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The roles of pyrite and calcite in the mobilization of arsenic and lead from hydrothermally altered rocks excavated in Hokkaido, Japan

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Abstract

This paper describes the enrichment of hydrothermally altered volcanic and sedimentary rocks with arsenic (As) and lead (Pb), and the effects of pyrite and calcite on the mobilities and release mechanisms of these toxic elements under oxic and anoxic conditions. Enrichment of the altered rock with As and Pb predominantly occurred in precipitated pyrite grains and not on the alumino-silicate minerals making up the matrix of the rock. Arsenic was incorporated in pyrite grains formed during alteration in both volcanic and sedimentary rocks, but Pb was only found in the pyrite grains of the volcanic rock samples. When in contact with water, altered volcanic rocks had acidic pH while altered sedimentary rocks had alkaline pH. The mobilities of both As and Pb from the altered rocks were enhanced at acidic and alkaline pH and a minimum was observed in the circumneutral pH under both oxic and anoxic conditions. The absence of O₂ retarded the oxidation of pyrite most notably in the alkaline region but not in the acidic and circumneutral pH. The absence of CO₂ increased the pH of samples with significant calcite content but did not affect those containing substantial amounts of pyrite. Increasing the CO₂ also had insignificant effect on the concentrations of As and Pb in the leachate. The mechanisms controlling the mobilization of As and Pb from these rocks like dissolution of soluble secondary minerals, pyrite oxidation and adsorption were all related to pyrite while the pH of the rock when in contact with water was controlled by pyrite and calcite. Thus, excavated waste rocks that have been altered can be grouped based on the relative abundance of pyrite and calcite and their pH when in contact with water.

Keywords: Pyrite; calcite; hydrothermal alteration; arsenic; lead; leaching

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1. Introduction

Arsenic (As) and lead (Pb) are elements well known for their toxicity when ingested, inhaled or absorbed through the skin. Arsenic-contaminated groundwater used for drinking and cooking is the cause of health problems such as keratosis and cancers reported in Bangladesh and the West Bengal area of India (Acharyya *et al.*, 2000; Akai *et al.*, 2004; Burgess *et al.*, 2010; Das *et al.*, 1996; Dowling *et al.*, 2002; Nickson *et al.*, 2000). Although the source of contamination is believed to be natural, cases elsewhere such as those reported in Thailand, Ghana and some areas of the United States were due to the improper disposal of contaminated soil or rock from mining activities (Bowell, 1993; Choprapawon and Rodcline, 1997; Welch *et al.*, 1999; Williams, 1997). Lead contaminations of soils have also been reported mostly from firing ranges and mining areas (Dermatas *et al.*, 2006; Rabinowitz, 2005). Chronic Pb poisoning causes paralysis of the peripheral motor nerves, anemia, kidney damage, abnormal fetal development, and abnormal neurological development and function (Hammond, 1977; Moore and Goldberg, 1985; Sengupta, 2002). Because of this, extensive studies have been made to understand the sources, speciation, mobilization and remediation of these toxic elements especially in soils and groundwater (Brannon and Patrick, 1987; Lowers *et al.*, 2007; Moore *et al.*, 1988; O'Day *et al.*, 2004; Peters and Blum, 2003; Savage *et al.*, 2000; Smedley and Kinniburgh, 2002).

Among the known sources of As and Pb, hydrothermally altered rocks are important especially in active volcanic regions like Japan. These rocks are enriched with metals like Pb, chromium (Cr), copper (Cu), gold (Au) and silver (Ag) due to the action of hydrothermal solutions/fluids. Hydrothermal solutions/fluids containing solutes derived from magmatic water or from the wall rock migrate upward and the changes in equilibrium with space and time allow the deposition of these solutes (Pirajno, 2009). Hydrothermally altered rocks have been extensively studied

especially in the context of ore formation, metal enrichment, and mineralization (Aiuppa *et al.*, 2006; Allen and Hahn, 1994; Franzson *et al.*, 2008; Halbach *et al.*, 1993; Horton *et al.*, 2001; Huston *et al.*, 1995; Marques *et al.*, 2010; Martin-Crespo *et al.*, 2004; Ostwald and England, 1977; Pokrovski *et al.*, 2007), but very few have investigated the potential leaching of toxic elements from these rocks when exposed to the environment.

We have been studying in detail the mobilities of As and Pb from these high-volume and low-content sources with the primary goal of developing an alternative method for their disposal. Special landfills are currently used, but this is very expensive and unsustainable due to the large volume of rocks excavated that other methods should be explored. One of the biggest challenges encountered in the design of an alternative disposal method is to establish a simple but robust classification scheme for these excavated altered rocks. This is of primary importance because of their complex properties and large spatial variations in chemistry and mineralogy caused by hydrothermal alteration. The number of potentially toxic elements in these rocks is also variable, that is, other harmful elements like chromium (Cr), selenium (Se) and boron (B) are also found together with As and Pb in some excavated rocks. Moreover, the leaching behaviors of these toxic elements differ from each other even under the same geochemical conditions so that classifying these rocks based on their properties and leaching behaviors would maximize the efficiency of countermeasures incorporated in the design of the new disposal method. Developing this scheme requires the evaluation of processes involved in the enrichment of the altered rocks with respect to the toxic elements and the mechanisms controlling their mobilization from different kinds of altered rocks. Our previous papers showed that the mechanisms controlling the release of As and Pb from altered rocks were largely dependent on the pH and oxidation-reduction potential (Eh) similar to

their observed behaviors in contaminated soils and groundwater (Igarashi *et al.*, 2008; Tabelin and Igarashi, 2009; Tabelin *et al.*, 2010; Tabelin *et al.*, 2012a).

The samples used in our previous paper (Tabelin and Igarashi, 2009) were all taken from borehole cores of a single tunnel project. In contrast, the samples used in this study were collected not only from borehole cores but also from the bulk excavated rocks of several tunnel projects. Moreover, our previous publications did not address the following issues: the important minerals incorporating As and Pb in altered rocks, comparison of the leaching behavior and release mechanisms of borehole core and bulk excavated rock samples under oxic conditions, the leaching behavior and release mechanisms of As and Pb under anoxic conditions in the entire pH range, the individual effects of O₂ and CO₂ on As and Pb leaching, and the influence of minor minerals like pyrite and calcite on the mobilization of these toxic elements from altered rocks. All of these issues were elucidated in this study and the collected data were used to develop a classification scheme to group excavated rocks based on their chemical and mineralogical properties as well as the dominant release mechanisms of As and Pb.

2. Materials and methods

2.1 Materials

Hydrothermally altered volcanic and sedimentary rock samples used in this study were obtained from tunnel projects around the island of Hokkaido, Japan. Altered andesite and tuff were collected from a former gold mine with epithermal gold-silver-copper deposit (location: Teine, Sapporo City) (Tabelin and Igarashi, 2009; Tatsuhara *et al.*, 2012) while the altered sedimentary rock samples were obtained from two different tunnel projects for roads: one was mainly slate hosting silicified veins located in the town of Kamikawa (Tabelin *et al.*, 2012a) and the other was composed of mudstone and sandstone with mild alteration from the town of Honbetsu near Obihiro City (Tabelin

et al., 2012b). Details of the tunnel projects and a brief geological history of the samples are summarized in Table 1. Borehole core samples were collected by diamond core drilling prior to the blasting of the tunnel while bulk excavated rocks were sampled from interim storage sites that had been exposed to the atmosphere for ca. 6 months. Samples collected from the bulk excavated rock were air dried, coarsely crushed and sieved while those from the borehole cores were directly crushed and sieved (<2 mm). Samples taken from the <2 mm fraction were further ground to <50 µm in preparation for the chemical and mineralogical analyses. Polished thin sections of the rock samples were also prepared for the microscopic, SEM-EDS and EPMA analyses.

2.2 Chemical and mineralogical analyses of the rock samples

The bulk chemical composition and mineral constituents of the rock samples were determined using X-ray fluorescence spectrometry (SPECTRO XEPOS, Rigaku Corporation, Japan) and X-ray diffraction spectrometry (MULTIFLEX, Rigaku Corporation, Japan), respectively. Both analyses were done using pressed powders of the sample. Loss on ignition (LOI) was measured by heating the samples (<2mm) inside a furnace for 1 hour at 750⁰C after oven drying at 110⁰C for 24 hours.

Light and electron microscopes were used to examine the properties of pyrite crystals found in the altered rocks. Polarized and reflected light microscopes (open and crossed Nicols) were used to analyze the kind of alteration that had affected the rocks. A scanning electron microscope (SEM) was further used to investigate the crystal morphology of pyrite grains found in the rock. The SEM also has energy dispersive X-ray fluorescence spectrometry (EDS) capabilities and was used to take elemental maps of rock surfaces (SSX-550, Shimadzu Corporation, Japan). The SEM was operated using a 30.0 kV accelerating voltage while the EDS maps were taken at 13,000 cps with 1000 ms time constant. An electron probe micro analyzer (EPMA) (JEOL Corporation, Japan) was utilized

to measure the relative abundance of As and Pb in pyrite. All EPMA analyses were conducted using a 2.0×10^{-8} A, 15 kV primary electron beam and ZAF correction was utilized for the quantification of trace elements. Iron (Fe), sulfur (S), As and Pb concentrations were quantified using peak counting times of 20 s for Fe and S, and 30 s for As and Pb. Background counting times of 5 s were used for Fe and S while background counting times of 10 s were used for As and Pb. $L\alpha$ X-ray lines were analyzed for As, $K\alpha$ X-ray lines for Fe and S, and $M\alpha$ X-ray lines for Pb. The detection limit of the EPMA using the above conditions is 100 ppm (0.01wt%).

2.3 Batch leaching experiments

Batch leaching experiments were conducted under oxic (*i.e.*, ambient) and anoxic (*i.e.*, inside a glove box) conditions by mixing 15 grams of sample (<2 mm) and 150 ml of prepared leachants. Hydrochloric acid (HCl) and sodium hydroxide (NaOH) solutions of varying concentrations were used as leachants. Deionized water (18 M Ω -cm) used in the preparation of the leachants was obtained from a Millipore Milli-Rx 12 α system (Millipore Corporation, USA). Triplicate leaching experiments were done for samples SR1 and VR4 under oxic conditions using deionized water as leachant to estimate deviations of the leachate chemistry due to the heterogeneity of the crushed rock samples. Two kinds of anoxic experiments were conducted: one was under pure argon (Ar) gas and the other under mixed CO₂ and Ar gases. In the anoxic experiments under pure Ar gas, the pH of the leachants was adjusted using reagent grade NaOH or HCl solution. On the other hand, only deionized water was used as leachant under mixed CO₂ and Ar gases (*i.e.*, 1, 3 or 30% of CO₂ gas with Ar gas as the remainder). To ensure that truly anoxic conditions exist in these experiments, leachant solutions were left in the glove box for ca. 10 minutes while the vacuum pump was still running. This resulted in the “boiling” of the solution and expulsion of dissolved gases. Also, dissolved O₂ (DO) was monitored prior to the anoxic experiments using an oxygen probe (Water

checker IWC-5, Sansyo Corporation, Taiwan) and were always 0% DO. Anoxic leaching experiments were conducted under different gas compositions to identify the individual effects of O₂ and CO₂ on the mobilities of As and Pb from the rock. In all of the leaching experiments, mixing was done at 120 rpm using a reciprocal shaker. The pH and Eh of the suspensions were measured after 24 hours followed by filtration through 0.45 µm Millex[®] sterile membrane filters (Millipore Corporation, USA). Filtrates collected were then analyzed for their toxic and major elements concentrations. The mixing time was fixed at 24 hours based on the study of Igarashi *et al.* (2008), which showed that apparent equilibrium was reached after 1 day of leaching for hydrothermally altered crushed rocks.

2.4 Chemical analysis of the leachate

The chemical analysis of the leachates was carried out using an inductively coupled plasma atomic emission spectrometer (ICP–AES) (ICPE–9000, Shimadzu Corporation, Japan). Arsenic concentrations less than 0.1 mg/L were analyzed using the hydride generation process coupled with the ICP-AES. In preparation for the hydride generation analysis, 30 ml of leachate was mixed with 15 ml of 12 M HCl, 2 ml of 20% potassium iodide (KI) solution, 1 ml of 10% ascorbic acid solution and diluted with deionized water to 50 ml. All chemicals used in the preparation and analysis were reagent grade. Lead concentrations less than 0.1 mg/L were analyzed using an ultrasonic aerosol generator connected to the ICP-AES. The concentrations of other coexisting ions (*e.g.*, Ca²⁺ and Fe) were also analyzed using ICP-AES including SO₄²⁻, which was the major species of S found in the leachate using anion chromatography (ICS – 90, Dionex Corporation, USA). The standard ICP-AES method (*i.e.*, without attachment) has a margin of error of ca. 2 – 3% while the more sensitive hydride generation process and ultrasonic aerosol generator have uncertainties of ca. 5%. The detection limits of the standard ICP-AES method for Ca, Fe, S, Pb, As, Al and Si are approximately

0.1, 2, 0.02, 9, 40, 20 and 4 $\mu\text{g/L}$, respectively. The hydride generation method for As has a detection limit of 0.1 $\mu\text{g/L}$ while the ultrasonic aerosol generator attachment has a detection limit of 1 $\mu\text{g/L}$.

2.5 Geochemical modeling

Thermodynamic calculations using geochemical models in conjunction with our experimental results could provide additional insights into the processes occurring in a rock-water-air system under equilibrium conditions. PHREEQC (Parkhurst and Apello, 1998) was used to calculate the saturation indices of several minerals important in the mobilization of As and Pb like Fe and Al-oxyhydroxides/oxides, carbonates and sulfates. It was also used to simulate ideal systems composed of one or more of the following minerals: calcite (CaCO_3), pyrite (FeS_2), gypsum ($\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$) and melanterite ($\text{FeSO}_4 \cdot 7\text{H}_2\text{O}$). In these simulations, the amounts of calcite and/or pyrite were assumed from their relative abundance in the rock (XRD results). For example, when calcite is present as a minor mineral and pyrite as a trace mineral the assumed amounts are 10 and 1 mol, respectively. Gypsum and melanterite were selected as representative soluble Ca and Fe-sulfates, respectively. In addition, PHREEQC was also used to simulate the effects of O_2 and CO_2 on the leachate chemistry of an ideal system of calcite-pyrite-soluble Ca and Fe-sulfates. Moreover, the RXN program of the Geochemist's workbench[®] (Bethke, 1992) was used to evaluate changes in the precipitation of Fe-oxyhydroxides and oxides with or without O_2 . Precipitation reactions without O_2 were calculated using N_2 in place of Ar because the aforementioned geochemical model did not include Ar gas used in this study.

