



HOKKAIDO UNIVERSITY

Title	Polycyclic Aromatic Compounds-mediated Electrochemical Reduction of Alkyl Mesylates
Author(s)	Senboku, Hisanori; Takahashi, Megumi; Fukuhara, Tsuyoshi et al.
Citation	Chemistry Letters, 36(2), 228-229 https://doi.org/10.1246/cl.2007.228
Issue Date	2007-02-05
Doc URL	https://hdl.handle.net/2115/70907
Type	journal article
File Information	cl.2007.228.pdf



Polycyclic Aromatic Compounds-mediated Electrochemical Reduction of Alkyl Mesylates

Hisanori Senboku,* Megumi Takahashi, Tsuyoshi Fukuhara, and Shoji Hara

Laboratory of Organic Reaction, Division of Chemical Process Engineering, Graduate School of Engineering, Hokkaido University, Sapporo 060-8628

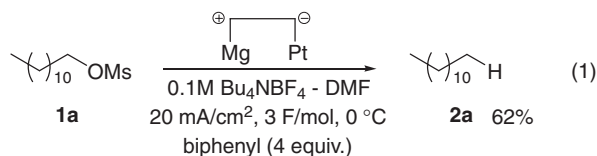
(Received October 30, 2006; CL-061275; E-mail: senboku@eng.hokudai.ac.jp)

Electrochemical reduction of alkyl mesylates was successfully carried out by using an undivided cell equipped with a Pt cathode and an Mg anode in the presence of biphenyl and *t*-BuOH. The reaction could proceed efficiently under mild conditions to give the corresponding alkanes in moderate to good yields. This procedure could also be applicable to chemoselective reduction of mesylates having functional groups such as epoxide, olefin, acetal, hydroxy, or cyano groups.

Reduction of alcohols to the corresponding alkanes is an important transformation in organic synthesis.^{1–3} One of simple and useful methods involves conversion of hydroxy groups to the corresponding alkyl sulfonates, such as tosylates or mesylates, followed by reduction with hydride. As a hydride reagent, LiAlH₄⁴ or LiBEt₃H^{5,6} is frequently used for this aim. Their high reactivity, however, sometimes causes serious troubles such as over-reduction of other functional groups in the substrates. Therefore, chemoselective and widely applicable reduction of alkyl sulfonates under mild conditions would be still desirable although several alternative methods for the reduction of alkyl tosylates or mesylates have been developed.^{7–13} On the other hand, electrochemical process can reduce various organic compounds under neutral and mild conditions.¹⁴ While electrochemical reduction of alkyl tosylates predominantly gave the corresponding alcohols,¹⁵ a similar reduction of alkyl mesylates gave the corresponding alkanes. Thus, electrochemical reduction of alkyl mesylates could take place efficiently by using a divided cell equipped with a lead cathode and a platinum anode to give the corresponding alkanes in good yields.¹² It was also reported that mesylates of α -hydroxy esters could be reduced to the corresponding esters by indirect electrochemical reduction using diphenyl diselenide in a divided cell.¹³ Although both electrochemical reactions could proceed under neutral and mild conditions to give the products in good yields, use of unfunctionalized metal as an electrode and a complicated divided cell could not be satisfactory from the view points of environmentally friendly organic synthesis. We tried electrochemical reduction by using a simple and easy-handled undivided cell equipped with a platinum cathode and a magnesium anode, and found that electrochemical reduction of alkyl mesylates could proceed efficiently in the presence of polycyclic aromatic compound such as naphthalene and biphenyl as an electron-transfer mediator to give the corresponding alkanes in moderate to good yields. In this communication, we wish to report polycyclic aromatic compounds-mediated electrochemical reduction of alkyl mesylates using an undivided cell.

A constant current electrolysis of a DMF solution of dodecyl mesylate **1a** containing 0.1 M Bu₄NBF₄ using an undivided cell equipped with a platinum plate cathode and a magnesium rod

anode gave a 35:65 mixture of dodecane **2a** and 1-dodecanol. Therefore, we tried the electrolysis in the presence of polycyclic aromatic compounds as an electron-transfer mediator. Other than benzene and toluene, we found naphthalene, phenanthrene, and *m*-terphenyl were effective for yielding **2a** selectively. Better results were obtained when the electrolysis of **1a** was carried out in the presence of biphenyl which exclusively gave **2a** in 62% yield under the reaction conditions described in eq 1 (Conditions A). Biphenyl could be recovered quantitatively after the reaction and the recovered biphenyl could also be reused as a mediator.



