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ARTICLE

Distinctive Stability of Free-standing Monolayer Clay Mineral Nanosheet under Transmission Electron Microscopy

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Among 2D materials, clay mineral nanosheets have been extensively studied owing to their specific features, such as high surface charge and large surface area. Recently, we reported a stable free-standing (without any surfactants or matrices) monolayer clay mineral, characterized via annular dark-field scanning transmission electron microscopy (ADF-STEM) at the atomic-scale. Herein, we demonstrated that the monolayer clay mineral exhibited the outstanding stability under electron beam irradiation compared to two- or three-layered nanosheets via electron diffraction analysis. In addition to its low thickness (~1 nm-thick), the absence of an interlayer space was the critical factor contributed to the distinctive stability of the monolayer clay mineral, compared to that of the two- or three-layered clay mineral. The findings here inspire further investigation in free-standing clay mineral using (S)TEM and also for a wide variety of nanomaterials which are strongly hydrated.

Introduction

Clay minerals are phyllosilicates, featuring physicochemical properties such as atomically flat surfaces with a large surface area, high surface charges, and high thermal stability. Utilizing these features, clay minerals have been used in various fields ranging from environmental applications (as absorbents, ion exchangers, and water decontamination) to fine chemical synthesis.^{1–3} Among the many types of clay minerals, smectite groups (e.g., saponite, montmorillonite) are frequently used owing to their moderate anionic charge density, imparting exfoliation (or stacking) ability to individual nanosheets in aqueous solutions as well as optical transparency in the visible light region in the exfoliated state at small particle sizes.^{4–7} Smectite possess the following chemical formula: $[(\text{Si}_{8-x}\text{Al}_x)(\text{Mg}_{6-y}\text{Al}_y)\text{O}_{20}(\text{OH})_4] \cdot (x-y)\text{Y}^+$, where Y represents the cations to balance the charges, and they possess layers consisting of a 2:1 ratio of tetrahedral and octahedral sheets, with a thickness of 0.96 nm. Completely exfoliated clay mineral, as a single layer, can be defined as a family of two-dimensional (2D) nanosheets.

Recently, characterization of 2D nanomaterials has greatly advanced owing to the atomic-scale imaging using aberration-corrected transmission electron microscopy (TEM) or scanning transmission electron microscopy (STEM).^{8–13} Furthermore, for materials with simple structures such as graphene and transition metal dichalcogenide nanosheets, single-atom and single-molecule observations within hybrid materials have been

achieved.^{14–21} However, (S)TEM imaging of free-standing monolayer clay minerals, which have a more complicated structure than their preceding few-atom thick materials, such as graphene, at the atomic-scale has not been achieved, except in our recent study.²² Once non-destructive direct imaging of clay mineral nanosheet or useful clay mineral-based organic-inorganic hybrid materials at the atomic scale is established, it will strongly promote their applicability such as selective adsorption for heterogeneous catalyst or sensing,²³ or photoenergy conversion.⁵

We recently achieved atomic-scale imaging of a free-standing (i.e., without any surfactants or supporting matrices) monolayer clay mineral by annular dark-field (ADF) STEM.^{22,24} The monolayer clay mineral was observed at the atomic-scale without any destruction, although bulk clay minerals are generally sensitive to the electron beam irradiation, as reported for cross-sectional imaging at the atomic-scale via (S)TEM.^{25–29} In this report, we demonstrate the outstanding stability of free-standing monolayer clay mineral by comparing the decrease of selected area electron diffraction (SAED) intensity against those of the two- or three-layered (2L or 3L) clay mineral. The critical dose was determined as the point at which the electron diffraction intensities were reduced to 1/e (corresponding to 37% reduction) of its initial value. In addition to the fascinating property of clay mineral as itself, clay mineral nanosheets can serve as a simple model for considering beam-damage mechanisms in oxide nanosheets without using any surfactants that are commonly required for many 2D materials in their synthetic or exfoliation processes.^{8,30} The results here revealed the crucial effect of interlayer spaces to the beam damage mechanism rather than the thickness among a few layers of clay mineral, which can also be useful for understanding the electron beam damage to a wide variety of nanomaterials that are strongly hydrated.