3. Results

3.1 Sample Characterization

The bulk chemical compositions of the altered rock samples are summarized in Table 2. This table shows that altered andesite and tuff from the Teine deposit have S contents greater than 1 wt% and As contents greater than 100 ppm, which is ca. 100 times higher than the average As content of igneous rocks (average As content of igneous rocks is ca. 1.5 ppm (Ure and Berrow, 1982)). In the altered sedimentary rocks, S and As contents are less than 0.5 wt% and 24 ppm, respectively and are significantly lower than those found in the altered volcanic rocks. Lead content is also high in the altered volcanic rock samples (>100 ppm), but was not detected in the altered sedimentary rock samples. Although both kinds of altered rocks are predominantly composed of silicate minerals (*i.e.*, quartz and/or feldspar), the altered volcanic rock samples contain pyrite as a minor mineral while the altered sedimentary rock samples have calcite as a minor mineral (Table 3). Galena (PbS) was only detected as a trace mineral in one of the altered volcanic rock samples (VR3).

3.2 Hydrothermal alteration of the sedimentary and volcanic rocks

Fractures running through the rock (sample SR3) and the changes that happened on the rock grains located along these fractures are clearly visible in Figure 1. Also, this figure shows pyrites formed inside and outside the fractured zone. Pyrite grains located outside the fractured zone were formed inside the calcified remains of a marine protozoan known as foraminifer (Figure 2). The crystals inside the calcareous protozoan shell are botryoidal (*i.e.*, globular external form resembling a bunch of grapes) with surfaces bounded by the foraminifer shell. On the other hand, smectite and illite are observed inside the fractured zone together with pyrite as illustrated in Figure 3. The morphology of pyrite located in the fractured area is framboidal with its distinct spherical equi-granular micro crystals of ca. 2 μm diameter that are closely packed together (Figure 3(b)). Our EPMA results show that this framboidal pyrite has incorporated As amounting to as much as 0.13 wt% (1,300 ppm) while those formed inside the foraminifer shell did not have measurable amounts of As in the

crystal (Table 4). The amounts of Pb in pyrite grains formed inside and outside the fractured zone were all less than the detection limit of the EPMA (Table 4). Also, illite and smectite surrounding the pyrite grains inside the fractured area have measurable amounts of As and Pb (Table 4).

Evidences of alteration were also observed in the altered volcanic rock samples as illustrated in Figure 4(a). Several altered plagioclase grains are visible together with alteration products like illite and pyrite. Pyrite crystals found in the altered volcanic rock samples have angles characteristic of hexagonal, cubic, or pyritohedron morphologies, but their complete shapes were not visible because the samples used were polished sections (Figure 4(b)). The SEM-EDS maps of one of these crystals (Figure 5) indicate that As is preferentially distributed into pyrite and not with the surrounding alumino-silicate minerals. It was also observed that the relative abundance of As in pyrite was higher around growth areas (Figure 5(c)). In Figure 5, the pictured pyrite grain is located on “top” of the alumino-silicate minerals, therefore the silicon (Si) and aluminum (Al) present inside the pyrite grain is probably a result of over-polishing and not an integral part of the pyrite crystal. The relative abundance of As and Pb in pyrite found in the altered volcanic rock are summarized in Table 5. The average As content of pyrite found in these rock samples is 2,800 ppm with an observed range of <100 ppm to over 10,000 ppm. On the other hand, the average Pb content is 1,300 ppm with a range of <100 to 2,600 ppm. Statistical analysis (F- and T-tests) showed that there were insignificant differences in the As content of pyrites found in altered volcanic and sedimentary rocks. Similarly, statistically significant differences in the As content of pyrites from VR1, VR2 and VR3 were not observed. In terms of the Pb content of pyrites, there were no statistical differences between the various volcanic rocks used in this study. Unfortunately, a similar statistical comparison of the Pb content of pyrites depending on rock type could not be made because of the

very low Pb content of pyrites found in altered sedimentary rock (Note: Details of the statistical analysis are provided as supplementary tables).

3.3 Effects of O₂ and CO₂ on arsenic and lead leaching

Triplicate leaching experiments of samples SR1 (altered sedimentary rock) and VR4 (altered volcanic rock) show that the standard deviations in pH values and concentrations of As, Ca²⁺, SO₄²⁻ and Fe were all very low (Tables 6 and 7), which indicate that experimental errors associated with the heterogeneity of the altered rocks were negligible and the samples were properly and adequately homogenized. The leachate chemistry of the altered rock samples without pH adjustment is summarized in Table 8. Altered rocks containing pyrite as a minor mineral (all of the altered volcanic rock samples) had acidic leachate pH while those with calcite as a minor mineral (all of the altered sedimentary rocks) had alkaline pH. The concentrations of As in the leachate of altered volcanic rock samples were higher (45.4 – 195 µg/L) than those from the altered sedimentary rocks (17.3 – 28.3 µg/L). The altered volcanic rocks also leached out Pb amounting to as much as 2.12 mg/L. In the altered volcanic rocks, the leachate of the sample taken from the bulk waste (VR4) had lower pH and significantly higher concentrations of Fe and SO₄²⁻ than the borehole core samples (VR1, VR2 and VR3). A similar trend was also observed in the leachate concentrations of Fe and SO₄²⁻ from borehole core (SR 2) and bulk (SR 1) samples of the altered sedimentary rock.

Table 9 illustrates the changes in the leachate chemistry of the altered volcanic rock as a function of gas composition. The pH of the leachate did not change dramatically with or without O₂ and CO₂, but the mobilization of As was enhanced from ca. 12 µg/L without O₂ to 102 µg/L with O₂. This eight-fold increase in As concentration was accompanied by significant increases in the Fe and SO₄²⁻ concentrations in the leachate (Table 7). In contrast, significant variations in Pb

concentrations with changing O₂ and CO₂ compositions were not observed in the leachate of the altered volcanic rock. The leachate pH of altered sedimentary rock samples containing calcite was dependent on CO₂ as illustrated in Table 10. Increasing the amount of CO₂ from 0 to 30% dramatically decreased the pH from 9.2 to 6.1 accompanied by a significant increase in Ca²⁺ concentration. However, the highest concentration of As in the leachate was observed under oxic conditions when both O₂ and CO₂ are present and not at higher CO₂ values.

3.4 Effects of pH on arsenic and lead leaching under oxic conditions

The leaching behavior of As in altered volcanic and sedimentary rocks across the entire pH range under oxic conditions is illustrated in Figure 6. The results show similar leaching behaviors in both altered rocks, that is, more As were released under alkaline and acidic conditions with minimum in the circumneutral pH range (6 – 8) and were consistent with our earlier published works (Tabelin and Igarashi, 2009; Tabelin *et al.*, 2010). With pH variation, leachates from the borehole core samples of the altered volcanic rock (VR1, VR2 and VR3) had higher concentrations of As than the bulk excavated rock samples (VR4) especially around the acidic and alkaline pH range (Figure 6(a)). In contrast, As concentrations in leachates from borehole core of the altered sedimentary rock (SR2) were lower than those collected from the bulk rock (SR1) in the acidic and alkaline pH range (Figure 6(b)). In the acidic region, concentrations of As in the leachates of altered volcanic and sedimentary rock samples were found around a similar range of values (altered sedimentary rocks: 60 – 400 µg/L; altered volcanic rocks: 100 – 700 µg/L). Strong positive ($R > 0.90$) and statistically significant ($p < 0.05$) correlations in the concentrations of Ca²⁺, Fe and As in both types of altered rocks were also observed under these conditions. In the alkaline region, concentrations of As in the leachates of the altered volcanic rocks were in the range of ca. 500 - 3000 µg/L, which were more than 20 times higher in comparison to those observed in the altered sedimentary rock samples

(range: 20 – 150 $\mu\text{g/L}$). The SO_4^{2-} curves of the altered volcanic rock increased with increasing pH similar to that of pure pyrite while those of the sedimentary rock samples decreased with pH (Figure 7). Moreover, SO_4^{2-} concentrations in the leachate of altered volcanic rocks were more than 10 times higher than those from the altered sedimentary rocks.

Figure 8 illustrates that the leaching behavior of Pb is also enhanced in the acidic and alkaline pH regions with minimum at the circumneutral pH range. However, the amount of Pb leached out appears to be more significant in the acidic than the alkaline region since as much as 2.39 mg/L of Pb was released at pH 2.6. Although the amount of Pb in the alkaline region was lower, it still amounted to 0.23 mg/L at pH 12. Also, Pb concentrations of leachates from the bulk rock were higher in the acidic region and slightly lower in the alkaline region than those from the borehole core samples. Similar to As, strong positive correlations ($R > 0.94$) that were statistically significant ($p < 0.025$) also existed between Ca^{2+} , Fe and Pb in the acidic region. Unfortunately, only altered volcanic rocks have significant concentrations of Pb so that a comparison of its leaching behavior depending on the type of rock could not be made.

3.5 Effects of pH on arsenic and lead leaching under anoxic conditions

Figure 9(a) illustrates the leaching behavior of As under anoxic conditions throughout the entire pH range from the altered volcanic rock samples. The leaching behavior of As from these rocks was similar with or without O_2 and CO_2 , that is, As leaching was enhanced in both acidic and alkaline pH regions. Anoxic conditions clearly reduced the mobilization of As in the alkaline region, but had minimal effect in the acidic and circumneutral pH regions. Similarly, the mobilization of Pb was pH dependent even under anoxic conditions as shown in Figure 9(b). The amount of Pb released from the rock samples also decreased in the alkaline region under anoxic conditions similar to that of As.

At pH less than 4, the concentrations of Pb under oxic and anoxic conditions were identical. However, between pH 4 – 6, the concentration of Pb in the leachate was conspicuously higher under anoxic condition. In this pH region, Fe concentrations were also significantly higher under anoxic than oxic conditions, which is identical to those observed with Pb (Figure 10(a)). Concentrations of SO_4^{2-} in the leachate also decreased dramatically under alkaline conditions in the absence of O_2 , but were similar under circumneutral and acidic pH (Figure 10(b)). Moreover, the decrease of As and Pb concentrations in the alkaline region under anoxic conditions occurred in conjunction with the decrease in SO_4^{2-} concentrations of the leachate.

The leaching behavior of As from altered sedimentary rocks under anoxic condition had the same U-shaped curve observed under oxic conditions as illustrated in Figure 11. However, there are three distinct regions in these curves where the concentrations of As changed significantly in the absence of O_2 and CO_2 . Region 1 (R1) is located in the acidic region at pH less than 4, region 2 (R2) is found in between pH 4 – 8 and region 3 (R3) is in the alkaline region at pH greater than 8. In R1 and R3, the amount of As released from the altered sedimentary rock decreased under anoxic conditions while in R2, the concentration of As in the leachate increased. The effect of anoxic conditions on the concentration of Fe was only observed at pH < 4 (Figure 12(a)). At higher pH values, low concentrations of Fe in the leachate made the differentiation difficult. Significant changes in the concentrations of Ca^{2+} and SO_4^{2-} were also observed as illustrated in Figures 12(b) and (c). Under anoxic conditions, the concentrations of Ca^{2+} at pH < 4 decreased but not at pH > 4. For SO_4^{2-} , anoxic conditions increased its concentration in the leachate at pH < 8 but decreased at pH > 8.

3.6 Geochemical modeling

Tables 11 and 12 show the saturation indices of Pb, Ca and Fe minerals in the leachate of samples VR1 (altered volcanic rock) and SR1 (altered sedimentary rock), respectively. The pe values used in the modeling were calculated from the measured Eh of the leachates. In the altered volcanic rocks, precipitations of Pb minerals like Pb(OH)_2 and hydrocerussite ($\text{Pb(CO}_3)_2(\text{OH})_2$) are thermodynamically possible under alkaline conditions while Fe oxides/oxyhydroxides precipitates may form at $\text{pH} > 3$. Calcite is predicted to precipitate in the leachates of the altered sedimentary rock under alkaline conditions together with the oxides and oxyhydroxides of Fe. In both types of rocks, the stability of Fe-oxyhydroxides/oxides, calcite, and Pb(OH)_2 decreases with increasing acidity. Furthermore, modeling results summarized in Table 13 show that the equilibrium constants (K) of Fe-oxyhydroxides/oxides precipitation reactions with a more inert gas (N_2) are lower in comparison to similar reactions with O_2 , which indicates that the extent of Fe-oxyhydroxides/oxides precipitation is minimized under anoxic conditions.

Table 14 shows the pH and concentrations of Ca^{2+} , Fe and SO_4^{2-} in several simulated systems and those measured in sample SR1. Simulated pH and concentrations of Ca^{2+} and SO_4^{2-} in the calcite-pyrite-gypsum-melanterite ideal system closely resembled those measured in sample SR1, which suggest that the pH of the altered rock-water system as well as the concentrations of Ca^{2+} and SO_4^{2-} in the leachate could be adequately explained by these four mineral phases. Inclusion of pyrite in the calcite-only system decreased the pH by 0.6 pH units. Similarly, addition of soluble Ca and Fe-sulfate minerals in the calcite-pyrite system further decreased the pH by ca. 0.4 pH units. In total, addition of pyrite and soluble minerals into the calcite-only system resulted in a pH decrease of ca. 1 pH unit. Contribution of calcite dissolution alone on the total concentration of Ca^{2+} is estimated at ca. 19%. However, this value roughly doubled to ca. 37% after the introduction of pyrite probably due to the increased dissolution of calcite to neutralize the acidity produced by

pyrite oxidation. Similarly, ca. 19% of the total SO_4^{2-} concentration of the leachate could be attributed to pyrite oxidation. This means that majority of the Ca^{2+} (63%) and SO_4^{2-} (81%) dissolved in the leachate originates from the dissolution of soluble Ca and Fe phases. Concentrations of Fe from the simulations were ca. 1 – 3 orders of magnitude higher than the experimental results, which could be attributed to the different chemical forms/species of Fe critical to its precipitation that were not included in the model.