During continuous study on scope and limitation of the present reaction using functionalized substrates, it was found that this reaction system was also applicable to the substrate having a hydroxy group¹⁶ (See below). Then, we tried the electrolysis in the presence of a proton source and the results are summarized in Table 1. As the best result, electrochemical reduction of mesylate **1a** in the presence of biphenyl (4 equiv.) and *t*-BuOH (10 equiv.) with 8 F/mol of electricity gave **2a** in 74% yield (Entry 7 in Table 1) (Conditions B). It is also noted that in this case, biphenyl could also be recovered in quantitative yields.

Electrochemical reduction of various alkyl mesylates **1** was carried out under "Conditions A (in the presence of biphenyl with 3 F/mol of electricity) and B (in the presence of biphenyl and *t*-BuOH with 8 F/mol of electricity),"¹⁷ and the results are summarized in Table 2. The present reaction was also effective

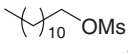
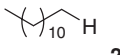
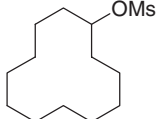
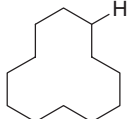
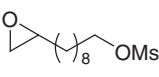
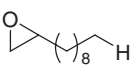
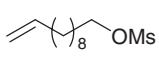
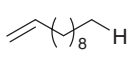
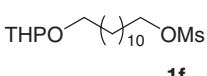
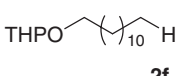
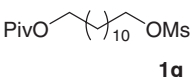
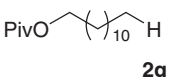
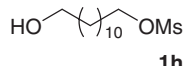
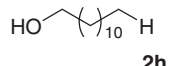
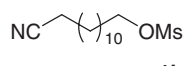
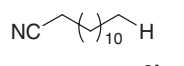
Table 1. Effect of an additive^a

Entry	Additive (equiv.)	Electricity /F mol ⁻¹	Yield ^b /%	Recovered 1a ^b /%
1	none	5	59	0
2	<i>t</i> -BuOH (2)	5	62	0
3	<i>t</i> -BuOH (4)	5	64	trace
4	<i>t</i> -BuOH (10)	5	61	17
5	<i>t</i> -BuOH (20)	5	57	20
6	<i>t</i> -BuOH (50)	5	56	28
7	<i>t</i> -BuOH (10)	8	74	0
8	<i>i</i> -PrOH (10)	8	66	trace
9	EtOH (10)	8	37	42
10 ^c	H ₂ O (10)	8	61	15

^aA constant current electrolysis (20 mA/cm²) of **1a** (1 mmol) in DMF (16 mL) containing 0.1 M Bu₄NBF₄ was carried out in the presence of biphenyl (4 equiv.) at 0 °C. ^bIsolated yield.

^c1-Dodecanol was obtained in 4% yield.

Table 2. Electrochemical reduction of alkyl mesylate **1**^a

$ \begin{array}{ccc} \text{R-OMs} & \xrightarrow[\text{biphenyl (4 equiv.), } t\text{-BuOH (10 equiv.)}]{\text{Mg}^+ \text{ Pt}^- \text{ 0.1M Bu}_4\text{NBF}_4 \text{ - DMF}} & \text{R-H} \\ \mathbf{1} & & \mathbf{2} \\ & \text{20 mA/cm}^2, 8 \text{ F/mol, } 0^\circ\text{C} & \end{array} $			
Entry	Substrate	Product	Yield/% ^b
1	 1a	 2a	74 (62)
2	 1b	 2b	62 (55)
3	 1c	 2c	68
4	 1d	 2d	69 (55)
5	 1e	 2e	63 (48)
6	 1f	 2f	76 (59)
7	 1g	 2g	92 (61)
8	 1h	 2h	75 (57)
9	 1i	 2i	84 (57)

^aA constant current electrolysis (20 mA/cm²) of **1** (1 mmol) in DMF (16 mL) containing 0.1 M Bu₄NBF₄ was carried out at 0°C in the presence of biphenyl (4 equiv.) and *t*-BuOH (10 equiv.) with 8 F/mol of electricity. ^bIsolated yields. In parentheses, the yields in the reaction under Conditions A are shown.

for reduction of secondary alkyl mesylate **1b** and cycloalkyl mesylate **1c** (Entries 2 and 3 in Table 2). It is also noted that several functional groups were inert in this reduction system. Thus, electrochemical reduction of alkyl mesylates having epoxide, olefin, acetal, pivalate, hydroxy, or cyano group was also achieved in moderate to good yields without change of their functional groups (Entries 4–9 in Table 2). It is difficult to reduce mesylates **1d**, **1g**, and **1i** to the corresponding alkanes **2d**, **2g**, and **2i** without changing their functional groups, respectively when the reduction was carried out by using LiAlH₄, a representative

reducing agent.