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[†]Electronic Supplementary Information (ESI) available: Additional relative SAED intensity decay for 2L Mt, and SAED patterns at difference dose rate. See DOI: 10.1039/x0xx00000x

Methods

Sample preparation. Montmorillonite (Mt), a typical anionic clay mineral, was acquired from Kunimine Industries (Kunipia F with the chemical structure corresponding to $\text{Na}_{0.35}\text{K}_{0.01}\text{Ca}_{0.06}(\text{Al}_{1.56}\text{Mg}_{0.33}\text{Fe}_{0.09}\text{Ti}_{0.01})(\text{Si}_{3.87}\text{Al}_{0.13})\text{O}_{10}(\text{OH})_2$ and the cation exchange capacity of 1.19 meq. g^{-1})^{31,32} and was purified using the following procedure. First, raw clay mineral dispersion (300 mL , 10 g L^{-1}) was placed in centrifuge bottles, which were then centrifuged at a rotation speed of 7000 rpm for 15 min . The obtained supernatant was collected and poured into a container with 1500 mL of ethanol. After 1 h of mixing at a temperature of 70°C , the resulting colloidal solution was filtered using a PTFE membrane with a pore size of $0.1 \mu\text{m}$ (Millipore). The produced cake was collected and dried under vacuum by rotary pump overnight. The dried cake was dispersed in water, and the dispersion was stored overnight to achieve complete exfoliation into nanosheets. A specimen for STEM/TEM was prepared by drop casting the aqueous solution of monolayer Mt on a holey carbon-coated Cu grid.

ADF-STEM observation. For the STEM observations for determination of the layer number, a Titan Cubed G2 equipped with spherical aberration correctors was used at an acceleration voltage of 300 kV . The convergence semi-angle of the incident probe was 21.4 mrad . The inner and outer detection semi-angle of ADF detector were 31.4 and 192 mrad , respectively. The incident probe current was set to $\sim 30 \text{ pA}$.

Time series of TEM-SAED. TEM-SAED observation was performed by using the Titan Cubed G2 at an acceleration voltage of 300 kV . The dose rate was set to $10 \text{ e}^- \text{ \AA}^{-2} \text{ s}^{-1}$ by monitoring the screen current and the exposure time was 3 s for each image acquisition. Note that there is no calibration from the screen current to the actual number of electrons.

Results and Discussion

Determination of the stacking number of free-standing Mt nanosheet by ADF-STEM observation.

The atomic structure of the unit layer structure for typical monolayer montmorillonite (1L Mt) is shown in Fig. 1a. The unit layer comprises 2:1 combination of tetrahedral silica and octahedral aluminum hydroxide sheets. The general chemical formula of the layered aluminosilicates, such as Mt, mica, and pyrophyllite, is $[\text{Si}_4\text{O}_8][\text{Al}_2\text{O}_2(\text{OH})_2]$. Two-thirds of the oxygen atoms in the octahedral sheet are supplied by silica tetrahedra, and the others are supplied by hydroxyl groups (indicated as yellow spheres in Fig. 1a). The isomorphous substitution of Al^{3+} to Mg^{2+} in the octahedral sheet produces anionic charges within the Mt structure. Fig. 1b shows the ADF-STEM image of the free-standing Mt nanosheets on a holey carbon film at low magnification. The acceleration voltage and the incident probe current were 300 kV and $\sim 30 \text{ pA}$, respectively. The free-standing Mt sample was prepared by pouring an aqueous dispersion of the exfoliated monolayer Mt onto a holey carbon-coated Cu grid. As shown in Fig. 1b, the free-standing Mt nanosheets were