For the altered volcanic rock, measured Ca^{2+} , Fe and SO_4^{2-} concentrations were successfully simulated using a pyrite-gypsum-melanterite system (Table 15). However, the simulated pH was ca. 1.2 pH units higher than the measured values, which is probably due to sulfuric acid produced by the atmospheric/partial oxidation of pyrite in the rock that is neglected in the simulation. The contributions of pyrite oxidation and dissolution of soluble Ca and Fe-sulfates to the total SO_4^{2-} concentration of the leachate were estimated at around 2 and 98%, respectively. Concentrations of As and Pb are not present in the simulated results because these elements are present in the rock mostly as impurities of pyrite rather than as distinct mineral phases. However, the individual contributions of pyrite, calcite and soluble phases to the leaching of As and Pb could be estimated from the simulated results of Ca^{2+} and SO_4^{2-} because the release of As and Pb from the rock is associated with minerals like pyrite

Both O_2 and CO_2 have strong effects on an ideal calcite-pyrite-gypsum-melanterite system as shown in Table 16. The pH of the leachate tends to decrease as the simulated system approaches oxic conditions, which is similar to the trends of the measured results (sample SR1). Similar trends were also observed in the Ca^{2+} and SO_4^{2-} concentration results of both simulation and experiment. These gases also strongly influenced the pH of an ideal pyrite-gypsum-melanterite system as shown in Table 17. The pH dramatically decreased from 6 under anoxic condition to 4 at oxic conditions.

This trend was also observed in the measured results (sample VR4), but the magnitude of the decrease was substantially lower. In contrast, concentrations of Ca^{2+} and SO_4^{2-} were not affected by O_2 and CO_2 in both the simulated and measured results.

4. Discussion

4.1 Arsenic and lead in hydrothermally altered rocks

The partly altered sedimentary rock sample in this study had both altered and unaltered zones with each zone containing pyrite grains. In the unaltered zone, the shell bounding the pyrite crystals was mostly composed of calcite and was located away from the fractured zone (Figures 1 and 2). Because this shell was still intact and the minerals surrounding it were markedly different from those in the altered zone, pyrite grains inside the shell were most likely formed during the sedimentation process with pyrite probably replacing the foraminifer and not due to hydrothermal alteration. In contrast, alteration probably occurred around the fractured area resulting in the formation of pyrite, illite, and smectite. Sedimentary rocks usually have layered structures that will cause variations of their mineral composition. However, this layered structure was not apparent in Figure 1 as evidenced by the identical mineral grain sizes and composition above and below the fractured area. Furthermore, the presence of illite and smectite only around the fractured zone supports our deduction that the rock underwent hydrothermal alteration. Illite and smectite have been used by several authors as indicators of hydrothermal alteration. De Carlo *et al.* (1983) and Dekov *et al.* (2008) pointed out that smectite can directly precipitate from hydrothermal fluids especially under alkaline conditions with low Si activity and Mg/Al ratio. On the other hand, illite can be an alteration product of smectite during hydrothermal annealing according to this sequence (Alt and Jiang, 1991; Kisch, 1983; Srodon and Eberl, 1984):

Smectite → random Illite/Smectite → ordered Illite/Smectite → Illite

This type of alteration where clay minerals (*e.g.*, illite and smectite) replace plagioclase is termed argillic alteration. Argillic alteration is characterized by the formation of clay minerals due to intense H^+ metasomatism and acid or base leaching at temperatures between 100 and 300°C (Pirajno, 2009).

On the other hand, the volcanic rock samples most likely underwent phyllic/sericitic alteration that resulted in the formation of illite and pyrite in partly altered grains of plagioclase (Figure 4 (a)). Phyllic/sericitic alteration is one of the most common types of hydrothermal alteration and is essentially due to the destabilization of feldspars in the presence of H^+ , OH^- , K, and S to form quartz, illite and pyrite, and simultaneously leaching out Na, Mg, and Fe in the process (Pirajno, 2009). The original rock (*i.e.*, the precursor rock) in Figure 4(a) is probably porphyritic based on the large plagioclase grains within the rock matrix.

The As content of pyrite formed under hydrothermal conditions was in the wt% values while those formed during sedimentation were surprisingly below the detection limit of the EPMA (Table 4). These results suggest that As-rich solutions like hydrothermal fluids when present during pyrite formation would facilitate As incorporation into pyrite. This interpretation is consistent with the reports of other authors who also emphasized the importance of an As-rich solution in the formation of arsenian pyrite in both hydrothermal systems and sedimentary deposits (Bostick and Fendorf, 2003; Kirk *et al.*, 2004; Lowers *et al.*, 2007; O'Day *et al.*, 2004; Savage *et al.*, 2000). Also, the relative abundance and variability of As incorporation into pyrite was higher than Pb. Arsenic was measured in almost all of the pyrite grains regardless of rock type while Pb was only measured in pyrite crystals found in altered volcanic rocks (Tables 4 and 5). This relatively higher abundance of

As in pyrite suggests that it is preferentially incorporated into pyrite in comparison to Pb.

According to Huston *et al.* (1995), the incorporation of Pb into pyrite, like Cu and Zn, is most likely through the inclusion of other sulfide minerals containing these elements (*e.g.*, galena, chalcopyrite and sphalerite) and not directly into the crystal lattice. In contrast, As, with its smaller atomic size and similar properties to S, could be incorporated directly into the pyrite lattice most probably as nonstoichiometric substitutions or as a metastable $\text{Fe}(\text{As,S})_2$ solid solution (Fleet *et al.*, 1989; Huston *et al.*, 1995).

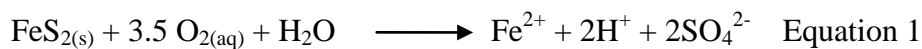
The altered volcanic and sedimentary rock samples used in this study underwent different types of alteration as discussed above, but the enrichment of the rock with As and Pb preferentially occurred in and around precipitated pyrite grains and not on the silicate and/or aluminate mineral matrix of the rock (Figure 5; Tables 4 and 5). These results suggest that pyrite is one of the primary sources of As and Pb in hydrothermally altered rocks regardless of the type of alteration, and this mineral would therefore strongly influence the mobilization of both elements when these rocks are exposed to the environment.

4.2 Soluble arsenic and lead-bearing phases, calcite dissolution and pyrite oxidation

Pyrite is one of the primary sources of As and Pb in hydrothermally altered rocks. However, altered rocks have complex chemical and mineralogical compositions, which mean that other constituent minerals could have significant effects on the mobilities of the toxic elements in these systems.

Minerals of particular importance include those that have strong acid/base buffering capacities like calcite and pyrite because of the highly pH-dependent leaching behavior of As and Pb from these rocks (Tabelin and Igarashi, 2009; Tabelin *et al.*, 2010, Tabelin *et al.*, 2012a). The altered volcanic and sedimentary rocks used in this study had acidic and alkaline pH values when in contact with

water, respectively, which could be attributed to the relative abundance of pyrite and calcite in the rock samples (Tables 3, 6, 14 and 15). The acidic pH of leachates from the altered volcanic rock samples indicates that pyrite oxidation is dominant and these rocks have very low acid buffering capacities (Tables 6 and 15). Pyrite oxidation generates acidity according to the following reaction:



On the other hand, the alkaline pH characteristic of the altered sedimentary rock samples was mainly due to calcite. The pH of a system controlled by carbonate minerals, such as calcite, will depend on whether the rock-water system is “open” or “closed” with respect to atmospheric CO₂. A system open to CO₂ generally has a lower pH between 7.0 and 8.3 while a closed system has a higher pH usually in the range of 7.6 – 10.1 (Appelo and Postma, 2005). The alkaline leachate pH of the altered sedimentary rocks also indicates that the strong buffering effect of calcite was more than sufficient in neutralizing the acidity produced by the oxidation of pyrite (Tables 8 and 14).

It was observed that leachate of samples taken from the bulk excavated rock had higher concentrations of Fe and SO₄²⁻ compared to those of the borehole cores (Table 6). The increase is particularly evident in the leachates of altered volcanic rocks where the concentrations of Fe and SO₄²⁻ from the bulk excavated rock were more than ten times higher than those in the borehole cores. These could be attributed to the atmospheric/partial oxidation of the bulk rock during excavation, transport and storage prior to sampling, which induced the formation of more soluble secondary phases especially on the surfaces of exposed pyrite grains. Pyrite oxidation starts rapidly upon exposure to the atmosphere, commencing with the oxidation of S²⁻ species (Schaufuss *et al.*, 1998). Oxidation is further enhanced in the presence of water or moisture that could strip reaction products on the pyrite surface exposing “fresh” sites for further oxidation. Atmospheric pyrite

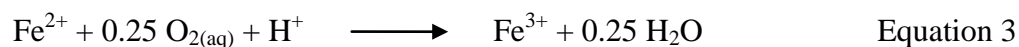
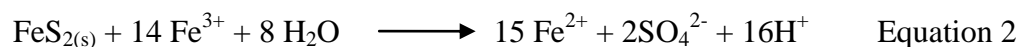
oxidation leads to the formation of Fe-sulfates, Fe oxyhydroxides/oxides, which are soluble especially under acidic conditions (De Donato *et al.*, 1993; Schaufuss *et al.*, 1998; Todd *et al.*, 2003). Dissolution of these soluble products of atmospheric/partial oxidation of pyrite enhanced the mobility of Pb but not As (Table 8). These differences could be attributed to the strong adsorption of As onto Fe-oxyhydroxides/oxides, which were also formed as a consequence of pyrite oxidation (Cornelis *et al.*, 2008; Dousova *et al.*, 2003; Dzombak and Morel, 1990; Gosh and Teoh, 1985; Lin and Puls, 2000; Wang and Mulligan, 2006). In contrast, Pb precipitation or adsorption onto Fe-oxyhydroxides/oxides was minimal especially in the acidic region, which could explain its substantial increase in the leachate of the bulk rock (Tables 8 and 11). Regardless of these seemingly opposing trends, majority of the As and Pb leached out from both volcanic and sedimentary altered rocks were supplied by the dissolution of these soluble Ca and Fe-sulfates under conditions close to equilibrium and without pH adjustment (Tables 6, 14 and 15). However, it should be noted that the contribution of these soluble phases to the As and Pb load of the leachate would precipitously decrease with time especially when the leachant (*i.e.*, water) is constantly renewed/transported (Tabelin *et al.*, 2012a, b).

4.3 Mechanisms controlling the mobilities of arsenic and lead under anoxic conditions

Gaseous O₂ and CO₂ strongly influenced the pH of the altered rock-water system and can therefore affect the mobilities of As and Pb. Increasing the CO₂ decreased the leachate pH and increased the dissolution of calcite in altered sedimentary rocks, but did not have any significant effect on the concentration of As in the leachate (Tables 10, 16 and 17). These results in conjunction with the low As content of calcite (Table 3) indicate that calcite found in the altered sedimentary rocks used in this study is not one of the primary sources of As. However, this might not be true for other hydrothermally altered rocks as noted by Horton *et al.* (2001). These authors reported that although

most of the calcite samples that they analyzed had low As contents, those located in fault zones in direct contact with the hydrothermal fluid were rich in As (mean As content = 81.2 ppm). The absence of O₂ significantly decreased the concentration of As in the leachate of the altered sedimentary rock even if the pH was higher. Similarly, the absence of O₂ in the altered volcanic rock sample dramatically decreased the leaching of As without any dramatic changes on the pH. These results were consistent with the SO₄²⁻ concentration decrease without O₂, which indicates that pyrite oxidation decreased in both types of rocks under anoxic conditions resulting in lower As concentrations in the leachate. These identical results observed in the two types of altered rocks suggest that pyrite oxidation and related oxidants (*e.g.*, O₂) are important in the mobilization of As from these sources (Tables 9 and 10).

Although O₂ is an essential oxidant of pyrite as explained by Equation 1, the extent of its effect on pyrite oxidation was pH dependent (Figures 9 and 10). In the altered volcanic rock samples, removal of O₂ had a significant effect on the amount of SO₄²⁻ released especially in the alkaline region (Figure 10). This decrease in SO₄²⁻ concentration under anoxic-alkaline conditions occurred together with that of As, which implies that the mobilization of As decreases under anoxic conditions because pyrite oxidation is minimized (Figure 10). In contrast, the effect of O₂ on As and SO₄²⁻ concentrations were minimal in the acidic and circumneutral pH. These results have two important implications. First, the effects of O₂ on pyrite oxidation under acidic and circumneutral pH is minimal. Second, a more powerful oxidant, probably Fe³⁺ ions, is available in these pH regions that overshadowed the effects of O₂ on pyrite oxidation. Ferric ion (Fe³⁺) and O₂ are two of the most important oxidants of pyrite (Holmes and Crundwell, 2000; McKibben and Barnes, 1986; Moses *et al.*, 1987; Moses and Herman, 1991; Nicholson *et al.*, 1988). Furthermore, O₂ is important in the cycling of Fe³⁺ ions as explained by the following chemical reactions:



When these two oxidants are present in a system, Fe^{3+} is more powerful than O_2 (Moses *et al.*, 1987). Aside from functioning as a direct oxidant of pyrite, O_2 has the additional role of oxidizing Fe^{2+} ions to Fe^{3+} ions as shown in Equation 3. However, Equation 3, which is known as the rate determining step of pyrite oxidation, is relatively slow in comparison to Equations 1 and 2 (Singer and Stumm, 1970). This means that both the direct (as oxidant of pyrite) and indirect (as oxidant of Fe^{2+}) effects of O_2 on pyrite oxidation are minimal in the acidic and circumneutral pH regions consistent with our results. Towards the alkaline region, the rate of Equation 3 increases dramatically as the pH increases, but under these conditions Fe^{3+} is also easily precipitated (Evangelou, 1995). Regardless of these very low concentrations of Fe^{3+} in the alkaline region, it is still an important oxidant of pyrite as reported by Moses *et al.* (1987). In addition, O_2 becomes more important to replenish the diminishing Fe^{3+} through $\text{Fe}^{3+}/\text{Fe}^{2+}$ recycling under alkaline conditions because of the very low concentrations of Fe^{3+} (Moses *et al.*, 1987). Therefore, our results illustrating the importance of O_2 in pyrite oxidation especially under alkaline conditions is consistent with these previously published works.

