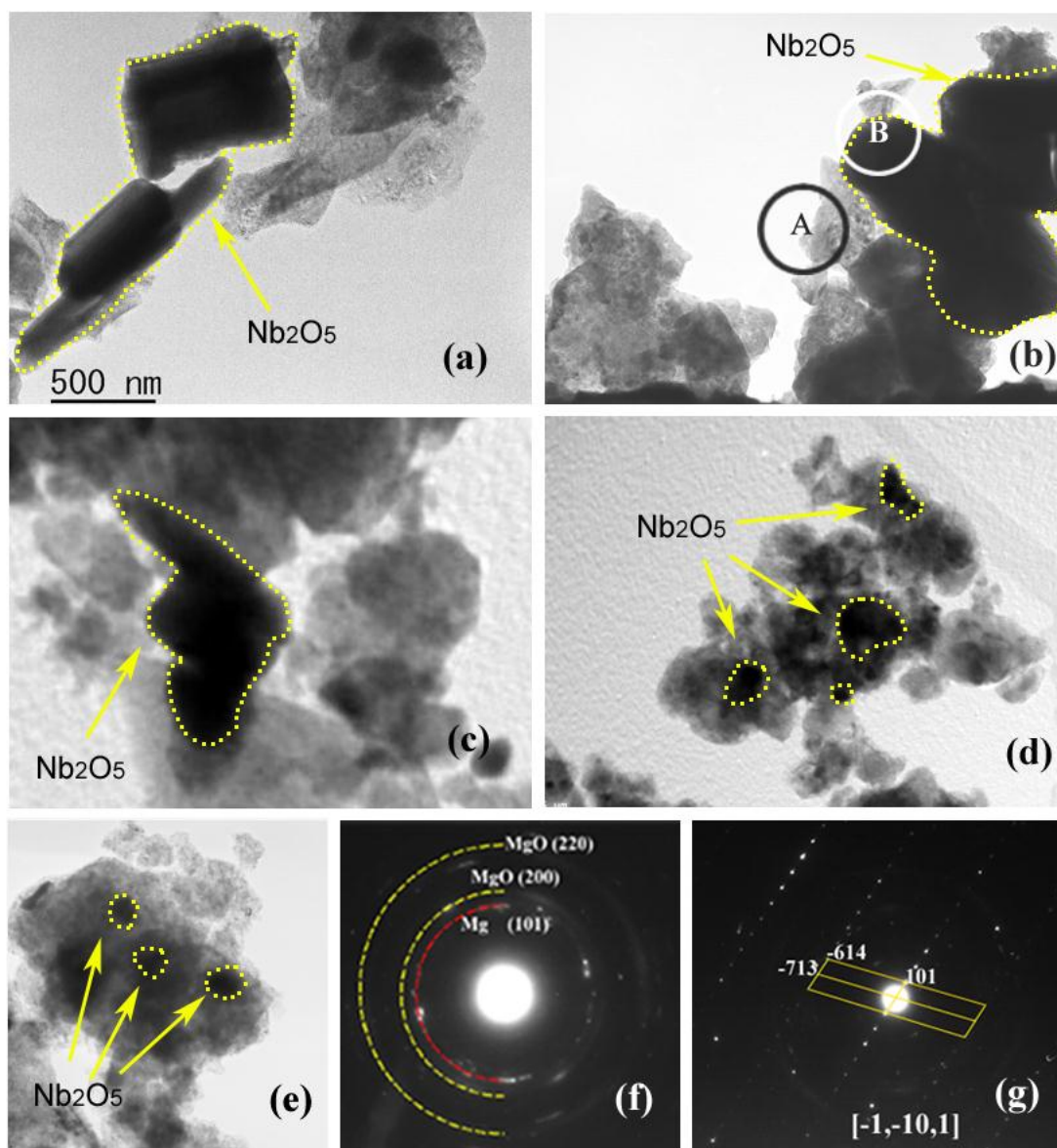


Figure 4

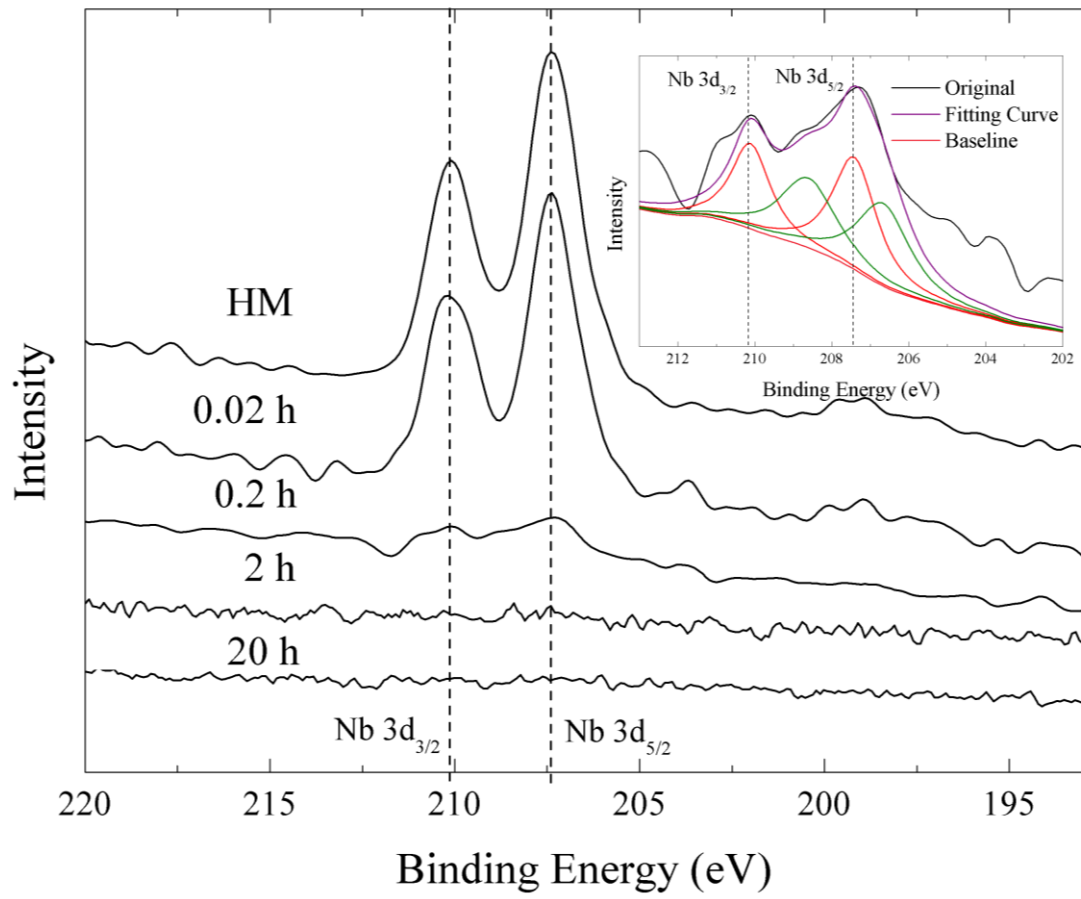


Figure 5