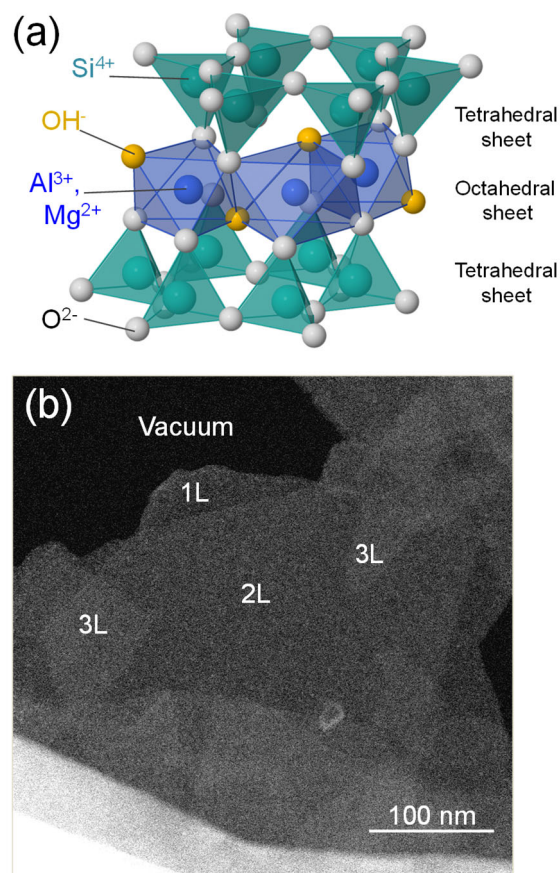


Fig. 1. (a) Atomic structure model for a unit layer of montmorillonite showing the octahedral alumina sheet sandwiched by two tetrahedral silica sheets. Hydrogen atoms at the tetrahedral sheet edge are omitted for clarity. (b) ADF-STEM image of free-standing Mt nanosheets deposited on a holey carbon grid, and the corresponding stacked-layer number from one (1L) to three (3L).

stacked horizontally. Hence, the electron beam was irradiated in a direction normal to the surface. The stacking number of the Mt nanosheets, ranging from 1L to 3L, was determined by the difference between their contrasts, as previously reported (Fig. 1b).²⁴ To further analyze the stability of free-standing Mt under electron beam irradiation, we determined the critical doses of 1L, 2L, and 3L Mt from the intensity reduction in their SAED patterns. Note that free-standing Mt, especially 1L Mt, were difficult to find in normal TEM, due to the ultrathin structure consisted with light elements. Therefore, the stacking number of Mt was determined via ADF-STEM, where the contrast positively correlated with the atomic number or thickness of the specimen.²⁴

The critical dose of electron beam damage as a function of the layer number.

Fig. 2 shows a series of SAED patterns for (a) 1L, (b) 2L, and (c) 3L Mt at initial exposure and cumulative electron doses of 3 , 15 , and $30 \times 10^3 \text{ e}^- \text{ \AA}^{-2}$. The acceleration voltage and the dwell for each pattern were 300 kV and 3 s , respectively. For the 1L Mt shown in Fig. 2a,