Similarly, the decrease in pyrite oxidation minimized Pb release from the altered volcanic rock under alkaline-anoxic conditions. In addition, the absence of O_2 increased the concentration of Pb in the leachate between pH 4 – 6 (Figure 9(b)). This Pb concentration increase coincided with the increase of Fe concentration in the leachate in the absence of O_2 (Figure 10(a)). Ferric hydroxide starts to precipitate at pH 3.5 while Fe^{2+} ions start to form oxyhydroxides/oxides at pH 4 depending on the Eh of the system. The results of PHREEQC also showed that Pb is present between pH 4 – 6

as Pb^{2+} ion and not as sulfate, oxide or hydroxide under both oxic and anoxic conditions. Thus, the observed increase in Pb concentration in the absence of O_2 in conjunction with Fe is probably due to the decrease in co-precipitation and/or adsorption of Pb onto Fe oxyhydroxides/oxides precipitates (Table 13).

In the altered sedimentary rock samples, the leaching behavior of As as a function of pH changed in three distinct regions under anoxic conditions (Figure 11). In the first region ($\text{pH} < 4$), the concentration of As in the leachate decreased in conjunction with the decrease of Fe and Ca^{2+} concentrations. These results indicate that the mobilization of As at acidic pH (< 4) from altered sedimentary rocks is controlled by the dissolution of Fe and Ca-bearing minerals containing As under both oxic and anoxic conditions. The next region ($\text{pH} 4 - 8$) illustrates an increase in the concentration of As in the absence of O_2 , which could be attributed to the reduced precipitation of Fe-oxyhydroxides/oxides that are important scavengers of As especially in these pH range (Tables 12 and 13). In the third region ($\text{pH} > 8$), the concentration of As in the leachate decreased because of the retardation of pyrite oxidation similar to that observed in the altered volcanic rocks rich in pyrite. Although the altered sedimentary rock samples contain only trace amounts of pyrite, the decrease of SO_4^{2-} concentration at pH greater than 8 in the absence of O_2 was substantial, which supports our interpretation of pyrite oxidation reduction. These results suggest that pyrite oxidation is an important As release mechanism in hydrothermally altered rocks even if they only contain trace amounts of pyrite.

Conceptualized models of the release mechanisms of As and Pb from altered rocks are illustrated in Figures 13 and 14, respectively. The mechanisms controlling the mobilization of As and Pb from altered rocks with pH under oxic and anoxic conditions include: dissolution, pyrite oxidation, adsorption/desorption and precipitation and are similar to those proposed under oxic

conditions. Atmospheric/partial oxidation of pyrite and the subsequent dissolution of soluble As and Pb-bearing phases like Ca and Fe-sulfates are the dominant leaching mechanisms of As and Pb from altered rocks. However, both As and Pb are “reactive” solutes whose mobilities are easily affected by changes in the pH of the altered rock-water system. In the acidic region, acid dissolutions of Ca and Fe-bearing minerals further contribute to the concentrations of As and Pb in the leachate. At the circumneutral pH, As and Pb mobilities dramatically decrease because of precipitation and adsorption reactions. Finally, As and Pb mobilities again increase in the alkaline region because of desorption/enhanced oxidation of pyrite. Pyrite oxidation increase in the alkaline region under abiotic conditions could be related to the availability of surface OH⁻ species that are involved in more rapid electron transfer at the pyrite surface (Evangelou, 1995).

4.4 Engineering implications

This study identified the roles of calcite and pyrite in the enrichment and leaching behaviors of As and Pb in excavated altered rocks. Pyrite is an important source of As and Pb in altered rocks and could directly influence the pH of the rock when in contact with water. On the other hand, the role of calcite in pH control was more important than as a source of As and Pb in our samples, but this may not apply to other hydrothermally altered rock deposits containing calcite. None the less, pH of the altered rock when in contact with water is dependent on the abundance of these two minerals relative to each other. When altered rocks contain significant amounts of pyrite but the calcite content is negligible, the pH is acidic. On the other hand, the pH would be alkaline when calcite is substantial in the rock even with trace amounts of pyrite. Thus, it is possible to categorize excavated hydrothermally altered rocks into pyritic/acidic or calcareous/alkaline rocks, depending on whether they contain significant sulfide minerals (*e.g.*, pyrite) or carbonate minerals (*e.g.*, calcite) and their pH in water. With this simple classification based on the mineralogy and pH of the excavated

altered rocks, countermeasures for their disposal can be created separately through the enhancement or retardation of the main mechanism/s that control the mobilities of As and Pb. Moreover, pyritic rocks should be disposed of immediately after excavation to prevent the partial oxidation of pyrite and the formation of soluble minerals that could lower the pH and enhance the mobilization of heavy metals like Pb. If this is not possible, the interim storage sites of these excavated rocks must be covered to minimize their exposure to moisture and rain. Our results also indicate that the best way to immobilize As and Pb is through the enhancement of adsorption and precipitation processes, respectively and these processes must be incorporated into the design of the alternative disposal method. Finally, because the mobilization mechanisms of As and Pb are strongly pH-dependent but do not dramatically change under oxic and anoxic conditions, pH of the altered rock-water system must be closely controlled during the actual disposal of the rock.

5. Conclusions

Hydrothermally altered rocks contain elevated amounts of As and Pb that are easily mobilized under various conditions encountered in the environment. Enrichment of these rocks with As and Pb predominantly occurred within and around precipitated pyrite crystals and not on the surrounding alumino-silicate minerals. However, the As and Pb contents of pyrite formed in these altered rocks varied depending on the inherent physical and chemical properties of the hydrothermal solution that caused the alteration. Dissolution of soluble Ca and Fe-phases, which are mostly products of the atmospheric/partial oxidation of pyrite, is the dominant leaching mechanism of As and Pb. However, changes in the pH of the rock-water system could increase the contributions of the other mechanisms like pyrite oxidation, adsorption/desorption and precipitation/co-precipitation reactions to the equilibrium concentration of As and Pb in the leachate. The leaching behaviors of As and Pb were both pH dependent with or without O₂ and CO₂. The absence of O₂ also minimized

the mobilization of As and Pb especially in the alkaline region because pyrite oxidation was retarded but not in the acidic and circumneutral pH. The pH and the leaching behaviors of As and Pb also depended on the relative abundance of pyrite and calcite in the rock. Thus, the classification scheme outlined in this paper using pyrite and calcite as “indicators” could be used to group together excavated rocks with complex mineral properties and from different locations. Moreover, the mechanisms of As and Pb release from altered rocks could be predicted based on the relative abundance of pyrite and calcite as well as the pH of the altered rock when in contact with water.

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References

- Acharyya, S.K., Lahiri, S., Raymahashay, B.C., Bhowmik, A., 2000. Arsenic toxicity of groundwater in parts of the Bengal basin in India and Bangladesh: the role of Quarternary stratigraphy and Holocene sea level-fluctuation. *Environmental Geology* 39(10), 1127-1137.
- Aiuppa, A., Avino, R., Brusca, L., Caliro, S., Chiodini, G., D’Alessandro, W., Favara, R., Federico, C., Ginevra, W., Inguaggiato, S., Longo, M., Pecoraino, G., Valenza, M., 2006. Mineral control of arsenic content in thermal waters from volcano-hosted hydrothermal systems: Insights from the island of Ischia and Phlegrean Fields. *Chemical Geology* 229, 313 – 330.
- Akai, J., Izumi, K., Fukuhara, H., Masuda, H., Nakano, S., Yoshimura, T., Ohfuji, H., Anwar, H.M., Akai, K., 2004. Mineralogical and geomicrobial investigations on groundwater arsenic enrichment in Bangladesh. *Applied Geochemistry* 19, 215-230.
- Allen, K.D., Hahn, G.A., 1994. Geology of Sunbeam and Grouse Creek gold-silver deposits, Yankee Fork mining district, Eocene Challis volcanic field, Idaho: a volcanic dome and volcanoclastic-hosted epithermal system. *Econ. Geol.* 89, 1964-1982.
- Alt, J.C., Jiang, W.T., 1991. Hydrothermally precipitated mixed-layer illite-smectite in recent massive deposits from the sea floor. *Geology* 19, 570-573.

Apello, C.A.J., Postma, D., 2005. *Geochemistry, Groundwater and Pollution*, 2nd ed., A.A. Balkema Publishers, London.

Bethke, C.M., 1992. *The Geochemist's Workbench – A User's guide to Rxn, Act2, Tact, React and Gtplot*. University of Illinois, Urbana, Illinois.

Bostick, B.C., Fendorf, S., 2003. Arsenite sorption on troilite (FeS) and pyrite (FeS₂). *Geochim. Cosmochim. Acta* 67, 909–921.

Bowell, R.J., 1993. Mineralogy and geochemistry of tropical rain forest soils: Ashanti, Ghana. *Chemical Geology* 106, 345-348.

Brannon, J.M., Patrick, W.H., 1987. Fixation, transformation and mobilization of arsenic in sediments. *Environ. Sci. Technol.* 21, 450-459.

Burgess, W.G., Hoque, M.A., Michael, H.A., Voss, C.I., Breit, G.N., Ahmed, K.M., 2010. Vulnerability of deep groundwater in the Bengal Aquifer System to contamination by arsenic. *Nature Geoscience* 3, 83-87.

Choprapawon, C., Rodcline, A., 1997. Chronic arsenic poisoning in Ronpibool Nakhon Sri Thammarat, the southern province of Thailand, in: Abernathy, C.O., Calderon, R.L., Chappell, W.R. (Eds.), *Arsenic Exposure and Health Effects*, Chapman and Hall, London, pp. 69-77.

Cornelis, G., Anette-Johnson, C., Van Gerven, T., Vandecasteele, C., 2008. Leaching mechanisms of oxyanionic metalloid and metal species in alkaline solid wastes: A review. *Applied Geochemistry* 23, 955-976.

Das, D., Samanta, S., Mandal, B.K., Chowdhury, T.R., Chanda, C.R., Chowdry, P.P., Basu, G.K., Chakraborti, D., 1996. Arsenic in groundwater in six districts of West Bengal, India. *Environmental Geochemistry and Health* 8, 5-15.

De Carlo, E.H., McMurtry, G. M., Yeh, H.W., 1983. Geochemistry of hydrothermal deposits from Loihi submarine volcano, Hawaii. *Earth and the Planetary Science Letters* 66, 438-449.

De Donato, P., Mustin, C., Benoit, R., Erre, R., 1993. Spatial distribution of iron and sulphur species on the surface of pyrite. *Applied Surface Science* 68, 81 – 93.

Dekov, V.M., Cuadros, J., Shanks, W.C., Koski, R.A., 2008. Deposition of talc – kerolite – smectite – smectite at seafloor hydrothermal vent fields: Evidence from mineralogical, geochemical and oxygen isotope studies. *Chemical Geology* 247, 171-194.

Dermatas, D., Menounou, N., Dadachov, D.O., Shen, M. G., Xu, X., Tsaneva, V., 2006. Lead leachability in firing range soils. *Environmental Engineering Science* 23(1), 88-101.

Dousova, B., Machovic, V., Kolousek, D., Kovanda, F., Dornicak, V., 2003. Sorption of As (V) species from aqueous solution. *Water Air and Soil Pollution* 149, 251-267.

- Dowling, C.B., Poreda, R.J., Basu, A.R., Peters, S.L., Aggarwal, P.K., 2002. Geochemical study of arsenic release mechanisms in the Bengal Basin groundwater. *Water Resources Research* 38(9), 1173.
- Evangelou, V.P., 1995. *Pyrite oxidation and its control*, CRC Press, New York.
- Fleet, M.E., MacLean, P.J., Barbier, J., 1989. Oscillatory-zoned As-bearing pyrite from strata-bound and stratiform gold deposits: An indicator of ore fluid evolution. *Economic Geology Monograph* 6, 356 – 362.
- Franzson, H., Zierenberg, R., Schiffman, P., 2008. Chemical transport in geothermal systems in Iceland: Evidence from hydrothermal alteration. *Journal of Volcanology and Geothermal Research* 173, 217-229.
- Ghosh, M.M., Teoh, R.S., 1985. Adsorption of arsenic on hydrous aluminum oxide. In Proc. 7th Mid-Atlantic Industrial Waste Conference, Lancaster, PA, 139-155.
- Halbach, P., Pracejus, B., Marten, A., 1993. Geology and mineralogy of massive sulfide ores from Okinawa trough, Japan. *Econ. Geol.* 88, 2210-2225.
- Hammond, P.B., 1977. Exposures of humans to lead. *Annu. Rev. Pharmacol. Toxicol.* 17, 197-214.
- Holmes, P.R., Crundwell, F.K., 2000. The kinetics of the oxidation of pyrite by ferric ions and dissolved oxygen: An electrochemical study. *Geochim. Cosmochim. Acta* 64, 263-274.
- Horton, T.W., Becker, J.A., Craw, D., Koons, P.O., Page Chamberlain, C., 2001. Hydrothermal arsenic enrichment in an active mountain belt: Southern Alps, New Zealand. *Chemical Geology* 177, 323 – 339.
- Huston, D.L., Sie, S.H., Sauter, G.F., Cook, D.R., Both, R.A., 1995. Trace elements in sulfide minerals from eastern Australian volcanic-hosted massive sulfide deposits: Part 1. Proton microprobe analyses of pyrite, chalcopyrite and sphalerite, and Part 2. Selenium levels in pyrite: Comparison with $\delta^{34}\text{S}$ values and implications for the source of sulfur in volcanogenic hydrothermal systems. *Econ. Geol.* 90, 1167-1196.
- Igarashi, T., Imagawa, H., Uchiyama, H., Asakura, K., 2008. Leaching behavior of arsenic from various rocks by controlling geochemical conditions. *Miner. Eng.* 21, 191-199.
- Imagawa, H., Igarashi, T., Uchiyama, H., Asakura, K., 2007. Leaching behavior of arsenic from altered igneous rocks by batch method. *Journal of the Mining and Materials Processing Institute of Japan* 123, 158-164. (Paper is written in Japanese with English abstract)
- Kirk, M.F., Holm, T.R., Park, J., Jin, Q., Sanford, R.A., Fouke, B.W., Bethke, C.M., 2004. Bacterial sulfate reduction limits natural arsenic contamination in groundwater. *Geology* 32, 953–956.
- Kisch, H.J., 1983. Mineralogy and petrology of burial diagenesis (burial metamorphism) and incipient metamorphism in clastic rocks, in: Larsen, G., Chilingar, G.V., (Eds.), *Diagenesis in sediments and sedimentary rocks 2*, Elsevier, Amsterdam, pp.289-493.