Cyclic voltammetry of **1a** and biphenyl was carried out. No reduction peak of **1a** was observed at >–3.2 V vs Ag/Ag⁺ and a reduction peak of biphenyl appeared at ca. –3.2 V vs Ag/Ag⁺. On the other hand, cyclic voltammetry of biphenyl in the presence of **1a** showed that the reduction current at –3.2 V vs Ag/Ag⁺ increased when the amount of **1a** increased. These results indicated that biphenyl would work as an electron-transfer mediator.

This work was financially supported by Grant-in-Aid for Scientific Research (C) (No. 15550082) from the Japan Society for the Promotion of Science.

References and Notes

- 1 S. W. McCombie, in *Comprehensive Organic Synthesis*, ed. by B. M. Trost, I. Fleming, Pergamon Press, **1991**, Vol. 8, pp. 811–833.
- 2 M. Yasuda, Y. Onishi, M. Ueba, T. Miyai, A. Baba, *J. Org. Chem.* **2001**, 66, 7741.
- 3 a) V. Gevorgyan, J.-X. Liu, M. Rubin, S. Benson, Y. Yamamoto, *Tetrahedron Lett.* **1999**, 40, 8919. b) V. Gevorgyan, M. Rubin, S. Benson, J.-X. Liu, Y. Yamamoto, *J. Org. Chem.* **2000**, 65, 6179.
- 4 L. A. Paquette, in *Encyclopedia of Reagents for Organic Synthesis*, ed. by L. A. Paquette, John Wiley and Sons Ltd., Chichester, **1995**, Vol. 5, pp. 3009–3014.
- 5 M. Zaidlewicz, H. C. Brown, in *Encyclopedia of Reagents for Organic Synthesis*, ed. by L. A. Paquette, John Wiley and Sons Ltd., Chichester, **1995**, Vol. 5, pp. 3180–3183.
- 6 LiEt₃BH generated in situ from Et₃B and lithium dimethylaminoborohydride was also applicable for the reduction of alkyl mesylates to the corresponding alkanes: S. Thomas, T. Huynh, V. Enriquez-Rios, B. Singaram, *Org. Lett.* **2001**, 3, 3915.
- 7 Y. Fujimoto, T. Tatsuno, *Tetrahedron Lett.* **1976**, 17, 3325.
- 8 F. Rolla, *J. Org. Chem.* **1981**, 46, 3909.
- 9 Y. Ueno, C. Tanaka, M. Okawara, *Chem. Lett.* **1983**, 795.
- 10 G. Radivoy, A. Alonso, M. Yus, *Tetrahedron* **1999**, 55, 14479.
- 11 G. Radivoy, A. Alonso, Y. Moglic, C. Vitale, M. Yus, *Tetrahedron* **2005**, 61, 3859.
- 12 T. Shono, Y. Matsumura, K. Tsubata, Y. Sugihara, *Tetrahedron Lett.* **1979**, 20, 2157.
- 13 T. Inokuchi, T. Sugimoto, M. Kusumoto, S. Torii, *Bull. Chem. Soc. Jpn.* **1992**, 65, 3200.
- 14 S. Torii, *Electroorganic Reduction Synthesis*, Kodansha Ltd., Tokyo, **2006**.
- 15 P. Yousefzadeh, C. K. Mann, *J. Org. Chem.* **1968**, 33, 2716.
- 16 There was a possibility that anion radical of biphenyl generated by one electron reduction would react with a proton source such as hydroxy group. Further one electron reduction followed by protonation would produce 1,4-dihydrobiphenyl, which could not work as an electron-transfer mediator.
- 17 General procedure for electrochemical reduction (Conditions B): A DMF solution (16 mL) of **1** (1 mmol), biphenyl (4 mmol), and *t*-BuOH (10 mmol) containing 0.1 M Bu₄NBF₄ was electrolyzed at 0°C with a constant current (20 mA/cm²) under a nitrogen. Undivided cell equipped with a platinum plate cathode (2 × 2 cm²) and a magnesium rod anode (3 mmφ) was used for electrolysis. After 8 F/mol of electricity was passed, electrolyzed solution was acidified with 1 M HCl and then extracted with ether. The ethereal solution was washed successively with H₂O and saturated brine. Dryness with MgSO₄ followed by evaporation gave a crude product, which was purified by column chromatography with SiO₂ to obtain **2**. Biphenyl could also be recovered quantitatively.