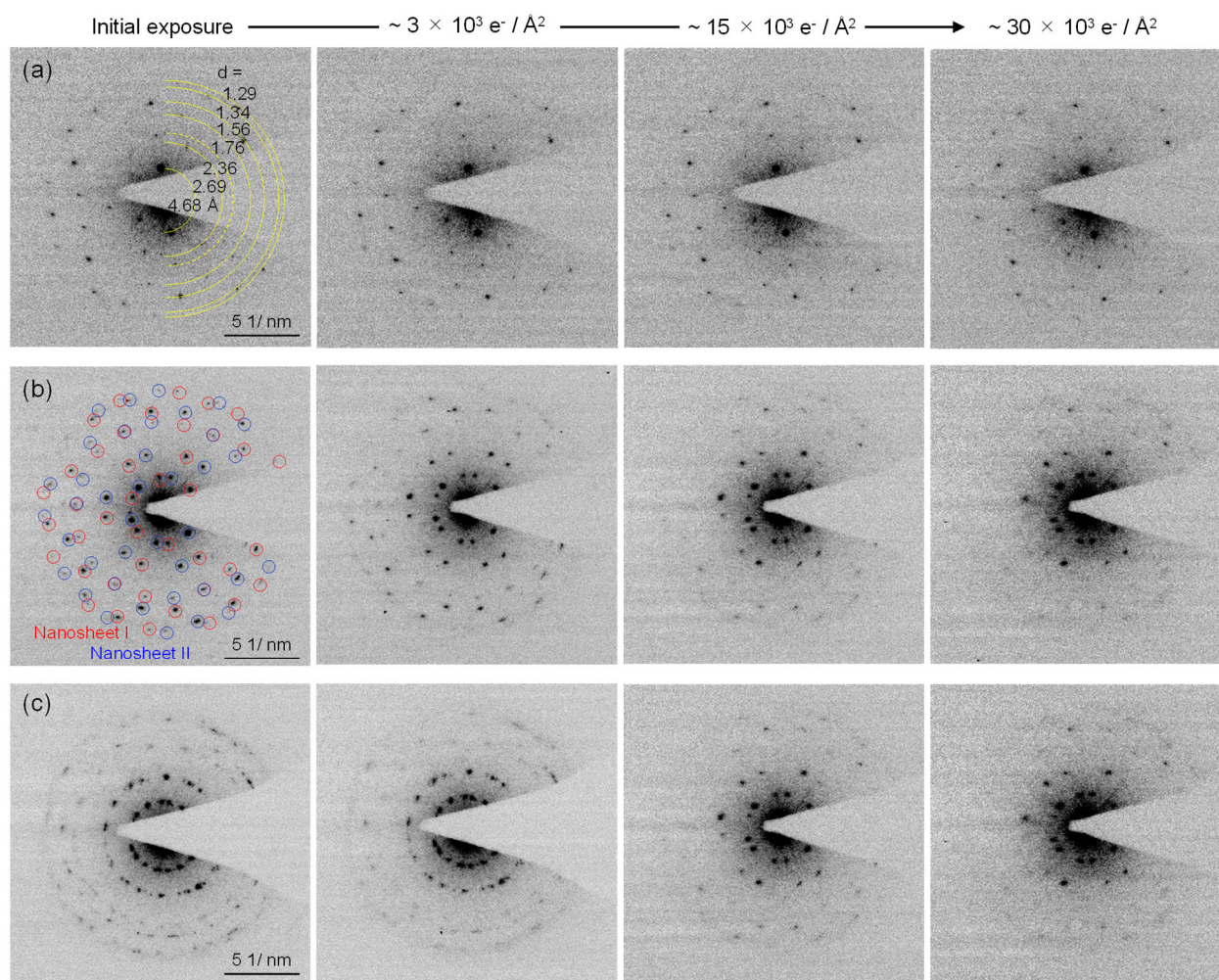


Fig. 2. Series of SAED patterns for (a) monolayer (1L), (b) two-layered (2L), and (c) three-layered (3L) Mt at initial exposure and cumulative electron doses of 3, 15, and $30 \times 10^3 \text{ e}^- \text{ Å}^{-2}$. The acceleration voltage and the electron dose rate were 300 kV and $10 \text{ e}^- \text{ Å}^{-2} \text{ s}^{-1}$, respectively. The dwell time for each pattern was set at 3 s. The in-plane distance (d) of 1L Mt was indicated by yellow lines in (a). Red and blue circles in panel b (left) denote diffraction from nanosheet I and II. Note that all the SAED patterns in the same line (a, b, and c) are scaled to the same display value range.

the diffraction spots remained clearly visible even after significant exposure to electrons, approximately $30 \times 10^3 \text{ e}^- \text{ Å}^{-2}$. The in-plane distance (d) of 1L Mt was indicated by yellow lines. A yellow dashed line, corresponding to the d value of 2.36 Å , was observed only for 1L Mt. The d values summarized in Table 1 are in good agreement with the values of ($hk0$) facets reported elsewhere.^{33–35} Fig. 2b and c show series of SAED images for 2L and 3L Mt at initial exposure and cumulative electron doses of 3, 15, and $30 \times 10^3 \text{ e}^- \text{ Å}^{-2}$. The red and blue circles in Fig. 2b denote diffraction from nanosheet I and II which was distinguished by comparing the pattern of 1L Mt nanosheet. Some of the spots corresponding to the d value of 1.76 Å in the two nanosheets were overlapped, and they were excluded from the intensity analysis discussed later. In contrast to the results of 1L Mt, the diffraction patterns of the 2L and 3L Mt nanosheets show an obvious change after a certain amount of exposure, from appearing as sharp spots to blurred rings, indicating the loss of their crystallinity.