Lin, Z., Puls, R.W., 2000. Adsorption, desorption and oxidation of arsenic affected by clay minerals and aging process. *Environmental Geology* 39(7), 753-759.

Lowers, H.A., Breit, G.N., Foster, A.L., Whitney, J., Yount, J., Uddin, M.N., Muneem, A.A., 2007. Arsenic incorporation into authigenic pyrite, Bengal Basin sediment, Bangladesh. *Geochim. Cosmochim. Acta* 71, 2699 – 2717.

Marques, J.M., Matias, M.J., Basto, M.J., Carreira, P.M., Aires-Barros, L.A., Goff, F.E., 2010. Hydrothermal alteration of Hercynian granites, its significance to the evolution of geothermal systems in granitic rocks. *Geothermics* 39, 152-160.

Martin-Crespo, T., Vindel, E., Lopez-Garcia, J.A., Cardellach, E., 2004. As-(Ag) sulphide veins in the Spanish central system: Further evidence for a hydrothermal event of Permian age. *Ore Geology Reviews* 25, 199-219.

McKibben M.A., Barnes H.L., 1986. Oxidation of pyrite in low temperature acidic solutions: Rate laws and surface textures. *Geochim. Cosmochim. Acta* 50, 1509-1520.

Moore, J.N., Ficklin, W.H., Johns, C., 1988. Partitioning of arsenic and metals in reducing sediments. *Environ. Sci. Technol.* 22, 432-437.

Moore, M.R., Goldberg, A., 1985. Health implications of the hematopoietic effect of lead, in: Mahaffey, K.R. (Ed.), *Dietary and Environmental Lead: Human Health Effects*, Elsevier, Amsterdam, pp. 260-314.

Moses, C.O., Herman, J.S., 1991. Pyrite oxidation at circumneutral pH. *Geochim. Cosmochim. Acta* 55, 471-482.

Moses, C.O., Nordstrom, D.K., Herman, J.S., Mills, A., 1987. Aqueous pyrite oxidation by dissolved oxygen and by ferric iron. *Geochem. Cosmochim. Acta* 54, 1561-1572.

Nicholson, R.V., Gillham, R.W., and Reardon, E.J., 1988. Pyrite oxidation in carbonate-buffered solution: 1. Experimental kinetics. *Geochim. Cosmochim. Acta* 52, 1077-1085.

Nickson, R.T., McArthur, K.M., Ravenscroft, J.M., Burgess, W.G., Ahmed, K.M., 2000. Mechanism of arsenic release to groundwater, Bangladesh and West Bengal. *Applied Geochemistry* 15, 403-413.

O'Day, P.A., Vlassopoulos, D., Root, R., Rivera, N., 2004. The influence of sulfur and iron on dissolved arsenic concentrations in the shallow subsurface under changing redox conditions, *Proceedings of the National Academy of Sciences* 101 [38], 13703-13708.

Ostwald, J., England, B.M., 1977. Notes on framboidal pyrite from Allandale, New South Wales, Australia. *Mineral. Deposita* 12, 111-116.

Parkhurst, D.L., Appelo, C.A.J., 1999. User's Guide to PHREEQC (Version 2) – A computer program for speciation, batch-reactions, one-dimensional transport, and inverse geochemical calculations. U.S. Geological Survey, Denver, CO.

- Peters, S.C., Blum, J.D., 2003. The source and transport of arsenic in a bedrock aquifer, New Hampshire, USA. *Applied Geochemistry* 18, 1773-1787.
- Pirajno, F., 2009. *Hydrothermal Processes and Mineral Systems*. Springer Science, The Netherlands.
- Pokrovski, G.S., Borisova, A.Y., Harrichoury, J.C., 2007. The effect of sulfur on vapor-liquid fractionation of metals in hydrothermal systems. *Earth and Planetary Science Letters* 266, 345-362.
- Rabinowitz, B.M., 2005. Lead isotopes in soils near five historic American lead smelters and refineries. *Science and Total Environment* 346, 138-148.
- Savage, K.S., Tingle, T.N., O'Day, P.A., Waychunas, G.A., Bird, D.K., 2000. Arsenic speciation in pyrite and secondary weathering phases, Mother Lode Gold District, Tuolumne County, California. *Applied Geochemistry* 15, 1219-1244.
- Schaufuss, A.G., Nesbitt, H.W., Kartio, I., Laajalehto, K., Bancroft, G.M., Szargan, R., 1998. Reactivity of surface chemical states on fractured pyrite. *Surface Science* 411 (3), 321-328.
- Sengupta, A.K., 2002. *Environmental Separation of Heavy Metals – Engineering Processes*. Lewis Publishers, London.
- Singer, P.C., Stumm, W., 1970. Acidic mine drainage: The rate-determining step. *Science* 167, 1121-1123.
- Smedley, P.L., Kinniburgh, D.G., 2002. A review of the source, behavior and distribution of arsenic in natural waters. *Applied Geochemistry* 17, 517-568.
- Srodon, J., Eberl, D.D., 1984. Illite, in: Bailey, S.W., (Ed.), *Micas: Reviews in Mineralogy Vol. 13*, Mineralogical Society of America, Washington D.C., pp.495-544.
- Tabelin, C.B., Igarashi, T., 2009. Mechanisms of arsenic and lead release from hydrothermally altered rock. *Journal of Hazardous Materials* 169, 980-990.
- Tabelin, C.B., Igarashi, T., Tamoto, S., 2010. Factors affecting arsenic mobility from hydrothermally altered rock in impoundment-type in situ experiments. *Miner. Eng.* 23, 238-248.
- Tabelin, C.B., Igarashi, T., Takahashi, R., 2012a. Mobilization and speciation of arsenic from hydrothermally altered rock in laboratory column experiments under ambient conditions. *Applied Geochemistry* 27, 326-342.
- Tabelin, C.B., Basri, A.H.M., Igarashi, T., Yoneda, T., 2012b. Removal of arsenic, boron and selenium from excavated rocks by consecutive washing. *Water, Air, & Soil Pollution* 223, 4153-4167.
- Tatsuhara, T., Arima, T., Igarashi, T., Tabelin, C.B., 2012. Combined neutralization-adsorption system for the disposal of hydrothermally altered excavated rock producing acidic leachate with hazardous elements. *Engineering Geology* 139-140, 76-84.

- Todd, E.C., Sherman, D.M., Purton, J.A., 2003. Surface oxidation of pyrite under ambient atmospheric and aqueous (pH = 2 to 10) conditions: Electronic structure and mineralogy from X-ray absorption spectroscopy. *Geochim. Cosmochim. Acta* 67(5), 881-893.
- Ure, A., Berrow, M., 1982. Chapter 3: The elemental constituents of soils, in: Bowen, H.J.M., (Ed.), *Environmental Chemistry*, Royal Society of Chemistry, London, pp. 94-203.
- Wang, S., Mulligan, C., 2006. Natural attenuation processes for remediation of arsenic contaminated soils and groundwater. *Journal of Hazardous Materials* 138(3), 459-470.
- Weise, R.G., Fyfe, W.S., 1986. Occurrences of iron sulfides in Ohio coals. *Int. J. Coal Geol.* 6, 251-276.
- Welch, A.H., Helsel, D.R., Focazio, M.J., Watkins, S.A., 1999. Arsenic in ground water supplies of the United States, in: Chappell, W.R., Abernathy, C.O., Calderon, R.L. (Eds.), *Arsenic Exposure and Health Effects*, Elsevier, Amsterdam, pp. 9-17.
- Williams, M., 1997. Mining-related arsenic hazards: Thailand case study, Summary Report, Brit. Geol. Surv. Tech. Rep. WC/97/49.

FIGURE CAPTIONS

- FIGURE 1** Photomicrographs taken with open (a) and crossed (b) Nicols showing the minerals present around and away from the zone of alteration in sample SR3 (altered sedimentary rock).
- FIGURE 2** Photomicrographs taken with open Nicols (reflected light) showing the minerals present in the unaltered zone of sample SR3 (altered sedimentary rock).
- FIGURE 3** Photomicrographs taken with open Nicols (polarized light) (a) and SEM (b) showing the minerals present in the zone of alteration and the morphology of pyrite in sample SR3 (altered sedimentary rock).
- FIGURE 4** Photomicrographs taken with open Nicol (polarized light) (a) and SEM (b) showing the minerals present in the altered volcanic rock samples and the morphology of pyrite.
- FIGURE 5** The elemental maps of Fe (a), S (b), As (c), Si (d), and Al (e) from the altered volcanic rock at a magnification of 40 μm modified with permission from the Journal of the Mining and Materials Processing Institute of Japan (Imagawa et al., 2007).
- FIGURE 6** As concentrations in the leachate as a function of pH under oxic conditions: altered volcanic rocks (a) and altered sedimentary rocks (b).
- FIGURE 7** SO_4^{2-} concentrations in the leachate as a function of pH under oxic conditions: altered volcanic rocks (a) and altered sedimentary rocks (b).
- FIGURE 8** Pb concentrations in the leachate of the altered volcanic rock samples as a function of pH under oxic conditions.
- FIGURE 9** As and Pb concentrations in the leachate of altered volcanic rocks under anoxic conditions: As concentration as a function of pH (a) and Pb concentration as a function of pH (b).
- FIGURE 10** Fe and SO_4^{2-} concentrations in the leachate of altered volcanic rocks under anoxic conditions: Fe concentration as a function of pH (a) and SO_4^{2-} concentration as a function of pH (b).
- FIGURE 11** As concentration in the leachate of altered sedimentary rocks under anoxic conditions.
- FIGURE 12** Fe, Ca^{2+} and SO_4^{2-} concentrations in the leachate of altered sedimentary rocks under anoxic conditions: Fe concentration as a function of pH (a), Ca^{2+} concentration as a function of pH (b) and SO_4^{2-} concentration as a function of pH (c).
- FIGURE 13** A conceptual model of the mechanisms of As release from altered rocks with pH.
- FIGURE 14** A conceptual model of the mechanisms of Pb release from altered rocks with pH.

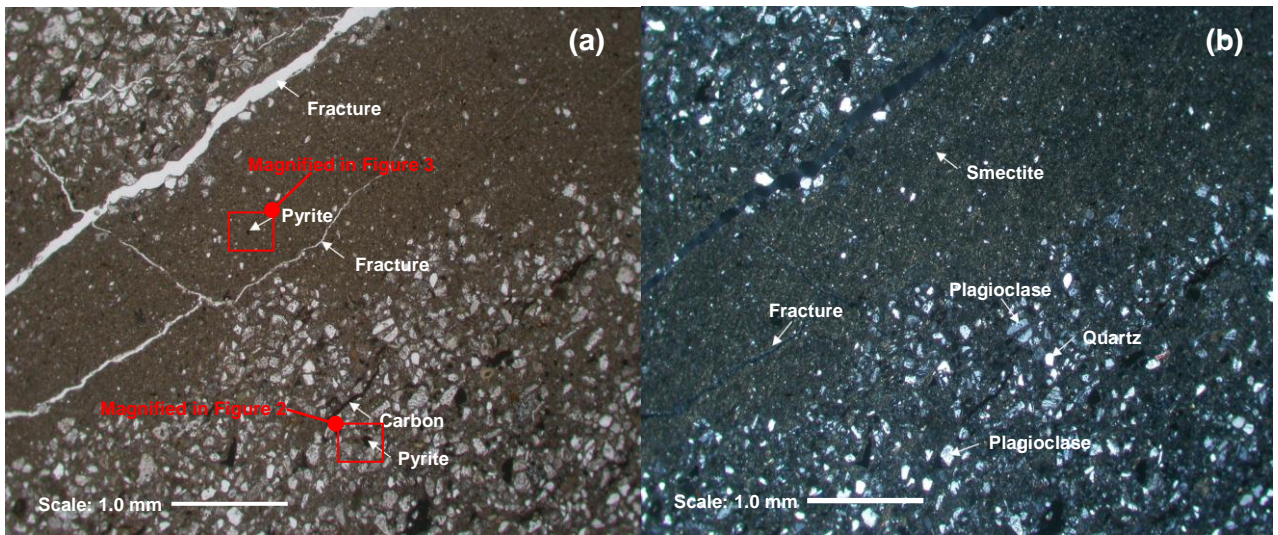


Figure 1

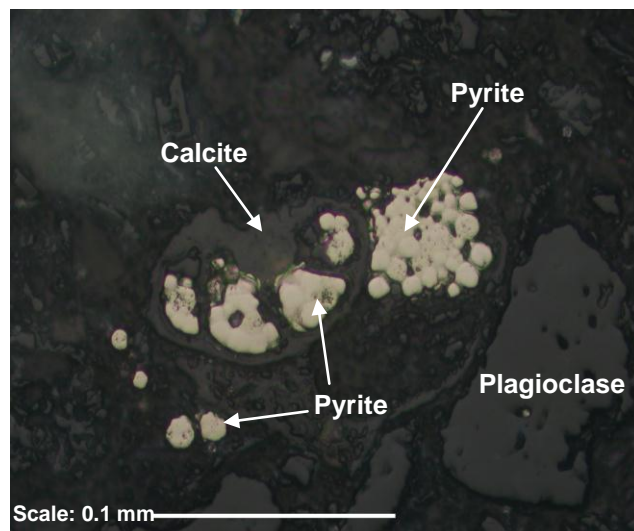


Figure 2

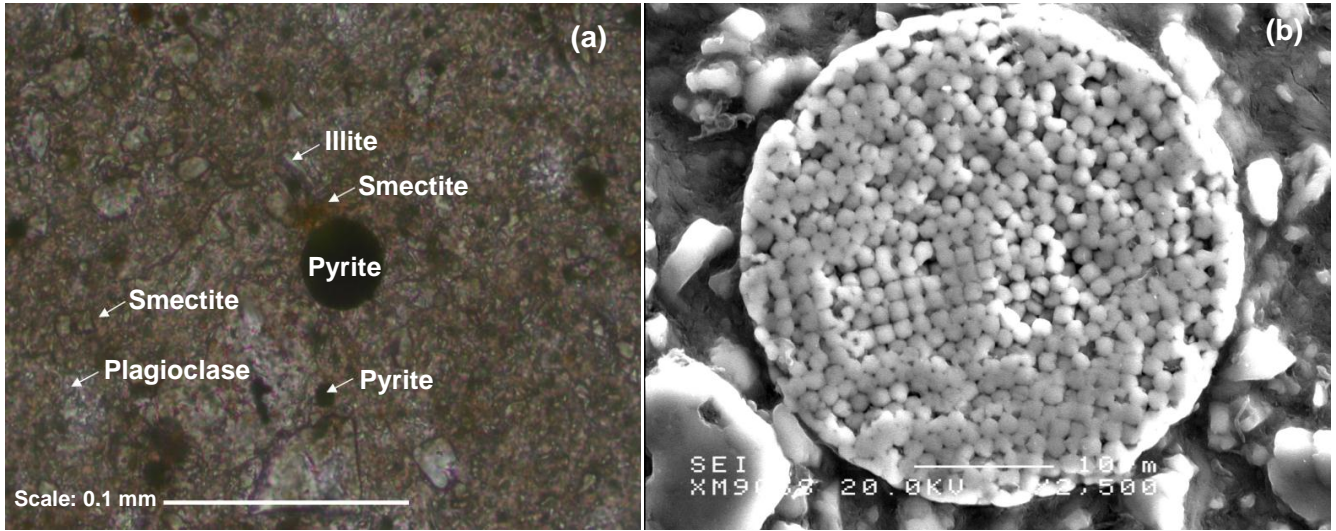


Figure 3

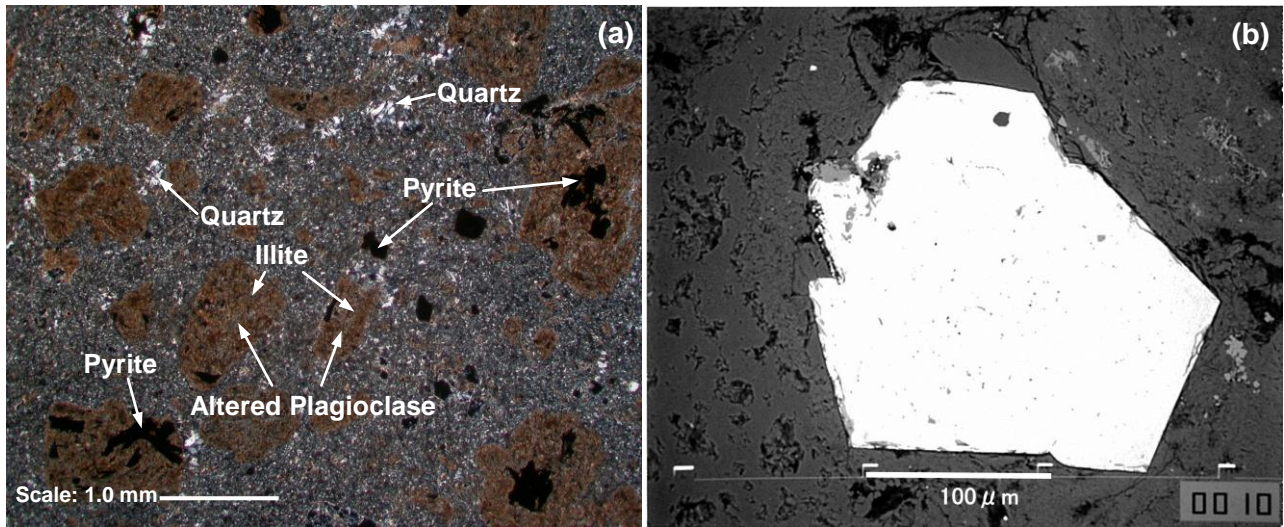


Figure 4

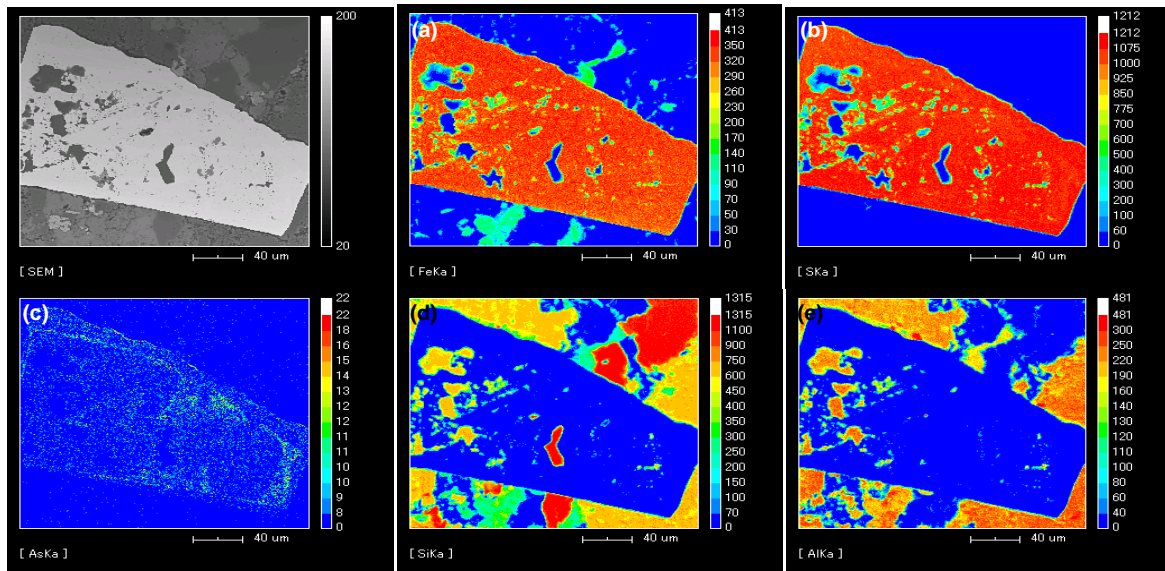


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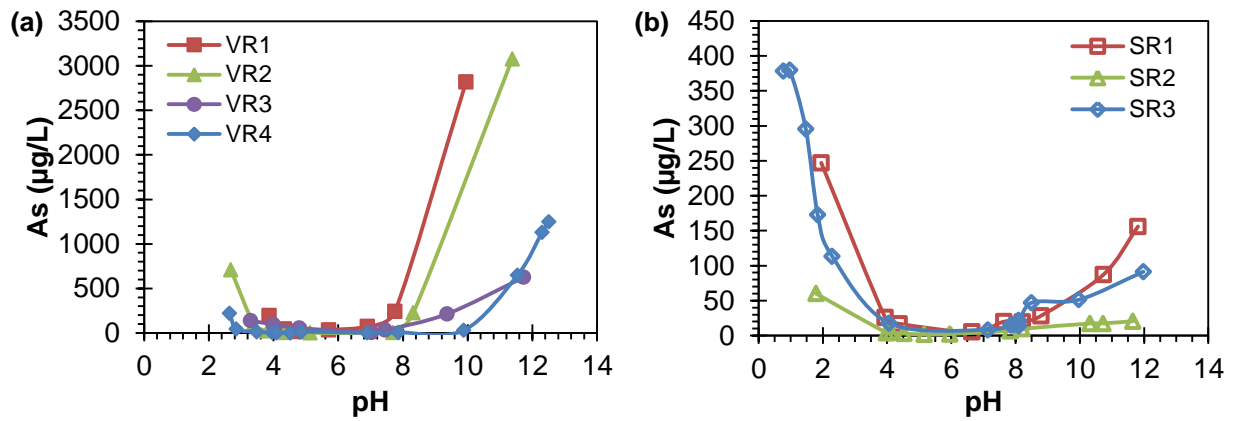


Figure 6

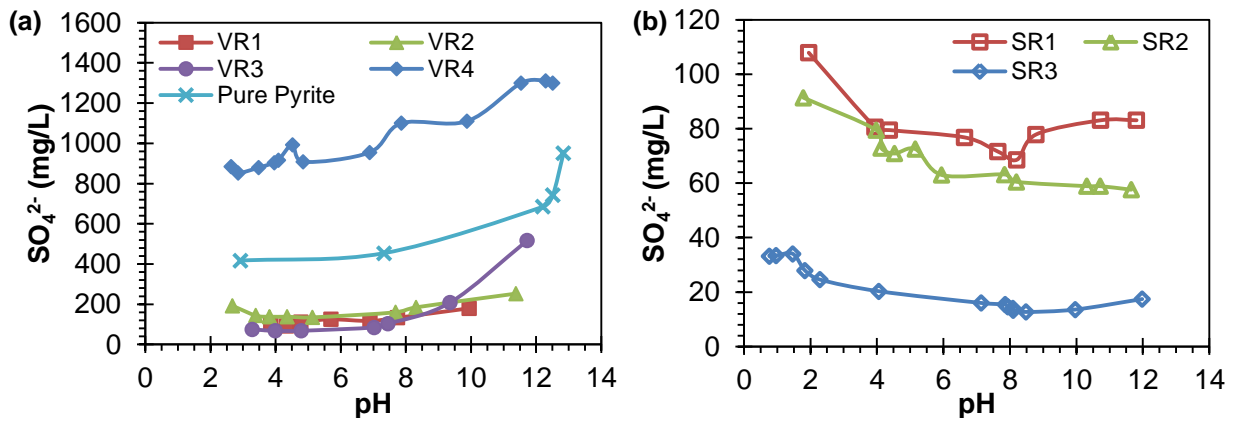


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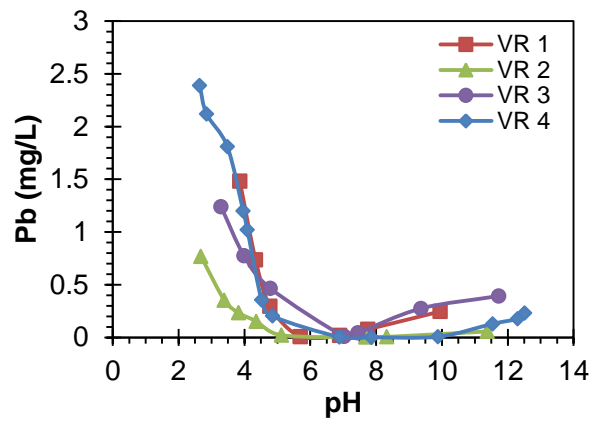