Table 1. The in-plane distance (d) of 1L Mt determined by SAED at initial exposure (Fig. 2a, left). The d values are in good agreement with the values of ($hk0$) facets reported previously.³⁰

($hk0$)	$d/\text{Å}$
11, 02	4.68
13, 20	2.69
22, 04	2.36
31, 15, 24	1.76
33, 06	1.56
26, 40	1.34
35, 17, 42	1.29

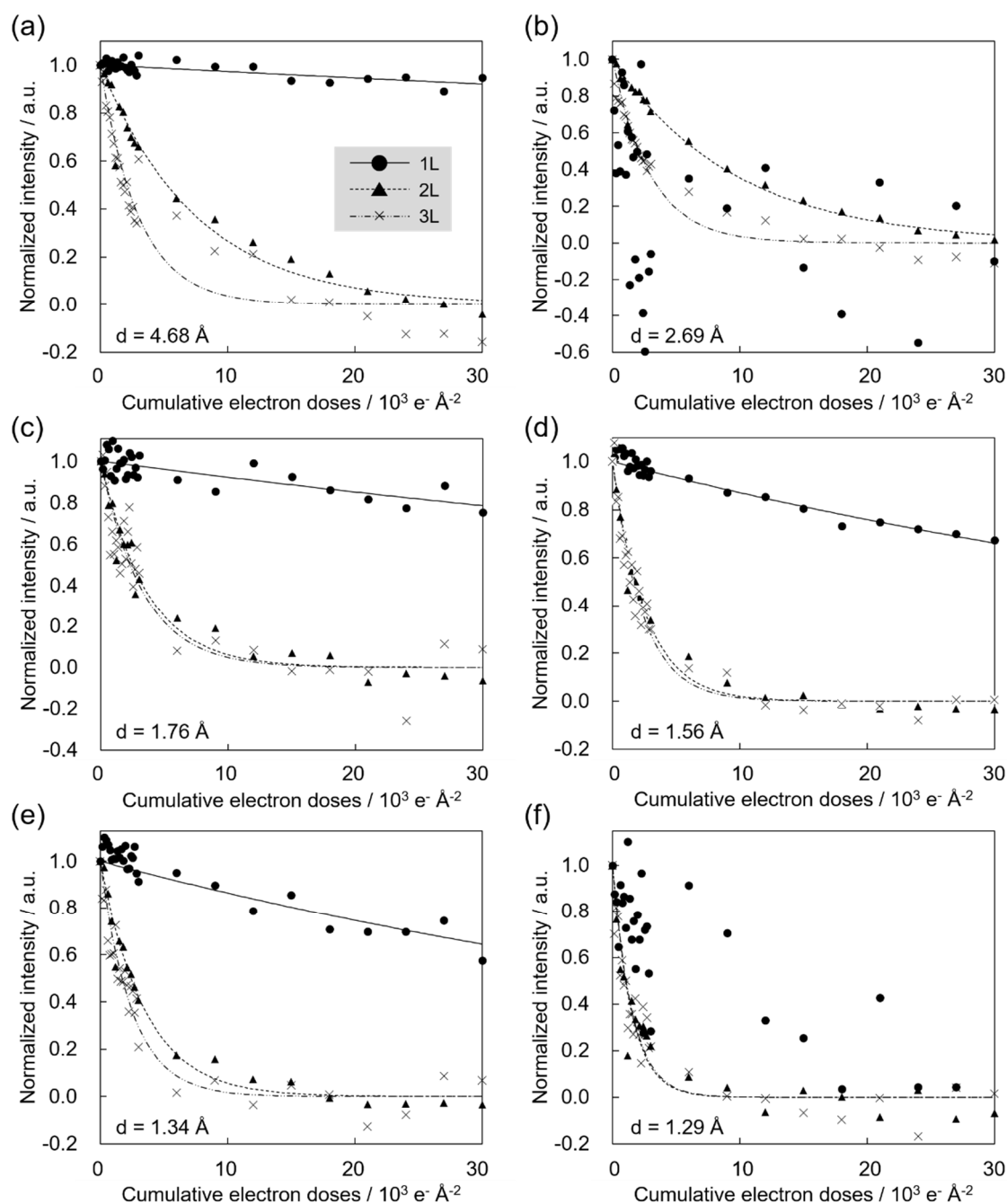


Fig. 3. Relative intensity decay for the d-spacing of (a) 4.68, (b) 2.69, (c) 1.76, (d) 1.56, (e) 1.34, (f) 1.29 Å, corresponding to the diffraction from (11, 02); (13, 20); (31, 15, 24); (33, 06); (26, 40); and (35, 17, 42) facets as a function of the cumulative electron doses for 1L (●), 2L (▲), and 3L Mt (×), respectively. The acceleration voltage and the electron dose rate were 300 kV and $30 \times 10^3 \text{ e}^- \text{ Å}^{-2} \text{ s}^{-1}$, respectively.