Figure 8

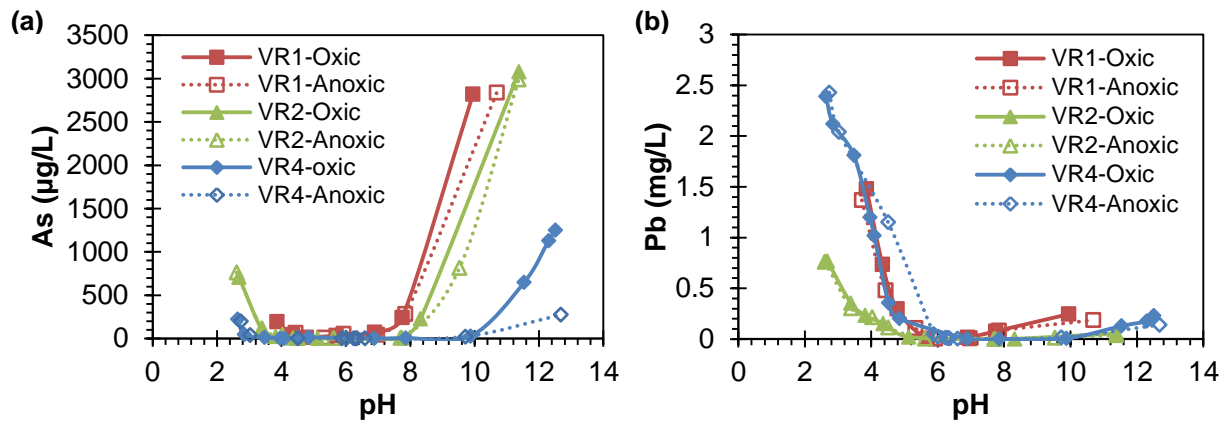


Figure 9

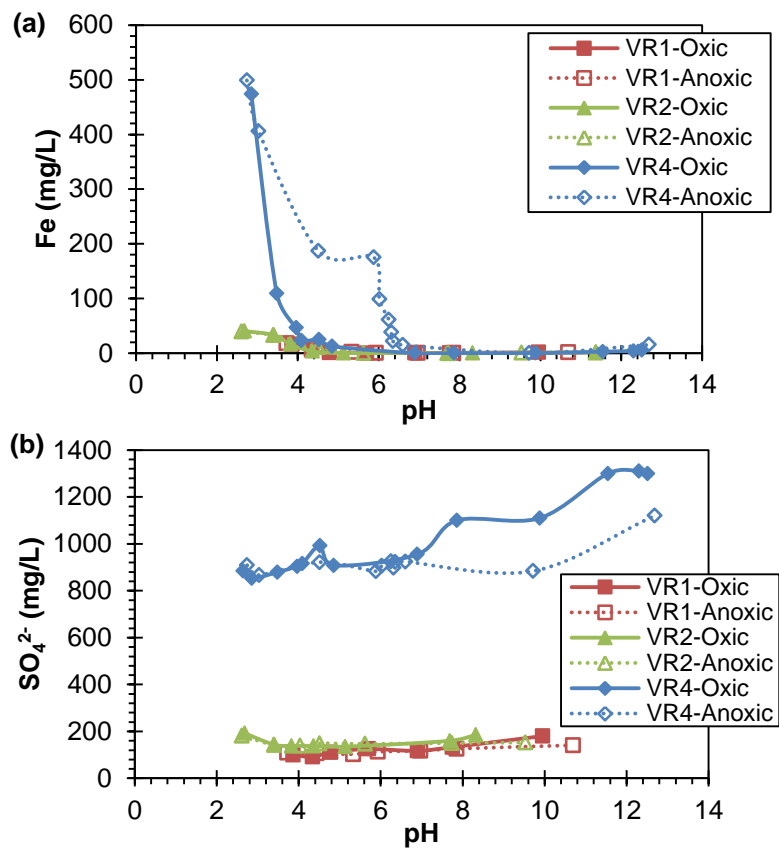


Figure 10

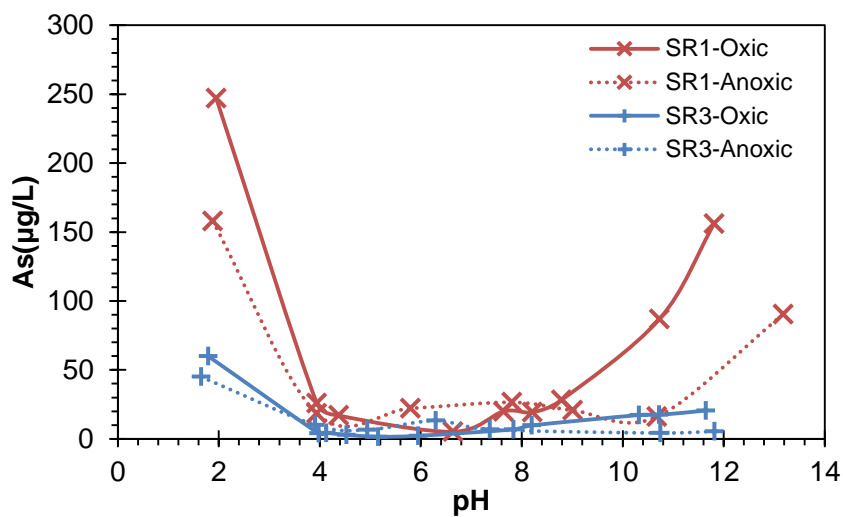


Figure 11

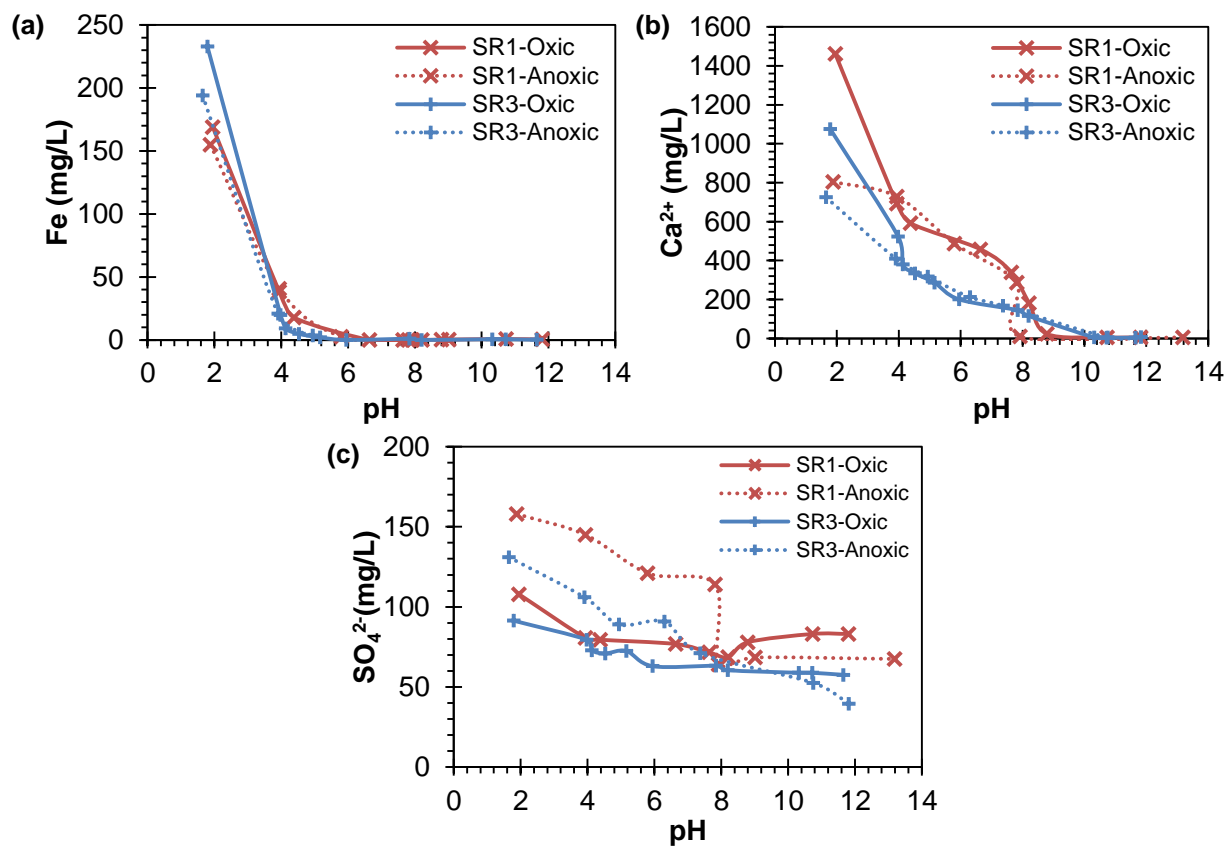


Figure 12

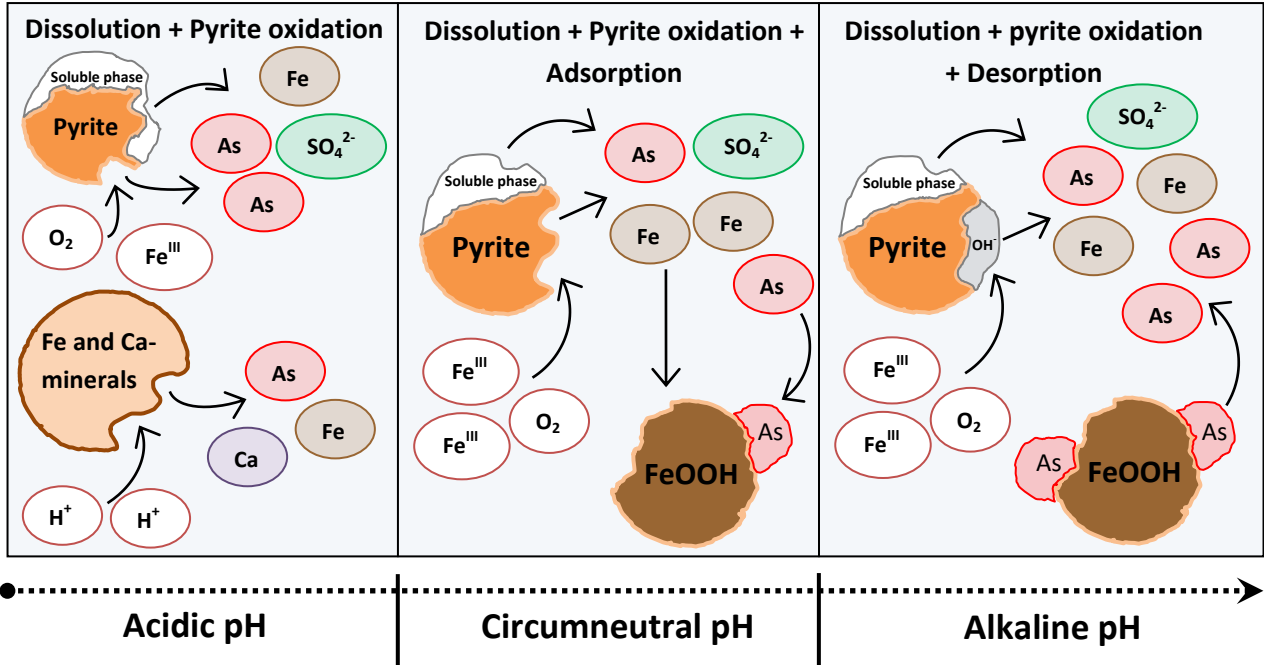


Figure 13

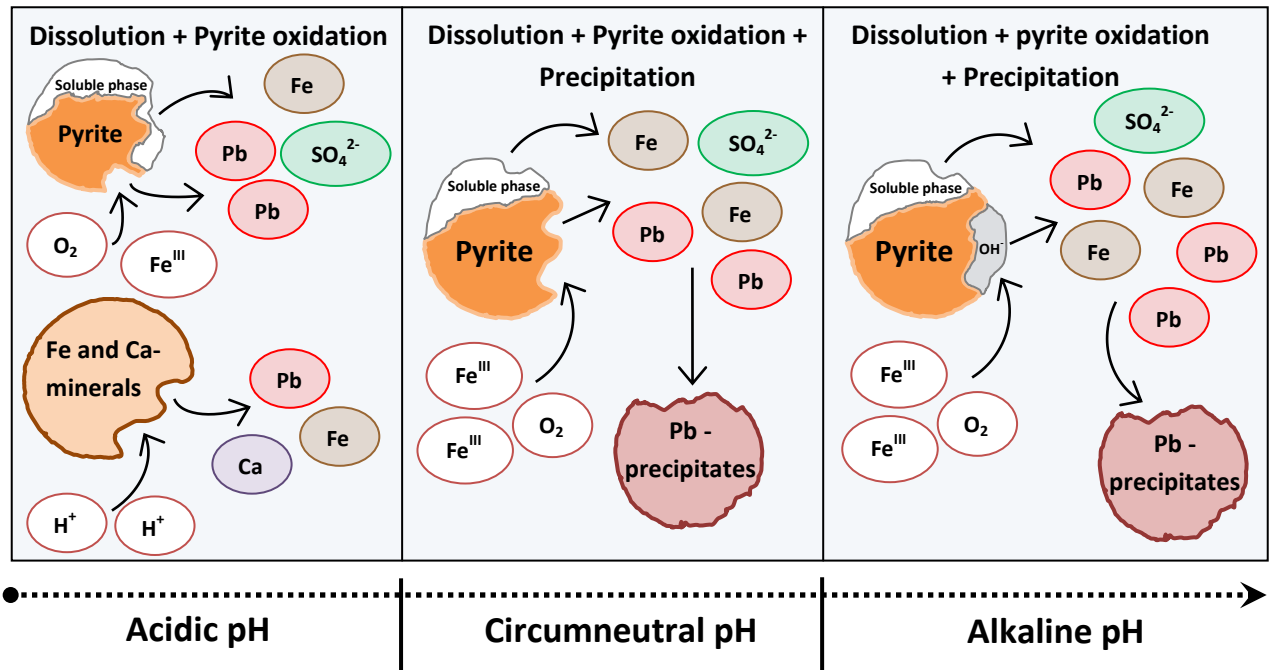


Figure 14

Table 1. Brief description of the tunnel projects where the altered rock samples were collected.

Sample	Type	Location	Purpose of excavation	Brief geological summary
VR 1	Borehole core	Teine, Hokkaido	Maintenance tunnel	Former mine with epithermal Au-Ag-Cu deposit
VR 2	Borehole core	Teine, Hokkaido	Maintenance tunnel	Former mine with epithermal Au-Ag-Cu deposit
VR 3	Borehole core	Teine, Hokkaido	Maintenance tunnel	Former mine with epithermal Au-Ag-Cu deposit
VR 4	Bulk excavated rock	Teine, Hokkaido	Maintenance tunnel	Former mine with epithermal Au-Ag-Cu deposit
SR 1	Bulk excavated rock	Kamikawa, Hokkaido	Road	Altered Mesozoic sedimentary rock composed mainly of slate from the Hidaka metamorphic belt
SR 2	Borehole core	Kamikawa, Hokkaido	Road	Altered Mesozoic sedimentary rock composed mainly of slate from the Hidaka metamorphic belt
SR 3	Bulk excavated rock	Obihiro, Hokkaido	Road	Partly altered mudstone/sandstone of marine origin from the Cretaceous period

Table 2. Chemical composition of rock samples used in the leaching experiments.

Sample	SiO ₂ (wt%)	Al ₂ O ₃ (wt%)	Fe ₂ O ₃ (wt%)	MnO (wt%)	MgO (wt%)	CaO (wt%)	Na ₂ O (wt%)	K ₂ O (wt%)	TiO ₂ (wt%)	P ₂ O ₅ (wt%)	S (wt%)	LOI (wt%)	Pb (ppm)	As (ppm)
VR1	77.6	13.5	2.85	0.06	0.06	0.03	0.24	1.37	0.47	0.06	1.82	5.91	439	1480
VR2	82.2	13.2	3.17	0.01	0.05	0.04	0.24	0.26	0.48	0.04	1.6	6.63	166	218
VR3	80.7	17.2	3.75	0.01	0.06	0.01	0.26	1.57	0.28	0.05	2.61	4.23	1910	101
VR4	63.3	6.88	16.0	0.02	0.05	0.05	0.01	0.10	0.25	0.06	10.6	12.7	375	150
SR1	58.7	14.4	6.22	0.07	3.49	3.31	1.31	3.22	0.82	0.13	0.2	6.26	BDL	23.6
SR2	63	12.1	5.43	0.1	2.85	1.59	0.5	3.02	0.41	0.04	0.27	2.86	BDL	10.5
SR3	65.8	15	6.4	0.07	2.81	2.6	1.7	3.54	0.72	0.17	0.48	6.77	BDL	6.9

*BDL: below detection limit

Table 3: Mineralogical composition of rock samples determined using XRD.