The outer spots, corresponding to smaller atomic spacing such as 1.34 or 1.29 Å, reduced rapidly, indicating their sensitivity to the electron beam irradiation.

Fig. 3 shows normalized intensity decay for the d-spacing of (a) 4.68, (b) 2.69, (c) 1.76, (d) 1.56, (e) 1.34, (f) 1.29 Å, corresponding to the diffraction from (11, 02); (13, 20); (31, 15, 24); (33, 06); (26, 40); and (35, 17, 42) facets as a function of the cumulative electron doses for 1L (●), 2L (▲), 3L Mt (×),

respectively. Black lines in Fig. 3 indicate exponential decay curves determined by the least-squares method. The critical doses for each curve, determined as the point at which the electron diffraction intensities reduced to $1/e$ (corresponding to 37% reduction) of their initial value, are summarized in Table 2. Fig. 3 was normalized to compare the critical doses of different stacked number of Mt nanosheets. An exponential approximation was first performed and the obtained constant

Table 2. The critical doses for 1L, 2L, 3L Mt, determined by exponential approximation in Fig. 3.

	d / Å	Critical dose ($10^3 \text{ e}^- \text{ Å}^{-2}$)		
		1L	2L	3L
(a)	4.68	363.02	7.55	3.02
(b)	2.69	-	9.64	2.99
(c)	1.76	124.36	4.08	3.31
(d)	1.56	72.60	2.58	2.35
(e)	1.34	68.29	3.48	2.46
(f)	1.29	-	1.49	1.54

Table 3. The critical doses for the two component nanosheets (noted as nanosheet I and II) in 2L Mt as indicated by red and blue circles in Fig. 2b. The values were determined by exponential approximation in Fig. S1.

d / Å	Critical dose / ($10^3 \text{ e}^- \text{ Å}^{-2}$)	
	Nanosheet I	Nanosheet II
4.68	11.14	7.02
2.69	24.35	6.61
1.76	2.51	5.16
1.56	2.44	1.97
1.34	3.29	3.9
1.29	1.79	1.07

term was set to zero, the data set was then normalized. Some negative values were found for the data set with a large variation relative to the absolute intensities. exponential approximation was first performed and the obtained constant term was set to zero, the data set was then normalized. Some negative values were found for the data set with a large variation relative to the absolute intensities. Note that normalization and exponential curve fitting for the d-spacing of 2.69 and 1.29 Å of 1L Mt was not performed correctly due to the large variation of the intensity. All critical doses for 1L Mt were significantly larger than those for 2L or 3L Mt ones. The critical dose for 2L Mt, however, was in the same order as that of 3L Mt. According to the results of 1L Mt the outer spots corresponding to smaller atomic spacing were reduced rapidly, indicating their sensitivity to the electron beam irradiation. For 2L and 3L Mt, the relationship between the critical dose and the d-spacing value was no longer observable. Additionally, Table 3 shows the critical doses for the two component nanosheets (noted as nanosheet I and II) in 2L Mt as indicated by red and blue circles in Fig. 2b. Since there was no significant difference in the critical dose of nanosheets I and II, the ordering between

nanosheets I and II against the incident direction of electron beam could not be determined from SAED analysis (See Fig. S1 in the Supplementary Information for the entire decay curves). These results clearly demonstrate that the free-standing 1L Mt nanosheet showed unexpected, distinct stability compared with 2L or more layered Mt under electron beam irradiation.

Mechanism of electron beam damage on Mt and the role of the interlayer space.

In free-standing Mt nanosheets during the TEM observation, multiple electron beam damage mechanisms can proceed simultaneously. The dominant mechanism of beam damage in insulating materials such as clay minerals is radiolysis damage caused by inelastic scattering—the transfer of energy from the incident electron to the electrons in the specimen. Other mechanism includes local heating or knock-on damage. The local heating also proceed the damage rate due to low thermal conductivity of clay minerals. Additionally, any hydrogens in the sample readily stripped by electron beam irradiation at the acceleration voltage of 300 kV, and the emitted hydrogens with high kinetic energy can cause knock-on damage in the structure.

Recently we demonstrated the stable ADF-STEM imaging of free-standing monolayer Mt at the atomic scale, and concluded that radiolysis damage were effectively suppressed due to the thin structure of the 1L Mt (0.96 nm) compared to the mean free path of secondary electrons in the oxides (ca. 10–20 nm).^{36–38} Radiolysis damage is caused by secondary electrons produced by inelastic scattering, and the possibility of secondary electrons escaping from the sample surface into the vacuum depends on their depth from the surface.³⁶ Therefore, if the sample is an atomically thin film or small particulate matter, the beam sensitivity decreases as the secondary electrons excited from near-surface atoms escape from the specimen to vacuum without causing damage.^{37–39} The suppression of electron damage by lowering the dimension of the sample (from 3D to 2D or 1D) is becoming commonplace for direct imaging of beam-sensitive materials by electron microscopy.^{15,39}

Note that since the escaped secondary electrons are not compensated by the neighboring electrons quickly in insulator, the specimen can be positively charged. For ionic materials such as clay minerals, this positive charge can induce an electric field in the specimen, which results in image drift or distortion, structural change, mass transport, or Coulomb explosion in some case. To avoid such explosive damage, we optimized the dose rate as $10 \text{ e}^- \text{ Å}^{-2} \text{ s}^{-1}$, which is mild condition that allows us to track the destruction process by SAED pattern. At high dose rate we could not get the fine SAED pattern even at the first scanning, which indicating the specimen was already destructed (See Fig. S2 in the Supplementary Information). Such drastic destruction might be caused by Coulomb explosion at high dose rate.

In the present report, however, only the 1L Mt showed the remarkably high stability against to the electron beam irradiation, despite 2L and 3L samples being thin in comparison with the mean free path of secondary electrons.^{36–38} Also, the

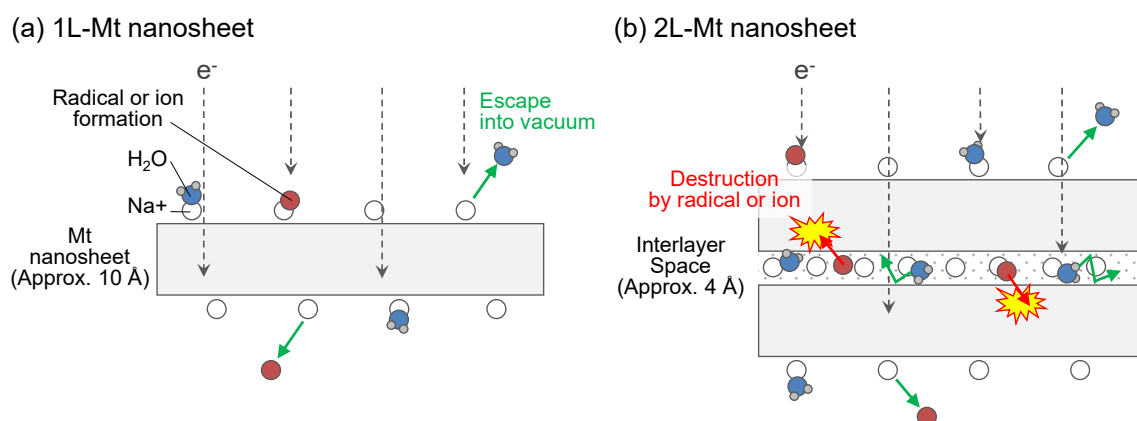


Fig. 4. A plausible schematic illustration for the influence of the existence of the interlayer space on the destruction of (a) 1L and (b) 2L Mt, representing samples that include interlayer space.