Sample	Qz	Pl	Chl	Cal	Py	Kao	Gal	Bir	Mus
VR1	+++				+	-			
VR2	+++				+	-			
VR3	+++				+	-	-		
VR4	+++				+	-	-	-	
SR1	+++	++	-	+	-	-			-
SR2	+++	++		+	-	-			-
SR3	+++	++	+	+	-				

+++; Major; ++: Moderate; +: Minor; -: Trace

Qz: Quartz; Pl: Plagioclase; Chl: Chlorite; Cal: Calcite; Py: Pyrite; Kao: Kaolinite; Gal: Galena; Bir: Birnessite; Mus: Muscovite

*Note: In samples with little or no volcanic glass (e.g., sedimentary rocks), major, moderate, minor and trace roughly represent >30%, 10-30%, 2-10% and <2%, respectively.

Table 4. EPMA data of As and Pb from minerals inside and outside the fractured area of the sedimentary rock (SR3).

Inside the fractured zone						
Mineral name	No.	S (wt%)	Fe (wt%)	As (wt%)	Pb (wt%)	Total
Pyrite	1	49.8	46.2	0.13	<0.01	96.1
	2	49.5	45.3	0.11	<0.01	94.9
	3	50.5	43.3	0.06	<0.01	93.9
Smectite	1	0.01	3.42	<0.01	0.04	3.47
	2	0.01	4.63	<0.01	<0.01	4.64
Illite	1	0.01	3.25	0.07	0.03	3.29
Outside the fractured zone						
Mineral name	No.	S (wt%)	Fe (wt%)	As (wt%)	Pb (wt%)	Total
Pyrite	1	50.9	46.7	<0.01	<0.01	97.6
	2	51	46.1	<0.01	<0.01	97.1
	3	51.2	45.7	<0.01	<0.01	96.9
Calcite	1	0.06	0.22	<0.01	0.03	0.31

Table 5. EPMA data of Fe, S, As and Pb in pyrite from altered volcanic rocks (VR1, 2 and 3).

No.	S (wt%)	S (at%)	Fe (wt%)	Fe (at%)	As (wt%)	As (at%)	Pb (wt%)	Pb (at%)	Total
1	54.2	67.1	45.9	32.6	0.55	0.28	0.26	0.04	100.9
2	53.4	66.7	46.3	33.1	0.38	0.20	0.07	0.01	100.1
3	53.6	67.0	46	32.9	0.14	0.08	0.18	0.04	99.9
4	52.9	66.1	46.3	33.3	1.07	0.56	0.08	0.02	100.3
5	53.7	66.8	46.6	33.2	0.09	0.04	0.11	0.02	100.5
6	54	66.9	46.5	33.1	<0.01	<0.01	0.15	0.03	100.6
7	53.7	66.5	46.7	33.5	<0.01	<0.01	0.07	0.01	100.5
8	52.7	66.6	45.2	32.9	0.97	0.53	0.19	0.04	99.1
9	54	66.9	46.6	33.1	0.01	0.01	0.19	0.04	100.8
10	54.2	66.8	46.9	33.2	0.02	0.01	0.09	0.02	101.2
11	54.1	66.8	47.1	33.2	0.02	0.01	0.09	0.02	101.3
12	53.9	66.3	47.2	33.5	0.19	0.12	0.15	0.03	101.4
13	52.4	67.1	44.4	32.7	0.35	0.19	<0.01	<0.01	97.15
14	51.4	66.2	45.6	33.7	0.03	0.02	0.12	0.02	97.15
15	54.2	66.4	47.8	33.6	0.01	0.01	0.11	0.02	102.1
16	54.3	66.3	47.9	33.6	0.12	0.06	0.09	0.02	102.4
Minimum	51.4	66.1	44.4	32.6	<0.01	<0.01	<0.01	<0.01	99.1
Maximum	54.3	67.1	47.9	33.7	1.07	0.56	0.26	0.04	101.4
Mean	53.5	66.7	46.4	33.2	0.28	0.15	0.13	0.02	100.4

Table 6. Leachate chemistry of sample SR1 without pH adjustment in triplicate runs

Trial	pH	As ($\mu\text{g/L}$)	Fe (mg/L)	Ca ²⁺ (mg/L)	SO ₄ ²⁻ (mg/L)
1	8.79	28.3	0.304	22.9	77.9
2	8.79	35	0.307	23.4	77.8
3	8.82	30.6	0.313	22.9	78.6
Mean	8.8	31.3	0.308	23.1	78.1
Standard deviation	0.017	3.40	0.005	0.289	0.436

Table 7. Leachate chemistry of sample VR4 without pH adjustment in triplicate runs

Trial	pH	As ($\mu\text{g/L}$)	Fe (mg/L)	Ca ²⁺ (mg/L)	SO ₄ ²⁻ (mg/L)
1	2.85	45.4	474	2.1	853
2	2.82	47.9	471	3	843
3	2.83	44.5	476	1.8	853
Mean	2.83	45.9	473	2.3	850
Standard deviation	0.015	1.76	2.52	0.62	5.77

Table 8. Leachate chemistry of the altered rocks without pH adjustment under oxic conditions.

Sample	pH	As ($\mu\text{g/L}$)	Pb (mg/L)	Fe (mg/L)	Ca ²⁺ (mg/L)	SO ₄ ²⁻ (mg/L)
VR1	3.86	195	1.48	16.2	2.62	100
VR2	3.39	102	0.353	33.3	3.03	144
VR3	3.99	94.7	0.773	8.59	2.4	67.8
VR4	2.85	45.4	2.12	474	2.1	853
SR1	8.79	28.3	-	0.304	22.9	77.9
SR2	8.49	46.8	-	0.113	17.3	12.7
SR3	10.32	17.25	-	0.648	5.18	58.9

*-: not measured

Table 9. Leachate chemistry of sample VR2 under oxic, anoxic and mixed CO₂/Ar gases.

Conditions	pH	As (µg/L)	Pb (mg/L)	Fe (mg/L)	Ca ²⁺ (mg/L)	SO ₄ ²⁻ (mg/L)
Anoxic	3.45	12.9	0.374	25.9	0.284	121
Oxic	3.39	102	0.353	33.3	3.03	144
1% CO ₂ ; 99% Ar gas	3.4	12.5	0.382	25.3	0.831	124
3% CO ₂ ; 97% Ar gas	3.48	12.7	0.375	24.8	0.646	134

Table 10. Leachate chemistry of sample SR2 under oxic, anoxic and mixed CO₂/Ar gases.

Conditions	pH	As (µg/L)	Fe (mg/L)	Ca ²⁺ (mg/L)	SO ₄ ²⁻ (mg/L)
Anoxic	9.21	14.8	0.101	9.44	11.7
Oxic	8.49	46.8	0.113	17.3	12.7
1% CO ₂ ; 99% Ar gas	7.29	3.8	0.028	33.7	15
3% CO ₂ ; 97% Ar gas	6.90	3.8	0.021	55.5	12.3
30% CO ₂ ; 70% Ar gas	6.10	3.35	0.012	108	12.7

Table 11. Calculated saturation indices for sample VR 1

Parameter	Leachate			
	1	2	3	4
pH (measured)	3.86	4.34	7.75	9.95
pe (calculated from Eh)	9.06	8.65	6.15	6.15
<i>Saturation indices</i>				
Gypsum (CaSO ₄ ·2H ₂ O)	-2.51	-2.62	-4.38	-5.60
Calcite (CaCO ₃)	-	-	-3.81	-2.75
Anglesite (PbSO ₄)	-0.39	-0.70	-2.96	-4.74
Cerussite (PbCO ₃)	-	-	-0.62	-0.12
Hydrocerussite (Pb(CO ₃) ₂ (OH) ₂)	-	-	-2.19	1.35
Pb(OH) ₂	-6.0	-5.33	-0.91	1.62
Fe(OH) ₃	-0.05	0.53	3.46	3.15
Goethite (FeOOH)	4.12	4.70	7.63	7.32
Hematite, alpha (Fe ₂ O ₃)	7.75	8.90	14.8	14.4
Magnetite, alpha (Fe ₃ O ₄)	1.48	3.14	11.0	7.89

Table 12. Calculated saturation indices for sample SR 1

Parameter	Leachate				
	1	2	3	4	5
pH (measured)	1.95	4.38	7.64	8.21	11.8
pe (calculated from Eh)	10.7	8.07	5.54	5.63	3.03
<i>Saturation indices</i>					
Gypsum (CaSO ₄ ·2H ₂ O)	-0.57	-0.67	-0.81	-0.98	-3.1
Calcite (CaCO ₃)	-	-	-0.12	0.94	1.29
Fe(OH) ₃	-3.22	0.52	2.12	2.09	0.71
Goethite (FeOOH)	0.94	4.68	6.29	6.25	4.88
Hematite, alpha (Fe ₂ O ₃)	1.39	8.87	12.1	12	9.29
Magnetite, alpha (Fe ₃ O ₄)	-7.8	3.64	7.73	5.77	1.84

Table 13. Chemical reactions and equilibrium constants at 25⁰C of hematite, magnetite and goethite in the presence of O₂ and N₂ using Geochemist's workbench[®] RXN program.

Chemical reactions	log K
2 Fe ²⁺ _(aq) + 2 H ₂ O + 0.5 O _{2(aq)} \longleftrightarrow Hematite (Fe ₂ O ₃) _(s) + 4 H ⁺	16.9
3 Fe ²⁺ _(aq) + 3 H ₂ O + 0.5 O _{2(aq)} \longleftrightarrow Magnetite (Fe ₃ O ₄) _(s) + 6 H ⁺	6.68
Fe ²⁺ _(aq) + 1.5 H ₂ O + 0.25 O _{2(aq)} \longleftrightarrow Goethite (FeOOH) _(s) + 2 H ⁺	7.99
2 Fe ²⁺ _(aq) + 1.8 H ₂ O + 0.4 NO ₃ ⁻ _(aq) \longleftrightarrow Hematite (Fe ₂ O ₃) _(s) + 3.6 H ⁺ + 0.2 N _{2(aq)}	15.4
3 Fe ²⁺ _(aq) + 2.8 H ₂ O + 0.4 NO ₃ ⁻ _(aq) \longleftrightarrow Magnetite (Fe ₃ O ₄) _(s) + 5.6 H ⁺ + 0.2 N _{2(aq)}	5.13
Fe ²⁺ _(aq) + 1.4 H ₂ O + 0.2 NO ₃ ⁻ _(aq) \longleftrightarrow Goethite (FeOOH) _(s) + 1.8 H ⁺ + 0.1 N _{2(aq)}	7.21

Table 14. pH and concentrations of Ca²⁺, Fe and SO₄²⁻ in simulated systems under equilibrium conditions, and measured values from sample SR1

Type of system	Ratio of minerals (mol)	pH	Ca ²⁺ (mol/L)	Fe (mol/L)	SO ₄ ²⁻ (mol/L)
altered sedimentary rock (SR1)	-	8.79	5.7 x 10 ⁻⁴	5.4 x 10 ⁻⁶	8.1 x 10 ⁻⁴
calcite only	10	9.84	1.2 x 10 ⁻⁴	-	-
calcite : pyrite	10 : 1	9.23	2.3 x 10 ⁻⁴	4.1 x 10 ⁻⁵	1.5 x 10 ⁻⁴
calcite : pyrite : gypsum	10 : 1 : 4.2 x 10 ⁻⁴	8.91	5.8 x 10 ⁻⁴	5.3 x 10 ⁻⁵	5.3 x 10 ⁻⁴
calcite : pyrite : gypsum : melanterite	10 : 1 : 4.2 x 10 ⁻⁴ : 3.2 x 10 ⁻⁴	8.83	6.2 x 10 ⁻⁴	2.8 x 10 ⁻³	8.1 x 10 ⁻⁴

Table 15. pH and concentrations of Ca^{2+} , Fe and SO_4^{2-} in simulated systems under equilibrium conditions, and measured values from sample VR4

Type of system	Ratio of minerals (mol)	pH	Ca^{2+} (mol/L)	Fe (mol/L)	SO_4^{2-} (mol/L)
altered volcanic rock (VR4)	-	2.85	5.7×10^{-5}	8.5×10^{-3}	9.0×10^{-3}
pyrite only	10	3.83	-	7.4×10^{-5}	1.5×10^{-4}
pyrite : gypsum	$10 : 6 \times 10^{-5}$	3.84	5.8×10^{-5}	7.4×10^{-5}	2.1×10^{-5}
pyrite : gypsum : melanterite	$10 : 6 \times 10^{-5} : 1.25 \times 10^{-2}$	4.05	4.1×10^{-5}	9.0×10^{-3}	9.0×10^{-3}

Table 16. Simulated pH concentrations of Ca^{2+} and SO_4^{2-} in a calcite-pyrite-gypsum-melanterite system under equilibrium conditions, and measured values from sample SR1

Parameters	Simulated			Measured
		Anoxic - oxic		Anoxic - oxic
P_{O_2}	10^{-34}	$10^{-1.36}$	$10^{-0.68}$	-
P_{CO_2}	10^{-175}	$10^{-7.0}$	$10^{-3.5}$	-
pH	9.23	9.15	8.83	9.01 - 8.79
Ca^{2+} (mol/L)	5.2×10^{-4}	5.4×10^{-4}	6.3×10^{-4}	4.6×10^{-4} - 5.7×10^{-4}
SO_4^{2-} (mol/L)	6.8×10^{-4}	7.1×10^{-4}	8.1×10^{-4}	7.1×10^{-4} - 8.1×10^{-4}

Table 17. Simulated pH and concentrations of Ca^{2+} and SO_4^{2-} in a pyrite-gypsum-melanterite system under equilibrium conditions, and measured values from sample VR4

Parameters	Simulated			Measured
		Anoxic - oxic		Anoxic - oxic
P_{O_2}	10^{-34}	$10^{-1.36}$	$10^{-0.68}$	-
P_{CO_2}	10^{-175}	$10^{-7.0}$	$10^{-3.5}$	-
pH	6	4.7	4.05	3.03 - 2.85
Ca^{2+} (mol/L)	4.1×10^{-5}	4.1×10^{-5}	4.1×10^{-5}	5.7×10^{-5} - 5.2×10^{-4}
SO_4^{2-} (mol/L)	9.0×10^{-2}	9.0×10^{-3}	9.0×10^{-3}	9.0×10^{-3} - 9.0×10^{-3}