critical doses for 2L and 3L Mt were of the same magnitude. According to the results, the absence of the interlayer space governs the difference in stability between 1L and multilayered Mt rather than the thickness. All surfaces in the 1L Mt were exposed to vacuum, whereas multilayered Mt included interlayer spaces (minimum of 4 Å under high vacuum^{40,41}) as shown in Fig. 4a and b, respectively. Although multiple damage mechanism would proceed simultaneously as described above, the drastic difference of critical dose between 1L and 2L Mt can be plausibly explained by the effect of the interlayer space to hold water molecules. Generally, it is practically impossible to completely remove the water from clay minerals. Even after drying by heat or vacuum, certain amount of the water molecules remains in clay minerals due to their high activation energy for dehydration from counter cations under well-dried condition.^{42–44} Kuligiewicz et al. demonstrated that the activation energy of dehydration for water molecules which coordinate the interlayer cations is 90–180 kJ/mol, while that for bulk water molecules in the interlayer, not directly coordinated to the interlayer cations, is 45–60 kJ/mol.⁴⁵ Note that the hydration number of sodium ion is decreased to 0.5–0.6 under well-dried condition.⁴⁵ When these water molecules irradiated by electron beam, excited water molecules, and other molecular species (i.e. $\text{H}^\bullet$, $\text{OH}^\bullet$, $\text{HO}_2^\bullet$, OH^- , H_3O^+ , H_2 , H_2O_2) are generated,^{46,47} or desorption from the coordinated cations occurs. For the beam damage of zeolite, an aluminosilicate, the lower water content was associated with higher stability during electron microscopy.^{48,49} The generated ions or radical species decrease the strength of the Si–O or Al–O bonds, followed by a change in the structure. The same mechanism is true for the beam damage in clay minerals. During the observation in the experiment here, water molecules and the generated ions and radical species are hardly released from the interlayer spaces into the vacuum owing to the physical limitation of escape routes to the edges (Fig. 4b); conversely, the same is easy in the monolayered sample, whose surfaces are completely exposed to vacuum, resulting in remarkable stability of 1L Mt. The results in Table 3, wherein nanosheets I and II in the 2L sample showed the same critical dose, indicate isotropic destruction from the interlayer space.

Furthermore, it is known that Al–OH in the octahedral site is relatively unstable, and dehydroxylation in the neighboring Al–OH sites and associated amorphization occurs in clay minerals under high temperature or irradiation.^{45,50–52} The generated ions and radical species can pass through the six-membered rings of the Si–O tetrahedron and attack the Al–OH sites to promote dehydroxylation, and thus change the crystal structures. Although we were not able to distinguish the adjacent Si and Al atoms in monolayer Mt, ADF–STEM imaging may help monitor the evolution of the dehydroxylation process empirically, which has been previously studied only by thermogravimetry, X-ray diffraction, Fourier-transform infrared spectroscopy, or DFT calculations.^{37,50,51}

Conclusions

Herein, we demonstrated the remarkable stability of free-standing monolayer Mt against the beam irradiation using electron diffraction technique, while 2L and 3L Mt showed similar low stabilities. The results suggested that the stability of 1L Mt was owing to the absence of interlayer space, as a site for hydration, rather than the thinness of the structure. Although complete dehydration of clay minerals was not achieved in the multilayered Mt, the monolayer Mt with entirely exposed surfaces to the vacuum enables the efficient release of water molecules imparting much less beam damage. Although we focused on Mt in this work as a good model for experimental demonstration owing to the moderate anionic charge density and the large particle size up to hundreds of nanometers, the results here can be applied for other clay minerals in principle because the hydration is a common phenomenon of clay minerals. The moderate anionic charge density of Mt eases the exfoliations into individual layers in aqueous dispersion, and relatively larger particle sizes (~ hundreds nm) that is comparable to the illuminated area in TEM have an advantage on avoiding the quick beam damage from edges of layers. Other common clay minerals such as mica (too high charge density) or saponite (sizes of only ~ tens nm) are therefore not the good model for this purpose. By using Mt we can thus focus on the

dependence of electron beam damage on the layer number, i.e. the existence of the interlayer space for the first time. The findings here encourage further investigation using (S)TEM not only for the free-standing monolayer clay minerals but also for a wide variety of nanomaterials, such as the ones that are strongly hydrated in their interlayer spaces or nanopores.

Conflicts of interest

The authors declare no competing financial interest.

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